

Description of the Chemical Engineering BSc Program

Name and address of the institute: **University of Debrecen, H-4032 Debrecen, Egyetem square 1., Hungary**

Responsible faculty: **Faculty of Science and Technology**

Launching date: **Sep 1st, 2006**

Head of the Program: **Prof. Dr. László Somsák, Full Professor**

1. Name of the BSc Program: **Chemistry**

2. Acquired degree level and specialization:

- degree level: **BSc (baccalaureus, bachelor)**
- specialization: **Chemist**

3. Area of the Program: **science**

4. Duration of studies: **6 semesters**

5. Number of required ECTS credits: **180**

- orientation: **balanced (theory-practice 40-60 %)**
- credits assigned for the thesis: **15 credits**
- minimum credits of the optional courses: **9 credits**

6. Classification of the Program by the uniform classification system: 442

7. Objectives and Perspectives, acquired professional competences

The aim of this study programme is the training of chemists possessing theoretical and practical knowledge in chemistry as well as satisfactory basic knowledge in related fields of science (e. g. mathematics, physics, informatics, biology) and at least one foreign language. The degree holders will be able to apply their knowledge in recognizing and solving first of all practical problems in chemical industrial production, in analytical, agricultural, and quality assurance laboratories, as well as in various fields of administration and environmental protection. The Chemistry Bachelor will have in depth knowledge to continue his/her studies in the second (MSc) cycle and will be able to gain further knowledge either individually or in any organized manner.

The Chemistry Bachelor will

- be familiar with the principles and practical applicability of the most important methods in the chemical laboratory;
- be able to communicate efficiently the results of his/her work for both professional or non-experienced audience also in a foreign language by using multimedia methods;
- be able to acquire new competences by way of further education.

The Chemistry Bachelor will be qualified for

- identification and independent solution of mainly practical problems and tasks in the production process of chemical industry, in academic and industrial research institutions, in laboratories with agrochemical, food industry, crop protection, quality assurance, medical analytical profiles;
- performing daily operation of chemical apparatus, carrying out routine measurements;
- responsible operation of large laboratory instruments;
- individual decision making in his/her special field;
- doing his/her job in a quality-conscious, success-oriented, and value-committed way.

Outcomes: Subject Knowledge

Aspect of chemistry	Treated in module/course unit
a) Major aspects of chemical terminology, nomenclature, conventions, units	General Chem., Organic Chem. I.-V., Physical Chem. I.-II.
b) The major types of chemical reactions and the main characteristics associated with them	General Chem., Organic Chem. I.-V., Inorganic Chem. I.-II.
c) The principles and procedures used in chemical analysis and the characterisation of chemical compounds	Analytical Chem. I.-III., Separation Techniques, Spectroscopy, Organic Chem. III.
d) The principal techniques of structural investigations, including spectroscopy	General Chem., Analytical Chem. I.-III., Spectroscopy, Organic Chem. III., Structural Chem., Physics I.-II.
e) The characteristics of the different states of matter and the theories used to describe them	General Chem., Physical Chem. I., Physics I.-II., Colloid Chem. I., Structural Chem.
f) The principles of thermodynamics and their applications to chemistry	General Chem., Physical Chem. I.-II., Physics I.-II.
g) The principles of quantum mechanics and their application to the description of the structure and properties of atoms and molecules	General Chem., Structural Chem., Physics I.-II., Mathematical Methods in Chem.
h) The kinetics of chemical change, including catalysis; the mechanistic interpretation of chemical reactions	Physical Chem. II., Reaction Kinetics and Catalysis, Organic Stereochem. and Reaction Mechanisms
i) The characteristic properties of elements and their compounds, including group relationships and trends within the Periodic Table	General Chem., Inorganic Chem. I.-II.
j) The structural features of chemical elements and their compounds, including stereochemistry	Inorganic Chem. I.-II., Organic Stereochem. and Reaction Mechanisms
k) The properties of aliphatic, aromatic, heterocyclic and organometallic compounds	Organic Chem. I.-VI.
l) The nature and behaviour of functional groups in organic molecules	Organic Chem. I.-VI.
m) Major synthetic pathways in organic chemistry, involving functional group interconversions and carbon-carbon and carbon-heteroatom bond formation	Organic Chem. I.-VI.
n) The relation between bulk properties and the	Physical Chem. I.-II., Macromolecular Chem.

properties of individual atoms and molecules, including macromolecules (both natural and man-made), polymers and other related materials	I.-II., Biological Macromolecules
o) The structure and reactivity of important classes of biomolecules and the chemistry of important biological processes.	Organic Chem. II.-III., Biochem. I.-III.

Outcomes: Abilities and Skills

1. Chemistry-related cognitive abilities and competences	Treated in module/course unit
1.1 Ability to demonstrate knowledge and understanding of essential facts, concepts, principles and theories relating to the subject areas identified above.	General Chem., Inorganic Chem. I.-II., Organic Chem. I.-VI., Physical Chem. I.-II., Physics I.-II., Chemical Technology I.-II.
1.2 Ability to apply such knowledge and understanding to the solution of qualitative and quantitative problems of a familiar nature.	General Chem., Inorganic Chem. I.-II., Organic Chem. I.-VI., Physical Chem. I.-II., Physics I.-II.
1.3 Competences in the evaluation, interpretation and synthesis of chemical information and data.	Inorganic Chem. I.-II., Organic Chem. I.-VI., Physical Chem. I.-II.
1.4 Ability to recognise and implement good measurement science and practice.	Physical Chem. I.-II., Analytical Chem. I.-III., Physics I.-II., Project, Thesis
1.5 Competences in presenting scientific material and arguments in writing and orally, to an informed audience.	Project, Thesis
1.6 Computational and data-processing skills, relating to chemical information and data.	Basic Chem. Informatics, Chem. Informatics
2. Chemistry-related practical skills	
2.1 Skills in the safe handling of chemical materials, taking into account their physical and chemical properties, including any specific hazards associated with their use	Inorganic Chem. I., Organic Chem. IV.-VI., Phys. Chem I.-II., Analytical Chem. II.-III.
2.2 Skills required for the conduct of standard laboratory procedures involved and use of instrumentation in synthetic and analytical work, in relation to both organic and inorganic systems	Inorganic Chem. I., Organic Chem. IV.-VI., Analytical Chem. II.-III.
2.3 Skills in the monitoring, by observation and measurement, of chemical properties, events or changes, and the systematic and reliable recording and documentation thereof.	Physical Chem. I.-II., Analytical Chem. I.-III.
2.4 Ability to interpret data derived from laboratory observations and measurements in terms of their significance and relate them to appropriate theory	Physical Chem. I.-II., Physics I.-II.
2.5 Ability to conduct risk assessments concerning the use of chemical substances and laboratory procedures	Inorganic Chem. I., Analytical Chem. II.-III., Visits at Chem. Companies

3. Chemistry-related practical skills

From the total of 180 credits 82 are allocated to compulsory practical courses inclusive the bachelor thesis. These involve seminars to help the students in strengthening calculational, numerical and computational skills, analyzing various spectra (24 credits), as well as practicals in the laboratory (48 credits). The bachelor thesis (15 credits) is prepared on the basis of experimental work carried out in research laboratories. There will be possibilities to complete the bachelor thesis outside of the university, especially with chemical companies. In these instances a close cooperation between the supervisors from the university and the company will be expected

Generic Competences

3.1 The capacity to apply knowledge in practice, in particular problem-solving competences, relating to both qualitative and quantitative information.

Structural Chem., General Chem., Green Chem., Mathematics I.-II., Physical Chem. I.-II., Physics I.-II., Chem. Experiment Demonstrations, Organic Chem. I.-III.

3.2 Numeracy and calculation skills, including such aspects as error analysis, order-of-magnitude estimations, and correct use of units.

General Chem., Structural Chem., Introduction to Chem. Calculations, Green Chem., Mathematics I.-II., Physical Chem. I.-II., Physics I.-II., Analytical Chem. I.

3.3 Information-management competences, in relation to primary and secondary information sources, including information retrieval through on-line computer searches.

Structural Chem., Mathematical Methods in Chem., Chem. Programming, Physics I.-II., Organic Chem. VI.

3.4 Ability to analyse material and synthesise concepts.

Green Chem., Structural Chem., Physics I.-II., Spectroscopy

3.5 The capacity to adapt to new situations and to make decisions.

Biochem. II., Green Chem., Mathematics I.-II.

3.6 Information-technology skills such as word-processing and spreadsheet use, data-logging and storage, subject-related use of the Internet.

Chemical Programming, Basic Chem. Informatics, Chem. Informatics, Physical Chem. II.

3.7 Skills in planning and time management.

Biochem. II., Organic Chem. V.-VI., Project, Thesis

3.8 Interpersonal skills, relating to the ability to interact with other people and to engage in team-working.

Project, Thesis

3.9 Communication competences, covering both written and oral communication, in one of the major European languages (English, German, Italian, French, Spanish) as well as in the language of the home country.

Project, Thesis

3.10 Study competences needed for continuing professional development. These will include in particular the ability to work autonomously.

Chem. Concepts and Misconceptions, Green Chem., Chem. Programming, Physical Chem. I., Reaction Kinetics and Catalysis, Project, Thesis

3.11 Ethical commitment

Topics in Modern Chem., Project, Thesis, Chem. Programming (ethics of computer use)

8. Characteristics of the Program

Scientific disciplines comprising the study program:

Foundation in sciences: 15-25 credits

Chemistry (general chemistry min. 8 credits, inorganic chemistry min 12 credits, analytical chemistry min. 12 credits, organic chemistry min. 12 credits, physical chemistry min. 12 credits, applied chemistry min. 8 credits): 80-120 credits

Specialized professional knowledge in the fields of inorganic, analytical, organic and physical chemistry as well as from interdisciplinary fields: 15-65-credits

9. Bachelor Thesis

The bachelor thesis that terminates the studies is based on independent work of the student supervised by a staff member of the Institute of Chemistry (IC) and appears as a written description of the activities done and results achieved in an extent of 20-30 pages. The aim of the thesis is to prove that the student is able to collect and critically evaluate information, to set an aim and solve tasks to reach it, evaluate observations and results obtained in a branch of chemistry or chemistry related research field. The thesis work must be done in a research group of the IC or with a supervisor outside of the institute approved by the IC. The thesis should be discussed in a department's seminar before submission and must be presented and defended at the final exam.

Formal requirements of the thesis as well as the evaluation of the thesis by the supervisor are detailed in a manual available from the homepage of the IC.

10. Final Exam

The final examination is an oral exam of ~30 min to be absolved in front of a Final Examination Committee (FEC) appointed by the dean of the faculty. The supervisor of the student is member of the committee.

Preconditions to start the final exam:

Acquired absolutorium

Submitted thesis

Submitted evaluation sheet for the thesis by the supervisor, with a minimum grade of pass (2).

Parts and evaluation of the Final Exam

1. Presentation of the thesis (~5 min) – mark on the 1-5 scale
2. Defence of the thesis based on questions raised by the members of the FEC (~5 min) – mark on the 1-5 scale
3. Presentation of a topic related to the thesis work (a list of 4-6 items is provided by the supervisor and one of them is selected by a member of the FEC, ~5 min) – mark on the 1-5 scale
4. Demonstration of the knowledge of chemistry by presenting a topic from the basic chemical subjects (list of items available in the homepage of the IC, ~15 min) – mark on the 1-5 scale

The final exam evaluation mark comprises of two marks obtained as averages of part 1. and 2 as well as of 3. and 4.

11. Evaluation of the diploma

The diploma grade is determined by averaging the above two marks of the final exam together with the weighted average of the overall studies at the program.

Qualification of the diploma according to the Education and Examination Rules and Regulation of the University of Debrecen:

excellent	4,81 – 5,00
very good	4,51 – 4,80
good	3,51 – 4,50
satisfactory	2,51 – 3,50
pass	2,00 – 2,50

11. Master degrees offered primarily after finishing the Chemistry BSc

chemistry MSc

chemical engineering MSc

environmental engineering MSc

environmental science MSc

STRUCTURE OF THE CURRICULUM IN ECTS CREDITS

Compulsory Courses							
Modul <i>Blocks of courses (suggested credits)</i> Courses Codes– credits (cr) Lecturer	Semester (teaching hours: lectures + seminars + (laboratory) practice; type of examination: e:exam, p:practice, t: term grade, s: signature)						Prerequisites
	1.	2.	3.	4.	5.	6.	
Science							
<i>Mathematics</i>							
Mathematics I. TTMBE0808_EN – 5 cr TTMBG0808_EN – 2 cr Zoltán Muzsnay	4e+3p+0						None
Mathematics II. TTMBE0809_EN – 3 cr TTMBG0809_EN – 2 cr Zoltán Muzsnay		2e+3p+0					TTMBE0808_EN TTMBG0808_EN Mathematics I. (lecture and seminar)
<i>Physics</i>							
Physics for Engineers I. TTFBE2111_EN – 3 cr Balázs Ujvári	(2+1)e+0						None
Physics for Engineers II. TTFBE2113_EN – 3 cr Balázs Ujvári		(2+1)e+0					TTFBE2111_EN Physics for Engineers I.
<i>Informatics</i>							
Basic Chemical Informatics TTKBG0901_EN – 2 cr Mihály Purgel	0+2p+0						None
Chemical Informatics TTKBG0902_EN – 2 cr Ákos Kuki		0+2p+0					TTKBG0901_EN Basic Chemical Informatics
<i>Általános tárgyak</i>							
Basic Economics and Management TTTBE0010_EN – 1 cr Mária Újhelyi	1e+0+0						None
Quality Management TTTBE0020_EN – 1 cr György Deák	1e+0+0						None
History and Structure of the EU TTTBE0030_EN – 1 cr Károly Teperics	1e+0+0						None
Environmental Science TTTBE0040_EN – 1 cr István Gyulai	1e+0+0						None
Core Subjects							
<i>General Chemistry</i>							
General Chemistry I. (lecture and seminar) TTKBE0101_EN – 4 cr Gábor Bellér TTKBG0101_EN – 3 cr Norbert Lihi	3e+3p+0						None Paralel registration for both courses is required
General Chemistry II. (laboratory practice) TTKBL0101_EN – 3 cr Norbert Lihi		0+0+3p					TTKBE0101_EN and TTKBG0101_EN General Chemistry I.

Compulsory Courses							
Modul <i>Blocks of courses</i> <i>(suggested credits)</i> Courses Codes – credits (cr) <i>Lecturer</i>	Semester (teaching hours: lectures + seminars + (laboratory) practice; type of examination: e:exam, p:practice, t: term grade, s: signature)						Prerequisites
	1.	2.	3.	4.	5.	6.	
<i>Inorganic Chemistry</i>							
Inorganic Chemistry I. TTKBE0201_EN – 3 cr <i>István Lázár</i>		2e+0+0					TTKBE0101_EN General Chemistry I.
Inorganic Chemistry II. TTKBE0202_EN – 3 cr <i>Péter Buglyó</i>			2e+0+0				TTKBE0201_EN Inorg. Chem. I. TTKBE0301_EN Org. Chem. I. TTKBE0401_EN Phys Chem. I.
Inorganic Chemistry III. TTKBL0201_EN – 5 cr <i>Péter Buglyó</i>			0+(1+4)p				TTKBL0101_EN Gen. Chem. II. TTKBE0201_EN Inorg. Chem. I. TTKBE0301_EN Org. Chem. I. TTKBE0401_EN Phys Chem. I.
Inorganic Chemistry IV. TTKBL0202_EN – 4 cr <i>Péter Buglyó</i>				0+(1+3)p			TTKBE0202_EN Inorg. Chem. II. TTKBL0201_EN Inorg. Chem. III.
<i>Physical Chemistry</i>							
Physical Chemistry I. (lecture and seminar) TTKBE0401_EN – 3 cr TTKBG0401_EN – 2 cr <i>Attila Bényei</i>		2e+2p+0					TTKBE0101_EN Gen. Chem. I. TTMBE0808_EN Mathematics I. TTFBE2111_EN Phys. for Eng. I.
Physical Chemistry II. (lecture and seminar) TTKBE0402_EN – 3 cr TTKBG0402_EN – 2 cr <i>Attila Bényei</i>			2e+2p+0				TTKBE0201_EN Inorg. Chem. I. TTKBE0301_EN Org. Chem. I. TTKBE0401_EN Phys. Chem. I.
Introduction to Physical Chemistry Measurements TTKBL0401_EN – 4 cr <i>Ferenc Krisztián Kálmán</i>			0+0+4p				TTKBL0101_EN Gen. Chem. II. TTKBE0201_EN Inorg. Chem. I. TTKBE0301_EN Org. Chem. I. TTKBE0401_EN Phys Chem. I.
Physical Chemistry III. TTKBE0403_EN – 3 cr <i>Noémi Nagy</i>				2e+0+0			TTKBE0402_EN Phys. Chem. II.
Physical Chemistry IV. TTKBE0404_EN – 5 cr <i>Attila Bényei</i>					(2+2)e+0		TTKBE0402_EN Phys. Chem. II.
Physical Chemistry V. TTKBL0402_EN – 5 cr <i>Ferenc Krisztián Kálmán</i>					0+0+4p		TTKBE0402_EN Phys. Chem. II. TTKBL0401_EN Intr. Phys. Chem. Meas.

Compulsory Courses							
Modul <i>Blocks of courses</i> <i>(suggested credits)</i> Courses Codes – credits (cr) <i>Lecturer</i>	Semester (teaching hours: lectures + seminars + (laboratory) practice; type of examination: e:exam, p:practice, t: term grade, s: signature)						Prerequisites
	1.	2.	3.	4.	5.	6.	
<i>Organic Chemistry</i>							
Organic Chemistry I. TTKBE0301_EN – 4 cr <i>László Juhász</i>		(2+1)e+0					TTKBE0101_EN Gen. Chem. I.
Organic Chemistry II. TTKBE0302_EN – 4 cr <i>László Juhász</i>			(2+1)e+0				TTKBE0201_EN Inorg. Chem. I. TTKBE0301_EN Org. Chem. I. TTKBE0401_EN Phys Chem. I.
Organic Chemistry III. TTKBE0303_EN – 3 cr <i>László Juhász</i>				2e+0+0			TTKBE0302_EN Org. Chem. II.
Organic Chemistry IV. TTKBG0301_EN – 1 cr <i>Attila Mándi</i> TTKBL0301_EN – 4 cr <i>Marietta Vágvölgyi Tóth</i>				0+1p+4p			TTKBL0101_EN Gen. Chem. II. TTKBE0302_EN Org. Chem. II.
Organic Chemistry V. TTKBL0302_EN – 7 cr <i>Marietta Vágvölgyi Tóth</i>					0+(2+4)p		TTKBL0301_EN Org. Chem. IV.
Biochemistry I. TTBBE2035_EN – 3 cr <i>Kerégyártó János</i>					2e+0+0		TTKBE0303_EN Org. Chem. III.
Biochemistry II. (laboratory practice) TTKBL0303_EN – 3 cr <i>Gyöngyi Gyémánt</i>						0+(1+2)p	TTBBE2035_EN Biochemistry I.
<i>Analytical Chemistry</i>							
Analytical Chemistry I. TTKBE0501_EN – 3 cr <i>Péter Buglyó</i>							TTKBE0201_EN Inorg. Chem. I. TTKBE0301_EN Org. Chem. I. TTKBE0401_EN Phys Chem. I.
TTKBG0501_EN – 2 cr <i>Csilla Kállay</i>			2e+2p+4p				TTKBG0101_EN Gen. Chem. I. and paralel registration to TTKBE0501_EN Anal. Chem. I.
TTKBL0501_EN – 4 cr <i>Csilla Kállay</i>							TTKBL0101_EN Gen. Chem. II. and paralel registration to TTKBE0501_EN Anal. Chem. I.
Separation Techniques I. TTKBE0502_EN – 1 cr <i>István Lázár</i>			1e+0+0				TTKBE0201_EN Inorg. Chem. I. TTKBE0301_EN Org. Chem. I. TTKBE0401_EN Phys. Chem. I.
Separation Techniques II. TTKBL0502_EN – 3 cr <i>István Lázár</i>				0+0+3p			TTKBL0201_EN Inorg. Chem. III. TTKBE0502_EN Separ. Tech. I.

Compulsory Courses							
Modul <i>Blocks of courses</i> <i>(suggested credits)</i> Courses Codes – credits (cr) <i>Lecturer</i>	Semester						Prerequisites
	(teaching hours: lectures + seminars + (laboratory) practice; type of examination: e:exam, p:practice, t: term grade, s: signature)						
	1.	2.	3.	4.	5.	6.	
Analytical Chemistry II. TTKBL0503_EN – 6 cr <i>Attila Gáspár</i>				0+0+6p			TTKBE0501_EN Anal. Chem. I. TTKBL0501_EN Anal. kém. I. (lab)
Spectroscopic Methods I. TTKBE0503_EN – 3 cr <i>Gyula Batta</i>				2e+0+0			TTKBE0302_EN Org. Chem. II. TTFBE2113_EN Phys. for Eng. II.
Spectroscopic Methods II. TTKBL0504_EN – 4 cr <i>László Juhász</i>						0+3p+0	TTKBE0503_EN Spectr. Meth. I.
<i>Applied Chemistry (≥12)</i>							
Chemical Technology I. TTKBE0601_EN – 3 cr TTKBG0601_EN – 1 cr <i>Lajos Nagy</i>				2e+1p+0			TTKBE0201_EN Inorg. Chem. I. TTKBE0301_EN Org. Chem. I. TTKBE0401_EN Phys. Chem. I.
Chemical Technology II. TTKBE0602_EN – 4 cr TTKBG0602_EN – 2 cr <i>Lajos Nagy</i>					3e+2p+0		TTKBE0601_EN TTKBG0601_EN Chem. Tech. I.
Macromolecular Chemistry TTKBE0611_EN – 3 cr TTKBG0611_EN – 1 cr <i>Sándor Kéki</i>						2e+1p+0	TTKBE0302_EN Org. Chem. II.
Environmental Technology TTKBE1114_EN – 3 cr <i>Dávid Rácz</i>						2e+0+0	TTKBE0602_EN TTKBG0602_EN Chem. Tech. II.
Special courses							
<i>Practical courses</i>							
Visit in Chemical Industries TTK BX0607_EN <i>Ákos Kuki</i>				1 week(s)			Parallel registration to TTKBE0601_EN Chem. Tech. I.
<i>Closing Block</i>							
BSc Thesis I. TTKBL0001_EN – 5 cr <i>László Somsák</i>					0+(2+3)p		Minimum of 110 fulfilled credits + courses determined by the supervisor
BSc Thesis II. TTKBL0002_EN – 10 cr <i>László Somsák</i>						0+0+10p	TTKBL0001_EN BSc Thesis I.

Optional Chemistry Courses (6 credits)							
Modul <i>Blocks of courses</i> <i>(suggested credits)</i> Courses Codes – credits (cr) <i>Lecturer</i>	Semester						Prerequisites
	(teaching hours: lectures + seminars + (laboratory) practice; type of examination: e:exam, p:practice, t: term grade, s: signature)						
	1.	2.	3.	4.	5.	6.	
Crystallography TTGBE5104_EN – 3 cr <i>Gábor Dobosi</i>	2e+0+0 fall semester						None
History of Chemistry TTKBE0007_EN – 3 cr <i>Ágnes Dávid</i>	2e+0+0 spring semester						TTKBE0101_EN Gen. Chem. I.

Optional Chemistry Courses (6 credits)							
Modul Blocks of courses (suggested credits) Courses Codes– credits (cr) Lecturer	Semester (teaching hours: lectures + seminars + (laboratory) practice; type of examination: e:exam, p:practice, t: term grade, s: signature)						Prerequisites
	1.	2.	3.	4.	5.	6.	
Special and Dangerous Materials TTKBE0204_EN – 3 cr <i>István Lázár</i>			2e+0+0 fall semester				TTKBE0201_EN Inorg. Chem. I. TTKBE0301_EN Org. Chem. I. TTKBE0401_EN Phys. Chem. I.
Environmental Chemistry I. TTKBE0417_EN – 3 cr <i>Mónika Kéri</i>			2e+0+0 fall semester				TTKBE0201_EN Inorg. Chem. I. TTKBE0301_EN Org. Chem. I. TTKBE0401_EN Phys. Chem. I.
The Basics of Liquid Chromatography – Pharmaceutical Application TTKBE0310_EN – 3 cr <i>László Krusper</i>				2e+0+0			TTKBE0501_EN Anal. Chem. I.
Computational Quantum Chemistry TTKBG0903_EN – 3 cr <i>Mihály Purgel</i>				0+2p+0 spring semester			TTMBE0809_EN TTMBG0809_EN Mathematics II. (lect.+ sem.) TTKBG0901_EN Basic. Chem. Inf.
Process Control I. TTKBG0612_EN – 4 cr <i>István Árpád</i>				(2+1)t+0 spring semester			TTKBG0902_EN Chem. Informatics
Unit Operations I. TTKBG0614_EN – 6 cr <i>Miklós Nagy</i>			(2+3)t+0 fall semester				TTKBE0201_EN Inorg. Chem. I. TTKBE0301_EN Org. Chem. I. TTKBE0401_EN Phys. Chem. I.
Unit Operations II. TTKBG0615_EN – 6 cr <i>Miklós Nagy</i>				(2+3)t+0 spring semester			TTKBG0614_EN Unit Operations I.
Unit Operations III. TTKBE0616_EN – 6 cr <i>Miklós Nagy</i>					(2+3)e+0		TTKBG0615_EN Unit operations II.
Applied Radiochemistry TTKBE0504_EN – 3 cr <i>Noémi Nagy</i>					2e+0+0		TTKBE0403_EN Phys. Chem. III.
NMR Operator Training I. TTKBL0004_EN – 2 cr <i>Gyula Batta</i>					0+0+2p		TTKBE0503_EN Spectr. Meth. I.
Biochemistry III. TTBBE0304_EN – 3 cr <i>Teréz Barna</i>						2k+0+0	TTBBE2035_EN Biokémia I.
Biocolloids TTKBE0405_EN – 3 cr <i>Levente Novák</i>			2e+0+0 spring semester				TTKBE0402_EN Phys. Chem. II.
Colloid Chemistry TTKBE0415_EN – 3 cr <i>Levente Novák</i>						2e+0+0	TTKBE0403_EN Phys. Chem. III.
Plastics and Processing II. TTKBE1213_EN – 2 cr <i>György Deák</i>						0+2p+0	TTKBE0603_EN Macromol. Chem.

Optional Chemistry Courses (6 credits)							
Modul <i>Blocks of courses</i> <i>(suggested credits)</i> Courses Codes – credits (cr) <i>Lecturer</i>	Semester (teaching hours: lectures + seminars + (laboratory) practice; type of examination: e:exam, p:practice, t: term grade, s: signature)						Prerequisites
	1.	2.	3.	4.	5.	6.	
Organic Chemistry Seminar I. TTKGB0311_EN -1 cr László Juhász		0+1p+0					TTKBE0101_EN Gen. Chem. I.
Organic Chemistry Seminar II. TTKGB0312_EN -1 cr László Juhász			0+1p+0				TTKBE0201_EN Inorg. Chem. I. TTKBE0301_EN Org. Chem. I. TTKBE0401_EN Phys. Chem. I.
Advanced Organic Chemistry Seminar TTKGB0313_EN -2 cr László Juhász				0+2p+0			TTKBE0302_EN Org. Chem. II.

Title of course: Mathematics I. Code: TTMBE0808	ECTS Credit points: 5
Type of teaching, contact hours - lecture: 4 hours/week - practice: - - laboratory: -	
Evaluation: exam	
Workload (estimated), divided into contact hours: - lecture: 56 hours - practice: - - laboratory: - - home assignment: 44 hours - preparation for the exam: 50 hours Total: 150 hours	
Year, semester: 1 st year, 1 st semester	
Its prerequisite(s): -	
Further courses built on it: TTMBE0809_EN, TTMBG0809_EN	

Topics of course
Sets. Real numbers. Complex numbers. Sequences and series. Convergence, limits. Real functions. Limit, continuity and differentiation of functions. Monotonicity, convexity, inflection. Approximation with polynomials, Taylor formula. Definition and calculation of definite, indefinite and improper integrals. Ordinary differential equations. Vector spaces. Matrices, operations with matrices. Determinants and properties; the matrix rank. Linear equation systems. Euclidean spaces and their transformations.
Literature
<i>Compulsory:</i> - <i>Recommended:</i> Thomas, Weir & Hass: Thomas' Calculus, K. A. Stroud: Calculus and Mathematical Analysis, K. A. Stroud: Engineering Mathematics, E. Mendelson: Schaum's 3000 Solved Problems in Calculus,
Course objective/intended learning outcomes
a) Knowledge - He/She has knowledge to solve problems on the field of natural processes, using natural sources, and understanding the chemical background of living and non-living systems. - He/She has the knowledge to test or measure chemical reactions, systems with scientific methods (including computational) under supervision.
b) Abilities

- He/She is able to use the previously obtained knowledge on the field of natural and anthropogenic studies to solve practical problems.

c) Attitude

- He/She is ready to discuss problems on the field of chemistry and other science with professionals.

d) Autonomy and responsibility

- He/She stands for his/her opinion or ideology in professional discussions.

Schedule:

1st week

Operations with sets, set algebra. Descartes product, relations, functions. Special functions: injectivity, surjectivity, bijectivity. The inverse of a function. Real numbers. Exact lower and upper bounds. Open and closed sets. Bolzano-Weierstrass theorem.

2nd week

Complex numbers. The algebraic structure of the set of complex numbers. The complex plane. Trigonometric form of complex numbers, multiplication, division, n-th power, n-th root.

3rd week

sequences. Convergence and limit of real sequences. Monotonous, bounded, convergent sequences, Cauchy's convergence criteria. Algebraic operations with convergent sequences. Squeezing theorem. The generalization of the notion of limit.

4th week

Series. The convergence of series. Arithmetic series and geometric series. The harmonic series. Leibniz type series. Convergence tests: ratio and root tests. Power series.

5th week

Limits and continuity of functions. Properties of continuous functions. Continuity of the composition and the inverse function. Special properties of continuous functions defined on an interval. Elementary functions.

6th week

Differentiation. The geometric meaning of the derivative. Rules of differentiation. Derivative of a function of a function: the chain rule. The derivative of the inverse function. Relationship of monotonicity and the derivative. Roll's theorem and Lagrange's theorem. Conditions for the existence of extreme values. Derivative of elemental functions.

7th week

Higher order derivatives. Convexity and the derivatives. Approximating with polynomials, Taylor formula. Conditions for the existence of extreme value.

8th week

Primitive functions, the indefinite integral. Integration methods. Definite integral. Basic properties of the definite integrals. Integration of a continuous functions. The Newton-Leibniz formula.

9th week

Improper integrals. Applications.

10th week

Ordinary differential equations. The solution of separable, homogeneous and linear differential equations.

11th week

Vector space. Linear dependent and independent system of vectors. Base, dimension. Subspace. Vector space generated by a set of vectors. Rank of a system. Linear maps.

12th week

Matrices, matrix algebra. Determinants and their calculation. The rank of a matrix. The inverse of a matrix. Matrix representation of linear maps.

13th week

System of linear equations. Homogeneous and inhomogeneous systems. Gauss elimination, Cramer rule. Applications.

14th week

Euclidean spaces. Inner product, standard, angle, distance. Schwarz and Minkowski inequality. Orthogonality. Orthogonal projection. Symmetrical and orthogonal transformations.

Requirements:

Only students who have the grade from the practical part can take part of the exam. The exam is written. The grade is given according to the following table:

Score	Grade
0-49	fail (1)
50-62	pass (2)
63-74	satisfactory (3)
75-86	good (4)
87-100	excellent (5)

Person responsible for course: Dr. Zoltán Muzsnay, associate professor, PhD

Lecturer: Dr. Zoltán Muzsnay, associate professor, PhD

Title of course: Mathematics I.

Code: TTMBG0808_EN

ECTS Credit points: 3

Type of teaching, contact hours

- lecture: -
- practice: 3 hours/week
- laboratory: -

Evaluation: exam

Workload (estimated), divided into contact hours:

- lecture: -
- practice: 42 hours
- laboratory: -
- home assignment: 48 hours
- preparation for the exam:

Total: 90 hours

Year, semester: 1st year, 1st semester

Its prerequisite(s): -

Further courses built on it: TTMBE0809_EN, TTMBG0809_EN

Topics of course

Sets. Real numbers. Complex numbers. Sequences and series. Convergence, limits. Real functions. Limit, continuity and differentiation of functions. Monotonicity, convexity, inflection. Approximation with polynomials, Taylor formula. Definition and calculation of definite, indefinite and improper integrals. Ordinary differential equations. Vector spaces. Matrices, operations with matrices. Determinants and properties; the matrix rank. Linear equation systems. Euclidean spaces and their transformations.

Literature

Compulsory: -

Recommended:

Thomas, Weir & Hass: Thomas' Calculus,

K. A. Stroud: Calculus and Mathematical Analysis,

K. A. Stroud: Engineering Mathematics,

E. Mendelson: Schaum's 3000 Solved Problems in Calculus,

Course objective/intended learning outcomes

a) Knowledge

- He/She has knowledge to solve problems on the field of natural processes, using natural sources, and understanding the chemical background of living and non-living systems.

- He/She has the knowledge to test or measure chemical reactions, systems with scientific methods (including computational) under supervision.

b) Abilities

- He/She is able to use the previously obtained knowledge on the field of natural and anthropogenic studies to solve practical problems.

c) Attitude

- He/She is ready to discuss problems on the field of chemistry and other science with professionals.

d) Autonomy and responsibility

- He/She stands for his/her opinion or ideology in professional discussions.

Schedule:

1st week

Operations with sets, set algebra. Descartes product, relations, functions. Special functions: injectivity, surjectivity, bijectivity. The inverse of a function. Real numbers. Exact lower and upper bounds. Open and closed sets. Bolzano-Weierstrass theorem.

2nd week

Complex numbers. The algebraic structure of the set of complex numbers. The complex plane. Trigonometric form of complex numbers, multiplication, division, n-th power, n-th root.

3rd week

sequences. Convergence and limit of real sequences. Monotonous, bounded, convergent sequences, Cauchy's convergence criteria. Algebraic operations with convergent sequences. Squeezing theorem. The generalization of the notion of limit.

4th week

Series. The convergence of series. Arithmetic series and geometric series. The harmonic series. Leibniz type series. Convergence tests: ratio and root tests. Power series.

5th week

Limits and continuity of functions. Properties of continuous functions. Continuity of the composition and the inverse function. Special properties of continuous functions defined on an interval. Elementary functions.

6th week

Differentiation. The geometric meaning of the derivative. Rules of differentiation. Derivative of a function of a function: the chain rule. The derivative of the inverse function. Relationship of monotonicity and the derivative. Roll's theorem and Lagrange's theorem. Conditions for the existence of extreme values. Derivative of elemental functions.

7th week

Higher order derivatives. Convexity and the derivatives. Approximating with polynomials, Taylor formula. Conditions for the existence of extreme value.

8th week

Test.

Primitive functions, the indefinite integral. Integration methods. Definite integral. Basic properties of the definite integrals. Integration of a continuous functions. The Newton-Leibniz formula.

9th week

Improper integrals. Applications.

10th week

Ordinary differential equations. The solution of separable, homogeneous and linear differential equations.

11th week

Vector space. Linear dependent and independent system of vectors. Base, dimension. Subspace. Vector space generated by a set of vectors. Rank of a system. Linear maps.

12th week

Matrices, matrix algebra. Determinants and their calculation. The rank of a matrix. The inverse of a matrix. Matrix representation of linear maps.

13th week

System of linear equations. Homogeneous and inhomogeneous systems. Gauss elimination, Cramer rule. Applications.

14th week

Test.

Requirements:

- for a signature

Participation at **practice classes** is compulsory. A student must attend the practice classes and may not miss more than three times during the semester. In case a student does so, the subject will not be signed and the student must repeat the course. Attendance at practice classes will be recorded by the practice leader. Being late is equivalent with an absence.

- for a grade

During the semester one test is written. The grade is given according to the following table:

Score	Grade
0-49	fail (1)
50-59	pass (2)
60-74	satisfactory (3)
75-84	good (4)
85-100	excellent (5)

Students can take a retake test in conformity with the EDUCATION AND

EXAMINATION RULES AND REGULATIONS.**Person responsible for course:** Dr. Zoltán Muzsnay, associate professor, PhD**Lecturer:** Dr. Zoltán Muzsnay, associate professor, PhD**Title of course:** Mathematics II.**Code:** TTMBE0809_EN**ECTS Credit points: 3****Type of teaching, contact hours**

- lecture: 2 hours/week
- practice: -
- laboratory: -

Evaluation: exam**Workload (estimated), divided into contact hours:**

- lecture: 28 hours
- practice: -
- laboratory: -
- home assignment: 22 hours
- preparation for the exam: 40 hours

Total: 90 hours

Year, semester: 1st year, 2st semester**Its prerequisite(s):** TTMBE0808_EN, TTMBG0808_EN**Further courses built on it:****Topics of course**

Functions of several variables. Limit value, continuity, differentiation. Total derivative, partial derivatives, directional derivative. Partial Differential Equations. Multiple Integral. Elements of vector analysis. Curves, surfaces. Vector Fields. Gradient, rotation, divergence. Line, surface and volume integrals. Stokes', Green's and Gauss' theorems. Probability. Conditional probability. Total probability theorem, Bayes' theorem. Independence of events. Random variables. Discrete and continuous random variables. Probability distribution, density function. Expected value, standard deviation. Elements of statistics.

Literature*Compulsory:* -*Recommended:****Thomas, Weir & Hass: Thomas' Calculus,******P. Sahoo: Probability and Mathematical Statistics******E. Mendelson: Schaum's 3000 Solved Problems in Calculus,*****Course objective/intended learning outcomes****a) Knowledge**

- He/She has knowledge to solve problems on the field of natural processes,

using natural sources, and understanding the chemical background of living and non-living systems.

- He/She has the knowledge to test or measure chemical reactions, systems with scientific methods (including computational) under supervision.

b) Abilities

- He/She is able to use the previously obtained knowledge on the field of natural and anthropogenic studies to solve practical problems.

c) Attitude

- He/She is ready to discuss problems on the field of chemistry and other science with professionals.

d) Autonomy and responsibility

- He/She stands for his/her opinion or ideology in professional discussions.

Schedule:

1st week

Rn: the n-dimensional Euclidean space. Sequences in R^n . Function of several variables with real and vector values.

2nd week

Limit and continuity of multivariable functions.

3rd week

Total derivative and partial derivatives of a multivariable functions. Chain rule. Inverse function theorem. The implicit function theorem.

4th week

Directional derivative. Gradient and its application. Extreme values of real functions of several variables.

5th week

Multiple integral. Calculation of multiple integral, successive integration. Integration in normal domains.

6th week

Partial differential equations and systems of differential equations. Basic definitions and examples. Some elementary examples and problems.

7th week

Elements of vector analysis. Curves, surfaces. Vector Fields. Gradient, rotation, divergence.

8th week

Line integral. Basic properties. Applications.

9th week

Surface integral. Volume integral. Basic properties. Stokes', Green's and Gauss' theorems.

10th week

Element of the probability theory. Conditional probability. Total probability theorem, Bayes' theorem. Independence of events.

11th week

Concept of random variables. Probability distribution. Discrete probability variables. Some special discrete probability distributions: Bernoulli distribution, Binomial distribution, Geometric distribution, Binomial, Hypergeometric, and Poisson distribution. Continuous probability distributions, density function. Some special continuous distribution: uniform, normal, and exponential distributions.

12th week

Expected value of random variables, Variance of random variables. Examples. Markov and

Chebychev inequality, the law of large numbers.

13th week

Two Random Variables. Bivariate discrete and continuous random variables. Covariance of bivariate random variables. Correlation and independence.

14th week

Element of statistics.

Requirements:

Only students who have the grade from the practical part can take part of the exam. The exam is written. The grade is given according to the following table:

Score	Grade
0-49	fail (1)
50-62	pass (2)
63-74	satisfactory (3)
75-86	good (4)
87-100	excellent (5)

Person responsible for course: Dr. Zoltán Muzsnay, associate professor, PhD

Lecturer: Dr. Zoltán Muzsnay, associate professor, PhD

Title of course: Mathematics II.

Code: TTMBG0809_EN

ECTS Credit points: 3

Type of teaching, contact hours

- lecture: -
- practice: 3 hours/week
- laboratory: -

Evaluation: exam

Workload (estimated), divided into contact hours:

- lecture: -
- practice: 42 hours
- laboratory: -
- home assignment: 48 hours
- preparation for the exam:

Total: 90 hours

Year, semester: 1st year, 2st semester

Its prerequisite(s): TTMBE0808_EN, TTMBG0808_EN

Further courses built on it:

Topics of course

Functions of several variables. Limit value, continuity, differentiation. Total derivative, partial derivatives, directional derivative. Partial Differential Equations. Multiple Integral. Elements of vector analysis. Curves, surfaces. Vector Fields. Gradient, rotation, divergence. Line, surface and volume integrals. Stokes', Green's and Gauss' theorems. Probability. Conditional probability. Total probability theorem, Bayes' theorem. Independence of events. Random variables. Discrete

and continuous random variables. Probability distribution, density function. Expected value, standard deviation. Elements of statistics.

Literature

Compulsory: -

Recommended:

Thomas, Weir & Hass: Thomas' Calculus,

P. Sahoo: Probability and Mathematical Statistics

E. Mendelson: Schaum's 3000 Solved Problems in Calculus,

Course objective/intended learning outcomes

a) Knowledge

- He/She has knowledge to solve problems on the field of natural processes, using natural sources, and understanding the chemical background of living and non-living systems.

- He/She has the knowledge to test or measure chemical reactions, systems with scientific methods (including computational) under supervision.

b) Abilities

- He/She is able to use the previously obtained knowledge on the field of natural and anthropogenic studies to solve practical problems.

c) Attitude

- He/She is ready to discuss problems on the field of chemistry and other science with professionals.

d) Autonomy and responsibility

- He/She stands for his/her opinion or ideology in professional discussions.

Schedule:

1st week

Rn: the n-dimensional Euclidean space. Sequences in Rn. Function of several variables with real and vector values.

2nd week

Limit and continuity of multivariable functions.

3rd week

Total derivative and partial derivatives of a multivariable functions. Chain rule. Inverse function theorem. The implicit function theorem.

4th week

Directional derivative. Gradient and its application. Extreme values of real functions of several variables.

5th week

Multiple integral. Calculation of multiple integral, successive integration. Integration in normal domains.

6th week

Partial differential equations and systems of differential equations. Basic definitions and

examples. Some elementary examples and problems.

7th week

Test.

Elements of vector analysis. Curves, surfaces. Vector Fields. Gradient, rotation, divergence.

8th week

Line integral. Basic properties. Applications.

9th week

Surface integral. Volume integral. Basic properties. Stokes', Green's and Gauss' theorems.

10th week

Element of the probability theory. Conditional probability. Total probability theorem, Bayes' theorem. Independence of events.

11th week

Concept of random variables. Probability distribution. Discrete probability variables. Some special discrete probability distributions: Bernoulli distribution, Binomial distribution, Geometric distribution, Binomial, Hyper-geometric, and Poisson distribution. Continuous probability distributions, density function. Some special continuous distribution: uniform, normal, and exponential distributions.

12th week

Expected value of random variables, Variance of random variables. Examples. Markov and Chebychev inequality, the law of large numbers.

13th week

Two Random Variables. Bivariate discrete and continuous random variables. Covariance of bivariate random variables. Correlation and independence.

14th week

Test. Element of statistics.

Requirements:

- for a signature

Participation at **practice classes** is compulsory. A student must attend the practice classes and may not miss more than three times during the semester. In case a student does so, the subject will not be signed and the student must repeat the course. Attendance at practice classes will be recorded by the practice leader. Being late is equivalent with an absence.

- for a grade

During the semester one test is written. The grade is given according to the following table:

Score	Grade
0-49	fail (1)
50-59	pass (2)
60-74	satisfactory (3)
75-84	good (4)
85-100	excellent (5)

Students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

Person responsible for course: Dr. Zoltán Muzsnay, associate professor, PhD

Lecturer: Dr. Zoltán Muzsnay, associate professor, PhD

Title of course: Physics for Engineers I

ECTS Credit points: 3

Code: TTFBE2111_EN	
Type of teaching, contact hours	
<ul style="list-style-type: none"> - lecture: 2 hours/week - practice: 1 hours/week - laboratory: - 	
Evaluation: exam	
Workload (estimated), divided into contact hours:	
<ul style="list-style-type: none"> - lecture: 28 hours - practice: 14 hours - laboratory: - - home assignment: 24 hours - preparation for the exam: 24 hours <p>Total: 90 hours</p>	
Year, semester: 1 st year, 1 st semester	
Its prerequisite(s): –	
Further courses built on it: TTFBE2113_EN, TTKBE0401_EN, TTKBG0401_EN,	

Topics of course
Physical quantities, standards, units. Kinematics in one dimension. Kinematics in three dimensions. Dynamics. Force laws. Ballistic motions. Center of mass, constrained motion. Collisions. Work and energy. Oscillations. Elasticity. Wave motion. Temperature.
Literature
<i>J.W. Jewett Jr, R.A. Serway: Physics for Scientists and Engineers</i>
Course objective/intended learning outcomes
<p>a) Knowledge</p> <ul style="list-style-type: none"> - He/she knows the fundamental principles of classical mechanics and thermodynamics. - He/she knows the most important experiments of classical physics. <p>b) Abilities</p> <ul style="list-style-type: none"> - He/she is able to apply the most important laws of kinematics and dynamics in solving simple computational tasks. <p>c) Attitude</p> <ul style="list-style-type: none"> - He/she can accept the laws of classical physics. - He/she is open to improve his/her abilities in natural sciences as well as in other areas. <p>d) Autonomy and responsibility</p> <ul style="list-style-type: none"> - He/she continuously improves himself/herself. - He/she is realistic about his/her work and places it in the context of other colleagues' work.

Schedule:
1 st week
Physical quantities, standards, units: definition of length, equivalence relations and classes, scales,

standards of length, time and mass, basic and derived physical quantities, units and prefixes in SI, physical dimensions, dimensional analysis

2nd week

Kinematics in one dimension: Cartesian, spherical and cylindrical coordinate systems, vectors, operations with vectors, position vector, position function, average and instantaneous speed, average and instantaneous acceleration in one dimension

3rd week

Kinematics in three dimensions: displacement vector and path, average and instantaneous velocity, average and instantaneous acceleration in three dimensions, circular motion, tangential and normal acceleration, angular velocity, angular acceleration, relative motions, Galilean transformation, Coriolis acceleration

4th week

Dynamics: Newton's first law, inertial frames, experimental laws of two-body interactions, inertial mass, momentum, conservation of momentum, Newton's second law, Newton's third law

5th week

Force laws: basic interactions in nature, the role of force laws in equations of motion, force law of gravitation, force law of electrostatic interaction between two point charges, force law of a charged particle moving in magnetic field, force law of an idealized spring, force law of friction, force law of drag forces

6th week

Ballistic motions: analytic solution of the equation of motion near the surface of the Earth, describing the path, calculating the parameters of the special points of the path, numerical solution of the equation of motion near the surface of the Earth

7th week

Center of mass, constrained motion: center of mass defined in the discrete and in the continuum limit, density, internal and external forces, constrained motion on a slope, constrained motion of a pendulum

8th week

Collisions: describing collisions in the center-of-mass and in the laboratory frame, elastic and inelastic collisions, kinetic energy, collisions in one dimension, special cases of one-dimensional collisions

9th week

Work and energy: work, work-energy theorem, work of the gravitational pull of the Earth, work of an idealized spring, power, potential energy, conservation of total mechanical energy, conservative and dissipative forces, potential energy of a body under the influence of an idealized spring, potential energy of a body under the influence of gravitation

10th week

Oscillations: analyzing the motion of a pendulum, simple harmonic oscillations, addition of two simple harmonic oscillations, Lissajous figures, damped oscillations, forced oscillations, coupled oscillations

11th week

Elasticity: tensile stress, shearing stress, uniform compression, relative deformation, Young's modulus, shear modulus, compression modulus, Hooke's law, elastic energy, elastic energy density

12th week

Wave motion: mechanical waves, transverse and longitudinal waves, one-dimensional wave motion in a stretched string, wave speed, wave function, wave equation, harmonic waves, wavelength, wave number, time period, energy transports in wave motion, kinetic and potential energy density of an

elastic medium, energy density current, intensity

13th week

Wave motion: multi-dimensional wave motion, wavefronts, spherical waves, plane waves, principle of linear superposition, interference, coherent waves, standing waves, sound waves, intensity, pitch and tone, fundamental frequency and overtones, diffraction, Huygens' principle, Huygens–Fresnel principle

14th week

Temperature: extensive and intensive quantities, thermal equilibrium, zeroth law of thermodynamics, empirical measuring scales, Celsius scale, Kelvin scale, triple-point temperature, Gay-Lussac's law, constant-volume gas scales, ideal gas

Requirements:

The course exam is a written examination. In the exam theoretical questions and practical problems must be answered and solved in 100 minutes. The evaluation of the exam occurs based on the following grading:

0–49 % → 1,

50–62 % → 2,

63–75 % → 3,

76–88 % → 4,

89–100 % → 5

Person responsible for course: Dr. Balázs Ujvári, assistant professor, PhD

Lecturer: Dr. Balázs Ujvári, assistant professor, PhD

Title of course: Physics for Engineers II Code: TTFBE2113_EN	ECTS Credit points: 3
Type of teaching, contact hours - lecture: 2 hours/week - practice: 1 hours/week - laboratory: -	
Evaluation: exam	
Workload (estimated), divided into contact hours: - lecture: 28 hours - practice: 14 hours - laboratory: - - home assignment: 24 hours - preparation for the exam: 24 hours Total: 90 hours	
Year, semester: 1 st year, 2 nd semester	
Its prerequisite(s): TTFBE2111_EN	
Further courses built on it: TTKBE0503_EN	

Topics of course

Geometrical optics. Wave properties of light. Electrostatics. Gauss' law. Electric potential. Capacitors. Electric current. Direct current circuits. Magnetic field. Sources of magnetic field. Solenoids, displacement current. Induction. LC and RLC circuits. Electromagnetic waves.

Literature

J.W. Jewett Jr, R.A. Serway: Physics for Scientists and Engineers

Course objective/intended learning outcomes**a) Knowledge**

- He/she knows the fundamental principles of electrodynamics and modern physics.
- He/she knows the most important experiments leading to the birth of modern physics.

b) Abilities

- He/she is able to apply the most important laws of electrodynamics and modern physics in solving simple computational tasks.

c) Attitude

- He/she can accept the laws of modern physics.
- He/she is open to improve his/her abilities in natural sciences as well as in other areas.

d) Autonomy and responsibility

- He/she continuously improves himself/herself.
- He/she is realistic about his/her work and places it in the context of other colleagues' work.

Schedule:*1st week*

Geometrical optics: law of reflection, law of refraction, total reflection, imaging by concave and convex mirrors, imaging by a single spherical refractive surface, imaging by converging and diverging thin lenses, lense distortions

2nd week

Wave properties of light: coherent light waves, interference, diffraction, Young's double-slit experiment, thin-film interference, single-slit diffraction, diffraction gratings

3rd week

Electrostatics: electric charge, insulators, conductors and semi-conductors, Coulomb's law, electric field, field vector, field lines, electric field of a point charge, electric dipoles, linear, surface and volume charge distributions

4th week

Gauss' law: electric flux through open and closed surfaces, Gauss' law and its applications, electric field of a uniformly charged infinite line, electric field of a uniformly charged infinite plane, electric charge of a uniformly charged spherical volume

5th week

Electric potential: comparison of the force laws of gravitational and electrostatical interactions, work done by electric field, potential energy, potential energy of two-body and many-body systems, potential, potential due to a single point charge and charge distributions

6th week

Capacitors: parallel-plate, cylindrical and spherical capacitors, capacitance, energy and energy density stored by the electrostatic field, capacitors with dielectrics, equivalent capacitance of

capacitors connected in parallel and series

7th week

Electric current: electric current, electric current density, resistance, resistivity, conductivity, differential and integral form of Ohm's law, temperature dependence of resistivity, electric power

8th week

Direct current circuits: equivalent resistance of resistors connected in parallel and series, ideal and non-ideal batteries, electromotive force, Kirchhoff's junction law, Kirchhoff's loop law, transient phenomena in RC circuits

9th week

Magnetic field: magnetic field, field vector, field lines, electric charge moving in magnetic field, Lorentz's force, cyclotron, magnetic force acting on a current-carrying conductor

10th week

Sources of magnetic field: Biot–Savart law, magnetic field of a current-carrying straight wire, magnetic force between two parallel conductors, definition of the unit of electric current, Ampere's law

11th week

Solenoids, displacement current: magnetic field of a solenoid, magnetic flux through open and closed surfaces, Gauss' law of magnetism, displacement current, Ampere–Maxwell law

12th week

Induction: induced electromotive force, Faraday's law of induction, Lenz's law, eddy currents, self-induction, inductance, transient phenomena in RL circuits

13th week

LC and RLC circuits: energy conditions in LC circuits, analogy to free harmonic oscillations of a mechanical system, energy conditions in RLC circuits, analogy to damped oscillations of a mechanical system

14th week

Electromagnetic waves: differential and integral form of Maxwell's equations, linearly polarized plane electromagnetic waves

Requirements:

The course exam is a written examination. In the exam theoretical questions and practical problems must be answered and solved in 100 minutes. The evaluation of the exam occurs based on the following grading:

0–49 % → 1,

50–62 % → 2,

63–75 % → 3,

76–88 % → 4,

89–100 % → 5

Person responsible for course: Dr. Balázs Ujvári, assistant professor, PhD

Lecturer: Dr. Balázs Ujvári, assistant professor, PhD

Title of course: Basic Chemical Informatics

Code: TTKBG0901_EN

ECTS Credit points: 2

Type of teaching, contact hours

- lecture: -

<ul style="list-style-type: none"> - practice: 2 hours/week - laboratory: -
Evaluation: exam
Workload (estimated), divided into contact hours: <ul style="list-style-type: none"> - lecture: - - practice: 28 hours - laboratory: - - home assignment: 16 hours - preparation for the exam: 16 hours Total: 60 hours
Year, semester: 1 st year, 1 st semester
Its prerequisite(s): -
Further courses built on it: TTKBG0902_EN

Topics of course
<ul style="list-style-type: none"> - Basic knowledge about Microsoft Office program packages - Preparing documents and presentations - Using Instant JChem and Marvin program packages for drawing molecular structures - Learning basic Linux commands - Using WinSCP Commander and Putty programs
Literature
<i>Compulsory:</i> https://support.office.com/
<i>Recommended:</i> https://chemaxon.com/products/instant-jchem https://www.opensuse.org/
Course objective/intended learning outcomes
a) Knowledge He/She has the knowledge to test or measure chemical reactions, systems with scientific methods (including computational) under supervision.
b) Abilities He/She is able to evaluate and discuss the laboratory measurements, and create a report about it.
c) Attitude He/She is ready to discuss problems on the field of chemistry and other science with professionals. He/She is open toward scientific and other postgradual education.

d) Autonomy and responsibility

He/She can make reasonable evaluations about his/her own work comparing to others to the same field.

Schedule:

1st week

MS Word. Professional word processing I. Formatting, symbols, formula. Formatting and editing documents: subscripts and superscripts, Greek letters. Reproducing chemical formula and reactions by MS Word.

2nd week

2Professional word processing II. Tables, title, table of contents, cover page, margins, orientation, bullets and numbering, page/section breaks

3rd week

Professional word processing III. Equations, figures, literature, remarks, header and footer: letterhead, page numbered form

4th week

Professional word processing IV. Chemistry add in Word

5th week

Power Point Presentation 1.

6th week

Power Point Presentation 2.

7th week

Spreadsheet I. Formatting, data management

8th week

Spreadsheet II. *xy* diagrams, linear fitting

9th week

Spreadsheet III. *xy* diagrams, non-linear curves

10th week

Drawing programs. Marvin Sketch

11th week

Data base. Instant JChem

12th week

Simulating programs. GaussView

13th week

Linux commands.

14th week

Exam

Requirements:

- for a signature

Attendance is recommended, maximum 3 absences are accepted.

- for a grade

MS Word document and presentation prepared during the semester. (50%)

Exam of MS Excel and Linux commands. (50%)

The final grade is based on the sum of the exam and the MS Office document and presentation.

The final grade is given according to the following table:

Score (%)	Grade
0-49	fail (1)
50-59	pass (2)
60-74	satisfactory (3)
75-89	good (4)
90-100	excellent (5)

If the score of the final grade is below 50%, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

Person responsible for course: Dr. Mihály Purgel, assistant professor, PhD

Lecturer: Dr. Mihály Purgel, assistant professor, PhD

Title of course: Chemical Informatics Code: TTKBG0902_EN	ECTS Credit points: 2
Type of teaching, contact hours - lecture: - - practice: 2 hours/week - laboratory: -	
Evaluation: practice grade	
Workload (estimated), divided into contact hours: - lecture: - - practice: 28 hours - laboratory: - - preparation for the tests: 30 hours Total: 58 hours	
Year, semester: 1 st year, 2 nd semester	
Its prerequisite(s): TTKBG0901_EN	
Further courses built on it: TTKBG0612_EN	

Topics of course
Application of spreadsheets: mathematical operations, equations, charts, curve fitting, least-squares fitting, numerical integration, numerical derivation, solving of nonlinear equations, solving of set of equations, linear regression, matrix operations, introductions to statistics.
Literature
<i>Recommended:</i> 1. Joan Preppernau, Joyce Cox and Curtis Frye. Microsoft® Office Home and Student 2007 Step by Step, Microsoft Press, 2007 2. Robert de Levi. Advanced Excel® for scientific data analysis, Oxford University Press, New York, 2004

3. Robert de Levi. How to Use Excel® in Analytical Chemistry: And in General Scientific Data Analysis, Cambridge University Press, Cambridge, 2004

Course objective/intended learning outcomes

a) Knowledge

- He/She has the knowledge to test or measure chemical reactions, systems with scientific methods (including computational) under supervision.

b) Abilities

- He/She is able to evaluate and discuss the laboratory measurements, and create a report about it.

c) Attitude

- He/She is able to represent his/her own personal scientific ideology toward professional and unprofessional groups.

- He/She is well aware about his/her propositions and its consequences.

d) Autonomy and responsibility

- He/She can make reasonable evaluations about his/her own work comparing to others to the same field.

Schedule:

1st week

Implementation of mathematical functions in the spreadsheet software. Plotting the result in *xy* scatter graphs.

2nd week

Solving calculation problems in chemistry by implemented mathematical functions.

3rd week

Numerical differentiation by spreadsheet software and its application for problem-solving in chemistry.

4th week

Numerical integration by spreadsheet software and its application for problem-solving in chemistry.

5th week

Regression, curve fitting

6th week

The application of interpolation for problem-solving in chemistry.

7th week

Solving nonlinear equations by spreadsheet software and its application for problem-solving in chemistry.

8th week

Solving nonlinear set of equations by spreadsheet software and its application for problem-solving in chemistry.

9th week

Matrix operations

10th week

Solving sets of linear equations by matrix operations.

11th week

Application of spreadsheets in combinatorics and probability.

12th week

Application of spreadsheets in statistics. Probability distributions.

13th week

Maxwell–Boltzmann molecular speed distribution for gases. Typical speeds.

14th week

Application of t-tests for problem-solving in chemistry.

Requirements:

- *for a signature*

Participation at the classes is compulsory. A student must attend the classes and may not miss more than three times during the semester. In case a student does so, the subject will not be signed and the student must repeat the course. Attendance at the classes will be recorded by the practice leader. Being late is equivalent with an absence. In case of further absences, a medical certificate needs to be presented. Missed classes should be made up for at a later date, to be discussed with the tutor. Active participation is evaluated by the teacher in every class. If a student's behavior or conduct doesn't meet the requirements of active participation, the teacher may evaluate his/her participation as an absence because of the lack of active participation in class.

- *for the practice grade*

The course ends with a test in the 14th week. The minimum requirement for the test is 50%. The grade is given according to the following table:

Score	Grade
0-49	fail (1)
50-59	pass (2)
60-74	satisfactory (3)
75-89	good (4)
90-100	excellent (5)

The students are allowed to retake the test once to improve their scores. Further improvement is in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

Person responsible for course: Dr. Ákos Kuki, associate professor, PhD

Lecturer: Dr. Ákos Kuki, associate professor, PhD

Title of course: Basic Economics and Business

Code: TTTBE0010-K1_EN

ECTS Credit points: 1

Type of teaching, contact hours

- lecture: 1 hours/week
- practice: -
- laboratory: -

Evaluation: exam

Workload (estimated), divided into contact hours:

- lecture: 14 hours
- practice: -
- laboratory: -
- home assignment: -
- preparation for the exam: 16 hours

Total: 150 hours

Year, semester: 1st year, 1st semester

Its prerequisite(s): -

Further courses built on it: -

Topics of course

The course reviews the main areas that are essential to understand the operation of a business organization. In addition to reviewing the basic concepts of economics and management, it focuses on strategic management and planning, marketing, asset management, quality management, project management, organizational behaviour, human resource management and finance.

Literature

Compulsory:

Robbins, Stephen P. – Judge, Timothy A. (2016): Essentials of Organizational Behavior, thirteenth edition, Global edition. Pearson Education Limited, Essex, England

Foster, S. Thomas – Ganguly, Kunal K. (2007): Managing Quality. Integrating the Supply Chain. Pearson, Prentice Hall

Hill, Charles W. L. – Jones, Gareth R. – Schilling, Melissa A. (2015): Strategic Management Theory. 11th edition, Cengage Learning

Lecture notes provided by the lecturers

Course objective/intended learning outcomes

a) Knowledge

- He/she knows the basic economics and business concepts, get acquainted with the most important functional areas of business organizations, such as strategic management, marketing, organizational behaviour, human resource management and finance.

- He/she possesses the knowledge whose application is necessary to solve basic chemical problems related to natural processes, natural resources, living and inanimate systems.

b) Abilities

- He/she is aware of the most important economic and management principles and the basic concepts and contexts of different functional areas.

- He/she can locate certain areas of business organizations, their relationship with one another, and with economic principles.

- He/she is able to see the functioning of different organizations in their contexts and to identify the challenging areas.

- He/she can apply in practice the basic concepts, tools and methods, known.

- He/she is able to use the acquired knowledge in solving basic practical (chemical laboratory, chemical, environmental, management and environmental) problems. Can gather and interpret relevant data, based on which he/she can form a well-founded opinion on social, scientific, or ethical issues.

c) Attitude

- The course facilitates students, to develop positive attitude towards the efficient and effective functioning of organizations. Later, as an organizational actor he/she will be open to successfully implement tasks, related to business organizations.

- Through his/her approach, open to wider professional cooperation, receptive to newer chemical

aspects of economics. He/she consciously assumes the ethical standards of his/her profession.

d) Autonomy and responsibility

- The course helps students to be responsible for economic and management problem-solving and to help the effective and efficient operation of organizations with independent, initiative behaviour.
- He/she takes on the scientific world view with responsibility in professional discussions and debates.

Schedule:

1st week (two hour)

Fundamental management terms, theories. Organizations. Key elements and common organizational designs.

2nd week -

3rd week (two hour)

Organizational behaviour and leadership

4th week -

5th week (two hour)

Fundamental economics and business terms. Quality assurance, maintenance, project management, innovation management.

6th week -

7th week (two hour)

Strategic management, business planning and marketing.

8th week -

9th week (two hour)

The most important elements of production and process management, the basic concepts of logistics.

10th week -

11th week (two hour)

Human resources management.

12th week -

13th week (two hour)

Fundamental terms of corporate finance

14th week

Final test

Requirements:

Attendance at lectures is recommended, but not compulsory.

Those students who write a successful final test on the 14th week can get an offered grade, based on the test. Those who do not participate, or fail, have to take a written exam during the examination period.

Person responsible for course: Dr. Mária Ujhelyi, associate professor, PhD

Lecturers: Dr. Mária Ujhelyi, associate professor, PhD; Dr. András István Kun, associate professor, PhD; Dr. Krisztina Dajnoki, head of institute, associate professor, habil, PhD; Dr. Ferenc Ede Buzás, scientific assistant, Dr. Andrea Szabó, assistant professor, PhD

Title of course: Quality Management

ECTS Credit points: 1

Code: TTTBE0020-K1_EN	
Type of teaching, contact hours	
<ul style="list-style-type: none"> - lecture: 1 hour/week - practice: - - laboratory: - 	
Evaluation: exam	
Workload (estimated), divided into contact hours:	
<ul style="list-style-type: none"> - lecture: 14 hours - practice: - - laboratory: - - home assignment: - - preparation for the exam: 16 hours <p>Total: 30 hours</p>	
Year, semester: 1 st year, 1 st semester	
Its prerequisite(s): -	
Further courses built on it: -	

Topics of course
History of quality management. Standards, standardization methods. ISO standard system and related standards. Basic principles of quality management. The EN ISO 9001:2015 standard. Application of TQM (Total Quality Management). The EN ISO 14001:2015 standard.
Literature
<p><i>Compulsory:</i></p> <ul style="list-style-type: none"> - The EN ISO 9001:2015, 14001:2015 and BS OHSAS 18001:2007 standards. - John Jeston, Johan Nelis: Business Process Management - Practical Guidelines to Successful Implementations (Elsevier, 2008) ISBN: 978-0-75-068656-3 - D. H. Stamatis: Failure Mode and Effect Analysis - FMEA from Theory to Execution (American Society for Quality, Quality Press, 2003) ISBN: 0-87389-598-3 <p><i>Recommended:</i></p> <ul style="list-style-type: none"> - Carl L. Pritchard: Risk Management Concepts and Guidance (CRC Press, Taylor & Francis Group, 2015) ISBN: 978-1-4822-5846-2
Course objective/intended learning outcomes
<p>a) Knowledge</p> <ul style="list-style-type: none"> - He/She has knowledge to solve problems on the field of natural processes, using natural sources, and understanding the chemical background of living and non-living systems. - He/She understands the progress and future trends in chemistry and chemical industries. <p>b) Abilities</p> <ul style="list-style-type: none"> - He/She is able to use the obtained chemical knowledge on the field of chemical, industrial, environmental sciences to solve the actual basic problems and prove the obtained solution. - He/She is able to apply those directives that necessary to operate instruments and control processes in a safe, cost effective way as well as avoid any problems causing health issues.

c) Attitude

- He/She is open toward scientific and other postgradual education.
- He/She is able to collaborate with other groups and capable to get insight to the chemical aspects of economy and environmental safety.

d) Autonomy and responsibility

- Under supervision he/she is responsible in collaboration with other professionals (especially from the field of environmental economy and safety)
- Following directions he/she can work without supervision considering all quality and safety rules.

Schedule:

1st week

The history of quality management.

2nd week

Overview of the quality management systems used in the world. US and EU standards.

3rd week

ISO 9001:2015: Customer focus and customer satisfaction, leadership

4th week

ISO 9001:2015: Involvement of people, process approach.

5th week

ISO 9001:2015: Systematic approach to management, continual improvement

6th week

ISO 9001:2015: Factual approach to decision making, mutually beneficial supplier relationship

7th week

Using the Cycle of PDCA (Plan-Do-Check-Act) For Quality Management.

8th week

ISO 14001:2015: Context of the organization, Leadership.

9th week

ISO 14001:2015: Planning, Support, Operation.

10th week

ISO 14001:2015: Performance evaluation, Improvement

11th week

Integrating ISO 9001 and ISO 14001.

12th week

TQM: Quality and team organization. Teams' thinking and communication.

13th week

TQM: Problem solving and decision making process. Leadership and Empowerment. Benchmarking.

14th week

TQM: Achieving quality by planning: QFD, Hoshin planning. Quality through improvement and control: SPC. Quality through design: Robust design.

Requirements:

- for a signature

Attendance at **lectures** is recommended, but not compulsory. Active participation is rewarded by the teacher in every class.

During the semester there is one end-term test in the 15th week for an offered grade (optional). Students have to sit for the tests.

- for a grade

The course ends in an **examination**.

The minimum requirement for the end-term test and the examination respectively is 50%. Based on the score of the test, the grade for the test and the examination is given according to the following table:

<u>Score</u>	<u>Grade</u>
0-49	fail (1)
50-61	pass (2)
62-74	satisfactory (3)
75-87	good (4)
88-100	excellent (5)

If the score the test is below 50, students can take a retake test in conformity with the Education and Examination Rules and Regulations.

-an offered grade:

it may be offered for students if the grade of both the essay and the end-term test is at least satisfactory (3). The offered grade is the average of them.

Person responsible for course: Dr. György Deák, associate professor, PhD

Lecturer: Dr. György Deák, associate professor, PhD

Title of course: History and Structure of the EU Code: TTTBE0030-K1	ECTS Credit points: 1
Type of teaching, contact hours - lecture: 1 hour/week - practice: - - laboratory: -	
Evaluation: exam	
Workload (estimated), divided into contact hours: - lecture: 14 hours - practice: - - laboratory: - - home assignment: - - preparation for the exam: 40 hours Total: 54 hours	
Year, semester: 1 st year, 1 st semester	
Its prerequisite(s): -	
Further courses built on it: -	

Topics of course

The aim of the course is to give an overall picture for the students of the history of the development of the Community and the operation of its institutional system. It also aims at introducing the students to the enlargement process and the most important cooperation areas. On the level of EU policies, the issues of agriculture, regional policy, Economic and Monetary Union and the Schengen Area are discussed. The primary goal is that the future diploma holders have realistic knowledge about the functioning of the European Union, and of the international background of the Hungarian EU membership.

Literature

1. Bergmann, Julian – Niemann, Arne (2013): Theories of European Integration and their Contribution to the Study of European Foreign Policy, *Paper prepared for the 8th Pan-European Conference on International Relations, Warsaw 2013. p22.*
2. Ott, Andrea – Vos, Ellen (eds.) (2009): Fifty Years of European Integration: Foundations and Perspectives. T.M.C. Asser Press, Springer. 480pp. ISBN: 978-90-6704-254-3

Official website: https://europa.eu/european-union/about-eu_en

Course objective/intended learning outcomes

a) Knowledge

- knows the development of the European integration process, the institutional and decision-making systems, the operation of the major common policies, the criteria and impacts of the EU membership
- knows the unique characteristics of the fifth enlargement, the Copenhagen criteria with their special features, and the geopolitical role of East Central Europe
- knows the principles applied during the formation of the institutional system; understands the attempts in order to achieve the classic democratic share of power together with its deficit; knows the elements taken from the European Coal and Steel Community and the essence of the reform process
- is aware of the need for creating common policies, their history, the reform attempts and the peculiarities of the accession of Hungary
- knows the factors leading to the migration phenomena; knows the attempts made by the Community and the Member States in crisis management
- knows the plans and potential scenarios concerning the future of the Community; sees the need for reforms, knows about the Brexit and can reflect on its effects

b) Abilities

- has the ability to analyse the relations in terms of the EU membership of Hungary, the impacts of the membership and the tasks generated by it
- has the ability to treat the basic relationships of general geographic disciplines together with the results of economics, political sciences and jurisprudence.
- has the ability to apply the acquired knowledge of geography to solve practical problems
- based on his/her basic knowledge of social sciences, he/she is able to interpret the spatial results of related disciplines

c) Attitude

- endeavours to get a deeper knowledge of the theories, and principles related to the European

Union and seeks for relations with own experience

- with the application of the knowledge obtained, aims to get closer knowledge, to describe and explain the observable social phenomena
- is open to scientific exchange of views, professional co-operation, and aims to find solutions for his duties in cooperation with the collaborators through regarding their thoughts and opinion, if possible
- is ready to increase knowledge and to continue studies on a higher level

d) Autonomy and responsibility

- sets up position and brings decisions autonomously relying on technical sources in the issues related to the European Union
- takes over the values of his/her profession with responsibility, in the process of arranging assignments co-operates with experts representing other specialty fields
- takes responsibility for the professional work carried out

Schedule:

1st week

History of the Integration. Integration theories, stages of integration around the world. Specific features of the European integration process before the Second World War. Impacts of the Second World War on the history of the cooperation. Predecessors, impacts of the European Coal and Steel Community (ECSC) on the foundation of the European Economic Community. Steps towards the European Union.

2nd week

Process of the enlargement of the organisation. Preconditions of the enrolment of new members. Events of the period prior to the First Enlargement (1973). Steps, principles, causes and consequences of the Enlargements. Relationships between the decision-making mechanism and the Enlargement.

3rd week

Specific features of the enlargements after the turn of the millennium. Transformation of East Central Europe, and the unique features of its membership. Copenhagen criteria, pre-accession funds, prolonged negotiation process. Brexit.

4th week

History and principles of the creation of the institutional system. Taking-over the institutional system of the European Coal and Steel Community. Tasks of the most important institutions, operational mechanism, democratic deficit. Reform process of the institutional system, concepts laid down in the Constitutional Treaty. Decision-making in the EU.

5th week

Agricultural policy. History of the development of the CAP. The most important tools and sources of the funds. Horizontal measures. Current state of the common agricultural policy and its expected future. Reform attempts in agriculture. Hungary and common agricultural policy. Sharing the fish stocks of the seas.

6th week

Regional policy in the European Union. History of the regional policy. Regionalism – regionalisation in the EU Member States. General features of the regional policy. NUTS nomenclature. Regional disparities in the Community. Funds and main objectives. Decision-making in regional policy. Hungary and the regional policy.

7th week

Economic and Monetary Union (EMU). History of the European monetary co-operation. The European Monetary System (EMS). Role of the Maastricht Treaty in the monetary co-operation. Stages on the development of the Monetary Union. Convergence criteria. The euro and the currency market. Hungary and the Monetary Union.

8th week

Judicial co-operation in the Community. Legal order in the European Union. Role of the primary EU legislation in the European Community. European Community justice. Institutions serving the needs of judicial co-operation.

9th week

History of co-operations in home affairs. Schengen Convention. Regulations related to crossing state borders. Border checks. Checks between state borders, migration policy.

10th week

External relations. Principles of the common foreign trade policy. Autonomous import and export regulation. Issues related to the impediment to trade. External relations: African, Caribbean and Pacific Group of States (ACP), Global Mediterranean Policy, associated countries.

11th week

EU Budget: revenue side. Components of the EU budget and recent changes in the proportions. History of the EU budget. Budget revenues: duties, value-added tax (VAT), gross national product (GNP) sources.

12th week

Expenditures: agricultural policy, structural funds, external aid, research and development, pre-accession assistances, administrative expenditures. Economic characteristics. Budget procedure.

13th week

Migration and the European Union. Theoretical background to the migration crisis in 2015 and its practical consequences. History of the migration routes and movements. Natural and social (political) causes contributing to the crisis situation.

14th week

Common vision for the European co-operation. Possible development paths in the future of the European Union. Federal Europe or Europe of Nations? Reform options. Problem-solving attempts. Brexit.

Requirements:

- for a signature

Attendance at **lectures** is recommended, but not compulsory.

- for a grade

The course ends in a **written examination**.

Person responsible for course: Dr. Károly Teperics, senior lecturer, PhD

Lecturer: Dr. Klára Czimre, senior lecturer, PhD

Environmental Science

Title of course: General Chemistry I. Code: TTKBE0101_EN	ECTS Credit points: 4
Type of teaching, contact hours - lecture: 3 hours/week	

- practice: - - laboratory: -
Evaluation: exam
Workload (estimated), divided into contact hours: - lecture: 42 hours - practice: - - laboratory: - - home assignment: - - preparation for the exam: 68 hours Total: 110 hours
Year, semester: 1 st year, 1 st semester
Its prerequisite(s): -
Further courses built on it: TTKBL0101_EN, TTKBE0201_EN, TTKBE0007_EN

Topics of course
History and development of chemistry and its relation to other natural sciences. Development of atomic and molecular theory. The structure of atom. Basics of radioactivity. Discovery of the periodic table and periodically changing properties. Introduction to quantum chemistry. Primary and secondary chemical bonds. Description of gaseous, liquid and solid states of matter. Phase changes. Chemical equilibrium. Acid-base theories. Basics of thermochemistry, reaction kinetics and electrochemistry.
Literature
<i>Compulsory:</i> - John McMurry, Robert C. Fay: Chemistry, 7th ed., Prentice Hall ISBN: 0321943171. - Darrell D. Ebbing: General Chemistry, 9th ed. Belmont, CA, ISBN: 1439049829 - James E. Brady, Gerard E. Humiston: General chemistry: principles and structure, 3rd ed., New York, Wiley, ISBN: 0471808164
Course objective/intended learning outcomes
a) Knowledge - He/She knows the basic qualitative and quantitative chemical principles, and the methods based on it. - He/She knows the main models and theories of chemical bonds and molecular structure based on scientific findings. - He/She can understand and communicate professionally on chemical subjects at his/her native language. - He/She understands the structure of atoms, molecules and ions; and interprets the periodical change in the physical and chemical properties based on structural concepts. - He/She knows the principles of stoichiometry and the main characteristics of chemical reactions and phase changes.
b) Abilities - He/She is able to understand the natural and anthropogenic chemical reactions, and capable for data collection, evaluation on these fields including the data mining from literature.

- He/She is able to use the previously obtained knowledge on the field of natural and anthropogenic studies to solve practical problems.
- He/She is capable of describing physical and chemical processes with regard to quantitative, qualitative and energetic aspects.
- He/She is able to classify chemical reactions.

c) Attitude

- He/She pursues to use his/her knowledge to understand and describe the relationship of chemical processes connected to human life.

d) Autonomy and responsibility

- He/She can make reasonable evaluations about his/her own work comparing to others to the same field.

Schedule:

1st week

Classification of natural sciences, history and development of chemistry. The concept of chemical change. The SI system of units, the most important physical quantities and units. Conservation of mass and energy. The law of definite proportions, the law of multiple proportions, law of combining gas volumes, Avogadro's law. Dalton's atomic theory. Relative atomic and molecular weights. Amount of substance and the definition of mole. Notations for elements and compounds, symbol, empirical formula, molecular formula, structure, isomerism.

2nd week

Valency and oxidation number. Oxidation number in inorganic compounds. Types of chemical reactions. Latin names of compounds. Experimental background of the atomic theory, discovery of the nucleus. Discovery and basic properties of subatomic particles (electron, proton, neutron). Isotopes.

3rd week

Types and properties of radioactive radiation. Laws of radioactive decay, decay series. Medical and other practical importance of radioactive isotopes. The mass defect. Einstein's equation on mass-energy equivalence. Nuclear energy, nuclear fission and fusion. Quantized changes in the energy states of atoms. The photon hypothesis. The Bohr model of the atom. Characteristics of electromagnetic radiation, atomic line spectra, X-ray radiation.

4th week

The dual nature of matter. Heisenberg's uncertainty principle. Schrödinger's equation and its application for the hydrogen atom. Quantum numbers and their importance. The shape of atomic orbitals. Characterization of polyelectronic atoms. Principles of the periodic table.

5th week

Electronegativity, ionization energy, electronaffinity, atomic and ionic radii and their change across the periodic table. The ionic bond. Calculation of the lattice energy. Metallic bonding.

6th week

The covalent bond. Basic characteristics of the molecular orbital (MO) theory and its application for diatomic molecules. The valence shell electron pair repulsion (VSEPR) model. The shape of molecules, bond angles, bond orders, hybridization. Polarity of covalent bonds, polar and nonpolar molecules.

7th week

Intermolecular forces. London forces, dipole-dipole interaction. Hydrogen bond and its importance in inorganic and organic chemistry. General characterization of molecular, ionic, metallic, and network atomic solids.

8th week

Classification and structure of chemical systems. General characterization of different states of matter. The kinetic molecular theory of gases, ideal and real gases. Gas laws: Boyle's law, Charles's law, the ideal gas law. Gas mixtures, partial pressure. General characterization of liquids, surface tension, viscosity. General characterization and classification of solids. Changes of state: melting, freezing, evaporation, condensation, sublimation.

9th week

Classification of multicomponent systems, properties of solutions and mixtures. Solubility and units of concentration. Vapor pressure, freezing and boiling point of solutions. Osmosis pressure. Determination of molecular weight. Phase diagrams, critical temperature and pressure. Thermodynamic temperature.

10th week

Basics of thermochemistry. Heat of reaction, Hess's law. The importance of heat of formation. Heat of reaction and bond energies. The direction of spontaneous chemical reactions: internal energy, enthalpy, free energy and entropy.

11th week

Dependence of reaction rates on concentrations and the temperature. Order of reactions. Activation energy. Catalysts, homogeneous and heterogeneous catalytic reactions. Enzymes. Photochemical processes. The equilibrium condition and the equilibrium constant. Possibilities to shift the composition of equilibria. Dependence of the equilibrium constant on temperature and pressure. Le Chatelier's principle.

12th week

Solubility equilibria, solubility product. Temperature dependence of solubility. Gas-liquid and liquid-liquid equilibria. Extraction. Different theories of acid-base reactions (Arrhenius, Brønsted, Lewis). Characterization of aqueous solutions, electrolytic dissociation. Strength of acids and bases. Super acids. Dissociation constant and degree of dissociation.

13th week

Self-ionization of water. Ionic product of water. The definition and calculation of pH. Amphoteric substances. Buffer solutions and acid-base indicators. Acid-base properties of salts. Complex ion equilibria. Pearson's hard-soft theory.

14th week

Basics of electrochemistry. Galvanic cells and the concept of electrode potential. Standard electrode potentials, oxidizing and reducing agents. Water as a redox system. Electrolysis, voltage needed in electrolytic cells, overvoltage. Quantitative laws of electrolysis. Galvanic cells and batteries.

Requirements:

- for a signature

Attendance at **lectures** is recommended, but not compulsory.

- for a grade

The course ends in an **examination**. The result of the examination determines the final grade.

The minimum requirement for the examination is 50%. Based on the score of the exam, the grade is given according to the following table:

Score	Grade
0-49	fail (1)
50-62	pass (2)
63-75	satisfactory (3)
76-87	good (4)
88-100	excellent (5)

If the case of failure, students can take retake exam(s) in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

Person responsible for course: Dr. Gábor Bellér, assistant professor, PhD

Lecturer: Dr. Gábor Bellér, assistant professor, PhD

Title of course: General Chemistry I. (seminar) Code: TTKBG0101_EN	ECTS Credit points: 3
Type of teaching, contact hours - lecture: - - practice: 4 hours/week - laboratory: -	
Evaluation: middle-term and final exams	
Workload (estimated), divided into contact hours: - lecture: - - practice: 44 hours - laboratory: - - home assignment: 26 hours - preparation for the exam: 20 hours Total: 90 hours	
Year, semester: 1 st year, 1 st semester	
Its prerequisite(s): -	
Further courses built on it: TTKBL0101_EN, TTKBG0501_EN	

Topics of course
The main objective of the seminar is to give the basic knowledge and background for students to solve general calculation problems strictly connected to the general chemistry laboratory practice: calculations connected to mass and volume measurements, concentration and its units, crystallization, acid-base and redox equilibria, balancing chemical equations.
Literature
<i>Compulsory:</i> - The collection of calculation problems will be available at the Department's home page (inorg.unideb.hu) <i>Recommended:</i> - Darrell Ebbing, Steven D. Gammon: General Chemistry 10 th edition - Darrell Ebbing, Steven D. Gammon: General Chemistry – Standalone book
Course objective/intended learning outcomes
a) Knowledge - He/She knows the basic qualitative and quantitative chemical principles, and the methods based on it.

- He/She has a basic chemical knowledge on describing simple chemical processes as well as on recognizing, organizing these in practice.

- He/She can understand and communicate professionally on chemical subjects at his/her native language.

b) Abilities

- He/She is able to use the previously obtained knowledge on the field of natural and anthropogenic studies to solve practical problems.

c) Attitude

- He/She is open toward scientific and other post gradual education.

d) Autonomy and responsibility

- He/She can make reasonable evaluations about his/her own work comparing to others to the same field.

Schedule: The seminar will be held in 11 weeks.

1st week

Determination of atomic weight, molecular weight, empirical formula, molecular formula, amount of substance. Determination of empirical formula based on weight percent composition and on elemental analysis.

2nd week

General introduction to the units of concentration. Interconversion of units. Calculation problems connected to solution preparation. Introduction of the SI system. Mass concentration, molarity, mass percent composition, molar percent composition.

3rd week

Review exercises concerning on the first two weeks. Interconversion of concentration units. Density measurements. Mixing equations. Theoretical background of crystallization. Exercises calculation problems of crystallization.

4th week

Theoretical backgrounds of gas and solids. Composition of solid and gas mixtures. Introduction to basic chemical equations. Stoichiometric calculations based on chemical equations. Preparation of salts, calculation of theoretical and percent yield. Dissolving of metal mixtures in acids.

5th week

Acid-base equilibria. Theory of acid-base reactions and titrations. Exercises based on acid-base titrations. Stoichiometric calculations based on chemical equations. Determination of molar weight based on titration results.

6th week

Review exercises in stoichiometry and concentration calculations.

7th week

Introduction to basic gas laws. Laboratory preparation of gases. Calculation problems connected to evolution of gases based on chemical equations.

8th week

Theory of redox reactions. Balancing of redox reactions. Calculations based on redox reactions. Preparation of salts from its metal. Review exercises in balancing of redox and acid-base reactions.

9th week

Definition of pH. Theoretical background of pH calculation. Introduction to water ionisation constants. Relationship between the K_w and H^+ . Calculation of pH of strong acids and strong bases.

10th week

Calculation of pH of weak acids and weak bases. Determination of dissociation rate. Theoretical background of buffer systems, buffer capacity. Calculation problems regarding the pH of buffer systems.

11th week

Electrochemical exercises. Fundamental of galvanic cells (Daniell cell). The concept of electromotive force, redox potential, standard redox potential. Nernst equation. Review exercises of pH calculations.

Requirements:

Students are required to write two general tests (after week 6 and after week 11) which are based on the course material for weeks 1-5 and 7-11, respectively. Each general test is worth 50 points. Grading is based on a five-level scale: 1 (fail), 2 (pass), 3 (average), 4 (good), 5 (excellent). The final course grade is given based on the results of these tests. The score from the general tests must be above 50 % to avoid a 'fail' final course grade. In order to pass the seminar, a student should collect minimum 50 points from the general tests. Students with 'fail' final course grade due to low test results can re-take once a comprehensive test exam in the examination period.

It is not allowed to miss any seminars. If a student misses two seminars even for any medical reasons, the student's lecture book won't be signed and she or he has to retake the course next year.

Person responsible for course: Dr. Norbert Lihi, assistant research fellow, PhD

Lecturer: Dr. Norbert Lihi, assistant research fellow, PhD

Title of course: General Chemistry II. (laboratory practice) Code: TTKBL0101_EN	ECTS Credit points: 3
Type of teaching, contact hours - lecture: - - practice: - - laboratory: 4 hours/week	
Evaluation: short tests, middle-term and final exams	
Workload (estimated), divided into contact hours: - lecture: - - practice: - - laboratory: 44 hours - home assignment: 32 hours - preparation for the exam: 14 hours Total: 90 hours	
Year, semester: 1 st year, 2 nd semester	
Its prerequisite(s): TTKBE0101_EN, TTKBG0101_EN	
Further courses built on it: TTKBL0201_EN, TTKBL0501_EN	

Topics of course

The objective of the laboratory practice is to introduce first-year students of different background to laboratory work, the use of basic laboratory equipment, simple laboratory operations and

measurements. In addition, students are expected to prepare certain simple chemicals and run various basic experiments to familiarize themselves with chemical laboratory work.

Literature

Compulsory:

- General chemistry laboratory practice (laboratory manual)

Recommended:

- Darrell Ebbing, Steven D. Gammon: General Chemistry 10th edition
- Darrell Ebbing, Steven D. Gammon: General Chemistry – Standalone book

Course objective/intended learning outcomes

a) Knowledge

- He/She knows and can use the properties of most important chemicals, lab equipment and basic laboratory processes. He/She can follow and recognize the requirements of safety instructions.
- He/She can understand and communicate professionally on chemical subjects at his/her native language.

b) Abilities

- Using his/her chemical knowledge and experiences He/She is able to create and execute experiments to demonstrate and prove the basic chemical phenomena.
- He/She is able to evaluate and discuss the laboratory measurements and create a report about it.

c) Attitude

- He/She is able to represent his/her own personal scientific ideology toward professional and unprofessional groups.
- He/She is open toward scientific and other postgradual education.

d) Autonomy and responsibility

- During laboratory work he/she is capable to get related to the basics of chemical problems, able to conclude the questions to the supervisor that might come up to the solutions or decisions.
- He/She can make reasonable evaluations about his/her own work comparing to others to the same field.

Schedule: The laboratory practice will be held in 11 weeks.

1st week

General introduction to the laboratory rules and laboratory work. Safety training. Introduction to laboratory pieces of equipment. The use of gas burners. Overview of pieces of the received laboratory equipment.

2nd week

Mass and volume measurements: weighing on analytical and standard laboratory balances; introduction to volume measurement devices (pipette, burette, volumetric flask). Calibration of volumetric measuring equipment (pipette or volumetric flask). Calculation the standard error between the measured and nominal values.

3rd week

Introduction to solution preparation: grinding, use of mortar, pestle, volumetric flask. Preparation of a standard solution from a crystalline salt. Introduction to a density measurement. The use of the pycnometer. Determination of the density of the prepared solution by the help of the pycnometer. Calculating the weight percent composition of the prepared solution.

4th week

Introduction to separation methods: decantation, centrifuging, filtration. Purification of solids.

Theoretical background heating, cooling and the use of hot water bath. Purification of a benzoic acid sample contaminated with sodium chloride. Preparation of a double salt from simple salts and basic laboratory procedures.

5th week

Writing the general mid-term test based on the studied material of the laboratory practice and seminar until week 4. Determination of the composition of mixture of potassium chloride and potassium chlorate. Review of different methods used to temperature measurements. Introduction to the measurements of melting point of the solid substances. Determination of the melting point of the purified benzoic acid sample. Determination of the contamination percentage of the purified benzoic acid sample.

6th week

Demonstration of acid-base titration. Preparation of a standard solution of NaOH. Concentration determination of the standard NaOH solution by acid-base titration. Determination of the molar weight of the recrystallized sample of benzoic acid by acid-base titration. Comparing the result with the literature value and calculating the standard error between the given and measured data. Purified benzoic acid due in.

7th week

Laboratory work with gases: introduction to the use of gas cylinders, simple gas generator, Kipp's apparatus. Studying the chemical and physical properties of gases. Demonstration of hydrogen preparation. The hydrogen explosion test. Preparation of oxygen in a laboratory gas generator and burning of sulphur in oxygen. Study of the observations during the reaction (oxidation product of sulphur). Determination of molecular weight based on the ideal gas law.

8th week

Practice the basic laboratory techniques considering the preparation of a salt. Preparation of salts from its metal. Studies of reactions involving gas formation and precipitation.

9th week

Quantitative study of a precipitation reactions to determine the stoichiometric composition of water insoluble precipitates using the method of continuous variation. Dependence of reaction rate of concentration of reactants. Studying the factor affecting the reaction rates. Determination of the reaction rate and the rate law of the studied reaction. Metal salts preparations due in.

10th week

Theoretical background of liquid-liquid extractions and demonstration of the separation techniques. Introduction to buffer systems, buffer capacity by studying a particular buffer system (acetic acid/acetate ion buffer; ammonium ion/ammonia buffer). Hydrolysis of salts to study the acid-base properties of ionic and covalent compounds in aqueous solutions or in reactions with water. Writing of the ionic equations based on the observed chemical reactions.

11th week

General test from week 5 to week 10. General introduction to electrochemistry. Study of redox reactions. Prediction of the direction of spontaneous processes based on standard potentials. Factors affecting the order of the deposition of different metals during electrolysis (study of Daniell cell). Return of the received pieces of laboratory equipment.

Requirements:

Each week the laboratory session begins with a short test (not more than 20 minutes) based exclusively on the preparatory material of that week and the previous week and the results of the experiments carried out the previous week. With each short test a student can collect 25 points. Altogether there are eight short tests during the semester. Students are also required to write two general tests (week 5 and week 11) which are based on the course material for weeks 1-4 and 5-10,

respectively. Each general test is worth 50 points. Grading is based on a five-level scale: 1 (fail), 2 (pass), 3 (average), 4 (good), 5 (excellent). The final course grade is given based on the results of these tests, the quality of the laboratory notes and the quality of laboratory work. The average score from both the short tests and the general tests must be above 50 % to avoid a 'fail' final course grade. In order to pass the laboratory practice, a student should collect minimum 100 points from the short tests and minimum 50 points from the general tests. Students with 'fail' final course grade due to inadequate laboratory work have to retake the course the next year. Students with 'fail' final course grade due to low test results can re-take a comprehensive test exam in the examination period.

Those students, whose results are lower than 25% either from the short test or from the general test, cannot write a final exam, they will receive a 'fail' final course grade.

It is not allowed to miss any laboratory practices/seminars. If a student misses one or two lab practices, medical certification is needed. If a student misses three lab practices/seminars even for any medical reasons, the student's lecture book won't be signed and she or he has to retake the course next year. It is not possible to miss short tests at the beginning of the laboratory practice. If a student misses more than two short tests, the laboratory practice will not be accepted for him or her. The students cannot miss either of the general tests, otherwise no signature and final grade is given to the student.

Person responsible for course: Dr. Norbert Lihi, assistant research fellow, PhD

Lecturer: Dr. Norbert Lihi, assistant research fellow, PhD

Title of course: Inorganic Chemistry I Code: TTKBE0201_EN	ECTS Credit points: 3
Type of teaching, contact hours - lecture: 2 hours/week	
Evaluation: examination	
Workload (estimated), divided into contact hours: - lecture: 28 hours - practice: - - laboratory: - - home assignment: - - preparation for the exam: 56 hours Total: 84 hours: 3 credit x 28 hours	
Year, semester: 1 st year, 2 nd semester	
Its prerequisite(s): TTKBE0101_EN	
Further courses built on it: TTKBE0202_EN, TTKBL0201_EN, TTKBE0501_EN, TTKBE0502_EN, TTKBL0502_EN, TTKBE0204_EN	

Topics of course

Literature

Compulsory:

1) N. N. Greenwood, A. Earnshaw: Chemistry of the Elements, 2nd Edition, 1997 (or later ed.)

Recommended:

2) Geoff Rayner-Canham, Tina Overton: Descriptive Inorganic Chemistry (5th Edition), W. H. Freeman and Company, New York, 2010, ISBN-13: 978-1-4292-2434-5 (or later edition)

3) Glen E. Rodgers, Descriptive Inorganic, Coordination and Solid-Phase Chemistry, (3rd Edition), Brooks/Cole, 2012, ISBN-13: 978-0-8400-6846-0 (or later edition)

Course objective/intended learning outcomes

a) Knowledge

- He/she fundamentally knows principles, procedures, properties, reactions and chemical processes of the p-group elements and inorganic chemistry related materials.

b) Abilities

- He/she is able to apply the most important terminology, theories, procedures of the given inorganic chemistry-related field when completing the relevant tasks.

- He/she is able to create fundamental models of inorganic chemical reactions, and processes.

c) Attitude

- He/she is open to learn and accept professional, technological improvement and innovation in his/her profession and convey it genuinely.

- He/she makes a decision in complex and unexpected cases by completely taking into account legal and ethical norms.

d) Autonomy and responsibility

- Even in unexpected decision-making situations he/she is capable of considering complex, fundamental questions from his/her professional field and elaborating them on the basis of the given sources.

- He/she confesses and represents the value system of the qualified chemists and chemical engineers profession with responsibility. He/she is open to critical remarks which are professionally well-founded.

Schedule:

1st week

Origin of the elements in the periodic table. Abundances of the elements in the Universe and on the Earth. Energy production by nuclear reactions. Development of atmosphere on Earth.

Major chemical forms of elements, grouping elements by their oxidation state. Production and exploitation methods of the chemical elements.

2nd week

Chemical, physical and atomic properties of the element hydrogen. Solubility of hydrogen in different materials. Hydrogen isotopes, their nuclear properties, nuclear spin isomers, practical application of isotopes and nuclear isomers. Electronic structure, oxidation number, and chemical reactivity, and major reactions of hydrogen. Laboratory and industrial production of hydrogen. Hydrogen as an environment-friendly fuel. Binary hydrides, their most important groups, properties. Hydrogen bonding and its role in the structure and properties of the materials. The most important hydrogen compounds and their use in the practice.

3rd week

Noble gases. General characterization, special physical and chemical properties. Electronic structure of noble gases. Correlation between electronic structure and chemical reactions. Noble gas compounds. The history of xenon compounds, structure, electronic structure, characteristic

reactions. Xenon oxides and oxoacids and their salts. Preparation and practical use. Separation of noble gases from natural sources.

General properties of the halogens. Physical properties, electronic structure, chemical reactivity, possible oxidation numbers. Interaction of the halogens with different solvents. Hydrate formation, chemical hydrolysis.

4th week

Characteristic chemical reactions of the halogens, interhalogen compounds, polyhalogenium ions, polyhalide anions. Structural aspects of interhalogen compounds, VSEPR theory to describe geometric structures. Halogen-containing minerals, natural resources. Biological role of halides. Laboratory scale and industrial production of the halogen elements. Most important groups of halides regarding their chemical bondings and lattices, physical properties.

5th week

Halogen-oxygen compounds, physical and chemical properties of halogen oxides, and methods of their synthesis, and practical uses. Halogen oxyacids and their salts. Oxidation numbers of the component atoms, laboratory scale and industrial productions. Chemical reactions of halogen oxides and oxoacids. Sterilization, drinking water treatment with halogen oxides and oxoacids.

6th week

Elements of the oxygen group. Electronic structure, physical and chemical properties, characteristic oxidation numbers. Allotropic forms of dioxygen. Structure of dioxygen, explanation of the magnetic properties. Solubility of oxygen in water and its biological role. *oxigéncsoport elemeinek előfordulása, általános jellemzése, solubility in water.*

Ozone, physical and chemical properties, formation of ozone in the high atmosphere. Ozone depletion, ozone hole in the arctic region. The role of ozone shield. Ozone precursors, chemicals that can destroy the ozone shield. Chemical reactions of ozone. Practical applications.

Sulfur, selenium, tellurium, allotropic forms, physical properties, oxidation numbers, electronic structures. Chemical reactivity of the elements. Acid-base properties of the sulfides. Laboratory scale and industrial production techniques of the elements. Biological role of the oxygen group elements and their compounds. Oxygen and sulfur cycles in the biosphere.

7th week

Binary hydrides of the oxygen group elements. Water, physical and chemical properties, its role in the life and the environment. Types of water in the nature. Gas hydrates. Water purification techniques, water hardness and water treatment. Water wars.

Hydrogen peroxide. Structure, electronic structure, characteristic physical and chemical properties, appearance and role in the living organisms. Synthesis of hydrogen peroxide, in the laboratory and in the industry. Practical uses of hydrogen peroxide.

Binary hydrides of sulfur, selenium, tellurium and lead. Stabilities, chemical properties, synthesis, toxicity, practical uses. Analytical system based on hydrogen sulfide.

8th week

Halides of the calcogenic elements. Synthesis of sulfur chlorides, their properties, reactivities and practical uses. Sulfur oxides, their structure, synthesis, physical and chemical properties, production in the industry. Sulfur-containing oxoacids and their salts: structure, properties, reactivities, practical uses. Peroxy sulfuric acids and S-S bond-containing sulfur oxoacids and their salts: structure, reactivity, preparation, practical uses.

Environmental concerns regarding the concentration of atmospheric sulfur dioxide: formation and effect of acid rain.

9th week

Elements of the nitrogen group: appearance, electronic structure, physical properties, allotrops,

chemical properties, oxidation states, hybridization. Synthesis and isolation of the elements. Industrial production, air liquifaction and fractionated distillation. Physical methods of nitrogen generation. Practical uses of the elements.

10th week

Hydrides of the nitrogen-group elements. Ammonia and hydrazine: composition, structure, electronic properties, molecular movements. Physical and chemical properties, reactivities, acid-base properties, redox states, characteristic chemical reactions. Synthesis of ammonia and hydrazine in the laboratory and in the industry. Haber-Bosch and Raschig processes. Practical uses of ammonia and hydrazine.

Halides and halogeno-complexes of the nitrogen-group elements. Composition, formation, structure, characteristic physical and chemical properties, reactivities. Practical uses.

Oxides and oxo-compounds of the nitrogen-group elements. Structure, formation, composition, physical and chemical properties. Electronic structure, spectral and magnetic properties. Laboratory-scale and industrial production, Ostwald synthesis. Acid-base properties.

Environmental and health issues of nitrogen oxides, role of NO in the human body.

11th week

Nitrogen and phosphorus oxoacids. Chemical composition, oxidation states, stabilities, physical properties, characteristic reactions, most important salts. Practical uses of nitric and phosphoric acids. Other oxides, oxoacids and oxoanions of other elements of the nitrogen group. Compounds with sulfur: sulfur nitrides, phosphorus sulfides, molecular structures, stabilities, physical and chemical properties, practical uses.

Elements of the carbon group. Electronic structure, oxidation states, hybridization, types of chemical bondings. Stereochemistry of carbon. Comparison of the structure of analogous carbon and silicon compounds.

12th week

Carbon allotrops, structural properties, characteristic physical and chemical properties. Natural carbon sources. Synthesis of carbon allotropes. Isotopes of carbon, stability, properties, practical uses, radiocarbon method. Silicon and other elements: natural sources, properties, synthesis, practical uses. Production and purification of semiconductor grade silicon and germanium. Tin and lead: allotropes, preparation/production, properties, toxicity, practical uses.

Comparison of the structure and stability, hydrolytic properties of the binary hydrides of the carbon group elements. Preparation of the hydrides, practical uses in the analytical chemistry.

Halides of the carbon group elements: Composition, hydrolysis, complex formation, geometry, nature of the bonds, redox properties and stabilities of the halides.

13th week

Oxides and oxoacids of carbon and silicon. Composition and electronic structure of carbon oxides and oxoacids. Binding modes and coordination chemistry of carbon monoxide, the most important carbonyl complexes. Properties, toxicity and environmental issues of carbon dioxide. Carbonic acid and their salts, carbonates in the nature. Greenhouse effect, increase of atmospheric carbon dioxide, climate changes, global warming, and the role of technical civilization. Silicic acids and silicates. Types of natural and synthetic silicates. Polymeric and 3D structures, basic types, appearance in the nature. Special silicon oxides and silicates, silica gels and aerogels. Oxides of tin and lead.

Carbon-nitrogen bond containing inorganic compounds: Cyanic acid and isocyanic acid and their salts. Thiocyanic acid and isothiocyanic acid and their salts. Properties, practical uses.

Carbon and silicon sulfides. Comparison of oxo and thio compounds. Thio-bases and thio-acids. Types of carbides, ionic, covalent and interstitial carbides. Properties, practical uses.

14th week

Elements of the boron group. Appearance, natural resources, most important minerals. Electronic structure, Lewis-acidity, physical and chemical properties, most important chemical reactions. Hybridization. Halides of the boron group elements. Properties, hydrolysis, complex formation, structure, practical uses. Industrial production of aluminum.

Binary and complex hydrides of boron group elements. Special structural characteristics and bonding mode of diborane: 2-electron-3-center binding mode. Synthesis, physical and chemical properties of hydrides and complex hydrides. Comparison of hydrolytic and thermal stabilities. Practical uses of the complex hydrides. Reduction, hydroboration. Polyhedral boron hydrides, structure, stability, carboranes. Boron oxides, boric acid, aluminum oxide and hydroxide. Practical use of aluminum oxide and high surface area alumina.

Requirements:

- for a signature

Attendance at **lectures** is recommended, but not compulsory.

- for a grade

The course ends in an **examination**.

The examination starts with a qualification test. The minimum requirement to qualify for the examination is: 60 score. Below score 60 Grade 1 (Fail) is given.

Score	Grade
0-59	fail (1)
60-100	qualified to the exam

The minimum requirement for the examination is 50 score. Based on the score, the grade for the examination is given according to the following table:

Score	Grade
0-49	fail (1)
50-62	pass (2)
63-75	satisfactory (3)
76-88	good (4)
89-100	excellent (5)

If the score of any test is below 60, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

Person responsible for course: Dr. István Lázár, associate professor, PhD

Lecturer: Dr. István Lázár, associate professor, PhD

Title of course: Inorganic Chemistry II

Code: TTKBE0202_EN

ECTS Credit points: 3

Type of teaching, contact hours

- lecture: 2 hours/week

Evaluation: examination

Workload (estimated), divided into contact hours:

- lecture: 28 hours

- practice: -
- laboratory: -
- home assignment: -
- preparation for the exam: 56 hours
Total: 84 hours

Year, semester: 2nd year, 1st semester

Its prerequisite(s): TTKBE0201_EN, TTKBE0301_EN, TTKBE0401_EN

Further courses built on it:

TTKBL0201_EN, TTKBE0501_EN, TTKBE0502_EN, TTKBL0502_EN, TTKBE0204_EN

Topics of course

Literature

Compulsory:

1) N. N. Greenwood, A. Earnshaw: Chemistry of the Elements, 2nd Edition, 1997 (or later ed.)

Recommended:

2) Geoff Rayner-Canham, Tina Overton: Descriptive Inorganic Chemistry (5th Edition), W. H. Freeman and Company, New York, 2010, ISBN-13: 978-1-4292-2434-5 (or later edition)

3) Glen E. Rodgers, Descriptive Inorganic, Coordination and Solid-Phase Chemistry, (3rd Edition), Brooks/Cole, 2012, ISBN-13: 978-0-8400-6846-0 (or later edition)

Course objective/intended learning outcomes

a) Knowledge

- He/she fundamentally knows principles, procedures, properties, reactions and chemical processes of the s-, d-, and f-group elements and inorganic chemistry related materials.

b) Abilities

- He/she is able to apply the most important terminology, theories, procedures of the given inorganic chemistry-related field when completing the relevant tasks.

- He/she is able to create fundamental models of inorganic chemical reactions, and processes.

c) Attitude

- He/she is open to learn and accept professional, technological improvement and innovation in his/her profession and convey it genuinely.

- He/she makes a decision in complex and unexpected cases by completely taking into account legal and ethical norms.

d) Autonomy and responsibility

- Even in unexpected decision-making situations he/she is capable of considering complex, fundamental questions from his/her professional field and elaborating them on the basis of the given sources.

- He/she confesses and represents the value system of the qualified chemists and chemical engineers profession with responsibility. He/she is open to critical remarks which are professionally well-founded.

Schedule:

1st week

General characterization of the metals, structure of the metals, metallic bond. Principles of band theory, conductors, semiconductors and insulators. Characteristic physical and chemical properties of the metals.

2nd week

Alkali metals: general characterization, physical and chemical properties, abundance, preparation and use. Hydrides, halogenides, oxides, hydroxides of alkali metals, salts formed with the most important oxoanions. Complexes of alkali metal ions, crown ethers and cryptands. Covalent compounds of the alkali metals.

3rd week

Alkali earth metals: general characterization, physical and chemical properties, abundance, preparation and use. Role of the alkali earth metals in the nature, biological effect of the metals and their ions. Special features of beryllium and its compounds. Hydrides, halogenides, oxides, hydroxides of alkali earth metals, salts formed with the most important oxoanions. Covalent compounds and complexes of the alkali metals.

4th week

General characterization of the transition (d-block) metals. Important trends in the change of electronic configuration, electronegativity, atomic and ionic radii for the elements in the d-block. Physical and chemical properties of the transition metals, their similarity. Abundance of d-block metals and general methods for the preparation of transition metals. Theoretical and practical aspects of the selection of reducing agents.

5th week

Basic terms in coordination chemistry, coordination number, geometry of complexes. Isomerism and nomenclature of complex compounds. Factors influencing the stability of complexes. Fundamentals of the Hard-Soft Acid-Base (HSAB) theory. Classification of complex compounds and ligands, mono- and multidentate ligands, σ -donor and π -acceptor ligands. Chelate- and macrocycle effect, their importance. Inert and labile complexes.

6th week

Fundamentals of the crystal field theory, interpretation of the colors and magnetic behaviour of the complex compounds. High and low spin complexes. Definition and importance of crystal field stabilization energy (CFSE). Types of transition metal hydrides and their practical importance. Classification of transition metal halogenides based on their composition, structure and binding types. Some important halogenides of the transition metals.

7th week

Oxides, hydroxides and oxoacids of transition metals. Classification of oxides based on their composition and binding types. Physical and chemical properties of the oxides, their acid-base and redox reactions. Methods for the preparation of oxides. Transition metal sulphides, their importance in the environment and analytical chemistry. Carbides. Simple complexes of the transition metals: hydroxido, halogenido and cyano complexes.

8th week

Metals of the titanium and vanadium group and their most important compounds. General characterization, trends in oxidation numbers, physical and chemical properties, occurrence, preparation and use. Industrial preparation of titanium, practical importance of the metal. Properties of titanium-dioxide and -tetrachloride, their derivatives. Properties of vanadium oxides and their derivatives.

9th week

Members of the chromium group, some important compounds. General characterization, trends in

oxidation numbers, physical and chemical properties, occurrence, preparation and use. Halogenides and coordination chemistry of chromium, molybdenum and tungsten. Oxides and their derivatives. Thermal stability, acid-base and redox reactions of the oxides. Formation trends and structure of the iso- and heteropolyacids.

10th week

Members of the manganese and iron groups, some important compounds. General characterization, trends in oxidation numbers, physical and chemical properties, occurrence, preparation and use. Manganese oxides, their derivatives and redox reactions of them. Industrial preparation of iron and steel. Oxides, halogenides and important complex compounds of iron, cobalt and nickel.

11th week

General characterization of the platinum group metals, trends in oxidation numbers, physical and chemical properties. Theoretical aspects of their preparation, some important practical use. Oxides and halogenides. Coordination chemistry of platinum group metal ions: oxidation states and practical use.

12th week

Members and the most important compounds of the copper and zinc groups. General characterization, trends in oxidation numbers, physical and chemical properties, occurrence, preparation and use. Practical importance of the metals, alloys. Oxides and halogenides. Chemical background of black and white photography. Important complex compounds. Environmental and biological role of the metals and their cations.

13th week

General characterization of the lanthanoides and actinoides, electronic configuration, oxidation numbers, physical and chemical properties, occurrence, preparation and use. Oxides and halogenides, important complex compounds. Physical and chemical properties of thorium and uranium, important compounds. Theoretical aspects of the use of nuclear power.

14th week

Fundamentals of bioinorganic chemistry. Classification of the elements based on their biological role. Metalloenzymes and their role. Fundamentals of the medicinal and environmental use of metal ions and their complexes. Classification of organometallic compounds. Definition of hapticity. Covalent organometallic compounds. Carbonyls, alkenes and cyclopentadiene compounds of the transition metals.

Requirements:

- for a signature

Attendance at **lectures** is recommended, but not compulsory.

- for a grade

The course ends in an **examination**.

The minimum requirement for the examination is 40 score. Based on the score, the grade for the examination is given according to the following table:

Score	Grade
0-39	fail (1)
40-55	pass (2)
56-70	satisfactory (3)
71-85	good (4)
86-100	excellent (5)

If the score of any test is below 40, students can retake the test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

Person responsible for course: Dr. Péter Buglyó, associate professor, PhD	
Lecturer: Dr. Péter Buglyó, associate professor, PhD	
Title of course: Inorganic Chemistry III. Code: TTKBL0201_EN	ECTS Credit points: 5
Type of teaching, contact hours - lecture: - - practice: - - laboratory: 4 hours/week	
Evaluation: practice grade	
Workload (estimated), divided into contact hours: - lecture: - - practice: - laboratory: 56 hours - home assignment: 112 hours - preparation for the exam: - Total: 168 hours	
Year, semester: 2 nd year, 1 st semester	
Its prerequisite(s): TTKBL0101_EN, TTKBE0201_EN, TTKBE0301_EN, TTKBE0401_EN	
Further courses built on it:	

Topics of course
Physical and chemical properties, laboratory preparation, use, biological role, identification of hydrogen and the most important p-block elements. The most important compounds of hydrogen and the p-block non-metals: chemical reactions, reactivity, preparation and use. Qualitative analysis of the most important related (oxo)anions with classical analysis. Synthesis and purification of compounds of the p-block elements.
Literature
1. Syllabus provided by the tutor 2. Willard, L.L. Inorganic Chemistry, Wadsworth Publ. Co., Belmont, 1998. 3. Vogel's Qualitative Inorganic Analysis, (ed. Gy. Svehla), Longmann, 2007.
Course objective/intended learning outcomes
a) Knowledge - He/she knows the most important physical and chemical properties, production, practical use, identification and compounds of hydrogen and the most important p-block elements. - He/she knows and uses the various chemicals, devices and methods in the chemical laboratory together with the safety knowledge.
b) Abilities - He/she is able to apply the most important terminology, theories, procedures relating the elements and their most important compounds when completing the relevant tasks. - He/she is able to find solutions for inorganic laboratory problems.
c) Attitude

- He/she is open to learn and accept inorganic chemical improvement and innovation in his/her profession and convey it genuinely.
 - He/she makes a decision in complex and unexpected decision cases by completely taking into account legal and ethical norms.
- d) Autonomy and responsibility
- Under supervision he/she is responsible in collaboration with other professionals (especially from the field of inorganic laboratory work and safety). -

Schedule:

1st week: Introductory guidance, safety regulations

2nd week: Hydrogen: preparation, physical and chemical properties, reactions with non-metals

3rd week: Group 17 elements: preparation, physical and chemical properties, reaction of chlorine with metals. Identification of the halide ions

4th week: Unknown sample with halide ions. Group 16 elements: oxygen and sulfur. Laboratory preparation, physical and chemical properties, reactions of non-metals and metals with oxygen. Ozone.

5th week: Important compounds of group 16 elements: preparation, physical and chemical properties, redox properties of hydrogen-peroxide. Sulfur oxides and oxoacids. Identification of sulphite and sulphate ions.

6th week: Unknown sample with group 17 and 16 anions. Group 15 elements: preparation, physical and chemical properties of nitrogen and phosphorus. Allotropic modifications of phosphorus. Ammonia, dissolution of alkaline metals in liquid ammonia. Identification of ammonia and ammonium ion.

7th week: Synthesis of various compounds consisting p-block elements. Group 15 halogenides and related compounds.

8th week: Nitrogen oxides and the appropriate oxoacids and salts. Preparation and reactivity of nitrogen-dioxide. Properties of nitric acid. Identification of nitrite and nitrate ion.

9th week: Phosphorous oxides and the appropriate oxoacids and salts. Properties of phosphoric acid. Formation, reactivity, redox features and identification of nitrite and nitrate ion.

10th week: Unknown sample with group 17-15 anions. Group 15 sulphides, their acid-base properties and analytical importance.

11th week: Group 14 elements: physical and chemical properties. Reactivity of lead and tin. Formation of silane. Carbon oxides: synthesis, their properties and important reactions. Cyanide and thiocyanate ions and their analytical importance.

12th week: Preparation of various compounds consisting p-block elements. The most important properties of tin and lead: oxides and sulphides.

13th week: Complex unknown sample with the most important anions. Group 13 elements: physical and chemical properties.

14th week: The most important aluminium and boron compounds. Acid-base character of boric acid. Identification of aluminium(III) and borate ions. Complex hydrides and their practical use.

Requirements:

- *for a signature*

Participation at practices is compulsory. Students must attend each practice during the semester. In case of absence(s), a medical certificate needs to be presented. Missed practice classes should be made up for at a later date, to be discussed with the tutor.

- *for a grade*

Grading is given by considering the following grades:

- the average grade of the short tests written at the beginning of the practices (an average of at least 2.0 is necessary to avoid a 'fail' final grade) (70 %)

- the average grade of the unknown samples (20 %)

- the average grade of evaluation of the quality and quantity of the prepared compounds and the laboratory notebook prepared by the student (10 %)

A final 'fail' mark can only be improved once if it is due theoretical insufficiency during the short tests.

Person responsible for course: Dr. Péter Buglyó, associate professor, PhD

Tutor: Dr. Péter Buglyó, associate professor, PhD

Title of course: Inorganic Chemistry IV.

ECTS Credit points: 4

Code: TTKBL0202_EN

Type of teaching, contact hours

- lecture: -

- practice: -

- laboratory: 3 hours/week

Evaluation: practice grade

Workload (estimated), divided into contact hours:

- lecture: -

- practice:

- laboratory: 42 hours

- home assignment: 84 hours

- preparation for the exam: -

Total: 126 hours

Year, semester: 2nd year, 2nd semester

Its prerequisite(s): TTKBE0202_EN, TTKBL0201_EN

Further courses built on it:

Topics of course

Physical and chemical properties, laboratory preparation, use, biological role, identification of p-, s-

and d-block metals. The most important compounds and complexes of the metals: chemical reactions, reactivity, preparation and use. Qualitative analysis of the most important metal ions with classical analysis. Synthesis and purification of various metal compounds.

Literature

1. Syllabus provided by the tutor
2. Willard, L.L. Inorganic Chemistry, Wadsworth Publ. Co., Belmont, 1998.
3. Vogel's Qualitative Inorganic Analysis, (ed. Gy. Svehla), Longmann, 2007.

Course objective/intended learning outcomes

a) Knowledge

- He/she knows the most important physical and chemical properties, production, practical use, identification and compounds of the most important s-, p- and d-block metals.
- He/she knows and uses the various chemicals, devices and methods in the chemical laboratory together with the safety knowledge.

b) Abilities

- He/she is able to apply the most important terminology, theories, procedures relating the elements and their most important compounds when completing the relevant tasks.
- He/she is able to find solutions for inorganic laboratory problems.

c) Attitude

- He/she is open to learn and accept inorganic chemical improvement and innovation in his/her profession and convey it genuinely.
- He/she makes a decision in complex and unexpected decision cases by completely taking into account legal and ethical norms.

d) Autonomy and responsibility

- Under supervision he/she is responsible in collaboration with other professionals (especially from the field of inorganic laboratory work and safety).

Schedule:

2 h introduction + 8 x 5 h practice

1st week: Introductory guidance, safety regulations (2 h)

2nd week: Group I metals: preparation, physical and chemical properties, reactions with non-metals. Formation of (per)oxides, hydroxides, chemical reactions. Crown ether complexes. Synthesis of calcium-hydride.

3rd week: Group II metals and important compounds: oxides, hydroxides, carbonates. Poorly soluble compounds of Group I and II metals. Unknown sample with Group I and II metal ions.

4th week: D-block metals: laboratory preparation, chemical properties. Formation and stability of various oxidation states of transition metal ions. General characterization of complexes: basic terms, stability and kinetic behaviour.

5th week: Important compounds of transition metals: preparation, physical and chemical properties. Halogenides, oxides and their derivatives. Acid-base character of the oxides. Iso- and heteropolyacids and the analytical use.

6th week: Preparation of various compounds consisting s-block and transition elements. Hard-soft classification of the metal ions and ligands.

7th week: Halogeno, amino and hydroxo complexes. Cyano and thiocyanato complexes and their analytical significance.

8th week: The Fresenius system and its practical use in the qualitative analysis of metal ions. Cation groups I-V.

9th week: Analysis of an unknown sample with selected p- and d-block metal ions.

Requirements:

- *for a signature*

Participation at practices is compulsory. Students must attend each practice during the semester. In case of absence(s), a medical certificate needs to be presented. Missed practice classes should be made up for at a later date, to be discussed with the tutor.

- *for a grade*

Grading is given by considering the following grades:

- the average grade of the short tests written at the beginning of the practices (an average of at least 2.0 is necessary to avoid a 'fail' final grade) (70 %)

- the average grade of the unknown sample(s), evaluation of the quality and quantity of the prepared compound(s) and the laboratory notebook (30 %)

A final 'fail' mark can only be improved once if it is due theoretical insufficiency during the short tests.

Person responsible for course: Dr. Péter Buglyó, associate professor, PhD

Tutor: Dr. Péter Buglyó, associate professor, PhD

Title of course: Physical Chemistry I

Code: TTKBE0401_EN

ECTS Credit points: 3

Type of teaching, contact hours

- lecture: 2 hours/week

- practice: -

- laboratory: -

Evaluation: exam

Workload (estimated), divided into contact hours:

- lecture: 28 hours

- practice: -

- laboratory: -

- home assignment: -

- preparation for the exam: 60 hours

Total: 88 hours

Year, semester: 1st year, 2nd semester

Its prerequisite(s): TTKBE0101_EN, TTFBE2111_EN, TTMBE0808_EN

Further courses built on it: TTKBE0402_EN, TTKBE0202_EN, TTKBL0201_EN, TTKBG0402_EN, TTKBL0401_EN, TTKBE0302_EN, TTKBE0501_EN, TTKBE0502_EN, TTKBE0601_EN, TTKBG0601_EN, TTKBE0204_EN, TTKBE0417_EN, TTKBG0614_EN, TTKBG0312_EN

Topics of course

The series of lectures are based on the topics of chemical thermodynamics and equilibrium studies. It reviews the fundamental relations of physical chemistry. The course help to build and strengthen the concepts of physical chemistry in the students' scientific view. In this way the basic concepts and phenomena learned, especially in the General Chemistry course (prerequisite) will be placed into more exact and mathematically well-established surrounding. Application of the approach of physical chemistry in chemical engineering is discussed through examples. The main chapters include: Description of gases. Laws of thermodynamics. Thermochemistry. Description of one component and multicomponent systems. Equilibrium.

Literature

Compulsory:

- P.W: Atkins, J. de Paula (2006): Atkins' Physical Chemistry 8th Edition, W.H. Freeman and Company, New York, ISBN: 0-7167-8759-8.
- J. M. Smith, H. C. van Ness, M. M. Abbott (2003): Introduction to Chemical Engineering Thermodynamics 6th Edition, McGraw-Hill, ISBN 0-07-240296-2 ISBN:978-0070494862.
- Lecture notes and teaching material available via the e-learning system.
- Leo Lue (2009): Chemical Thermodynamics, Leo Lue and Ventus Publishing, ISBN: 9788776814977. Can be downloaded from bookboon.com

Recommended:

- A. J. Fletcher (2012): Chemistry for Chemical Engineers, ISBN: 9788740302493. Can be downloaded from bookboon.com
- Other corresponding books from bookboon.com

Course objective/intended learning outcomes

a) Knowledge

- He/She has a basic physical chemistry knowledge on describing simple chemical processes as well as on recognizing, organizing these in practice.
- He/She has a knowledge to solve problems on the field of physical chemistry, using natural sources, and understanding the chemical background of living and non living systems.
- He/She can understand and communicate professionally on subjects of physical chemistry in English.

b) Abilities

- He/She is able to use the previously obtained knowledge on the field of physical chemistry studies to solve practical problems
- He/she is able to create fundamental models of engineering systems and processes.

c) Attitude

- He/She is ready to discuss problems on the field of chemistry and other science with

professionals.

- He/She is able to represent his/her own personal scientific ideology toward professional and unprofessional groups.

d) Autonomy and responsibility

- He/She stands for his/her opinion or ideology in professional discussions

Schedule:

1st week Ideal and real gases

Concepts: Ideal gas and the state equation. Isotherm, isobar and isochor changes. Pressure of gases and its statistical interpretation. Mixture of ideal gases, molar fraction and partial pressure Dalton's law. Real gases, isotherms. Compressibility factor. Real gases and van der Waals equation. Molecular explanation of the pressure and volume correction. The virial equation. List of mathematical tools used during the course. Phenomenological and statistical approach. SI system and units. Dimension analysis. The 0th law of thermodynamics

2nd week The 1st law of thermodynamics

Concepts: Description and formulation of 1st law of thermodynamics. Internal energy and molecular explanation. Work and energy. Volume and other work. Heat. Enthalpy. Internal energy of ideal gases. Joule expansion experiment. Enthalpy and internal energy of real gases. Joule-Thomson experiment and Joule-Thomson coefficient. Liquefaction of gases. . Conservative force fields in physics. Exact differential.

3rd week Thermochemistry

Concepts: Thermochemical equations. Standard state. Reaction heat and its thermodynamic definition. Hess' law. Enthalpy of formation and enthalpy of combustion. Experimental determination of reaction enthalpy. Heat capacity. Temperature dependence of heat capacity. Kirchoff's law.

4th week 2nd law of thermodynamics

Concepts: Description and formulation of the 2nd law. Definition of entropy in thermodynamics and statistical definition. The entropy change of the system and the surrounding during reversible and irreversible isotherm, expansion of ideal gases. Entropy change of adiabatic processes. Transformation of heat into work. Efficiency. Heat engines, refrigerators, heat pumps. Temperature as integral dividend.

5th week 3rd law of thermodynamics

Concepts: Entropy and molar heat capacity. Heat capacity at extreme low temperatures. Absolute zero degree. Description and formulation of the 3rd law. Temperature dependence of entropy. Absolute and standard entropy. Standard reaction entropy. Comparison of phenomenological and statistical approach.

6th week Potential functions in thermodynamics

Concepts: Unification of the 1st and 2nd laws. Maximum useful work and its molecular explanation Free energy (Helmholtz) and free enthalpy (Gibbs) Potential function and their properties. Direction of spontaneous processes. Equilibrium in closed and open systems. Equilibrium and steady state.

7th week Chemical potential

Concepts: Chemical potential and its calculation one component and multicomponent systems. Gibbs–Duhem equation. Chemical potential in two component gas and liquid mixtures, ideal and real solutions Raoult's law and Henry's law. Fugacity and activity and its thermodynamic importance. Choice of standard state. The fundamental equation.

<i>8th week</i>	Thermodynamics of one component systems Concepts: Phase and component. Types of phase transitions. Application of chemical potential in the description of equilibrium of multiple phase one component systems. Phase stability and phase transition. Clapeyron's and Clausius–Clapeyron equation. Liquid-vapour systems, evaporation, boiling, enthalpy of evaporation, boiling point, saturated vapour pressure, entropy of evaporation. Trouton's law and phase diagram. Phase diagram of CO ₂ and water.
<i>9th week</i>	Thermodynamics of two component mixtures and dilute solutions Concepts: Ideal and real mixtures. Partial molar quantities. Partial molar volume and its determination. Thermodynamics of mixing. Excess functions of mixing, enthalpy and entropy of mixing. Colligative properties: melting point depression, boiling point elevation and osmosis. Practical importance and applications of colligative properties.
<i>10th week</i>	Mixture of volatile components Concepts: Vapour pressure of liquid mixtures. Vapour pressure and composition, boiling point-composition equilibrium plots for ideal and real mixtures. Distillation, azeotropic mixtures. Distribution equilibrium, Vapour pressure of non-miscible liquids. Steam distillation.
<i>11th week</i>	Phase rule Concepts: component, phase, degree of freedom. Phase rule. Phase diagram of partially miscible liquids. Eutectics, phase diagram of two component solids. Cooling of two component mixtures. Three component systems and their presentation in triangle diagram.
<i>12th week</i>	Thermodynamic equilibrium in reactive systems. Concepts: Chemical equilibrium. Reaction free energy. Exergonic and endergonic processes. Equilibrium constant. Standard reaction enthalpy and its relation to equilibrium constant and chemical potentials. Determination of equilibrium constant from thermodynamic data. Types of equilibrium constant: K_p , K_x , K_a . Reaction quotient and equilibrium constant.
<i>13th week</i>	Effect of parameters on chemical equilibrium Concepts: Dynamic nature of equilibrium, Le-Chatelier principle. Effect of pressure and temperature on the equilibrium constant, van't Hoff equation. Effect of addition of reactants and products. Practical applications.
<i>14th week</i>	Chemical equilibrium in various systems. Concepts: Types of equilibria: one step, multiple step, parallel, consecutive equilibrium. Equilibrium in homogeneous systems: acid-base, redox and stepwise equilibrium. Dissociation in solution and gas phase, equilibrium of reaction systems. Thermodynamics of ATP. Heterogeneous equilibrium, solubility product, decomposition of solids, adsorption of gases on solids. Buffers. pH scale and calculation of pH. Haber process.

Requirements:

- for a signature

Attendance at **lectures** is recommended, but not compulsory.

During the semester there is a written end-term test in the 14th week. Students have to sit for the tests. The material of the test is the same as the exam. All questions cover several parts of the topics of the lectures and the sub-questions are scored according to the given points.

- for a grade

The course ends in a **written or oral examination**. Based on the result of the examination questions scored according to pre-set maximum points for each sub-questions. The type of the examination (written or oral) is the choice of the student.

The minimum requirement for the examination is 60%. Based on the score of the tests separately, the grade for the tests and/or the examination is given according to the following table:

Score	Grade
0-59	fail (1)
60-69	pass (2)
70-79	satisfactory (3)
80-89	good (4)
90-100	excellent (5)
<p>If the score of any test is below 60, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.</p> <p><i>-an offered grade:</i></p> <p>It may be offered for students on the basis of the result of the end-term test if the grade is at least satisfactory (3).</p>	
<p>Person responsible for course: Dr. Attila Béneyei, associate professor, PhD</p>	
<p>Lecturer: Dr. Attila Béneyei, associate professor, PhD</p>	

<p>Title of course: Physical Chemistry I. Code: TTKBG0401_EN</p>	<p>ECTS Credit points: 2</p>
<p>Type of teaching, contact hours</p> <ul style="list-style-type: none"> - lecture: - - practice: 2 hours/week - laboratory: - 	
<p>Evaluation: practice</p>	
<p>Workload (estimated), divided into contact hours:</p> <ul style="list-style-type: none"> - lecture: - hours - practice: 28 hours - laboratory: - - home assignment: 30 hours - preparation for the exam: - <p>Total: 58 hours</p>	
<p>Year, semester: 1st year, 2nd semester</p>	
<p>Its prerequisite(s): TTKBE0101_EN, TTFBE2111_EN, TTMBE0808_EN, parallel registration to TTKBE0401_EN</p>	
<p>Further courses built on it: -</p>	

<p>Topics of course</p>
<p>The problem solving classes are based on the topics of the lectures in the field of chemical thermodynamics and equilibrium studies. Calculations are made for better understanding the fundamental relations of physical chemistry. The course help to build and strengthen the concepts of physical chemistry in the students' scientific view. In this way the basic concepts and phenomena learned, especially in the General Chemistry course (prerequisite) will be placed into more exact and mathematically well-established surrounding. Application of the approach of physical chemistry in chemical engineering is discussed through examples. The main chapters</p>

include: Description of gases. Laws of thermodynamics. Thermochemistry. Description of one component and multicomponent systems. Equilibrium.

Literature

Compulsory:

- P.W. Atkins, J. de Paula (2006): Atkins' Physical Chemistry 8th Edition, W.H. Freeman and Company, New York, ISBN: 0-7167-8759-8.
- J. M. Smith, H. C. van Ness, M. M. Abbott (2003): Introduction to Chemical Engineering Thermodynamics 6th Edition, McGraw-Hill, ISBN 0-07-240296-2 ISBN:978-0070494862.
- List of problems, their solutions and other teaching material available via the e-learning system.
- L. Lue (2009): Chemical Thermodynamics, Leo Lue and Ventus Publishing, ISBN: 9788776814977. Can be downloaded from bookboon.com

Recommended:

- A. J. Fletcher (2012): Chemistry for Chemical Engineers, ISBN: 9788740302493. Can be downloaded from bookboon.com
- Other corresponding books from bookboon.com

Course objective/intended learning outcomes

a) Knowledge

- He/She has a basic physical chemistry knowledge on describing simple chemical processes as well as on recognizing, organizing these in practice.
- He/She has a knowledge to solve problems on the field of physical chemistry, using natural sources, and understanding the chemical background of living and non living systems.
- He/She can understand and communicate professionally on subjects of physical chemistry in English.

b) Abilities

- He/She is able to use the previously obtained knowledge on the field of physical chemistry studies to solve practical problems
- He/she is able to create fundamental models of engineering systems and processes.

c) Attitude

- He/She is ready to discuss problems on the field of chemistry and other science with professionals.
- He/She is able to represent his/her own personal scientific ideology toward professional and unprofessional groups.

d) Autonomy and responsibility

- He/She stands for his/her opinion or ideology in professional discussions

Schedule:

1st week Ideal and real gases

Problem solving and calculations in the following topics: Ideal gas and the state equation. Isotherm, isobar and isochor changes. Pressure of gases and its statistical interpretation. Mixture of ideal gases, molar fraction and partial pressure Dalton's law. Real gases, isotherms. Compressibility factor. Real gases and van der Waals equation. Molecular explanation of the pressure and volume correction. The virial equation. List of mathematical tools used during the course. Phenomenological and statistical approach. SI system and units. Dimension analysis. The 0th law of thermodynamics

2nd week The 1st law of thermodynamics

Problem solving and calculations in the following topics: Description and formulation of 1st law of thermodynamics. Internal energy and molecular explanation. Work and energy. Volume and other work. Heat. Enthalpy. Internal energy of ideal gases. Joule expansion experiment. Enthalpy and internal energy of real gases. Joule-Thomson experiment and Joule-Thomson coefficient. Liquefaction of gases. . Conservative force fields in physics. Exact differential.

3rd week Thermochemistry

Problem solving and calculations in the following topics: Thermochemical equations. Standard state. Reaction heat and its thermodynamic definition. Hess' law. Enthalpy of formation and enthalpy of combustion. Experimental determination of reaction enthalpy. Heat capacity. Temperature dependence of heat capacity. Kirchoff's law.

4th week 2nd law of thermodynamics

Problem solving and calculations in the following topics: Description and formulation of the 2nd law. Definition of entropy in thermodynamics and statistical definition. The entropy change of the system and the surrounding during reversible and irreversible isotherm, expansion of ideal gases. Entropy change of adiabatic processes. Transformation of heat into work. Efficiency. Heat engines, refrigerators, heat pumps. Temperature as integral dividend.

5th week 3rd law of thermodynamics

Problem solving and calculations in the following topics: Entropy and molar heat capacity. Heat capacity at extreme low temperatures. Absolute zero degree. Description and formulation of the 3rd law. Temperature dependence of entropy. Absolute and standard entropy. Standard reaction entropy. Comparison of phenomenological and statistical approach.

6th week Potential functions in thermodynamics

Problem solving and calculations in the following topics: Unification of the 1st and 2nd laws. Maximum useful work and its molecular explanation Free energy (Helmholtz) and free enthalpy (Gibbs) Potential function and their properties. Direction of spontaneous processes. Equilibrium in closed and open systems. Equilibrium and steady state.

7th week Chemical potential

Problem solving and calculations in the following topics: Chemical potential and its calculation one component and multicomponent systems. Gibbs–Duhem equation. Chemical potential in two component gas and liquid mixtures, ideal and real solutions Raoult's law and Henry's law. Fugacity and activity and its thermodynamic importance. Choice of standard state. The fundamental equation.

8th week Thermodynamics of one component systems

Problem solving and calculations in the following topics: Phase and component. Types of phase transitions. Application of chemical potential in the description of equilibrium of multiple phase one component systems. Phase stability and phase transition. Clapeyron and Clausius–Clapeyron equation. Liquid-vapour systems, evaporation, boiling, enthalpy of evaporation, boiling point, saturated vapour pressure, entropy of evaporation. Trouton's law and phase diagram. Phase diagram of CO₂ and water.

9th week Thermodynamics of two component mixtures and dilute solutions

Problem solving and calculations in the following topics: Ideal and real mixtures. Partial molar quantities. Partial molar volume and its determination. Thermodynamics of mixing. Excess functions of mixing, enthalpy and entropy of mixing. Colligative properties: melting point depression, boiling point elevation and osmosis. Practical importance and applications of colligative properties.

10th week Mixture of volatile components

Problem solving and calculations in the following topics: Vapour pressure of liquid mixtures. Vapour pressure and composition, boiling point-composition equilibrium plots for ideal and real

mixtures. Distillation, azeotropic mixtures. Distribution equilibrium, Vapour pressure of non-miscible liquids. Steam distillation.

11th week Phase rule

Problem solving and calculations in the following topics: component, phase, degree of freedom. Phase rule. Phase diagram of partially miscible liquids. Eutectics, phase diagram of two component solids. Cooling of two component mixtures. Three component systems and their presentation in triangle diagram.

12th week Thermodynamic equilibrium in reactive systems.

Problem solving and calculations in the following topics: Chemical equilibrium. Reaction free energy. Exergonic and endergonic processes. Equilibrium constant. Standard reaction enthalpy and its relation to equilibrium constant and chemical potentials. Determination of equilibrium constant from thermodynamic data. Types of equilibrium constant: K_p , K_x , K_a . Reaction quotient and equilibrium constant.

13th week Effect of parameters on chemical equilibrium

Problem solving and calculations in the following topics: Dynamic nature of equilibrium, Le-Chatelier principle. Effect of pressure and temperature on the equilibrium constant, van't Hoff equation. Effect of addition of reactants and products. Practical applications.

14th week Chemical equilibrium in various systems.

Problem solving and calculations in the following topics: Types of equilibria: one step, multiple step, parallel, consecutive equilibrium. Equilibrium in homogeneous systems: acid-base, redox and stepwise equilibrium. Dissociation in solution and gas phase, equilibrium of reaction systems. Thermodynamics of ATP. Heterogeneous equilibrium, solubility product, decomposition of solids, adsorption of gases on solids. Buffers. pH scale and calculation of pH. Haber process.

Requirements:

- for a signature

Participation at **practice classes** is compulsory. A student must attend the practice classes and may not miss more than three times during the semester. In case a student does so, the subject will not be signed and the student must repeat the course. A student can't make up any practice with another group. Attendance at practice classes will be recorded by the practice leader. Being late is equivalent with an absence. In case of further absences, a medical certificate needs to be presented. Missed practice classes should be made up for at a later date, to be discussed with the tutor. Students are required to bring calculators or computers pencil and ruler to each practice class. Active participation is evaluated by the teacher in every class. If a student's behaviour or conduct doesn't meet the requirements of active participation, the teacher may evaluate his/her participation as an absence because of the lack of active participation in class.

During the semester there are two tests: the mid-term test in the 8th week and the end-term test in the 15th week. Students have to sit for the tests. The problems to be solved and calculated are highly analogous with the ones made available in e-learning system. Scoring system is also provided, i.e. possible maximum points for the given problem.

- for a grade

The course ends with signature and mark. The mark is based on the result of the two tests scored according to pre-set maximum points for each sub-problems.

The minimum requirement for the mark is 60%, based on the score of the tests separately, the grade for the tests is given according to the following table:

Score	Grade
0-59	fail (1)
60-69	pass (2)
70-79	satisfactory (3)

80-89	good (4)
90-100	excellent (5)
If the score of any test is below 60, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.	
Person responsible for course: Dr. Attila Bényei, associate professor, PhD	
Lecturer: Dr. Attila Bényei, associate professor, PhD	

Title of course: Physical Chemistry II. Code: TTKBE0402_EN	ECTS Credit points: 3
Type of teaching, contact hours - lecture: 2 hours/week - practice: - - laboratory: -	
Evaluation: exam	
Workload (estimated), divided into contact hours: - lecture: 28 hours - practice: - - laboratory: - - home assignment: - - preparation for the exam: 60 hours Total: 88 hours	
Year, semester: 2 nd year, 1 st semester	
Its prerequisite(s): TTKBE0401_EN, TTKBE0201_EN, TTKBE0301_EN	
Further courses built on it: TTKBE0403_EN, TTKBE0404_EN, TTKBL0402_EN, TTKBE0405_EN	

Topics of course
The series of lectures are based on the topics of electrochemistry and reaction kinetics It reviews the fundamental relations of physical chemistry. The course help to build and strengthen the concepts of physical chemistry in the students' scientific view. Application of the approach of physical chemistry in chemical engineering is discussed through examples. The main chapters include: Homogeneous and heterogeneous equilibrium electrochemistry. Transport processes. Kinetics of homogeneous and heterogeneous reactions.
Literature
<i>Compulsory:</i> - P.W: Atkins, J. de Paula (2006): Atkins' Physical Chemistry 8th Edition, W.H. Freeman and Company, New York, ISBN: 0-7167-8759-8. - H. S. Fogler (2011): Elements of Chemical Reaction Engineering, 4 th Edition, Prentice Hall, ISBN: 0-13-047394-4. ISBN: 9780130473943. - R.M. Pashley, M. E. Karaman: Applied Colloid and Surface Chemistry. ISBN 13 978-0-470-

86882-9(HB) Teaching material is available via the e-learning system

- R. Kandiyoti (2009): Fundamentals of Chemical Reaction Engineering, R. Kandiyoti and Ventus Publishing, ISBN: 9788776815103. Can be downloaded from bookboon.com

Recommended:

-R. Kandiyoti (2009): Fundamentals of Chemical Reaction Engineering- Examples, R. Kandiyoti and Ventus Publishing, ISBN: 9788776815127. Can be downloaded from bookboon.com

Course objective/intended learning outcomes

a) Knowledge

- He/She has a mathematical and scientific background to understand processes in chemical and chemistry related industries.

b) Abilities

- He/She capable to apply the learned methods, models and planning's of chemical technology and chemical processes through calculations.

c) Attitude

- He/She makes effort to improve and apply the practical methods with new results and experiences.

- He/She can collaborate with other people and discuss their opinions in problem solving processes before making new decisions.

d) Autonomy and responsibility

- He/She tends to establish new solutions and technologies

Schedule:

1st week Homogeneous equilibrium electrochemistry. Thermodynamics of electrolyte solutions
Concepts: Thermodynamic functions of ions. Standard state of ions. Activity in electrolyte solutions. Mean activity coefficient and its experimental determination. Debye-Hückel limiting law. Ionic strength. Solubility equilibria of salts and effect of ionic strength on the solubility. Calculation of solubility from thermodynamic tables Ostwald's law of dilution.

2nd week Heterogeneous equilibrium electrochemistry. Thermodynamics of electrodes.
Concepts: Processes on electrodes. Main types of electrodes, gas electrodes, first and second kind electrodes, redox electrodes. Dependence of electrode potential on concentrations. Nernst equation. Standard electrode potential. Measurement of pH, glass electrode. Redox potentials and their application. Half reactions. Electrochemical series. Electrolysis, industrial applications.

3rd week Thermodynamics of galvanic cells
Concepts: Electrodes and galvanic cells. Diffusion potential and its elimination. Types of batteries. Chemistry of various batteries. Thermodynamics of batteries. The connection between electromotive force and reaction free enthalpy. Thermodynamic parameters from electrochemistry measurement.

4th week Transport processes

Concepts: Flux. Diffusion. Fick I and Fick II law. Stokes-Einstein equation. Diffusion equation and their solutions. Convection, diffusion and chemical reaction. Heat conductance. Viscosity. Uniform discussion of transport processes.

5th week Movement of ions in electrolyte solutions. Conductance of electrolytes

Concepts: Measurement of conductivity and conductance. Molar conductivity and its dependence on concentration. Conductivity of weak and strong electrolytes. Kohlrausch's law. Independent migration of ions. Transference number and its determination. Interaction among moving ions. □

6th week Reaction kinetics. Rate of chemical reactions. Rate law of chemical reactions

Concepts: Definition of reaction rate. Experimental methods to determine reaction rates. Fast reaction kinetics. Flow, relaxation and other techniques. Types of reactors. Rate equation, rate coefficient and order of reaction. Experimental methods to determine rate equation. Methods to evaluate experimental results.

7th week Kinetics of simple reactions

Concepts: Formal kinetics. Rate equation of first and second order reactions. Integral forms. Third order reactions. Formal kinetics of equilibrium. Consecutive reactions. Rate determining step. Half-life methods.

8th week Complex reaction systems

Concepts: Elementary reactions and molecularity. Simplification of reaction rate determination, flooding or isolation. Steady state and pre-equilibrium. Unimolecular reactions and their Lindemann-Hinshelwood □ mechanism. Enzyme reactions, Michaelis-Menten mechanism.

9th week Reaction encounters

Concepts: Basic steps of chain reactions: initiation, propagation, branching, termination. Formation of hydrogen halogenides. Thermal and chain explosion, explosion limits. The Hinshelwood-Semenov mechanism. Catalysis, formal kinetic description and energetics of catalysis. Homogeneous and heterogeneous catalytic systems Autocatalysis and chemical feedback. Continuous and open reactors.

10th week Collision theory of chemical reactions

Concepts: Temperature dependence of rate coefficient, Arrhenius equation. Activation energy. Collision theory, its basic assumptions. Interpretation and calculation of pre-exponential factor. Steric factor, the anchoring mechanisms, Diffusion driven and activation energy driven reactions.

11th week The activated complex theory of chemical reactions

Concepts: The history of development of activated complex theory and the basic assumptions of the theory. Activated complex and its concentration, experimental evidences. Statistical mechanics in the activated complex theory. Thermodynamic approach in the activated complex theory. Activation free enthalpy activation enthalpy and entropy. Determination of activation parameters Non-thermal activation. Basics of photochemistry, industrial applications.

12th week Processes on solid surfaces

Concepts: Structure of solids and surfaces. Physisorption and chemisorption, their properties and differentiation. Isotherms, Langmuir □ and BET □ isotherms, basic assumptions of the models. Adsorption enthalpy. Basic steps of surface processes, possible rate determining step. Heterogeneous catalysis, the Langmuir-Hinshelwood □ and Eley-Rideal mechanisms. Heterogeneous catalytic processes in the chemical industry. Solid-liquid interface in electrochemistry. Basics of dynamic electrochemistry.

13th week Physical chemistry of colloid

Concepts: Introduction to the nature of colloidal systems, types of colloidal systems. The concept of surface tension. Wetting and spreading. Curved surfaces. Electric double layer, electrokinetic potential. The colloid stability.

14th week Application of colloids, nanoparticles

Concepts: Coherent incoherent systems. The basics of rheology. Liphobic colloids: aerosols, liosols, xerosols. Applications of colloids: nanoparticles, emulsions, suspebsions foams.

Liphilic colloids: association and macromolecular systems. The theory of surfactants and cleaning

Requirements:

- for a signature

Attendance at **lectures** is highly recommended, but not compulsory.

During the semester there is a written end-term test in the 14th week. Students have to sit for the tests. The material of the test is the same as the exam. All questions cover several parts of the topics of the lectures and the sub-questions are scored according to the given points.

- for a grade

The course ends in a **written or oral examination**. Based on the result of the examination questions scored according to pre-set maximum points for each sub-questions. The type of the examination (written or oral) is the choice of the student.

The minimum requirement for the examination is 60%. Based on the score of the tests separately, the grade for the tests and/or the examination is given according to the following table:

Score	Grade
0-59	fail (1)
60-69	pass (2)
70-79	satisfactory (3)
80-89	good (4)
90-100	excellent (5)

If the score of any test is below 60, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

-an offered grade:

It may be offered for students on the basis of the result of the end-term test if the grade is at least satisfactory (3).

Person responsible for course: Dr. Attila Bényei, associate professor, PhD

Lecturer: Dr. Attila Bényei, associate professor, PhD

Title of course: Physical Chemistry II.

Code: TTKBG0402_EN

ECTS Credit points: 2

Type of teaching, contact hours

- lecture: -
- practice: 2 hours/week
- laboratory: -

Evaluation: practice

Workload (estimated), divided into contact hours:

- lecture: -
- practice: 28 hours
- laboratory: -
- home assignment: 30 hours
- preparation for the exam: -

Total: 58 hours

Year, semester: 2nd year, 1st semester

Its prerequisite(s): TTKBE0401_EN, TTKBE0201_EN, TTKBE0301_EN, parallel registration to TTKBE0402_EN

Further courses built on it: -

Topics of course

The problem solving classes are based on the topics of the lectures in the field of electrochemistry, reaction kinetics, and colloid chemistry. Calculations are made for better understanding the fundamental relations of physical chemistry. The course help to build and strengthen the concepts of physical chemistry in the students' scientific view. In this way the basic concepts and phenomena learned, especially in the General Chemistry course (prerequisite) will be placed into more exact and mathematically well-established surrounding. Application of the approach of physical chemistry in chemical engineering and industry is discussed through examples. The main chapters include: Homogeneous and heterogeneous equilibrium electrochemistry. Transport processes. Kinetics of homogeneous and heterogeneous reactions. Physical chemistry of colloids.

Literature

Compulsory:

- P.W. Atkins, J. de Paula (2006): Atkins' Physical Chemistry 8th Edition, W.H. Freeman and Company, New York, ISBN: 0-7167-8759-8.
- H. S. Fogler (2011): Elements of Chemical Reaction Engineering, 4th Edition, Prentice Hall, ISBN: 0-13-047394-4. ISBN: 9780130473943.
- R. M. Pashley, M. E. Karaman: Applied Colloid and Surface Chemistry. ISBN 13 978-0-470-86882-9(HB)
- List of problems, their solutions and other teaching material available via the e-learning system.
- R. Kandiyoti (2009): Fundamentals of Chemical Reaction Engineering, R. Kandiyoti and Ventus Publishing, ISBN: 9788776815103. Can be downloaded from bookboon.com

Recommended:

- R. Kandiyoti (2009): Fundamentals of Chemical Reaction Engineering- Examples, R. Kandiyoti and Ventus Publishing, ISBN: 9788776815127. Can be downloaded from bookboon.com

Course objective/intended learning outcomes

a) Knowledge

- He/She has a mathematical and scientific background to understand processes in chemical and chemistry related industries.

b) Abilities

- He/She capable to apply the learned methods, models and planning's of chemical technology and chemical processes through calculations.

c) Attitude

- He/She makes effort to improve and apply the practical methods with new results and experiences.
- He/She can collaborate with other people and discuss their opinions in problem solving processes before making new decisions.

d) Autonomy and responsibility

- He/She tends to establish new solutions and technologies .

Schedule:

1st week Homogeneous equilibrium electrochemistry. Thermodynamics of electrolyte solutions
Problem solving and calculations in the following topics: Thermodynamic functions of ions. Standard state of ions. Activity in electrolyte solutions. Mean activity coefficient and its experimental determination. Debye-Hückel limiting law. Ionic strength. Solubility equilibria of salts and effect of ionic strength on the solubility. Calculation of solubility from thermodynamic tables Ostwald's law of dilution.

2nd week Heterogeneous equilibrium electrochemistry. Thermodynamics of electrodes.
Problem solving and calculations in the following topics: Processes on electrodes. Main types of electrodes, gas electrodes, first and second kind electrodes, redox electrodes. Dependence of electrode potential on concentrations. Nernst equation. Standard electrode potential. Measurement of pH, glass electrode. Redox potentials and their application. Half reactions. Electrochemical series. Electrolysis, industrial applications.

3rd week Thermodynamics of galvanic cells
Problem solving and calculations in the following topics: Electrodes and galvanic cells. Diffusion potential and its elimination. Types of batteries. Chemistry of various batteries. Thermodynamics of batteries. The connection between electromotive force and reaction free enthalpy. Thermodynamic parameters from electrochemistry measurement.

4th week Transport processes
Problem solving and calculations in the following topics: Flux. Diffusion. Fick I. and Fick II. law. Stokes-Einstein equation. Diffusion equation and their solutions. Convection, diffusion and chemical reaction. Heat conductance. Viscosity. Uniform discussion of transport processes.

5th week Movement of ions in electrolyte solutions. Conductance of electrolytes
Problem solving and calculations in the following topics: Measurement of conductivity and conductance. Molar conductivity and its dependence on concentration. Conductivity of weak and strong electrolytes. Kohlrausch's law. Independent migration of ions. Transference number and its determination. Interaction among moving ions.

6th week Reaction kinetics. Rate of chemical reactions. Rate law of chemical reactions
Problem solving and calculations in the following topics: Definition of reaction rate. Experimental methods to determine reaction rates. Fast reaction kinetics. Flow, relaxation and other techniques. Types of reactors. Rate equation, rate coefficient and order of reaction. Experimental methods to determine rate equation. Methods to evaluate experimental results.

7th week Kinetics of simple reactions
Problem solving and calculations in the following topics: Formal kinetics. Rate equation of first and second order reactions. Integral forms. Third order reactions. Formal kinetics of equilibrium. Consecutive reactions. Rate determining step. Half-life methods.

8th week Complex reaction systems
Problem solving and calculations in the following topics: Elementary reactions and molecularity. Simplification of reaction rate determination, flooding or isolation. Steady state and pre-equilibrium. Unimolecular reactions and their Lindemann-Hinshelwood mechanism. Enzyme reactions, Michaelis-Menten mechanism.

9th week Reaction encounters
Problem solving and calculations in the following topics: Basic steps of chain reactions: initiation, propagation, branching, termination. Formation of hydrogen halogenides. Thermal and chain explosion, explosion limits. The Hinshelwood-Semenov mechanism. Catalysis, formal kinetic

description and energetics of catalysis. Homogeneous and heterogeneous catalytic systems Autocatalysis and chemical feedback. Continuous and open reactors.

10th week Collision theory of chemical reactions

Problem solving and calculations in the following topics: Temperature dependence of rate coefficient, Arrhenius equation. Activation energy. Collision theory, its basic assumptions. Interpretation and calculation of pre-exponential factor.

Steric factor, the anchoring mechanisms, Diffusion driven and activation energy driven reactions.

11th week The activated complex theory of chemical reactions

Problem solving and calculations in the following topics: The history of development of activated complex theory and the basic assumptions of the theory. Activated complex and its concentration, experimental evidences. Statistical mechanics in the activated complex theory. Thermodynamic approach in the activated complex theory. Activation free enthalpy activation enthalpy and entropy. Determination of activation parameters Non-thermal activation. Basics of photochemistry, industrial applications.

12th week Processes on solid surfaces

Problem solving and calculations in the following topics: Structure of solids and surfaces. Physisorption and chemisorption, their properties and differentiation. Isotherms, Langmuir and BET isotherms, basic assumptions of the models. Adsorption enthalpy. Basic steps of surface processes, possible rate determining step. Heterogeneous catalysis, the Langmuir–Hinshelwood and Eley–Rideal mechanisms. Heterogeneous catalytic processes in the chemical industry. Solid-liquid interface in electrochemistry. Basics of dynamic electrochemistry.

13th week Physical chemistry of colloid

Problem solving and calculations in the following topics: Introduction to the nature of colloidal systems, types of colloidal systems. The concept of surface tension. Wetting and spreading. Curved surfaces. Electric double layer, electrokinetic potential. The colloid stability.

14th week Application of colloids, nanoparticles

Problem solving and calculations in the following topics: Coherent incoherent systems. The basics of rheology. Liphobic colloids: aerosols, liosols, xerosols. Applications of colloids: nanoparticles, emulsions, suspensions, foams.

Liphilic colloids: association and macromolecular systems. The theory of surfactants and cleaning

Requirements:

Participation at **practice classes** is compulsory. A student must attend the practice classes and may not miss more than three times during the semester. In case a student does so, the subject will not be signed and the student must repeat the course. A student can't make up any practice with another group. Attendance at practice classes will be recorded by the practice leader. Being late is equivalent with an absence. In case of further absences, a medical certificate needs to be presented. Missed practice classes should be made up for at a later date, to be discussed with the tutor. Students are required to bring calculators or computers pencil and ruler to each practice class. Active participation is evaluated by the teacher in every class. If a student's behaviour or conduct doesn't meet the requirements of active participation, the teacher may evaluate his/her participation as an absence because of the lack of active participation in class.

During the semester there are two tests: the mid-term test in the 8th week and the end-term test in the 15th week. Students have to sit for the tests. The problems to be solved and calculated are highly analogous with the ones made available in e-learning system. Scoring system is also provided, i.e. possible maximum points for the given problem.

- for a grade

The course ends with signature and mark. The mark is based on the result of the two tests scored according to pre-set maximum points for each sub-problems.

The minimum requirement for the mark is 60%, based on the score of the tests separately, the grade for the tests is given according to the following table:

Score	Grade
0-59	fail (1)
60-69	pass (2)
70-79	satisfactory (3)
80-89	good (4)
90-100	excellent (5)

If the score of any test is below 60, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

Person responsible for course: Dr. Attila Bényei, associate professor, PhD

Lecturer: Dr. Attila Bényei, associate professor, PhD

Title of course: Introduction to Physical Chemistry Measurements Code: TTKBL0401_EN	ECTS Credit points: 4
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Type of teaching, contact hours

- lecture: -
- practice: -
- laboratory: 4 hours/week

Evaluation: practice

Workload (estimated), divided into contact hours:

- lecture: -
- practice: -
- laboratory: 56 hours
- home assignment: 64 hours
- preparation for the exam: -

Total: 120 hours

Year, semester: 2nd year, 1st semester

Its prerequisite(s): TTKBL0101_EN, TTKBE0201_EN, TTKBE0301_EN, TTKBE0401_EN

Further courses built on it: TTKBL0402_EN

Topics of course

This course is intended to teach students how to perform basic physical chemistry measurements and to get familiar with measurement planning, data processing and report writing. The tasks detailed here contain mainly spectrophotometry, pH-potentiometry, gas-volumetric measurements, polarimetry, electrolysis, conductometry etc.

Set of measurements:

101. Measuring densities by pycnometer, composition of a binary mixture
102. Measuring the heat capacities of metals by calorimetry
103. Measuring electrical conductivity of solutions

104. Measuring the concentration of a coloured solute by spectrophotometry
105. Determination of NaHCO₃ content of a solid sample by gas volumetry
106. pH-metric titration curves of hydrochloric and acetic acids
107. Study of Cooling Curve
108. Study of electrolysis
109. Mutarotation of glucose measured by polarimetry
110. Measuring electromotive force of a galvanic cell
111. Refractometry and viscosimetry
112. Determination of enthalpy of dissolution
113. Investigation of redox electrodes
114. Conductometry
115. Reaction rate of decomposition of H₂O₂ measured by gas volumetry
116. Investigation of buffers
117. Electrochemical investigation of corrosion
118. Distillation of an alcohol-water mixture

Literature

- Laboratory notes and additional teaching materials available via the e-learning system.
- P.W. Atkins, J. de Paula: Atkins, Physical Chemistry 8th Edition, W.H. Freeman and Company, New York, ISBN: 0-7167-8759-8, 2006
- K. Ósz, A. Bényei: Physical Chemistry Laboratory Measurements (for students of Pharmacy, Chemistry and Chemical Engineering). Debreceni Egyetemi Kiadó, ISBN: 978-963-318-143-0, 2011

Course objective/intended learning outcomes

a) Knowledge

- He/She knows the basic qualitative and quantitative chemical principles, and the methods based on it.
- He/She knows and can use the properties of most important chemicals, lab equipments and basic laboratory processes.
- He/She can follow and recognize the requirements of safety instructions.
- He/She has knowledge to solve problems on the field of natural processes, using natural sources, and understanding the chemical background of living and non living systems.

b) Abilities

- He/She is able to use the previously obtained knowledge on the field of natural and anthropogenic studies to solve practical problems.
- He/She is able to use the previously learned paradigms, theories and laws on the field of chemical, natural science to plan, execute and evaluate laboratory experiments.
- He/She is able to use his/her chemical knowledge and experiences.
- He/She is able to create and execute experiments to demonstrate and prove the basic chemical phenomena.

c) Attitude

- He/She is ready to discuss problems on the field of chemistry and other science with professionals.
- He/She is open toward scientific and other post-gradual education.

d) Autonomy and responsibility

- Under supervision he/she is responsible in collaboration with other professionals (especially from the field of environmental economy and safety).
- He/She shows responsibility to gain enough experience to be involved in making decision for labs or industrial plants.

Schedule: One of the measurements listed above (**Topics of course**) per week except the 1st practice (introduction, general information and safety training).

Requirements:

Participation on the laboratory practice is compulsory. The measurements and knowledge of the associated theory are marked and an overall mark will be given. Safety training (1st week) is mandatory before the first lab practice (2nd week). Everybody should work individually according to the pre-set schedule (which will be provided on the 1st week). Lab practices are 4 hours long every week (from the 2nd until the 14th weeks). Being late or failed mark on the written test from the appropriate measurement is equivalent with an absence. In accordance with the regulations of University of Debrecen, attendance is compulsory with the exception of health or family problems (the reason of absence should be certified). In this case, the students should agree with the teacher on replacement dates for the missed experiments.

Requirements for the grade:

The measurements (regularly) and written tests (occasionally) according to the knowledge of the associated theory are marked and the overall mark will be given based on these.

- At least 11 notebooks of the measurements (from the 13) have to be marked as “pass (2)” or better for the successful completion.

- The minimum requirement for the written tests is 60%. Based on the score of the tests separately, the grade for the tests is given according to the following table:

Score	Grade
0-59	fail (1)
60-69	pass (2)
70-79	satisfactory (3)
80-89	good (4)
90-100	excellent (5)

If the average of written tests is below 60% the best grade for the course can be only “pass (2)” in any other cases the final mark is given with weighted average by means of the mark of the written tests and notebooks in 1 to 2 ratio.

Person responsible for course: Dr. Ferenc K. Kálmán, assistant professor, PhD

Lecturer: Dr. Ferenc K. Kálmán, assistant professor, PhD

Title of course: Physical chemistry III.

Code: TTKBE0403_EN

ECTS Credit points: 3

Type of teaching, contact hours

- lecture: 2 hours/week
- practice: -
- laboratory: -

Evaluation: exam

Workload (estimated), divided into contact hours:

- lecture: 28 hours
- practice: -
- laboratory: -

<ul style="list-style-type: none"> - home assignment: 22 hours - preparation for the exam: 40 hours <p>Total: 90 hours</p>
Year, semester: 2 nd year, 2 nd semester
Its prerequisite(s): TTKBE0402_EN
Further courses built on it: TTKBE0504_EN, TTKBE0415_EN

<p>Topics of course</p> <ul style="list-style-type: none"> - Basic properties of interfaces. - Adsorption. - Electric double layer. - Kinetics of heterogeneous reactions. - Heterogeneous catalysis. - Dynamic electrochemistry. - Practical applications of electrochemistry. - Definition, discovery, application of radioactivity. - Parts, structure of atomic nucleus, stable and radioactive nuclei. - Kinetics of radioactive decay. - Mechanism and type of radioactive decay. - Interaction of radiation with matter. - Nuclear reactions, nuclear energy production. - Chemical and biological effects of radiation. - Detection and measurement of radiation. - Environmental radioactivity.
<p>Literature</p> <p><i>Compulsory:</i></p> <ul style="list-style-type: none"> - Atkins, P.W. 1990. Physical Chemistry, Oxford University Press, Oxford. - Kónya, J., Nagy N.M., 2012, 2018. Nuclear and Radiochemistry, Elsevier, Oxford. - Choppin, G.R., Liljenzin, J-O., Rydberg, J. Ekberg, C., 2013. Radiochemistry and Nuclear Chemistry, 4th Edition, Elsevier, Amsterdam.
<p>Course objective/intended learning outcomes</p> <p>a) Knowledge</p> <p>- He/She has a basic chemical knowledge on interfaces and radioactivity, formation and properties of interfaces, methods for studying interfaces, adsorption processes of gases and liquids, formation and properties of electric double layer, kinetics of interfacial processes, mechanism of heterogeneous catalysis, formulas describing electrode processes under dynamic conditions, kinetics and types of radioactive decay, interactions of radiation with matter, basic process of nuclear energy production, abundance of radioactive isotopes in the environment, effect of radioactivity on biological systems, detection and measurement of radiation.</p> <p>b) Abilities</p> <p>- He/She is able to use the obtained chemical knowledge on the field of interfaces, adsorption, electrode reactions, heterogeneous catalysis, atomic nuclei, radioactivity, and interaction of</p>

radiation with matter, to solve the actual basic problems and prove the obtained solution.

- He/She is able to argue on scientific problems by his/her knowledge.

c) Attitude

- He/She is open toward scientific and other postgradual education.

d) Autonomy and responsibility

- He/She can make reasonable evaluations about his/her own work comparing to others to the same field.

- He/She can evaluate his/her co-workers work's responsibly in both laboratory and industrial environment, and report it to his/her chief.

Schedule:

1st week

Formation and properties of interfaces, methods for studying interfaces. Interfacial microscopic and macroscopic properties, surface analytical methods

2nd week

Interfacial thermodynamics: adsorption of gases on solid surface, adsorption isotherms, determination of surface area. Decrease of surface energy by adsorption, quantitative description of the process

3rd week

Solid/liquid interfaces, electric double layer. Surface excess concentration on solid/liquid interfaces, the role of interfacial electric properties

4th week

Kinetics of interfacial reactions. Heterogeneous catalysis. Steps of heterogeneous reactions, rate-determining step. Applications of heterogeneous catalysis

5th week

Dynamic electrochemistry. Rate of charge transport, activation free energy, relations of current and voltage (Erdey-Grúz and Volmer theory), exchange current, overpotential, polarization. Definitions and relations on electrode reactions.

6th week

Effects determining the rate of charge transfer, the influence of transport on kinetics of electrode reaction: diffusion, migration, and convection. Diffusion current, diffusion limit. Effects determining the electron transfer, selection of potential determining process

7th week

Electrochemistry in practice, electrolysis, voltage sources, industrial electrochemical processes, corrosion and passivity. Application of electrochemistry

8th week

Discovery of radioactivity, consequences. Properties and constituents of nucleus. Stable and radioactive nuclei. Isobar nuclei. Cause and result of radioactive decay. Radioactivity is a natural process. Scientific and practical consequences of the discovery of radioactivity. Stability/radioactivity of atomic nuclei, decay type are determined by the ratio of protons to neutrons.

9th week

Kinetics of radioactive decay. Simple radioactive decay. Branching decay. Successive decay. Radioactive equilibria: secular and transient equilibria. Natural decay series. Formulas expressing the kinetics of radioactive decay. Radioactive equilibria in nature.

10th week

Radioactive nuclei. Types of radioactive decay. Alpha, beta decays, electron capture, isomeric

transition (gamma radiation). Spontaneous fission. Interaction of radiation with matter. Probability, of the interactions. Interaction of alpha radiation with matter. Types of radioactive decay, emitted particles and photons. Energy release. General aspects of radiation-matter interactions.

11th week

Interaction of beta radiation with matter: ionization, Brehmsstrahlung, Cherenkov radiation, annihilation (positron emission tomography), back-scattering, absorption, self-absorption. Interaction of gamma and X-ray radiation with matter: Compton scattering, photoelectric effect, pair formation. General aspects of the interaction of beta radiation and high energy electromagnetic radiation, respectively, with matter.

12th week

Nuclear reactions, Conservation rules, kinetics. Nuclear reactions with neutrons. Nuclear reactions with charged particles. Nuclear energy production. Fission reaction with thermal neutrons. Main parts of nuclear reactors. Breeder reactors. Transformation of atomic nuclei. Basic reaction of nuclear energy production. Structure of nuclear reactors.

13th week

Environmental aspects of nuclear energy production, disposal of nuclear waste.

Detection and measurement of nuclear radiation. Detectors, electric units. Ionization, scintillation, semiconductor detectors. Imaging of radiation. Positive and negative impacts of nuclear energy production. Detection and measurement of radiation.

14th week

Dosimetry. Irradiation, absorbed, effective doses. The effect of nuclear radiation on living organisms: physical, chemical, biological effects, Radiolysis of water. Dose limits. Natural and artificial radionuclide in the environment. Effect of radiation on living organisms. Sources and quantity of environmental radioactivity.

Requirements:

- for a signature

Attendance at **lectures** is recommended, but not compulsory.

- for a grade

The course ends in an **examination**. Based on the examination, the exam grade is given according to the following table:

Score	Grade
0-59	fail (1)
60-69	pass (2)
70-79	satisfactory (3)
80-89	good (4)
90-100	excellent (5)

If the score of the examination is below 60, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

-an offered grade:

it may be offered for students if they write a test on the 14th week and the score of it is at least 60%. The offered grade is calculated as the exam grade (see above).

Person responsible for course: Dr. Noémi Nagy, professor, DSc

Lecturer: Dr. Noémi Nagy, professor, DSc

Dr. István Bányai, professor, DSc

Title of course: Physical Chemistry IV. Code: TTKBE0404_EN	ECTS Credit points: 5
Type of teaching, contact hours - lecture: 2 hours/week - practice: 2 hours/week - laboratory: -	
Evaluation: exam	
Workload (estimated), divided into contact hours: - lecture: 28 hours - practice: 28 hours - laboratory: - - home assignment: 40 hours - preparation for the exam: 50 hours Total: 146 hours	
Year, semester: 3 rd year, 1 st semester	
Its prerequisite(s): TTKBE0402_EN	
Further courses built on it: -	

Topics of course
The series of lectures are based on the topics of computational chemistry. Both theoretical outlook on the basics of quantum mechanics and their application in simple computational problems are discussed. The main topics include Symmetry properties and their description by group theory. Basics of wave mechanics, wave functions and operators. The different formulations of Schrödinger's equation and methods to solve them. Importance of rotational, vibrational and electron-energies in chemistry. Electronic and magnetic properties of particles. Diffraction methods as information sources from the structure of particles. Schrödinger equation: wavefunction. Hartree-Fock Theory. Density Functional Theory: electron density. Structural analysis. Understanding chemical bond theories In the seminars computational practices help the deeper understanding the theory. Simple quantum chemical calculations are made.
Literature
<i>Compulsory:</i> - P.W. Atkins, J. de Paula (2006): Atkins' Physical Chemistry 8 th Edition, W.H. Freeman and Company, New York, ISBN: 0-7167-8759-8. - Lecture notes and teaching material available via the e-learning system - http://employees.oneonta.edu/viningwj/Chem111/Chapters9-10.pdf - https://fns.uniba.sk/fileadmin/prif/chem/kag/Bakalar/vch_noga/GEN_INORG_CHEM03-04.pdf
<i>Recommended:</i> - http://folk.uio.no/helgaker/talks/la.pdf - http://www.chem.uci.edu/~lawm/9-28.pdf
Course objective/intended learning outcomes
a) Knowledge - He/She knows the main models and theories of chemical bonds and molecular structure based

on scientific findings.

- He/She has a basic chemical knowledge on describing simple chemical processes as well as on recognizing, organizing these in practice.

- He/She has a knowledge to solve problems on the field of natural processes, using natural sources, and understanding the chemical background of living and non-living systems.

- He/She has the knowledge to test or measure chemical reactions, systems with scientific methods (including computational) under supervision.

b) Abilities

- He/She is able to understand the natural and anthropogenic chemical reactions, and capable for data collection, evaluation on these fields including the data mining from literature.

- He/She is able to use the previously obtained knowledge on the field of natural and anthropogenic studies to solve practical problems.

- He/She is able to argue on scientific problems by his/her knowledge.

c) Attitude

- He/She pursues to use his/her knowledge to understand and describe the relationship of chemical processes connected to human life.

- He/She is ready to discuss problems on the field of chemistry and other science with professionals.

- He/She is committed learn or get insights into new competence or ideology.

d) Autonomy and responsibility

- He/She stands for his/her opinion or ideology in professional discussions.

- He/She can make reasonable evaluations about his/her own work comparing to others to the same field.

Schedule:

1st week Properties of symmetries and point groups

Concepts: Symmetry elements and their operators. Symmetry calculation at operator level
Symmetry properties and point group of a body. Character tables and their use. Physical quantities, operators and eigenvalues.

Calculations: Interactive learning of symmetry elements.

2nd week Schrödinger equation: operators and eigenvalues

Concepts: Basics of quantum mechanics. Physical definitions and operators. Operators of energy, and momentum. The concept and use of eigenvalues and equations.

Calculations: Computational possibilities, the role of supercomputers.

3rd week Electron structure of hydrogen atom, quantum numbers

Concepts: Polar coordinates. New forms of Schrödinger equation. Radial and angular part of the solution. Quantum numbers.

Calculations: Particle in the box. Solution of the Schrödinger equation for hydrogen atoms

4th week Hartree-Fock theory: approximations

Concepts: Approximations in computational chemistry. Pauli principle. Born-Oppenheimer approximation. Hierarchy of approximations. Hartree-Fock model. One nucleus and multiple electron systems. Term symbols and their information content and use.

Calculations: Hartree-Fock method in practice. Limitations.

5th week Geometry and structural parameters of molecules: VSPER theory

Concepts: VSPER theory. Structure of covalent molecules. Hybridization. Molecule geometry and hybridization.

Calculations: Use of symmetry considerations in VSPER theory.

<i>6th week</i>	Molecular Orbitals: LCAO-MO theory
Concepts: Molecular orbitals. Linear combination of atomic orbitals. Use of symmetry and point groups in description of molecular orbitals.	
Calculations: Software resources based on LCAO-MO theory.	
<i>7th week</i>	Density Functional Theory: electron density analysis
Concepts :Density functional theory and its use in understanding material properties. Meaning of abbreviations for the level of theory.	
Calculations: DFT calculation basics.	
<i>8th week</i>	Frequency and rotation
Concepts: Microwave and infrared spectra. Rotational and vibrational energies of simple molecules.	
Calculation: Use of character tables in vibrational and Raman spectroscopy.	
<i>9th week</i>	Dipole moment, magnetic properties
Concepts: Dipole moment, polarizability, refractive index, dispersion and their interconnection. Magnetic properties of materials, paramagnetic, diamagnetic and ferromagnetic systems Magnetic susceptibility. Basics of NMR and EPR. MRI, biological and medical applications.	
Calculations: Simple quantum chemical calculations. Use of softwares.	
<i>10th week</i>	Excitation states: spectroscopy
Concepts: Excitation methods and excitons as particles. Quenching. Practical applications	
Calculations: Simple quantum chemical calculations. Use of softwares.	
<i>11th week</i>	Diffraction methods I
Concepts: Symmetry in the solid state. Space groups. Basics of electron, neutron and X-ray diffraction. The limitations of the methods and their application in solving chemical structural problems	
Calculations: Interactive learning of space group notation. and symmetry elements in the solid states.	
<i>12th week</i>	Diffraction methods II
Concepts: The information content of diffraction studies. Possibilities for the experimental determination of electron density	
Calculations: Use of crystallographic databases.	
<i>13th week</i>	Spectroscopic method
Concepts: General description of spectroscopic methods. Applicability of the methods in structure determination. Oscillator strength and experimental consequences.	
Calculations: Spectroscopy of iodine.	
<i>14th week</i>	Secondary interactions
Concepts: List and energetics of secondary interactions. The possibility to predict solid state structures.	
Calculations: Simple quantum chemical calculations. Use of softwares.	

Requirements:

- for a signature

Attendance at **lectures** is recommended, but not compulsory.

Participation at **practice classes** is compulsory. A student must attend the practice classes and may not miss more than three times during the semester. In case a student does so, the subject will not be signed and the student must repeat the course. A student can't make up any practice with another group. Attendance at practice classes will be recorded by the practice leader. Being late is equivalent with an absence. In case of further absences, a medical certificate needs to be presented. Missed practice classes should be made up for at a later date, to be discussed with the tutor.

Students are required to bring the drawing tasks and drawing instruments of the course to each practice class. Active participation is evaluated by the teacher in every class. If a student's behaviour or conduct doesn't meet the requirements of active participation, the teacher may evaluate his/her participation as an absence because of the lack of active participation in class.

-for a grade

The course ends in an **examination**. The exam can be **written or oral**, choice of the student. In any case one more comprehensive and 3 smaller questions should be answered from the topics of the course. Scoring system is provided and percentage is calculated.

The minimum requirement for the examination is 60%. The grade of the examination is given according to the following table:

Score(%)	Grade
0-59	fail (1)
60-69	pass (2)
70-79	satisfactory (3)
80-89	good (4)
90-100	excellent (5)

If the score of any test is below 60, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

-an offered grade:

It may be offered for students on the basis of a written test (not compulsory) sit at the last week of the semester. Students can perform literature search and can submit a written report of that. Possible topics are given during the semester. The report should be at least 10 A4 pages and graded by the lecturers. The offered grade should be at least satisfactory (3).

Person responsible for course: Dr. Attila Bényei, associate professor, PhD

Lecturer: Dr. Attila Bényei, associate professor, PhD
Dr. Mihály Purgel, assistant professor, PhD

Title of course: Physical Chemistry V. Code: TTKBL0402_EN	ECTS Credit points: 5
Type of teaching, contact hours - lecture: - - practice: - - laboratory: 4 hours/week	
Evaluation: practice	
Workload (estimated), divided into contact hours: - lecture: - - practice: - - laboratory: 56 hours - home assignment: 94 hours - preparation for the exam: -	

Total: 150 hours

Year, semester: 3rd year, 1st semester

Its prerequisite(s): TTKBL0401_EN, TTKBE0402_EN

Further courses built on it: -

Topics of course

The main goal of this laboratory measurement is to get more detailed knowledge in the field of basic methodology of physical chemistry and perform basic laboratory practices. It gives the chance for the students to prove theories of physical and colloidal chemistry by their own hands and to determine physical and colloidal chemical quantities.

Topics of measurements:

- determination of material characteristic data
- determination of thermodynamic parameters
- study of amphoteric and complex ions
- electrochemical measurements
- reaction kinetics measurements
- investigation of photochemical reaction
- adsorption on a solid-liquid interface, determination and analysis of an isotherm
- critical micelle concentration of association colloids
- solubilization capacity of surfactants
- determination of surface tension and the Gibbs isotherm
- rheological investigation of different samples
- determination of the isoelectric point of isolabile proteins
- isotope dilution analysis
- backscattering of β radiation
- measurement of γ spectra

Set of measurements:

201. Determination of heat of combustion by using a bomb calorimeter
202. Thermodynamic quantities by measuring the temperature dependent EMF
203. Determination of partial molar volumes by measuring densities
204. Determination of the enthalpy and entropy of vaporization of liquids
205. Redox potentials from potentiometric titrations
206. Investigation of Kohlrausch's law
207. Determination of activity coefficient for concentration galvanic cell
208. Determination of diffusion coefficient by layered ("schlieren") method
209. Study of the photochemical degradation of tris(oxalato)iron(III) complex
210. Determination of protonation constants of an indicator
211. Study of the iodine-iodide equilibrium
212. Dissociation constant of weak acids measured by conductometry
213. Dissociation equilibria of amphoteric, determination of isoelectric pH
214. Study of stepwise complex formation
215. Decomposition kinetics of Kalmopyrin
216. Acid catalysed hydrolysis of saccharose
217. Kinetics of a second order reaction: hydrolysis of esters

- 218. Determination of activation energy
- 219. Initial rates and activation energy of the iodine clock
- 220. Adsorption at solid-liquid interfaces, construction and analysis of an isotherm
- 221. Solubilization capacity of surfactants
- 222. Determination of surface tension, the Gibbs isotherm.
- 223. Rheological investigation of different samples.
- 224. Determination of the CMC value of association colloids
- 225. Determination of the isoelectric point of isolabile proteins
- 226. Isotope dilution analysis: quantitative determination of a stable isotope by a radioactive isotope
- 227. Backscattering of β radiation by backscattering media of different thickness, composition and concentration
- 228. Gamma spectrometry: measurement and interpretation of γ spectra

Literature

- Laboratory notes and additional teaching materials available via the e-learning system.
- P. W. Atkins, J. de Paula: Atkins' Physical Chemistry 8th Edition, W.H. Freeman and Company, New York, ISBN: 0-7167-8759-8, 2006
- Katalin Ösz, A. Bényei: Physical Chemistry Laboratory Measurements (for students of Pharmacy, Chemistry and Chemical Engineering). Debreceni Egyetemi Kiadó, ISBN: 978-963-318-143-0, 2011
- G.T. Barnes, I.R. Gentle: Interfacial Science. Oxford UP. ISBN 0-a19-a927882-a2, 2005
- R. M. Pashley: Applied Colloid & Surface Chemistry. Wiley&Sons, ISBN 0-a470-a86883-aX, 2004
- T. Cosgrove: Colloid science. Blackwell Publishing ISBN:978-a14051-a2673-a1, 2005.

Course objective/intended learning outcomes

a) Knowledge

- He/She knows and can use the properties of most important chemicals, lab equipments and basic laboratory processes.
- He/She can follow and recognize the requirements of safety instructions.
- He/She has the knowledge to test or measure chemical reactions, systems with scientific methods (including computational) under supervision.

b) Abilities

- He/She is able to use the previously learned paradigms, theories and laws on the field of chemical, natural science to plan, execute and evaluate laboratory experiments.
- He/She is able to evaluate and discuss the laboratory measurements, and create a report about it.
- He/She is able to use the obtained chemical knowledge on the field of chemical, industrial, environmental sciences to solve the actual basic problems and prove the obtained solution.

c) Attitude

- It is important to him/her to work sufficiently and consciously in order to decrease the environmental load.

d) Autonomy and responsibility

- He/She can make reasonable evaluations about his/her own work comparing to others to the same field.

Schedule: One of the measurements listed above (**Topics of course**) per week except the 1st

practice (introduction, general information and safety training).

Requirements:

Participation on the laboratory practice is compulsory. The measurements and knowledge of the associated theory are marked and an overall mark will be given. Safety training (1st week) is mandatory before the first lab practice (2nd week). Everybody should work individually according to the pre-set schedule (which will be provided on the 1st week). Lab practices are 4 hours long every week (from the 2nd until the 14th weeks). Being late or failed mark on the written test from the appropriate measurement is equivalent with an absence. In accordance with the regulations of University of Debrecen, attendance is compulsory with the exception of health or family problems (the reason of absence should be certified). In this case, the students should agree with the teacher on replacement dates for the missed experiments.

Requirements for the grade:

The measurements (regularly) and written tests (occasionally) according to the knowledge of the associated theory are marked and the overall mark will be given based on these.

- At least 11 notebooks of the measurements (from the 13) have to be marked as “pass (2)” or better for the successful completion.

- The minimum requirement for the written tests is 60%. Based on the score of the tests separately, the grade for the tests is given according to the following table:

Score	Grade
0-59	fail (1)
60-69	pass (2)
70-79	satisfactory (3)
80-89	good (4)
90-100	excellent (5)

If the average of written tests is below 60% the best grade for the course can be only “pass (2)” in any other cases the final mark is given with weighted average by means of the mark of the written tests and notebooks in 1 to 2 ratio.

Person responsible for course: Dr. Ferenc K. Kálmán, assistant professor, PhD

Lecturer: Dr. Ferenc K. Kálmán, assistant professor, PhD

Title of course: Organic Chemistry I.

Code: TTKBE0301_EN

ECTS Credit points: 4

Type of teaching

- lecture: 2 hours/week
- practice: 1 hours/week
- laboratory: -

Evaluation: exam

Workload (estimated)

- lecture: 28 hours
- practice: 14 hours
- laboratory: -

<p>- home assignment: 14 hours - preparation for the exam: 40 hours Total: 96 hours</p>
<p>Year, semester: 1st year, 2nd semester</p>
<p>Its prerequisite(s): General Chemistry I. TTKBE0101_EN</p>
<p>Further courses built on it: TTKBE0202_EN, TTKBL0201_EN, TTKBE0402_EN, TTKBG0402_EN, TTKBL0401_EN, TTKBE0302_EN, TTKBE0501_EN, TTKBE0502_EN, TTKBE0601_EN, TTKBG0601_EN, TTKBE0204_EN, TTKBE0417_EN, TTKBG0614_EN, TTKBG0312_EN, MFVGE31V03_EN, TTKBE1111_EN</p>

<p>Topics of course</p> <ul style="list-style-type: none"> • Review the basic of organic chemistry basics • Types and theories of chemical bonds • Review the acid-base theories • Basic concepts of isomerism and stereochemistry. • Classification of organic chemical reactions. • Functional groups and the basics of organic nomenclature. • The structure, nomenclature, synthesis and reactions of alkanes, alkenes, alkenes, alkynes, mono- and polycyclic, homo- and heteroaromatic hydrocarbons.
<p>Literature</p> <p><i>Compulsory:</i></p> <ol style="list-style-type: none"> 1. Course material, concept and task collection for lectures, seminars in the e-learning system. <p><i>Recommended:</i></p> <ol style="list-style-type: none"> 2. John McMurry: Organic Chemistry (8th Edition), ISBN-10: 0840054440 ISBN-13: 9780840054449, 2012, Brooks/Cole 3. Jonathan Clayden, Nick Greeves, and Stuart Warren: Organic Chemistry (Second Edition), ISBN: 978-0-19-927029-3; 2012, Oxford University Press 4. Francis A. Carey: Organic Chemistry (4th Edition), ISBN 0-07-290501-8; 2000, The McGraw-Hill Companies, Inc. 5. Leroy G. Wade: Organic Chemistry (8th Edition), ISBN-10: 0321768140; 2012, Pearson 6. T. W. Graham Solomons, Craig Fryhle, Organic Chemistry (10th Edition), ISBN-10: 0470556595; 2009, Wiley & Sons 7. Herbert Meislich, Estelle Meislich, Jacob Sharefkin - 3000 Solved Problem in Organic Chemistry (1994)
<p>Course objective/intended learning outcomes</p> <p>a) Knowledge He/She knows the basic concepts and theories which are necessary to understand and interpret structure and reactivity of organic compounds (chemical bond, hybridization, resonance theory, isomerism) He/she knows the structure, physical and chemical properties and synthetic methods of saturated, unsaturated and aromatic hydrocarbons and He/She can apply these knowledges to solve chemical problems.</p> <p>b) Abilities - He/she knows and is able to apply his/her knowledge to solve simple tasks on the field of</p>

hydrocarbons.

- He/she is able to participate in professional communication on the field of structure, and chemical transformation of hydrocarbons.

- He/she is able to expand and/or develop his/her knowledge from the natural products.

c) Attitude

- He/she is open to getting new, scientifically proven knowledge on the subject, but to reject unsubstantiated or possibly misleading claims

d) Autonomy and responsibility

- He/she is able to independently perform the tasks of the course with professional guidance, and he/she can interpret and evaluate of the results obtained.

Schedule:

1st week

The definition and brief history of organic chemistry. Overview of the basic general chemical concepts needed for this subject. A brief summary of the theories of the chemical bond: the shared electron pair model, the valence bond model. Covalent and ionic bonds. The basics of LCAO-MO theories, types of atomic and molecular orbitals. Bi- and polycentric molecular orbitals, delocalization.

2nd week

VB theory, resonance structures and rules of their writing. Hybridization. Electron shift phenomena, inductive and mesomeric effects, conjugation and hyperconjugation. Secondary bonds, intermolecular interactions, hydrogen bond, dipole-dipole, dipole-induced dipole interactions.

3rd week

Description of functional groups in organic compounds. An overview of the most important organic compound groups based on their functional groups. The effect of functional groups on the electron structure of compounds.

4th week

The basic nomenclature systems in organic chemistry: common or trivial names and systematic nomenclature. Basic rules to generate systematic names of organic compounds; substitutive and functional class nomenclature. The rules to generate the names the groups derived from hydrocarbons. The rules to generate the name of unbranched and branched (saturated and unsaturated) hydrocarbons. Elemental reactions. Definitions of transition state, intermediates, Gibbs energy, kinetic and thermodynamic parameters of chemical reactions.

5th week

Multi-step reactions (consecutive reactions), intermediates. Hammond postulate. Parallel (competitive) reactions. Thermodynamic and kinetic control. Reactivity and selectivity. Reagents and reactive intermediates. Classification of organic chemical reactions based on attack agent and type of the reaction. Brønsted and Lewis acid-base theory, "hard" and "soft" acids and bases.

6th week

Basics of stereochemistry: characterization of constitutional, conformational and configuration isomers. Chirality, types of chiral molecules. The concept of enantiomers and diastereomers, general comparison of their chemical and physical properties. Absolute and relative configuration. Optical activity. The representation of organic molecules. The absolute configuration of chiral compounds, Fischer and Cahn-Ingold-Prelog convention. The role of chirality in drug chemistry.

7th week

Characterization of the structures of alkanes and cycloalkanes. Review their conformational and physical properties. Chemical properties of alkanes, radical substitution, chain reaction. Statistical

and regioselective halogenation and interpretation based on radical stability in alkane halogenation.

8th week

Sulphonation, sulphochlorination, nitration and oxidation of alkanes. The basic petrochemical processes (pyrolysis, cracking, isomerization) and their industrial significance. The most important natural sources and the synthetic methods of alkanes.

9th week

The characterization of the structure of alkenes, cycloalkenes, di- and polyenes. The hindered rotation: characterization of E / Z isomers. Synthesis of alkenes, cycloalkenes. Physical and chemical properties of alkenes and cycloalkenes. Electrophilic and radical addition reactions and practical significance. Interpretation of the regioselectivity of the addition reactions; the Markovnikov rule.

10th week

Types of polymerization. Substitution in allylic position, interpretation of the stability of allylic intermediates. Oxidation of alkenes. Addition of conjugated dienes, partial and complete addition. 1,2 and 1,4 addition and its interpretation based on kinetic and thermodynamic control. Diels-Alder cycloaddition.

11th week

Characterization of the structure of alkynes and their physical properties. The stability and synthesis of alkynes. Chemical transformations of alkynes: C-H acidity, addition reactions and their significance. The role of acetylene in the chemical industry, coal-based chemical industry.

12th week

The concept and the interpretation of aromaticity. Neutral and charged homo and heteroaromatic systems. The type and mechanism of the most important aromatic electrophilic substitution reactions (halogenation, nitration, sulphonation, Friedel-Crafts acylation and alkylation).

13th week

The S_{EAr} reactions of substituted benzene derivatives –the reactivity and regioselectivity. Classification of substituents and interpretation of their effect on reactivity and regioselectivity.

14th week

Electrophilic substitution reactions of five- and six-membered heteroaromatic base compounds. Addition reactions of monocyclic aromatic hydrocarbons. Reactions of aromatic hydrocarbons containing alkyl substituents, the stability of benzyl-type reactive intermediates. Most important representatives of polycyclic aromatic hydrocarbons.

Requirements:

- for a signature

Attendance at **lectures** is recommended, but not compulsory.

Attendance at seminars is compulsory. A student may not miss the seminar more than three times during the semester. In case of further absences, a medical certificate needs to be presented. In case a student does not do this, the subject will not be signed, and the student must repeat the course.

- for a grade

The course ends in an **examination**.

The exam grade is the result of the written exam.

The minimum requirement for the examination respectively is 50%. The grade for the written exam is given according to the following table:

Score	Grade
0-49	fail (1)
50-62	pass (2)
63-75	satisfactory (3)

76-87	good (4)
88-100	excellent (5)
If the score of any test below 50, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.	
Person responsible for course: László Dr Juhász, associate professor, PhD, dr. habil.	
Lecturer: Éva Juhászné Dr. Tóth, senior lecturer, PhD; Krisztina Dr. Kónya, senior lecturer, PhD	

Title of course: Organic chemistry II. Code: TTKBE0302_EN	ECTS Credit points: 4
Type of teaching, contact hours - lecture: 2 hours/week - practice: - - laboratory: -	
Evaluation: term mark	
Workload (estimated), divided into contact hours: - lecture: 28 hours - practice: 14 - laboratory: - - home assignment: 14 hours - preparation for the exam: 40 Total: 96 hours	
Year, semester: 2 nd year, 1 st semester	
Its prerequisite(s): TTKBE0201_EN, TTKBE0301_EN, TTKBE0401_EN	
Further courses built on it: TTKBE0303_EN, TTKBE0611_EN, TTKBE1212_EN, TTKBL1212_EN, TTKBE0503_EN, TTKGB0313_EN	

Topics of course
Systematical overview the structure, physical, chemical properties of hydrocarbons possessing heteroatoms as halogenated hydrocarbons, organometallic derivatives, alcohols, phenols, ethers and their thio analogues; amines, nitro derivatives, diazonium salts, aldehyde, ketones, carboxylic acids and their derivatives, derivatives of carbonic acid
Literature
<i>Compulsory:</i> 1. Course material, concept and task collection for lectures, seminars in the e-learning system.
<i>Recommended:</i> 2. J. G. Smith: Organic Chemistry, 5 th Edition, 2016, McGraw Hill; ISBN-13: 9780077354725 3. J. McMurry: Organic Chemistry, 8 th Edition, 2012, Brooks/Cole; ISBN-13: 9780840054449

4. J. Clayden, N. Greeves, and S. Warren: Organic Chemistry, 2nd Edition, 2012, Oxford University Press; ISBN-13: 9780199270293
5. F. A. Carey: Organic Chemistry, 4th Edition, 2000, The McGraw-Hill Companies; ISBN-13: 9780072905014
6. L. G. Wade: Organic Chemistry, 8th Edition, 2012, Pearson; ISBN-13: 9780321768148
7. T. W. G. Solomons, C. Fryhle, Organic Chemistry, 10th Edition, 2009, Wiley & Sons; ISBN-10: 0470556595
8. H. Meislich, E. K. Meislich, J. Sharefkin: 3000 Solved Problems in Organic Chemistry, 1st Edition, 1994, McGraw-Hill Companies; ISBN-13: 978-0070564244

Course objective/intended learning outcomes

a) Knowledge

- He/she knows the structure, physical and chemical properties and synthetic methods of the most important organic compounds possessing heteroatoms (halogenated hydrocarbons, organometallic derivatives, alcohols, phenols, ethers and their thio analogues; amines, nitro derivatives, diazonium salts, aldehyde, ketones, carboxylic acids and their derivatives, derivatives of carbonic acid). He/she is able to apply his/her knowledge to solve simple tasks on the field of this compounds. He/she knows the application and practical significance of these derivatives., and their applicabilities

b) Abilities

- He/she is able to participate in professional communication on the field of structure, and chemical transformation of this field.
 - He/she is able to expand and/or develop his/her knowledge from the natural products.

c) Attitude

- He/she is open to getting new, scientifically proven knowledge on the subject, but to reject unsubstantiated or possibly misleading claims

d) Autonomy and responsibility

- He/she is able to independently perform the tasks of the course with professional guidance, and he/she can interpret and evaluate of the results obtained.

Schedule:

1st week

Classification of halogenated hydrocarbons, characterization of their structure and physical properties. The effect of the structure of the hydrocarbon skeleton, and the quality of the halogen on the strength of the C-Hlg bond and reactivity. Synthesis of halogenated hydrocarbons.

2nd week

Reactions of halogenated hydrocarbons. Interpretation of decreased, normal and high reactivity of halogenated hydrocarbons. Nucleophilic substitution and elimination of halogenated hydrocarbons. Interpretation of the mechanism of these reaction (S_N1, S_N2; α- and β-elimination; E1, E2 and E1cB). Reaction of halogenated compounds with metals.

3rd week

The basics of chemistry of organometallic compounds. Their bonding system, the term "umpolung". Synthesis and reactivity of organometallic compounds. Organometallic compounds as nucleophiles and carbanion equivalents. C-C bond formation with organometallic reagents: Grignard compounds and their application. Synthesis and interconversion of organometallic

compounds, transmetallation.

4th week

Classification and characterization of hydroxyl derivatives of hydrocarbons (alcohols, phenols) and their thio analogues. Interpretation of their physical properties derived from their bonding system. The acid-base properties of alcohols, phenols and thio analogues. Preparation of alcohols, ethers, phenols and thio analogues.

5th week

Alcohols and phenols as nucleophiles: alkylation, acylation, formation of sulphonate and inorganic esters; acid catalyzed transformations of alcohols (conversion of alcohols to halogenated derivatives, elimination reactions). Oxidation of alcohols and phenols. The characterization of ethers; synthesis and cleavage of ethers. Characterization of the special ether derivatives: epoxides, semi-acetals, acetals and enol ethers. Cumene-based phenol synthesis.

6th week

Overview of the organic compounds possessing C-N single bond. Classification of amines and characterization of their bonding systems. Interpretation of their physical properties derived from their bonding system. Synthesis of aliphatic and aromatic amines; industrial methods.

7th week

Review and interpretation of basicity of amines. Chemical transformation of amines: alkylation, acylation of amino group. Synthesis of sulfonamide and reaction with nitric acid. Oxidation of the amines. S_EAr reactions of anilines.

8th week

Characterization of nitro compounds: the bonding system, interpretation of electron-withdrawing effect and C-H acidity. Synthesis of nitro compounds. Preparation of diazonium salts, reactions of diazonium salts and their practical significance. Azo compounds and their industrial significance.

9th week

Classification and characterization of oxo compounds: the bonding system and stability of carbonyl group. Physical properties of oxo compounds. Acid-base properties of aldehydes and ketones: acidity of the α -hydrogen, keto-enol tautomerism. Synthesis of aldehydes and ketones.

10th week

Reactions of aldehydes and ketones. Nucleophilic addition with O-, S-, N- and C-nucleophiles, the reversibility of the additions. Condensation reactions. Oxidation and reduction. Reactions on α -carbon; aldol dimerization, α -halogenation. Nucleophilic addition reactions of α,β -unsaturated oxo compounds.

11th week

Classification of carboxylic acids and their derivatives, description and comparison of their bonding systems. Stability and reactivity of the carboxylic acid derivatives. Physical properties and synthesis of carboxylic acids.

12th week

Review and interpretation of the acid-base properties of carboxylic acids and their derivatives (O-H, N-H and C-H acidity). Interconversion of the carboxylic acid derivatives, acyl nucleophilic substitution. Reductive transformations of carboxylic acid derivatives, transformation of their carbon skeleton.

13th week

β -Dicarbonyl and β -oxo-carboxylic acid derivatives, C-H acidity and basic of enolate chemistry: formation of carbon-carbon bond, malonic ester, acetoacetic ester and cyanoacetic ester syntheses.

14th week

Substituted (halogenated, hydroxy and oxo) carboxylic acid derivatives and their interconversion. Synthesis and interconversion of carbonic acid derivatives and their major representatives. Practical significance of carbonic acid derivatives.

Requirements:

- for a signature

Attendance at **lectures** is recommended, but not compulsory.

Attendance at seminars is compulsory. A student may not miss the seminar more than three times during the semester. In case of further absences, a medical certificate needs to be presented. In case a student does not do this, the subject will not be signed, and the student must repeat the course.

- for a grade

The course ends in an **examination**.

The exam grade is the result of the written exam.

The minimum requirement for the examination respectively is 50%. The grade for the written exam is given according to the following table:

Score	Grade
0-49	fail (1)
50-62	pass (2)
63-75	satisfactory (3)
76-87	good (4)
88-100	excellent (5)

If the score of any test below 50, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

Person responsible for course: László Dr.Juhász, associate professor, PhD, Habil.

Lecturer: Éva Juhászné Dr. Tóth, senior lecturer, PhD; Krisztina Dr.Kónya, senior lecturer, PhD

Title of course: Organic chemistry III.

Code: TTKBE0303_EN

ECTS Credit points: 3

Type of teaching, contact hours

- lecture: 2 hours/week
- practice: -
- laboratory: -

Evaluation: term mark

Workload (estimated), divided into contact hours:

<ul style="list-style-type: none"> - lecture: 28 hours - practice: - - laboratory: - - home assignment: 14 hours - preparation for the exam: 40 <p>Total: 96 hours</p>
Year, semester: 2 nd year, 2 st semester
Its prerequisite(s): TTKBE0302_EN
Further courses built on it: TTBBE2035_EN

<p>Topics of course</p> <p>Characterization of the building blocks of biomacromolecules (peptides and proteins, carbohydrates, nucleic acids, lipids) that form biological structures. Description and characterization of the most important biochemical reactions. Characterization of the structure of the biomacromolecules. Overview of the chemical and instrumental methods which can be used for the structure elucidation of these type of compounds. Review the basic of their information storage and storage capacity, the relationship between structure and function. Chemical properties of their monomers and synthesis of biopolymers. The structure and biological effect/function of some other significant natural compounds (isoprenoids, flavonoids, alkaloids, antibiotics, vitamins, porphyrinase compounds).</p>
<p>Literature</p> <p><i>mpulsory:</i></p> <ol style="list-style-type: none"> 1. Course material, concept and task collection for lectures, seminars in the e-learning system. <p><i>commended:</i></p> <ol style="list-style-type: none"> 2. J. G. Smith: Organic Chemistry, 5th Edition, 2016, McGraw Hill; ISBN-13: 9780077354725 3. C. Stan Tsai: Biomacromolecules, John Wiley & Sons, New Jersey (2007) 4. A. Miller-J. Tanner: Essentials of Chemical Biology, John Wiley & Sons, Chichester (2008) 5. P. M. Dewick: Medicinal Natural Products: A Biosynthetic Approach, 3rd Edition. John Wiley & Sons, Chichester (2009)

Course objective/intended learning outcomes

a) Knowledge

- He/she knows the structure, physical and chemical properties and synthetic methods of monomers of biomacromolecules.
- He/she knows the structure and function of biomacromolecules.
- He/she knows the basic structure and function of other secondary metabolites.

b) Abilities

- He/she is able to participate in professional communication on the field of structure, and chemical transformation of this field.
- He/she is able to expand and/or develop his/her knowledge from the natural products.

c) Attitude

- He/she is open to getting new, scientifically proven knowledge on the subject, but to reject unsubstantiated or possibly misleading claims

d) Autonomy and responsibility

- He/she is able to independently perform the tasks of the course with professional guidance, and he/she can interpret and evaluate of the results obtained.

Schedule:

1st week

Primary and secondary metabolism. Classification of natural compounds. Types of biological structural materials, general characterization. Common features of the synthesis of biopolymers: group protection, activation, coupling reactions, requirements for protective groups, orthogonality

2nd week

Structure, synthesis and chemical properties of amino acids. Characterization of α -amino acids which are forming protein/peptides. Structure and determinations of peptides. Determination of amino acid sequence by chemical and enzymatic methods, possibility of automation.

3rd week

Synthesis of peptides. The basic protecting groups and activation methods for peptide synthesis. Solid phase synthesis, automation. The occurrence, classification and functions of proteins. Levels of protein structure: primary, secondary, tertiary and quaternary structures, structure formation. Structure and function relationship.

4th week

Classification, structure and nomenclature of carbohydrates. Basic configuration and conformational conditions of monosaccharides. Most important chemical properties of monosaccharides: mutarotation, transformation of oxo group and hydroxyl groups, synthesis of glycosides.

5th week

Most important representatives of di- and oligosaccharides (sucrose, maltose, cellobiose, lactose, cyclodextrins), factors determining their structure. Synthesis of di- and oligosaccharides, basic protecting groups and activation methods.

6th week

Derivatives of Peptides / proteins and low molecular weight carbohydrates: peptidoglycans, glycoproteins, their biological significance. The carbohydrate code.

7th week

Polysaccharides (cellulose, chitin, starch, glycogen, pectin, mucopolysaccharides).

Polysaccharides as structural materials and reserve nutrients. Derivatives of polysaccharides and proteins (proteoglycans). The industrial significance of polysaccharides.

8th week

Classification and characterization of nucleic acids, their building blocks. Synthesis of nucleosides and nucleotides. Primary, secondary and tertiary structure and biological function of DNA and RNA. The genetic code. Information content of the nucleotide, amino acid and carbohydrate code and their correlation. Nucleotide coenzymes.

9th week

Classification and characterization of lipids, their structure, their biological role. Basics of the biosynthesis of fats, phospho- and glycolipids.

10th week

Isoprenes, terpenoids and carotenoids. The basics of their biosynthesis, and most important representatives of terpenoids. The chemical background of vision. Structure, classification of steroids, basics of their biosynthesis, their major representatives and their biological function.

11th week

Classification and structure of phenylpropanoids. The chemical synthesis of their basic skeletons. Structure and biological significance of flavonoids.

12th week

Classification of alkaloids and structure and function of their most important representatives. Alkaloids as drugs and medicines.

13th week

Definition of symbiosis, antibiosis. Definition and classification of antibiotics: β -lactam, amino acid or peptide, glycoside type antibiotics, polycyclic antibiotics. Preparation of antibiotics: fermentation, semi-synthetic and synthetic derivatives. The most important mode of action of antibiotics.

14th week

The structure, biosynthesis and biological role of porphyrins. Structure, biological role and metabolism of chlorophyll and hemoglobin. Classification of vitamins, their structure, their natural sources and their biological functions.

Requirements:

- for a signature

Attendance at **lectures** is recommended, but not compulsory. A student may not miss the lecture more than three times during the semester. In case of further absences, a medical certificate needs to be presented. In case a student does not do this, the subject will not be signed and the student must repeat the course.

- for a grade

The course ends in an oral exam in the exam period.

Person responsible for course: László Dr. Juhász, associate professor, PhD, dr. habil

Lecturer: László Dr. Juhász, associate professor, PhD, dr. habil.

Title of course: Organic Chemistry IV (seminar)

Code: TTKBG0301_EN

ECTS Credit points: 1

Type of teaching, contact hours

<ul style="list-style-type: none"> - lecture: - - practice: 1 hours/week (in blocks for 7 weeks and 2 hours/week) - laboratory: -
Evaluation: term mark
Workload (estimated), divided into contact hours: <ul style="list-style-type: none"> - lecture: - - practice: 14 - laboratory: - - preparation for the mid-term tests: 6 hours Total: 20 hours
Year, semester: 2 nd year, 1 st semester
Its prerequisite(s): TTKBE0302_EN
Further courses built on it: -

Topics of course <ul style="list-style-type: none"> - Basic stereochemical concepts (constitution, conformation, configuration, relative configuration, absolute configuration, enantiomers, diastereomers, epimers) - Constitutional isomerism - Conformational isomerism - <i>Cis / trans, E / Z</i> isomers - Projective formulas - D / L markers - CIP convention - Axial chirality, allene isomerism, atrop isomerism
Literature <p><i>Compulsory:</i></p> <ol style="list-style-type: none"> 1. Slides of the lectures by László Juhász: Organic Chemistry I. and Organic Chemistry II. <p><i>Recommended:</i></p> <ol style="list-style-type: none"> 1. IUPAC stereochemistry recommendations: http://old.iupac.org/reports/provisional/abstract04/BB-prs310305/Chapter9.pdf 2. J. P. Riehl: Mirror-Image Asymmetry - An Introduction to the Origin and Consequences of Chirality, John Wiley & Sons, 2010, Hoboken, New Jersey. 3. E. L. Eliel, S. H. Wilen: Stereochemistry of Organic Compounds, Wiley, New York, 1994.
Course objective/intended learning outcomes <p>a) Knowledge They know the main models and theories of chemical bonds and molecular structure based on scientific findings. They have knowledge to solve problems on the field of natural processes, using natural sources, and understanding the chemical background of living and non-living systems. They possess knowledge of the major types of isomers, projection formulas used in organic chemistry, stereochemical markers, central and axial chirality.</p> <p>b) Abilities</p>

They are able to use the previously obtained knowledge on the field of natural and anthropogenic studies to solve practical problems. They are able to use the obtained chemical knowledge on the field of chemical, industrial, environmental sciences to solve the actual basic problems and prove the obtained solution. They are able to identify the types of major isomers based on their acquired knowledge and practice, identify and determine chirality elements, apply projection formulas in organic chemistry. Based on the acquired knowledge, they are able to comprehend, interpret and solve practical problems of organic chemical reactions affecting stereochemistry.

c) Attitude

They are ready to discuss problems on the field of chemistry and other science with professionals. They represent authentically the stereochemical representations and the official IUPAC guidelines of stereochemistry. Through the acquisition of multiple projection formulas and markers, they are open to accommodate further stereochemical markers used in narrower field of organic chemistry. They are open to use computers in solving stereochemical problems. They are open for professional discussions to solve problems.

d) Autonomy and responsibility

They stand for their opinion or ideology in professional discussions. They can make reasonable evaluations about their own work comparing to others to the same field. They compare their results with those of similar professional positions (students). They apply the acquired stereochemical knowledge individually to solve specific problems and take responsibility for the results.

Schedule:

1st week

Review of stereochemical concepts. Constitution, conformation, configuration (relative and absolute), enantiomers, diastereomers, epimers.

Examples of constitutional isomers. Conformational isomers. Conformation of ethane, butane, open-chain compounds. Cyclohexane ring conformation. Conformation of cyclic compounds with examples.

2nd week

Cis / trans, E / Z isomery with examples.

Projective formulas, Fischer projection. Examples.

3rd week

Mid-term test 1.

D / L markers for amino acids, open chain and ring forms of sugars. Examples.

4th week

CIP convention is simpler (with one chirality center) and more complicated cases (in larger systems, with more chirality centers). Examples.

5th week

Practice.

Mid-term test 2.

6th week

CIP convention in special cases. Phantom atoms, *E / Z* isomers and CIP convention, bridged ring systems. Examples.

Axial chirality. Allen-isomerism. Atrop isomerism. Examples.

7th week

Practice.

Mid-term test 3.

Requirements:

- *for a signature*

Attendance of all practical classes.

- *for a grade*

Class performance (33%)

Mid-term tests (67%)

Based on the sum of the 3 tests and the class performance the practical grade is calculated.

The final grade is given according to the following table:

Score (%)	Grade
0-49	fail (1)
50-59	pass (2)
60-74	satisfactory (3)
75-89	good (4)
90-100	excellent (5)

If the score of the final grade is below 50%, students can not take a retake the tests in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

-*an offered grade:*

-

Person responsible for course: Dr. Attila Mándi, researcher / assistant professor, PhD

Lecturer: Dr. Attila Mándi, researcher / assistant professor, PhD

Title of course: Organic chemistry IV. Code: TTKBL0301_EN	ECTS Credit points: 4
Type of teaching, contact hours - lecture: - - practice: - - laboratory: 4 hours/week	
Evaluation: term mark	
Workload (estimated), divided into contact hours: - lecture: - - practice: - - laboratory: 56 hours - home assignment: 21 hours - preparation for the exam: - Total: 77 hours	
Year, semester: 2 nd year, 2 nd semester	
Its prerequisite(s): TTKBL0101_EN, TTKBE0302_EN	
Further courses built on it: TTKBL0302_EN	

Topics of course

The aim of the course is to enable students to become familiar with the theoretical background of

basic organic chemistry laboratory techniques and to learn how to apply them in practice, to deepen the theoretical knowledge gained in organic chemistry lectures and to understand the reactivity of functional groups by synthesizing simple preparations on a semi-micro scale and by carrying out test tube reactions. The other goal is to provide students with the right material knowledge and to understand and apply cleaning and identification techniques as typical organic chemistry activities.

Literature

Compulsory:

1. L. Juhász: Organic Laboratory Techniques and Manuals for Pharmacist Students, Debrecen, 2009
2. J. R. Mohrig, D. G. Alberg, G. E. Hofmeister, P. F. Schatz, C. Noring Hammond: **Laboratory Techniques in Organic Chemistry** (Supporting Inquiry-Driven Experiments), **4th edition**, W. H. Freeman and Company. ISBN-13: 978-1-4641-3422-7.

Recommended:

3. H. Meislich, E. K. Meislich, J. Sharefkin: 3000 Solved Problems in Organic Chemistry, 1st Edition, 1994, McGraw-Hill Companies; ISBN-13: 978-0070564244
4. R. O. C. Norman, J. M. Coxon: Principles of Organic Synthesis, 3rd Edition, 1993, Blackie Academic & Professional, Glasgow, UK; ISBN-13: 9780751401264
5. J. McMurry: Organic Chemistry, 8th Edition, 2012, Brooks/Cole; ISBN-13: 9780840054449
6. J. Clayden, N. Greeves, S. Warren: Organic Chemistry, 2nd Edition, 2012, Oxford University Press; ISBN-13: 9780199270293
7. F. A. Carey: Organic Chemistry, 4th Edition, 2000, The McGraw-Hill Companies; ISBN-10: 0072905018
8. J. G. Smith: Organic Chemistry, 5th Edition, 2016, McGraw Hill; ISBN-13: 9780077354725

Course objective/intended learning outcomes

a) Knowledge

- He/She knows and can use the properties of most important chemicals, lab equipment and basic laboratory processes.
- He/She can follow and recognize the requirements of safety instructions.
- He/She knows and applies basic organic chemistry laboratory techniques in organic chemistry and reactions typical of different types of compounds.

b) Abilities

- Using his/her chemical knowledge and experiences He/She is able to create and execute experiments to demonstrate and prove the basic chemical phenomena.
- He/She is able to evaluate and discuss the laboratory measurements and create a report about it.
- He/She is able to select and apply the necessary basic laboratory techniques during the preparatory work.
- He/She is able to participate in professional communication on operations and reactions in preparative organic chemistry.
- He/She is able to determine the different types of compounds by using test tube reactions.
- He/She is able to expand/improve his/her knowledge.

c) Attitude

- It is important to him/her to work sufficiently and consciously in order to decrease the environmental load.
- He/She is able to represent his/her own personal scientific ideology toward professional and unprofessional groups.
- He/She is open to get new, scientifically proven knowledge on the subject, but to reject unsubstantiated or possibly misleading claims.

d) Autonomy and responsibility

- During laboratory work he/she is capable to get related to the basics of chemical problems, able to conclude the questions to the supervisor that might come up to the solutions or decisions.
- He/She is able to operate the most important technological instruments of chemical industries.
- He/She is able to perform the tasks independently with professional management.
- He/She is able to interpret the results obtained and evaluate them realistically and analyse and solve the problems that arise.

Schedule:

1st week

Introduction: Timetable and requirements. Receiving of laboratory equipment and list of tasks.

Safety education.

Presentation of the device for recrystallization.

Presentation of gravity and vacuum filtration equipment.

Description of the operation of the rotary vacuum evaporator.

Recrystallization of acetanilide from water.

2nd week

Short written test.

Presentation of thin layer chromatography (TLC).

Presentation of determination of melting point.

Check of the purity of the compound recrystallized in previous practice by melting point and TLC.

Calculation of the yield of recrystallization.

Recrystallization of benzanilide from methanol.

Check of the purity of the recrystallized benzanilide by TLC.

3rd week

Short written test.

Description of liquid-liquid extraction.

Control the purity of the compound recrystallized in previous practice by melting point. Calculation of the yield of recrystallization.

Use of liquid-liquid extraction to separate m-dinitrobenzene and m-nitroaniline. Checking the success of the separation using TLC.

4th week

Short written test.

Presentation of equipment used for distillation at atmospheric and reduced pressure.

Distillation of acetone from KMnO_4 at atmospheric pressure.

Distillation of water in vacuum.

5th week

Short written test.

Presentation of steam distillation.

Isolation of S-(+)-Carvone from caraway and preparation of its 2,4-dinitrophenylhydrazone derivative.

6th week

Short written test.

Preparation of 4-chlorobenzoic acid and 4-chlorobenzyl alcohol. Check of the purity of the products using TLC and melting point measurement.

7th week

Short written test.

Presentation of a device used in reaction with three-necked round bottom flasks.

Preparation of benzamide and recrystallization of the product from water.

Preparation of iodoform.

8th week

Short written test.

Description of column chromatography.

Determination of melting point of iodoform, and calculation of the yield.

Check of the purity of benzamide by TLC and melting point measurement.

Separation of the mixture of acetanilide and m-dinitrobenzene by column chromatography.

9th week

Short written test.

Preparation of acetylsalicylic acid and purification of the product by recrystallization.

Check of the purity of the product by TLC and melting point measurement.

Calculation of the yield.

10th week

Short written test.

Identification of hydrocarbons and organic halides using test tube reactions.

Reaction of hydrocarbons with bromine.

Reaction of hydrocarbons with bromine in the presence of UV light.

Friedel-Crafts test of aromatic hydrocarbons.

Baeyer test of unsaturated hydrocarbons.

Beilstein and alcoholic silver nitrate test of organic halides.

Identification of unknown compounds.

Preparation of 2,6-dibenzylidene-cyclohexanone (test tube variant). Check of the purity of the products by TLC and melting point measurement. Calculation of the yield.

11th week

Short written test.

Identification of hydroxyl derivatives of hydrocarbons using test tube reactions.

Solubility of alcohols and phenols.

Determination of order of substitution of the carbon carrying the OH group by Lucas probe.

Oxidation of alcohols with Jones reagent.

Reaction of diols or polyols with copper(II) ions.

Reaction of phenols and enols with iron(III) ions.

Iodoform test of 2-alkanols.

Identification of unknown compounds.

Preparation of benzoic acid (test tube reaction).

12th week

Short written test.

Identification of amino derivatives of hydrocarbons using test tube reactions.

The Hinsberg test.

Reactions of amines with nitrous acid.

The Rimini reaction of aliphatic primary amines

Complex formation of amine with Cu(II) ions.

Identification of unknown compounds.

Preparation of benzotriazole (test tube variant)

13th week

Short written test.

Identification of oxo compounds using test tube reactions.

Detection of aldehydes with 2,4-dinitrophenylhydrazine test.

Oxidation of aldehydes by neutral potassium permanganate solution.

Oxidation of oxo compounds by Jones reagent.

Reaction of oxo compounds with Tollens reagent.

Iodoform test of oxo compounds.

Identification of unknown compounds.

Preparation of cyclohexanone (test tube variant)

14th week

Short written test.

Performing missed identification tasks (melting point measurement, TLC), yield calculation.

Cleaning and handovering of equipments.

Present the synthesized products to the instructor.

Evaluation.

Requirements:

Attendance at laboratory practice is mandatory.

Before starting the laboratory work, students must write a short written test on their theoretical organic chemistry and practical knowledge as well as on the safety rules about the previous laboratory practice (15-20 minutes).

On the one hand, the term mark consists of the marks obtained for the identification of the unknowns and on the other hand the marks written before the practice, which are closely related to the laboratory exercises carried out the week before (15-20 minutes). Of course, a prerequisite for successful laboratory practice is the synthesis of all preparations.

The final grade will be determined based on the average of the grades of tasks. A weighted average of the grades of subtasks will be calculated in the following manner:

- Short written test (65%)
- *Activity in laboratory practice* (15%)
- Identification of unknown compounds (20%)

Final grade: excellent (5): 90%; good (4): 80%; satisfactory (3): 65%; pass (2): 50%; fail (1): below 50%.

Person responsible for course: Marietta Vágvölgyiné Dr. Tóth, associate professor, PhD, Habil.

Lecturer: Marietta Vágvölgyiné Dr. Tóth, associate professor, PhD, Habil.

Title of course: Organic chemistry V.

Code: TTKBL0302_EN

ECTS Credit points: 7

Type of teaching, contact hours

- lecture: -
- practice: 2 hours/week
- laboratory: 4 hours/week

Evaluation: term mark

Workload (estimated), divided into contact hours:

- lecture: -
- practice: 28 hours

- laboratory: 56 hours
- home assignment: 32 hours
- preparation for the exam: -
Total: 116 hours

Year, semester: 3rd year, 1st semester

Its prerequisite(s): TTKBL0301_EN

Further courses built on it: -

Topics of course

The aim of the course is to enable students to learn new methods of organic chemical synthesis, learn their practical implementation and master the use of literature.

Students will get an individual list of tasks including ten organic compounds to be synthesized and a literature search. The execution of the tasks and the order of their implementation are planned by the students.

Literature

Compulsory:

1. J. R. Mohrig, D. G. Alberg, G. E. Hofmeister, P. F. Schatz, C. Noring Hammond: **Laboratory Techniques in Organic Chemistry** (Supporting Inquiry-Driven Experiments), **4th edition**, W. H. Freeman and Company. ISBN-13: 978-1-4641-3422-7.

Recommended:

1. E. K. Meislich, H. Meislich, J. Sharefkin: 3000 Solved problems in Organic Chemistry, McGraww-Hill INC, 1994.
2. P. Wyatt, S. Warren: Organic Synthesis: Wiley: Chischester, 2007
3. M. B. Smith: Organic Synthesis, 3rd Ed., McGraw-Hill: New York, 2008
4. F. A. Carey, R. J. Sundberg: Advanced Organic Chemistry, 3rd Ed., Part B, Plenum Press: New York-London, 1990
5. M. B. Smith, J. March: Advanced Organic Chemistry, 6th Ed., Wiley: New Jersey, 2007
6. R. C. Larock: Comprehensive Organic Transformations, 2nd Ed., Wiley: New York, 1999
7. R. O. C. Norman, J. M. Coxon: Principles of Organic Synthesis, 3rd Ed., Blackie Academic & Professional: London, 1994

Course objective/intended learning outcomes

a) Knowledge

- He/She has a basic chemical knowledge on describing simple chemical processes as well as on recognizing, organizing these in practice.
- He/She knows and can use the properties of most important chemicals, lab equipment and basic laboratory processes. He/She can follow and recognize the requirements of safety instructions.

b) Abilities

- Using his/her chemical knowledge and experiences He/She is able to create and execute experiments to demonstrate and prove the basic chemical phenomena.

- He/She is able to evaluate and discuss the laboratory measurements, and create a report about it.

c) Attitude

- It is important to him/her to work sufficiently and consciously in order to decrease the environmental load.
- He/She is ready to discuss problems on the field of chemistry and other science with professionals.

d) Autonomy and responsibility

- He/She is able to operate the most important technological instruments of chemical industries.
- He/She can evaluate his/her co-workers work's responsibly in either laboratory or industrial environment, and report it to his/her chief.

Schedule:

1st week

Introduction:

Timetable and requirements. Receiving of laboratory equipment and list of tasks. Safety education.

Preparative work:

Preparation and purification of a selected organic compound from the individual list. Determination of physical properties (melting point or boiling point) and purity (TLC, R_f). Calculation of yield.

2nd week

Short written test:

Topics: Substitution reactions (SN, SR) and their synthetic applications. Nomenclature. Safety rules.

Preparative work:

Preparation and purification of a selected organic compound from the individual list. Determination of physical properties (melting point or boiling point) and purity (TLC, R_f). Calculation of yield.

3rd week

Short written test:

Topics: Synthetic applications of electrophilic addition reactions. Nomenclature. Safety rules.

Preparative work:

Preparation and purification of a selected organic compound from the individual list. Determination of physical properties (melting point or boiling point) and purity (TLC, R_f). Calculation of yield.

4th week

Short written test:

Topics: Elimination reactions, Hofmann rule. Nomenclature. Basic laboratory techniques: heating, cooling. Safety rules.

Preparative work:

Preparation and purification of a selected organic compound from the individual list. Determination of physical properties (melting point or boiling point) and purity (TLC, R_f). Calculation of yield.

5th week

Short written test:

Topics: Diazotization, reactions of diazonium salts. Filtration techniques. Safety rules.

Preparative work:

Preparation and purification of a selected organic compound from the individual list. Determination of physical properties (melting point or boiling point) and purity (TLC, R_f). Calculation of yield.

6th week

Short written test:

Topics: Formation of functional groups on aromatic core. Nomenclature. Crystallization. Safety rules.

Preparative work:

Preparation and purification of a selected organic compound from the individual list. Determination of physical properties (melting point or boiling point) and purity (TLC, R_f). Calculation of yield.

7th week

Short written test:

Topics: Preparation and reaction of organometallic compounds. Nomenclature. Chromatography (TLC, column chromatography). Safety rules.

Preparative work:

Preparation and purification of a selected organic compound from the individual list. Determination of physical properties (melting point or boiling point) and purity (TLC, R_f). Calculation of yield.

8th week

Short written test:

Topics: Nucleophilic addition reactions. Nomenclature. Extraction methods. Safety rules.

Preparative work:

Preparation and purification of a selected organic compound from the individual list. Determination of physical properties (melting point or boiling point) and purity (TLC, R_f). Calculation of yield.

9th week

Short written test:

Topics: β -Dicarbonyl compounds and their synthetic applications. Nomenclature. Distillation at atmospheric pressure. Safety rules.

Preparative work:

Preparation and purification of a selected organic compound from the individual list.

Determination of physical properties (melting point or boiling point) and purity (TLC, R_f).
Calculation of yield.

10th week

Short written test:

Topics: Phase transfer catalysis. Vacuum distillation. Safety rules.

Preparative work:

Preparation and purification of a selected organic compound from the individual list.
Determination of physical properties (melting point or boiling point) and purity (TLC, R_f).
Calculation of yield.

11th week

Short written test:

Topics: Complex organic chemical problem solving (test). Nomenclature. Steam distillation.
Safety rules.

Preparative work:

Preparation and purification of a selected organic compound from the individual list.
Determination of physical properties (melting point or boiling point) and purity (TLC, R_f).
Calculation of yield.

12th week

Short written test:

Topics: Complex organic chemical problem solving (test). Nomenclature. Grouping of solvents
and their effects on organic chemical reactions. Safety rules.

Preparative work:

Preparation and purification of a selected organic compound from the individual list.
Determination of physical properties (melting point or boiling point) and purity (TLC, R_f).
Calculation of yield.

13th week

Short written test:

Topics: Complex organic chemical problem solving (test). Nomenclature. Safety rules.

Preparative work:

Preparation and purification of a selected organic compound from the individual list.
Determination of physical properties (melting point or boiling point) and purity (TLC, R_f).
Calculation of yield.

14th week

Performing missed identification tasks (melting point measurement, TLC), yield calculation.

Last occasion to present the synthesized products to the instructor.

Cleaning and handovering of equipments.

Evaluation.

Requirements:

Attendance at laboratory practice is mandatory.

Before starting to prepare a selected compound, students must give an oral report on their
theoretical organic chemistry and practical knowledge as well as on the safety rules.

The synthetic work can only be started after a successful discussion.

Minimum requirements for signing the course:

- Syntheses and characterizations of the selected 10 organic compounds.
- Sufficient level of the discussion (pass, (2)) for each preparation.
- Minimum level of the written test: at least 50% of the overall score.
- Presentation of the result of the literature search within the given time.

In case of failure of any subtask, the practice ends with a poor (1) grade.

The final grade will be determined based on the average of the grades of tasks. A weighted average of the grades of subtasks will be calculated in the following manner:

- Activity in laboratory practice, discussion (40%)
- Short written test (50%)
- Literature search (10%)

Final grade: excellent (5): 90%; good (4): 75%; satisfactory (3): 60%; pass (2): 50%; fail (1): below 50%.

Person responsible for course: Marietta Vágvölgyiné Dr. Tóth, associate professor, PhD, Habil.

Lecturer: Marietta Vágvölgyiné Dr. Tóth, associate professor, PhD, Habil.

Title of course: Biochemistry I. Code: TTBBE2035_EN	ECTS Credit points: 3
Type of teaching, contact hours - lecture: 2 hours/week - seminar: - - laboratory: -	
Evaluation: exam	
Workload (estimated), divided into contact hours: - lecture: 28 hours - practice: - - laboratory: - - home assignment: 28 hours - preparation for the exam: 40 hours Total: 96 hours	
Year, semester: 3 st year, 1 st semester	
Its prerequisite(s): TTKBE0303_EN	
Further courses built on it: TTKBL0303_EN, TTBE0304_EN	

Topics of course

Molecular design of life. Protein structure and function. Oxygen-transporting proteins: Myoglobin and Hemoglobin. Carbohydrates. Glycoconjugates. Glycobiology. Introduction to biological membranes. Enzymes. Metabolism: basic concepts and design. Glycolysis. Gluconeogenesis. Cori

cycle. Citric acid cycle. Oxidative phosphorylation. The pentose phosphate pathway. Glycogen metabolism. The coordinated control of synthesis and breakdown. Fatty acid metabolism. Oxidation of fatty acids and unsaturated fatty acids. Synthesis of ketone bodies. Biosynthesis of fatty acids. Digestion of proteins. Amino acid degradation. The urea cycle. The link between the urea and the citric acid cycle. The fates of the carbon skeletons of amino acids.

Literature

Compulsory:

- Lubert Stryer, Biochemistry, W. H. Freeman and Company, New York, 2002, ISBN 1-7167-4684-0.

Recommended:

- Glycoscience-Chemistry and Chemical Biology, (Eds: B. Fraser-Reid, K. Tatsua, J. Thiem) 2001, Springer-Verlag, Berlin

- Essentials of glycobiology (Eds: A.Varki, R. Cummings, J. Esko, H. Freeze, G. Hart, J. Marth, 1999, Cold Spring Harbor, New York, ISBN 0-87969-559-5)

Course objective/intended learning outcomes

a) Knowledge

-He/she fundamentally knows the basic principles of biochemistry.

-He/she expansively knows the areas of metabolism, bioregulation, and glycobiology.

b) Abilities

-He/she is able to apply the most important terminology, theories of the given biochemical field when completing the relevant tasks.

-He/she is able to understand scientific publications in the field of biochemistry.

c) Attitude

-He/she is open to learn and accept professional, technological improvement and innovation in his/her profession and convey it genuinely.

-He/she makes a decision in complex and unexpected decision cases by completely taking into account legal and ethical norms.

d) Autonomy and responsibility

-Even in unexpected decision-making situations he/she is capable of considering complex, fundamental questions from his/her professional field and elaborating them on the basis of the given sources.

-He/she is open to critical remarks which are professionally well-founded.

Schedule:

1st week: Introduction to Biochemistry. Molecular design of life. Amino acids. Peptides. Primary, secondary, tertiary, quaternary structures.

2nd week: Determination of peptide structures. Protein structure and function. Oxygen-transporting proteins: Myoglobin and Hemoglobin.

3rd week: Carbohydrates. Biological role of carbohydrates. Monosaccharides, disaccharides. Polysaccharides. Glycoconjugates. Glycobiology.

4th week: Introduction to biological membranes. Lipids. Classification and functions of lipids. Neutral fats, oils and waxes. The major classes of membrane lipids. Membrane models.

5th week: Enzymes. Classification. Coenzymes. Mechanism of enzyme action. Control of enzyme activity.

6th week: The kinetic properties of enzymes. The Michaelis-Menten model. Graphic evaluation of the kinetic parameters. Inhibition of enzyme activity. Diagnostic importance of enzymes.

7th week: Metabolism: basic concepts and design. Purine and pyrimidine bases, nucleosides and nucleotides. cAMP, ATP. Nucleotide coenzymes. Metabolism of carbohydrates. Glycolysis. The fate of pyruvate. Entry of fructose and galactose into glycolysis.

8th week: Gluconeogenesis. Cori cycle. The pentose phosphate pathway.

9th week: Citric acid cycle. Pyruvate dehydrogenase complex. The citric acid cycle is a source of biosynthetic precursors. Control of the citric acid cycle.

10th week: Oxidative phosphorylation. The four enzyme complexes of the respiratory chain. Synthesis of ATP. The ATP yield of the complete oxidation of glucose.

11th week: Glycogen metabolism. Glycogen degradation and synthesis. The coordinated control of synthesis and breakdown.

12th week: Fatty acid metabolism. Oxidation of fatty acids and unsaturated fatty acids. Energetics of fatty acid oxidation. Synthesis of ketone bodies.

13th week: Biosynthesis of fatty acids. The elongation cycle. Biosynthesis of cholesterol.

14th week: Digestion of proteins. Amino acid degradation. Transamination and oxidative deamination. The urea cycle. The link between the urea and the citric acid cycle. The fates of the carbon skeletons of amino acids.

Requirements:

- for a signature

Attendance at **lectures** is recommended, but not compulsory.

During the semester there are two tests: the mid-term test in the 8th week and the end-term test in the 15th week. Students have to sit for the tests

- for a grade

The course ends in an **examination**. Based on the average of the grades of the designing tasks and the examination, the exam grade is calculated as an average of them:

- the average grade of the two designing tasks
- the result of the examination

The minimum requirement for the mid-term and end-term tests and the examination respectively is 60%. Based on the score of the tests separately, the grade for the tests and the examination is given according to the following table:

Score	Grade
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0-59	fail (1)
60-69	pass (2)
70-79	satisfactory (3)
80-89	good (4)
90-100	excellent (5)

If the score of any test is below 60, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

-an offered grade:

it may be offered for students if the average grade of the two designing tasks is at least satisfactory (3) and the average of the mid-term and end-term tests is at least satisfactory (3). The offered grade is the average of them.

Person responsible for course: Dr. János Kerékgyártó, senior research fellow, CSc, PhD

Lecturer: Dr. János Kerékgyártó, senior research fellow, CSc, PhD

Title of course: Biochemistry II.

Code: TTKBL0303_EN

ECTS Credit points: 3

Type of teaching, contact hours

- lecture: -
- practice: 1 hours/week
- laboratory: 2 hours/week

Theoretical seminar (2 h), solving independent tasks (use of database, use of simulation program), practical laboratory work (4 h), evaluation and interpretation of results.

Evaluation:

Assessment methods:

An assessment carried out with written examinations at the end of semester.

Written examinations are used during the semester from the theoretical and practical part.

The ratings are not checked by a second examiner.

The examination papers are marked with name.

There is not an examination board.

Workload (estimated), divided into contact hours:

- lecture: -
- practice: 14 hours
- laboratory: 28 hours
- home assignment: 40 hours
- preparation for the exam: -

Total: 82 hours

Year, semester: 3nd year, 2nd semester

Its prerequisite(s): TTKBE2035_EN

Further courses built on it: -

Topics of course

Enzymes and mechanisms of enzyme action. Stability of enzymes, the influence of the reaction conditions on enzymatic activity. The Michaelis-Menten model for the kinetic properties of enzymes. Definition, significance and determination of K_M and v_{max} . Specific inhibition of enzymes and determination of the type of inhibition. Regulation of enzymes with allosteric interaction or covalent modification.

Preparation, activity measurement and kinetic investigation of some oxidoreductases and hydrolases.

Literature

Compulsory:

- Kandra Lili: Biokémiai gyakorlatok

Recommended:

- J. M. Berg, J. L. Tymoczko, L. Stryer: Biochemistry V. edition (W. H. Freeman and Co. 2002. ISBN 0-7167-4684-0)

- A. Cornish-Bowden: Fundamentals of enzyme kinetics, 3. reprint (Portland Press, 2002, ISBN 1 85578 072 0)

Course objective/intended learning outcomes

Objective of the course

To provide practical skills in biochemistry especially in enzymology and in characterization of biomolecules.

a) Knowledge

- He/she has the basic biochemical knowledge that will enable to describe basic life processes.
- He/she knows and is able to use the materials, tools and methods used in biochemical laboratories and has the relevant safety engineering knowledge.
- He/she possesses the knowledge whose application is necessary to solve the chemical problems of the living systems.
- He/she understands the essential biochemical, enzyme terminology in the mother tongue.

b) Abilities

- He/she is able to apply the most important terminology, theories, procedures of the given biochemistry field when completing the relevant tasks.
- He/she is able to create fundamental models of engineering systems and processes.

c) Attitude

- He/she seeks out to know the processes in the living organism and to describe their laws.
- He/she is environmentally conscious during his/her laboratory work.
- He/she is open to professional co-operation with professionals working in biochemistry.
- He/she is committed to acquiring new competencies.

d) Autonomy and responsibility

- During his laboratory work, he is able to consider basic professional issues independently, and can produce relevant compilations that can serve as a basis for decisions.
- He/she correctly evaluates the results of his / her own work and compares them with the results of his/her colleagues.
- Assesses the own work and the work of his / her colleagues responsibly during their laboratory activities.

Schedule: practices - 2 hours/week, laboratory - 5 hours/week, two independent tasks

1st week

Labor safety education. Semester schedule. Theory: The concept, structure and grouping of enzymes. Parameters influencing the speed of enzyme reactions. Occurrence, function, structure and activity of lipase enzyme.

2nd week

Laboratory practice: Extraction of lipase enzyme and determination of its activity.

3rd week

Enzyme activity measurement, reaction rate measurement for enzyme reactions. Enzyme structure and function relationship. Coenzymes, prosthetic groups. Enzyme regulation. The occurrence, function and structure of the catalase enzyme. Hem is a prosthetic group. Generation of hydrogen peroxide in living organisms, FADH₂ coenzyme, superoxide dismutase. Enzyme databases, molecular modelling.

4th week

Laboratory practice: Extraction of catalase enzyme from plant tissue, measurement of activity.

5th week

The mechanism of enzyme activity. Structural analysis of proteins. How can we develop an enzyme activity measurement method? The function and significance of the amylase enzyme, its mechanism of action and its activity. Definition and calculation of the subsite map.

6th week

Laboratory practice: Study of starch and oligosaccharide hydrolysis catalysed by amylase enzyme

7th week

Overview of the virtual laboratory program. Enzyme assays to **investigate the effects of pH, time, amount of enzyme, incubation temperature and substrate concentration on the activity of different enzymes. Students can also investigate the effects of adding different inhibitors, as well. The students carry out the tasks independently at home.**

8th week

Kinetics of enzymatic reactions, inhibition types. Methods for determining kinetic constants. Computer evaluation of enzyme kinetic measurements. The Grafit enzyme kinetic program. Function of emulsion beta-glucosidase, method of measuring activity.

9th week

Laboratory practice: Determination of kinetic parameters of almond emulsin beta-glucosidase. Enzyme and substrate concentration dependence of reaction rate. Determination of enzyme kinetic parameters K_M and v_{max} and inhibition assay.

10th week

Presentation and discussion of results obtained from a search for a given enzyme in the protein and enzyme databases.

11th week

End term test

Requirements:

- for a signature

Participation at **practice and laboratory classes** is compulsory. A student must attend the practice classes and may not miss more than one times during the semester. In case a student does so, the subject will not be signed and the student must repeat the course. A student can't make up any practice with another group. Attendance at practice classes will be recorded by the practice leader. Being late is equivalent with an absence. In case of further absences, a medical certificate needs to be presented. Missed practice classes should be made up for at a later date, to be discussed with the tutor.

During the semester there are tests in every second week as a part of the practice, which are

mandatory.

Students have to submit all the two tasks (database search and virtual laboratory) as a minimum on a sufficient level.

- *for a grade*

The course is evaluated based on the tests, designing tasks and the lab notebooks. The grade is calculated as an average.

The minimum requirement is 60%. The grade for the practice is given according to the following table:

Score	Grade
0-59	fail (1)
60-69	pass (2)
70-79	satisfactory (3)
80-89	good (4)
90-100	excellent (5)

If the score of any test is below 60, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

Person responsible for course: Dr. Gyöngyi Gyémánt, associate professor, PhD, habil

Lecturer: Dr. Gyöngyi Gyémánt, associate professor, PhD, habil

Title of course: Analytical Chemistry I. Code: TTKBE0501_EN	ECTS Credit points: 3
Type of teaching, contact hours - lecture: 2 hours/week	
Evaluation: examination	
Workload (estimated), divided into contact hours: - lecture: 28 hours - practice: - - laboratory: - - home assignment: - - preparation for the exam: 56 hours Total: 84 hours: 3 credit x 28 hours	
Year, semester: 2 nd year, 1 st semester	
Its prerequisite(s): TTKBE0201_EN, TTKBE0301_EN, TTKBE0401_EN	
Further courses built on it: TTKBL0512_EN	

Topics of course
Literature
<i>Compulsory:</i>

- 1) Syllabus provided by the tutor
- 2) Daniel C. Harris: Quantitative Chemical Analysis, 7th Ed., 2007, Freeman and Co.
- 3) Vogel's Qualitative Inorganic Analysis, (ed. Gy. Svehla), Longmann, 2007

Course objective/intended learning outcomes

a) Knowledge

- He/she knows the basic terms, the application fields of analytical chemistry.
- He/she knows the basics of the theory and practice of various analytical methods.
- He/she knows the basics of various separation techniques.
- He/she knows the practical application of various atomic spectrometric, UV-Vis and pH-potentiometric methods.

b) Abilities

- He/she is capable of setting up and carrying out analytical measurements based on classical titrimetric methods. He/she understands the details of the analytical application of chemical equilibria.
- He/she is familiar with the problems associated with measurements, uncertainties and errors.

c) Attitude

- He/she is open to learn and accept professional, technological improvement and innovation in his/her profession and convey it genuinely.
- He/she makes a decision in complex and unexpected cases by completely taking into account legal and ethical norms.

d) Autonomy and responsibility

- Even in unexpected decision-making situations he/she is capable of considering complex, fundamental questions from his/her professional field and elaborating them on the basis of the given sources.
- He/she confesses and represents the value system of the qualified chemists and chemical engineers profession with responsibility. He/she is open to critical remarks which are professionally well-founded.

Schedule:

1st week

Introduction to analytical chemistry. Measurements. Basic equations of equilibrium calculations.

2nd week

Acids and bases, acid-base theories. The Broensted equation. Buffers.

3rd week

Basic terms related to titrations. Practice of acid-base titrations.

4th week

Basics of complexometry. Complexometric titrations.

5th week

Solubility equilibria. Precipitation titrations, argentometry.

6th week

Redoxi equilibria. Permanganometry.

7th week

Chromatometry. Bromatometry. Iodometry.

8th week

Simple separation techniques I. Gravimetry.

9th week

Simple separation techniques II. Extraction.

10th week

Chromatographic separations and techniques.

11th week

Classification of instrumental analytical methods. Evaluation of analytical chemical results.

12th week

Spectroscopy I. Atomic spectroscopy.

13th week

Spectroscopy II. UV-Vis spectroscopy.

14th week

Potentiometry and conductometry.

Requirements:

- for a signature

Attendance at **lectures** is recommended, but not compulsory.

- for a grade

The course ends in an **examination**.

The minimum requirement for the examination is 40 score. Based on the score, the grade for the examination is given according to the following table:

Score	Grade
0-39	fail (1)
40-55	pass (2)
56-70	satisfactory (3)
71-85	good (4)
86-100	excellent (5)

If the score of any test is below 40, students can retake the test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

Person responsible for course: Dr. Péter Buglyó, associate professor, PhD

Lecturer: Dr. Péter Buglyó, associate professor, PhD

Title of course: Analytical Chemistry I (seminar)

Code: TTKBG0501_EN

ECTS Credit points: 2

Type of teaching, contact hours

- lecture: -
- practice: 2 hours/week
- laboratory: -

Evaluation: practice grade

Workload (estimated), divided into contact hours:

- lecture: -
- practice: 28 hours
- laboratory: -
- home assignment: 32 hours
- preparation for the exam: -

Total: 60 hours

Year, semester: 2nd year, 1st semester

Its prerequisite(s): TTKBG0101_EN, and parallel registration to TTKBE0501_EN

Further courses built on it: -

Topics of course

Calculations involved acid-base, redox, argentometric and complexometric equilibrium and titrations

Literature

Compulsory:

Study Aids on the website

Recommended:

Daniel C. Harris: Quantitative Chemical Analysis

R. Kellner, J.-M. Mermet, M. Otto, H. M. Widner: Analytical Chemistry, Wiley, 1997

Course objective/intended learning outcomes

a) Knowledge

- He/She knows the basic qualitative and quantitative chemical principles, and the methods based on it

b) Abilities

- He/She is able to use the previously obtained knowledge on the field of natural and anthropogenic studies to solve practical problems.

c) Attitude

- He/She is able to represent his/her own personal scientific ideology toward professional and unprofessional groups.

d) Autonomy and responsibility

- Under supervision he/she is responsible in collaboration with other professionals (especially from the field of environmental economy and safety)

Schedule:

1st week

Calculations in acid-base systems: Simple problems about pH calculations (revision). Quantitative description of solutions containing monobasic acids and bases.

2nd week

Buffers in acid-base chemistry. Titration curves, calculation of final results from experimental data.

3rd week

Di- and polybasic acids and bases, ampholytes (illustration with evaluating the titration curve of a sample of phosphoric acid). Problems based on acid-base titrations.

4th week

Problems based on acid-base titrations. Calculation of equivalence points, indicator selection. Calculations for planning titration-based methods, and of distribution curves of species.

5th week

Practice, consultation.

6th week

Test I.

7th week

Complex formation equilibria. The concept and calculation of conditional stability constants.

8th week

Calculations connected to complexometric titration methods.

9th week

Quantitative description of redox equilibria.

10th week

Calculations based on redox titration methods.

11th week

Quantitative description of precipitation equilibria. Solubility product and solubility.

12th week

Effects of pH and the excess of precipitating ion on solubility. Problems based on precipitation reactions and precipitation-based titrimetric methods.

13th week

Practice, consultation.

14th week

Test II.

Requirements:

- *for a signature*

Attendance at **seminars** is recommended, but not compulsory.

During the semester there are two tests: the mid-term test in the 6th week and the end-term test in the 14th week. Students have to sit for the tests.

- *for a grade*

If the score of any test is below 10 points of the maximal 50 points, the grade is fail (1).

The grade is calculated from the results of the tests.

%	Grade
0-49	fail (1)
50-61	pass (2)
62-74	satisfactory (3)
75-86	good (4)
87-100	excellent (5)

Students with fail (1) final course grade due to low test results can take once a comprehensive test exam during the examination period, where pass (2) can be obtained.

Person responsible for course: Dr. Csilla Kállay, senior research fellow, PhD

Lecturer: Dr. Csilla Kállay, senior research fellow, PhD

Title of course: Analytical Chemistry I (seminar)

Code: TTKBL0501_EN

ECTS Credit points: 2

Type of teaching, contact hours

- lecture: -
- practice: -
- laboratory: 4 hours/week

Evaluation: practice grade
Workload (estimated), divided into contact hours: - lecture: - - practice: - - laboratory: 56 hours - home assignment: 64 hours - preparation for the exam: - Total: 120 hours
Year, semester: 2 nd year, 1 st semester
Its prerequisite(s): TTKBL0101_EN, and parallel registration to TTKBE0501_EN
Further courses built on it: -

Topics of course
Acid-base, redox, argentometric and complexometric titrations
Literature
<i>Compulsory:</i> Study Aids on the website <i>Recommended:</i> Daniel C. Harris: Quantitative Chemical Analysis R. Kellner, J.-M. Mermet, M. Otto, H. M. Widner: Analytical Chemistry, Wiley, 1997
Course objective/intended learning outcomes
a) Knowledge - He/She knows the basic qualitative and quantitative chemical principles, and the methods based on it - He/She knows and can use the properties of most important chemicals, lab equipment and basic laboratory processes. He/She can follow and recognize the requirements of safety instructions.
b) Abilities - He/She is able to use the previously learned paradigms, theories and laws on the field of chemical, natural science to plan, execute and evaluate laboratory experiments. - He/She is able to evaluate and discuss the laboratory measurements, and create a report about it.
c) Attitude - It is important to him/her to work sufficiently and consciously in order to decrease the environmental load.
d) Autonomy and responsibility - During laboratory work he/she is capable to get related to the basics of chemical problems, able to conclude the questions to the supervisor that might come up to the solutions or decisions.

Schedule: 1 st week Introduction to the Quantitative Analytical Chemistry Laboratory. Laboratory Safety Information.

Review of lab equipment.

2nd week

Preparation of ~0.1 M HCl titrant (250 ml).

Determination of the exact concentration of the HCl titrant solution using potassium hydrogen carbonate stock solution.

Preparation of ~0.1 M NaOH titrant by the Sørensen (500 ml) and determination of its exact concentration.

3rd week

Determination of borax content of a solid sample (unknown sample).

Simultaneous determination of sulfuric acid and boric acid in a mixture (unknown sample).

4th week

Determination of oxalic acid (unknown sample).

Determination of $\text{Na}_2\text{S}_2\text{O}_3$ by measuring the acid formed in the oxidation reaction of $\text{Na}_2\text{S}_2\text{O}_3$ with bromine.

5th week

Determination of ascorbic acid active ingredient content of vitamin C tablet (unknown sample).

Determination of the composition of KCl-KBr mixture using 0.05 M silver nitrate stock solution (unknown sample).

Preparation of 0.02 M potassium bromate titrant (250.00 ml).

6th week

Determination of the exact concentration of the potassium permanganate titrant solution using sodium oxalate stock solution.

Determination of ferrous oxalate by permanganometric titration (unknown sample).

Determination of hydrogen peroxide (unknown sample).

7th week

Preparation of 0.02 M sodium thiosulfate titrant (250 ml) and determination of its exact concentration using 0.003 M potassium iodate stock solution.

Determination of iodide ion (unknown sample).

8th week

Redetermination of the exact concentration of the prepared 0.02 M sodium thiosulfate titrant

Determination of copper(II) (unknown sample).

9th week

Preparation of 0.01 M Na_2EDTA titrant solution (250.00 ml).

Simultaneous determination of calcium(II) and magnesium(II) ions (unknown sample).

Determination of Bi(III) (unknown sample).

10th week

Simultaneous determination of copper(II) and zinc(II) ions (unknown sample).

11th week

Quantitative description of precipitation equilibria. Solubility product and solubility.

12th week

Determination of Al(III) (unknown sample).

13th week

Lab equipment return.

14th week

Evaluation

Requirements:

- for a signature

Participation at practice classes is compulsory. A student must attend the practice classes and may not miss more than one during the semester. In case a student does so, the subject will not be signed and the student must repeat the course. Being late is equivalent with an absence. In case of absences, a medical certificate needs to be presented. Missed practice classes should be made up for at a later date, to be discussed with the tutor.

Each week the laboratory session begins with a short test (not more than 15 minutes) based exclusively on the preparatory material and calculations of that week.

Students are required to determine “unknown samples”.

- for a grade

The grade is calculated from the results of the tests (50%) and the unknown samples (50%). Both averages have to be to be minimum 2.00 in order to successfully complete the course. Otherwise the final grade will be fail (1). Students with fail (1) final course grades thanks to unacceptable test results can take once a comprehensive test exam during the examination period.

Person responsible for course: Dr. Csilla Kállay, senior research fellow, PhD

Lecturer: Dr. Csilla Kállay, senior research fellow, PhD

Title of course: Separation Techniques I Code: TTKBE0502_EN	ECTS Credit points: 1
Type of teaching, contact hours - lecture: 1 hours/week	
Evaluation: examination	
Workload (estimated), divided into contact hours: - lecture: 14 hours - practice: - - laboratory: - - home assignment: - - preparation for the exam: 28 hours Total: 42 hours: 1 credit x 42 hours	
Year, semester: 2 nd year, 1 st semester	
Its prerequisite(s): TTKBE0201_EN, TTKBE0301_EN, TTKBE0401_EN	
Further courses built on it: TTKBL0502_EN	

Topics of course
Basic concepts, theoretical and practical aspects, carry-out and use of fundamental laboratory and industrial scale separation processes. Set-up, major components and basic operation principles of modern analytical instruments using separation methods in their working methods.
Literature
<i>Compulsory:</i> 1) Separation process principles: chemical and biochemical operations / J. D. Seader, Ernest J. Henley, D. Keith Roper.—3rd ed. 2011, ISBN 978-0-470-48183-7, John Wiley & Sons, Inc.

2) Modern analytical chemistry / David Harvey. — 1st ed., 2000, ISBN 0-07-237547-7, The McGraw-Hill Companies, Inc.

Recommended:

3) Modern HPLC for practicing scientists / by Michael W. Dong., 2006, John Wiley & Sons, Inc., Hoboken, New Jersey, ISBN-13: 978-0-471-72789-7

4) Modern size-exclusion liquid chromatography / André M. Striegel et al., 2nd ed., 2009 by John Wiley & Sons, Inc., ISBN 978-0-471-20172-4

5) Modern practice of gas chromatography., 4th ed. / edited by Robert L. Grob, Eugene F. Barry. 2004 by John Wiley & Sons, Inc., ISBN 0-471-22983-0

6) Affinity Chromatography Methods and Protocols, 2nd Ed., Ed. by Michael Zachariou, 2008, Humana Press, a part of Springer Science+Business Media, LLC, ISBN: 978-1-58829-659-7

7) Gel Electrophoresis of Proteins A Practical Approach, 3rd Edition, B. D. Hames, Oxford University Press, 1998, ISBN 0-19-963641-9

Course objective/intended learning outcomes

a) Knowledge

- He/she fundamentally knows principles, procedures, properties, reactions and chemical processes required to perform laboratory and industrial separation processes, and control and use instrumental analytical techniques.

b) Abilities

- He/she is able to apply the most important terminology, theories, procedures of the given separation-related or analytical field when completing the relevant tasks.

- He/she is able to create fundamental models of separation processes and perform instrumental analytical investigations.

c) Attitude

- He/she is open to learn and accept professional, technological improvement and innovation in his/her profession and convey it genuinely.

- He/she makes a decision in complex and unexpected cases by completely taking into account legal and ethical norms.

d) Autonomy and responsibility

- Even in unexpected decision-making situations he/she is capable of considering complex, fundamental questions from his/her professional field and elaborating them on the basis of the given sources.

- He/she confesses and represents the value system of the qualified chemists' or chemical engineers' profession with responsibility. He/she is open to critical remarks which are professionally well-founded.

Schedule:

1st week

Basic concepts of separation processes. Removal of solvents from a mixture by different techniques: partial evaporation, batch and continuous mode thin film evaporators and concentrators. Partial evaporation by a rotating and circulating gas flow. Centrifugal evaporators.

2nd week

Partial freezing, removal of frozen solvent crystals, increase of concentration. Spray drying, freeze drying. Instruments of spray drying, practical use of spray drying for the production of drugs, and foods. Instruments of freeze drying, laboratory scale to industrial production. Freeze-dried food production and use.

3rd week

Separation of solid mixtures by physical methods: air flow sedimentation, selective dissolution, magnetic separation, flotation. Separation by solubilities, fractionated crystallization, crystallization. Stoke's law. Removal of solids from liquids and gases: sedimentation, centrifugation, cyclons, ultracentrifugation. Separation of emulsions by special centrifuges. Gas separation and isotopes enrichments with gas centrifuges.

4th week

Filtration: basic concepts, formation and role of filter cakes. The good laboratory practice of filtration. Removal of dust from gas streams, industrial sack-type filteres, filter candles, electrostatic dust collectors. Types of filter media, filter papers, filter membranes. Filtration apparatuses. Vacuum filtration, pressure filtration. Tangential filtration.

5th week

Extraction: liquid-liquid liquid-solid and liquid-gas processes. From laboratory scale to undustrial liquid-liquid extractors, the role of density, practical uses. Basic rules of extraction, disribution coefficients, selectivites, design of an extraction scheme. Soxhlet extractors, heated and non-heated types. Solid phase extraction (SPE) and solid phase microextraction (SPME), use of SPME in sample preparation. Osmosis, dialysis, reverse osmosis instruments and their use in drinking water production. Membrane dialysis, separation of molecules by size, medical application, hemodialysis.

6th week

General aspects and types of different chromatographic techniques. Grouping of techniques by the dimension of the separating medium. Layer chromatographies: paper chromatography (PC), thin layer chromatography (TLC). Basics of TLC: tools, chambers, separation modes, geometry, types of layers, calculations, visualization and evaluation methods. Computer aided analysis of TLC and HPTLC plates. Two-dimensional TLC.

7th week

Gas chromatography 1: Definition, basics of intruments. Sample preparation for chromatographic analysis: concentration, dissolution, filtration, extraction, head-space sampling, SPME, derivatization, adsorption. General setup, gas supply system , rotating and robot arm sample holders, injectors. The inlet: the key role of rapid sample evaporation.

8th week

Gas chromatography 2: Types of inlets, oven, temperature control, gas chromatography detectors (FID, ECD, MS). Types of analytes that can be measured by the given detectors. Working principles of FID? ECD and MS detectors. Preparative gas chromatography. Web communication within and outside of laboratories. 2D-gas chromatography (2D-GCxGC).

9th week

High pressure liquid chromatography (HPLC) 1. Basic principles, structure, potential fields of applications. Separation mechanisms and separation modes. Most important structural units and components of the HPLC instrument. Solvent supply system, degass station. Role of degassing, different degassing modes. Graadient formation unit. HPLC pumps, workin g principles, types, role of depulser. Major types of HPLC columns. Stationary phases, normal phase and reversed phase.

10th week

HPLC detectors, their working principles, structure, mode of use. (UV-Vis, scanning UV-Vis, dioade array, refractive index, fluorescence, evaporative light scattering, and mass spectrometry detectors). Isocratic and gradient elutions. Characterization of the chromatograms. Preparative HPLC.

11th week

Low pressure chromatography. Traditional, classic column chromatography, dry column

chromatography, flash chromatography. Basic operating techniques, limits of separations, hardware requirements, manual mode and instrumentation.

12th week

Affinity chromatography. General principles, hardware requirements, special interaction between the stationary phase and the analytes. Elution of the analytes. Operation in column mode and in the batch mode.

13th week

Gel chromatography. Basic principles, working concepts. Dead volume, gel volume, exclusion limit, penetration. Measurement of the bed volume, separation of large molecules. Bed making, conditioning. Separation of smaller molecules in organic solvent gel system. Characterization of the gel chromatograms, calculation of the molar mass.

14th week

Gel electrophoresis. Basic principles, translation of ions within a gel by the external electric potential. Types of gel materials, their use in the separation of proteins and nucleic acids. Vertical and horizontal electrophoresis chamber, gel casting, use of the comb. Loading of samples. Development of the gel. Visualization of the gel electroferograms, blotting. Computer aided evaluation and documentation.

Requirements:

- for a signature

Attendance at **lectures** is recommended, but not compulsory.

- for a grade

The course ends in an **examination**.

The minimum requirement for the examination is 50 score. Based on the score, the grade for the examination is given according to the following table:

Score	Grade
0-49	fail (1)
50-62	pass (2)
63-75	satisfactory (3)
76-88	good (4)
89-100	excellent (5)

If the score of any test is below 60, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

Person responsible for course: Dr. István Lázár, associate professor, PhD

Lecturer: Dr. István Lázár, associate professor, PhD

Title of course: Separation Techniques II

Code: TTKBL0502_EN

ECTS Credit points: 3

Type of teaching, contact hours

- laboratory: 3 hours/week, organized in six blocks in the semester

Evaluation: examination

Workload (estimated), divided into contact hours:

- lecture: -

<ul style="list-style-type: none"> - practice: - - laboratory: 42 hours - home assignment: 14 hours - preparation for the exam: - <p>Total: 56 hours: 3 credit x 18 hours</p>
Year, semester: 2 nd year, 2 nd semester
Its prerequisite(s): TTKBL0201_EN, TTKBE0502_EN
Further courses built on it:

Topics of course
Basic concepts, theoretical and practical aspects, carry-out and use of fundamental laboratory analytical separation and identification processes. Basic operation principles of modern analytical instruments using separation methods in their working methods. Use of analytical instruments and techniques to separate and identify components of complex samples.
Literature
<p><i>Compulsory:</i></p> <ol style="list-style-type: none"> 1) Specific handouts, each provided for the given laboratory practice. 2) Separation process principles: chemical and biochemical operations / J. D. Seader, Ernest J. Henley, D. Keith Roper.—3rd ed. 2011, ISBN 978-0-470-48183-7, John Wiley & Sons, Inc. 3) Modern analytical chemistry / David Harvey. — 1st ed., 2000, ISBN 0-07-237547-7, The McGraw-Hill Companies, Inc. <p><i>Recommended:</i></p> <ol style="list-style-type: none"> 4) Modern HPLC for practicing scientists / by Michael W. Dong., 2006, John Wiley & Sons, Inc., Hoboken, New Jersey, ISBN-13: 978-0-471-72789-7 5) Modern size-exclusion liquid chromatography / André M. Striegel et al., 2nd ed., 2009 by John Wiley & Sons, Inc., ISBN 978-0-471-20172-4 6) Modern practice of gas chromatography., 4th ed. / edited by Robert L. Grob, Eugene F. Barry. 2004 by John Wiley & Sons, Inc., ISBN 0-471-22983-0 7) Affinity Chromatography Methods and Protocols, 2nd Ed., Ed. by Michael Zachariou, 2008, Humana Press, a part of Springer Science+Business Media, LLC, ISBN: 978-1-58829-659-7 8) Gel Electrophoresis of Proteins A Practical Approach, 3rd Edition, B. D. Hames, Oxford University Press, 1998, ISBN 0-19-963641-9 9) Thin-Layer Chromatography, 4th Edition, by Joseph Sherma, Bernard Fried, 1999, Marcel Dekker Inc., New York, Basel, ISBN: 0-8247-0222-0
Course objective/intended learning outcomes
<p>a) Knowledge</p> <ul style="list-style-type: none"> - He/she fundamentally knows principles, procedures, properties, reactions and chemical processes required to perform analytical laboratory separation processes, and control and properly use classic and modern instrumental analytical techniques. <p>b) Abilities</p> <ul style="list-style-type: none"> - He/she is able to apply the most important terminology, theories, procedures of the given separation-related analytical field when completing the relevant tasks. - He/she is able to create fundamental models of separation processes and perform instrumental

analytical investigations.

c) Attitude

- He/she is open to learn and accept professional, technological improvement and innovation in his/her profession and convey it genuinely.

- He/she makes a decision in complex and unexpected cases by completely taking into account legal and ethical norms.

d) Autonomy and responsibility

- Even in unexpected decision-making situations he/she is capable of considering complex, fundamental questions from his/her professional field and elaborating them on the basis of the given sources.

- He/she confesses and represents the value system of the qualified chemists' or chemical engineers' profession with responsibility. He/she is open to critical remarks which are professionally well-founded.

Schedule:

1st block

Gas chromatography.

2nd block

High pressure liquid chromatography.

3rd block

Thin layer chromatography.

4th block

Gel permeation chromatography.

5th block

Low pressure liquid chromatography.

6th block

Radioisotopes separation.

Requirements:

- for a signature

Attendance at all of the laboratory practice blocks and fulfilment of the required experiments/tasks, preparing and presenting a valid lab report.

- for a grade

The course ends in a **term mark**.

Each block of practice ends in a mark. A mean value of all marks received in the course are calculated. The minimum requirement for the examination is a mean value of 2.00. Based on the mean value, the grade for the term mark is given according to the following table:

Mean value	Grade
0-1.99	fail (1)
2.00-2.74	pass (2)
2.75-3.49	satisfactory (3)
3.50-4.24	good (4)
4.25-5.00	excellent (5)

If the mean value is below 2.00, students retake the laboratory practice in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

Person responsible for course: Dr. István Lázár, associate professor, PhD

Lecturer: Dr. István Lázár, associate professor, PhD

Title of course: Instrumental analysis II

ECTS Credit points: 6

Code: TTKBL0503_EN

Type of teaching, contact hours

- lecture: -
- practice: -
- laboratory: 6-8 hours/week

Evaluation: practice grade

Workload (estimated), divided into contact hours:

- lecture: -
 - practice:
 - laboratory: 78 hours
 - home assignment: 102 hours
 - preparation for the exam: -
- Total: 180 hours

Year, semester: 2nd year, 2st semester

Its prerequisite(s): TTKBE0501, TTKBL0501

Further courses built on it:

Topics of course

The series of laboratory practices are based on the topics of different instrumental analysis like electrophoresis, atomic spectrometry, electroanalysis, validation, spectroscopic methods (IR, UV/vis, X-ray). The instrumental laboratories are connected to the topics of the Instrumental Analysis lecture and the Separation Techniques lecture and laboratory practice.

Literature

1. Daniel C. Harris: Quantitative Chemical Analysis, 7th Ed., 2007, Freeman and Co.H.H.
2. Willard, L.L. Merritt, J.A. Dean, F.A. Settle: Instrumental methods of Analysis, Wadsworth Publ. Co., Belmont, 1988.
3. Douglas A. Skoog, Donald M. West, F. James Holler, Stanley R. Crouch: Fundamentals of Analytical Chemistry, 8th. ed., 2004, Brooks/Cole
4. Syllabuses provided by the tutor.

Course objective/intended learning outcomes

a) Knowledge

- He/she fundamentally knows principles and means of instrumental analysis, sample pretreatment, data evaluation and validation of the measurements.
- He/she expansively knows the operating principles of the analytical instruments, auxiliary devices.

b) Abilities

- He/she is able to apply the most important terminology, theories, procedures of the given

instrumental analysis field when completing the relevant tasks.

- He/she is able to find solutions for the analytical problems.

c) Attitude

- He/she is open to learn and accept professional, analytical improvement and innovation in his/her profession and convey it genuinely.
- He/she makes a decision in complex and unexpected decision cases by completely taking into account legal and ethical norms.

d) Autonomy and responsibility

- Under supervision he/she is responsible in collaboration with other professionals (especially from the field of analytical and environmental economy and safety).
- He/she confesses and represents the value system of the instrumental analytical profession with responsibility. He/she is open to critical remarks which are professionally well-founded.

Schedule:

1st week: Introductory guidance, accident protection (2h)

2nd week: Gel electrophoresis (DNA analysis) (6h)

3rd week: Chromatographic purification methods for proteins (6h)

4th week: High Performance Liquid Chromatography II (8h)

5th week: Mass spectrometry (ESI, MALDI, CE-MS) (8h)

6th week: Evaluation of chromatograms (8h)

7th week: Gas chromatography – mass spectroscopy (6h)

8th week: IR spectroscopy (6h)

9th week: Atomic spectroscopy (8h)

10th week: UV-vis spectroscopy (6h)

11th week: pH-metry (6h)

12th week: Final test (2h)

Requirements:

- for a signature

Participation at practices is compulsory. A student must attend every practices during the semester. Attendance at practices will be recorded by the practice leader. Being late is equivalent with an absence. In case of further absences, a medical certificate needs to be presented. Missed practice classes should be made up for at a later date, to be discussed with the tutor. Active participation is evaluated by the teacher in every class. If a student's behavior or conduct doesn't meet the requirements of active participation, the teacher may evaluate his/her participation as an absence because of the lack of active participation in class.

- for a grade

Grading is given by the average of 3 separate grades:

- the average grade of the short tests written at the beginning of the instrumental analysis lab practices (an average of at least 2.0 is necessary to avoid a 'fail' final grade)
- the average grade of evaluation of the analytical data measured by the instrument, the laboratory notebook prepared by the student and final discussion/conclusion made between the student and the supervisor at the end of the lab practice (an average of at least 2.0 is necessary to avoid a 'fail' final grade)
- the grade of the final test

The grade of the final test is calculated according to the following table:

Score	Grade
0-59	fail (1)
60-69	pass (2)
70-79	satisfactory (3)
80-89	good (4)
90-100	excellent (5)

Person responsible for course: Dr. Attila Gáspár, professor, DSc

Lecturer Dr. Attila Gáspár, professor, DSc

Title of course: Spectroscopic methods I. Code: TTKBE0503_EN	ECTS Credit points: 3
Type of teaching, contact hours - lecture: 2 hours/week	
Evaluation: exam	
Workload (estimated), divided into contact hours: - lecture: 28 hours - home assignment: 22 hours - preparation for the exam: 40 hours Total: 90 hours:	
Year, semester: 3 rd year, 2 nd semester	
Its prerequisite(s): TTKBE0302_EN, TTFBE2113_EN	
Further courses built on it: TTKBL0504_EN, TTKBL0004_EN	

Topics of course

Modern chemical analytics is based on different branches of spectroscopy. The series of lecture are based on the topics of Nuclear Magnetic Resonance, Mass Spectrometry (MS), Infrared Spectroscopy (IR) and Ultraviolet/Visible Spectroscopy (UV).

It reviews: the fundamental relations of the angular momentum and nuclear magnetism, the connections between magnetic field and nuclear magnetisation, the selection rule for NMR and the resonance condition. After that it deals with connections between electron density shielding and chemical shifts; scalar spin-spin coupling, Karplus relationship, first order spectrum (weak coupling), first order rules, second-order spectrum ("strong" coupling), ¹³C NMR. In addition, theory and practice of optical and mass-spectroscopy is covered.

Literature:

1. Andrew Derome, Modern NMR Techniques for Chemistry Research, Pergamon, ISBN-10: 0080325149

2. Timothy D.W. Claridge, High-Resolution NMR Techniques in Organic Chemistry, Elsevier, ISBN: 9780080999869
3. Neil Jacobsen, NMR Spectroscopy Explained, Wiley, ISBN-10: 0471730963
4. R.M.Silverstein, F.X.Webster: "Spectrometric Identification of Organic Compounds", Wiley, 1998.
5. F.W.McLafferty: „Interpretation of mass spectra”, W.A.Benjamin, INC, New York, 1967
6. J.R.Chapman: „Practical Organic Mass Spectrometry”, Wiley, 1995
7. E.Pretsch, J.T.Clerc: „Interpretation of Organic Compounds”, VCH, 1997

Course objective/intended learning outcomes

a) Knowledge: - He/She knows the basic qualitative and quantitative chemical principles, and the methods based on it. He/She knows the main models and theories of chemical bonds and molecular structure based on scientific findings. Understanding the basic principles of pulse Fourier NMR method. Knows the representative ^1H and ^{13}C chemical shifts of functional groups, has information about the use of homo and heteronuclear spin-spin couplings and NOE-s.

b) Abilities: - He/She is able to use the previously obtained knowledge on the field of natural and anthropogenic studies to solve practical problems. Able to analyse first order high resolution ^1H NMR spectra, active knowledge on expected chemical shift ranges both for ^1H and ^{13}C spectra. Able to combine and evaluate the results of NMR, MS and IR spectra simultaneously.

c) Attitude:

- He/She is ready to discuss problems on the field of chemistry and other science with professionals.

- He/She is open toward scientific and other postgradual education.

- He/She is committed learn or get insights into new competence or ideology.

- During his/her work he/she committed to apply the quality concerns including the new assurances.

Willing to cope with the structure elucidation problem of unknown compounds. -

d) Autonomy and responsibility: - He/She can make reasonable evaluations about his/her own work comparing to others to the same field. Capable to predict or interpret NMR, MS and IR spectra of compounds. Critically evaluates spectra, recognizes possible contradictions and takes responsibility for final decision on the anticipated chemical structure.

Schedule:

1st week **Basics of NMR:** Magnetic dipoles in external B_0 field, nuclear Zeeman effect, selection rules, transition frequency, populations, Boltzmann distribution, bulk magnetisation, vector model. B_1 radiofrequency excitation, CW and pulse-Fourier spectrometer schemes. NMR active nuclei. Fields of applications: solid-state NMR, MRI, tomography in material science, relaxation for drug quality control and oil research.

2nd week **NMR parameters:** Spin-lattice (T_1) and spin-spin (T_2) relaxation. The nuclear Overhauser effect. Chemical shielding, chemical shift, ppm scale. Factors influencing chemical shifts. Indirect scalar spin-spin couplings. Splitting patterns of multiplets, multiplicity rules. Karplus curves for determining dihedral angles.

3rd week **Analysis of high resolution NMR spectra 1.** : ^1H spin system labelling rules based on molecular structure. First order analysis of ^1H NMR spectra. Strong couplings and their impact. Integration of ^1H NMR spectra, rules for quantitative NMR.

4th week **Analysis of high resolution NMR spectra 2.** : Interpretation of homo-and heteronuclear NOE data. Basic types of ^{13}C NMR spectra: broadband ^1H -decoupled, j-modulated attached proton test, gated decoupling for heteronuclear couplings, and inverse-gated decoupling for quantitative ^{13}C NMR.

5th week **Practicing organic molecule structure elucidation by NMR 1.**: ^1H NMR: Major factors

influencing proton chemical shifts: electronegative substituents, neighboring anisotropic shielding, H-bonds. Acids, aldehydes, aromatics, alkenes, aliphatics. Analyzing aromatic ring substitution patterns. Alcohols, ketones.

6th week **Practicing organic molecule structure elucidation by NMR 2.:** ¹³C NMR: Signal multiplicities in uncoupled spectra. Predicting the number of carbons from decoupled spectra. The carbon NMR chemical shift correlation chart. Assigning the ¹³C NMR spectra of aromatics, alcohols, ketones and aliphatics. Interpreting signal intensities in usual, decoupled and in "quantitative" ¹³C NMR.

7th week **NMR written TEST**

8th week Electromagnetic radiation, ranges and energy of electromagnetic radiation. Conditions for generating infrared spectra. Rotational and vibrational spectra. Characteristic group frequencies, characteristic vibrational frequencies. Overtone frequencies. Typical ranges of chemical vibrations and their dependence on binding energy and binding stability.

9th week IR spectra of alkanes, alkenes, alkynes and aromatic compounds. Alcohol identification, the effect of hydrogen bond on the IR spectrum of alcohols. Intra- and intermolecular effects affecting the C-O vibration of the carbonyl group. IR spectra of carboxylic acids and carboxylic acid derivatives.

10th week Absorption spectra (UV, IR, Raman) of molecules. The Beer-Lambert Law and its Analytical Applications. Electron excitation transitions. Maximum places and ϵ values of the UV transitions of chromophores. Selection rules. The Jablonski diagram. Frank-Condon principle, bathochromic, hypsochromic, hypochromic and hyperchromic shifts. The effect of conjugation, steric hindrance on chromophores.

11th week Conformation and geometry of polyene systems. Effect of solvent polarity on UV spectra. The basic concepts of mass spectrometry. The main ionization techniques of organic mass spectrometry. Ionization of molecules.

12th week General fragmentation and fragmentation of the molecular ion: mass spectrum. The advantages and disadvantages of ionization methods. Main parts of mass spectrometer. Optimal technical requirements for sample input aspects and multicomponent samples. Ion sources, EI ion source, CI ion source.

13th week Molecular ionization: ESI ion source, APCI ion source. The types of Mass analyzers. The Resolution. Signal Processing: detectors.

14th week Basic concepts of organic mass spectrometry, mol peak, molecular ion. The nitrogen rule, natural isotopes. General aspects of the interpretation of mass spectra. Main fragmentation processes: α -, benzyl, allyl cleavage. The McLafferty rearrangement. Generic mass spectrometry of different class of organic compounds

15th week **MS & IR written TEST**

Requirements:

- *for a signature*

Attendance at **lectures** is highly recommended (not compulsory) since interactive evaluation of test problems are parts of the lectures.

Students have to **submit all the two designing tasks** as scheduled minimum on a sufficient level.

During the semester there are two tests: the mid-term test in the 7th week and the end-term test in the 14th week. Students have to sit for the tests

- *for a grade*

The course ends in an **examination**. Based on the average of the grades of the designing tasks and the examination, the exam grade is calculated as an average of them:

- the average grade of the two designing tasks
- or the result of the oral examination

The minimum requirement for the mid-term and end-term tests and the examination respectively is

60%. Based on the score of the tests separately, the grade for the tests and the examination is given according to the following table:

Score	Grade
0-59	fail (1)
60-69	pass (2)
70-79	satisfactory (3)
80-89	good (4)
90-100	excellent (5)

If the score of any test is below 60, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

-an offered grade:

it may be offered for students if the average grade of the two designing tasks is at least satisfactory (3) and the average of the mid-term and end-term tests is at least satisfactory (3). The offered grade is the average of them.

Person responsible for course: Dr. Gyula Batta, professor, PhD

Lecturers: Dr. Gyula Batta, professor, PhD & Dr. Attila Kiss PhD, associate professor

Title of course: Spectroscopy methods II. Code: TTKBL0504_EN	ECTS Credit points: 4
Type of teaching, contact hours - lecture: - - practice: 3 hours/week - laboratory: -	
Evaluation: term mark	
Workload (estimated), divided into contact hours: - lecture: - practice: 42 hours - laboratory: - - home assignment: 20 hours - preparation for the exam: Total: 62	
Year, semester: 3 rd year, 2 nd semester	
Its prerequisite(s): TTKBE0503	
Further courses built on it: -	

Topics of course
- Spectroscopic characterization of alkanes, alkenes and alkynes. - Spectroscopic characterization of aromatic compounds. - Spectroscopic characterization of halogen-containing compounds. - Spectroscopic characterization of alcohols, phenols and ethers. - Spectroscopic characterization of amino, nitro and azide derivatives - Spectroscopic characterization of oxo compounds. - Spectroscopic characterization of sulfur-containing compounds. - Spectroscopic characterization of oxo compounds.

- Spectroscopic characterization of carboxylic acids and carboxylic acid derivatives.

Literature

Compulsory:

1. L D Field, S Sternhell, J R Kalman, Organic Structures from Spectra, 5th edition, Wiley, 2013
2. E. Pretsch, P. Bühlmann M. Badertscher, Structure Determination of Organic Compounds; 4th edition, Springer-Verlag, 2009
3. R. M. Silverstein, F. X. Webster, D. J. Kiemle, D. L. Bryce, Spectrometric Identification of Organic Compounds, 8th edition, Wiley, 2014

Course objective/intended learning outcomes

a) Knowledge

He/she knows the most important information of basic spectrum (UV, IR, MS, NMR) which can be using for structure determination of organic compounds. He/she is able to use the different spectroscopic methods collectively to solve simple spectroscopic problems and to elucidate the structure of simple compounds.

b) Abilities

- He/she is able to identify and interpret the basic information of the spectra in the structure elucidation.

- He is able to participate in professional communication on the field of structure determination of organic compounds.

-He is able to expand and/or develop his/her knowledge of spectroscopy.

c) Attitude

- He/she is open to getting new, scientifically proven knowledge on the subject, but to reject unsubstantiated or possibly misleading claims.

d) Autonomy and responsibility

- He/she is able to independently perform the tasks of the course with professional guidance, and he/she can interpret and evaluate of the results obtained.

Schedule:

1st week

Repeat the theoretical background of spectroscopic methods and review the general information content of each spectrum.

2nd week

Spectroscopic characterization of alkanes, alkenes and alkynes, identification of their constitutional and configuration isomers.

3rd week

Spectroscopic characterization of aromatic compounds and identification of their substitution pattern.

4th week

Spectroscopic characterization of halogenated derivatives.

5th week

Spectroscopic characterization of alcohols, phenols and ether derivatives.

6th week

Spectroscopic characterization of amino, nitro and azido derivatives.

7th week

Mid-term test.

8th week

Spectroscopic characterization of oxo compounds.

9th week

Spectroscopic characterization of thiols, sulphoxides, sulfones and sulfonic acids.

10th week

Spectroscopic characterization of carboxylic acids and esters.

11th week

Spectroscopic characterization of carboxylic amides and acid anhydrides and substituted carboxylic acids.

12th week

Complex structure identification of organic compounds I.

13th week

Complex structure identification of organic compounds II.

14th week

End-term test.

Requirements:

- *for a signature*

Attendance at **practice** is compulsory.

A student must attend the practice classes and may not miss more than two times during the semester (but not at the week of 7th and 14th), and a medical certificate needs to be presented. In case a student does so, the subject will not be signed, and the student must repeat the course. A student can't make up any practice with another group. Being late is equivalent with an absence. If the student does not write the tests (at the week of 7th and 14th), they must to write at the 1st week of the exam session otherwise fail the course.

- *for a grade*

The term mark is based on the average of the grades of written tests.

The minimum requirement for the mid-term and end-term tests respectively is 50%. Based on the score of the tests separately, the grade for the tests is given according to the following table:

Score	Grade
0-49	fail (1)
50-65	pass (2)
66-80	satisfactory (3)
80-89	good (4)
90-100	excellent (5)

If the score of any test is below 50, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

Person responsible for course: László Dr. Juhász, associate professor, PhD, dr. habil

Lecturer: László Dr. Juhász, associate professor, PhD, dr. habil, Tünde Zita Dr. Illyés, senior lecturer, PhD

Title of course: Chemical Technology I.

Code: TTKBE0601_EN

ECTS Credit points: 3

Type of teaching, contact hours

- lecture: 2 hours/week

- practice: -

- laboratory: -
Evaluation: exam
Workload (estimated), divided into contact hours: - lecture: 28 hours - practice: - - laboratory: - - preparation for the exam: 60 hours Total: 88 hours
Year, semester: 2 nd year, 2 nd semester
Its prerequisite(s): TTKBE0201_EN, TTKBE0301_EN, TTKBE0401_EN
Further courses built on it: TTKBE0602_EN, TTKBG0602_EN, Parallel registration to TTKBX0607_EN

Topics of course
Theory and equipments of basic unit operations. Transportation of fluids, mixing, filtering, fluidization. Heat transfer, mass transfer. Reactors.
Literature
<i>Recommended:</i> 1. Warren McCabe, Julian Smith, Peter Harriott, Unit Operations of Chemical Engineering, McGraw Hill, New York, 2005. 2. Peter Atkins, Physical Chemistry, Oxford University Press, Oxford, 2009. 3. Louis Theodore, R. Ryan Dupont, Kumar Ganesan, Unit Operations in Environmental Engineering, Scrivener Publishing LLC, Beverly, MA, 2017
Course objective/intended learning outcomes
a) Knowledge - He/She has knowledge to solve problems on the field of natural processes, using natural sources, and understanding the chemical background of living and non-living systems. - He/She understands the progress and future trends in chemistry and chemical industries.
b) Abilities - Using his/her chemical knowledge and experiences He/She is able to create and execute experiments to demonstrate and prove the basic chemical phenomena.
c) Attitude - He/She is ready to discuss problems on the field of chemistry and other science with professionals. - He/She is able to collaborate with other groups and capable to get insight to the chemical aspects of economy and environmental safety.
d) Autonomy and responsibility - He/She is able to operate the most important technological instruments of chemical industries. - He/She stands for his/her opinion or ideology in professional discussions. - Under supervision he/she is responsible in collaboration with other professionals (especially from the field of environmental economy and safety)

Schedule:*1st week*

Classification of unit operations

2nd week

Hydrostatics, hydrodynamics, Navier-Stokes equation, Pascal-rule. Industrial applications.

3rd week

Fluid flow types, Bernoulli equation, principle, pressure drop in pipes.

4th week

Pumps in chemical engineering, cavitation, pump characteristics.

5th week

Mixing, filtering

6th week

Fluidization

7th week

Heat transfer, Fourier equations, heat exchangers.

8th week

Vapor-liquid equilibrium, batch distillation.

9th week

Flash distillation.

10th week

Multi stage distillation, McCabe, Thiele method.

11th week

Extraction

12th week

Absorption

13th week

Adsorption

14th week

Chemical reactors

Requirements:

Attendance at lectures is recommended, but not compulsory.

The course ends with an exam. The minimum requirement for the exam is 50%. The grade is given according to the following table:

Score	Grade
0-49	fail (1)
50-59	pass (2)
60-74	satisfactory (3)
75-89	good (4)
90-100	excellent (5)

The students can retake the exam in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

Person responsible for course: Dr. Lajos Nagy, associate professor, PhD**Lecturer:** Dr. Ákos Kuki, associate professor, PhD

Title of course: Chemical Technology I. Code: TTKBG0601_EN	ECTS Credit points: 1
Type of teaching, contact hours - lecture: - - practice: 1 hour/week - laboratory: -	
Evaluation: practice grade	
Workload (estimated), divided into contact hours: - lecture: - - practice: 14 hours - laboratory: - - preparation for the exam: 16 hours Total: 30 hours	
Year, semester: 2 nd year, 2 nd semester	
Its prerequisite(s): TTKBE0201_EN, TTKBE0301_EN, TTKBE0401_EN	
Further courses built on it: TTKBE0602_EN, TTKBG0602_EN, Parallel registration to TTKBX0607_EN	

Topics of course
Calculations of basic unit operations. Transportation of fluids, mixing, filtering, fluidization. Heat transfer, mass transfer. Reactors.
Literature
<i>Recommended:</i> 1. Warren McCabe, Julian Smith, Peter Harriott, Unit Operations of Chemical Engineering, McGraw Hill, New York, 2005. 2. Peter Atkins, Physical Chemistry, Oxford University Press, Oxford, 2009. 3. Louis Theodore, R. Ryan Dupont, Kumar Ganesan, Unit Operations in Environmental Engineering, Scrivener Publishing LLC, Beverly, MA, 2017
Course objective/intended learning outcomes
a) Knowledge - He/She has knowledge to solve problems on the field of natural processes, using natural sources, and understanding the chemical background of living and non-living systems. - He/She understands the progress and future trends in chemistry and chemical industries.
b) Abilities - Using his/her chemical knowledge and experiences He/She is able to create and execute experiments to demonstrate and prove the basic chemical phenomena.
c) Attitude

- He/She is ready to discuss problems on the field of chemistry and other science with professionals.

- He/She is able to collaborate with other groups and capable to get insight to the chemical aspects of economy and environmental safety.

d) Autonomy and responsibility

- He/She is able to operate the most important technological instruments of chemical industries.

- He/She stands for his/her opinion or ideology in professional discussions.

- Under supervision he/she is responsible in collaboration with other professionals (especially from the field of environmental economy and safety)

Schedule:

1st week

Calculations in the field of hydrostatics, Pascal-rule.

2nd week

Fluid flow types, Reynolds number.

3rd week

Bernoulli equation

4th week

Pressure drop in pipes.

5th week

Pump power calculations

6th week

Pump duty point calculation

7th week

Calculations in the field of filtering, mixing.

8th week

Calculations in the field of fluidization.

9th week

Calculation of heat exchangers

10th week

Calculation of *flash* distillation.

11th week

Distillation, operating lines, *Short-cut* method.

12th week

Distillation, McCabe, Thiele method.

13th week

Calculations in the field of extraction.

14th week

Calculations in the field of absorption and adsorption.

Requirements:

- *for a signature*

Participation at the classes is compulsory. A student must attend the classes and may not miss more than three times during the semester. In case a student does so, the subject will not be signed and the student must repeat the course. Attendance at the classes will be recorded by the practice leader. Being late is equivalent with an absence. In case of further absences, a medical certificate needs to

be presented.

- *for the practice grade*

The course ends with a test in the 14th week. The minimum requirement for the test is 50%. The grade is given according to the following table:

Score	Grade
0-49	fail (1)
50-59	pass (2)
60-74	satisfactory (3)
75-89	good (4)
90-100	excellent (5)

The students are allowed to retake the test once to improve their scores. Further improvement is in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

Person responsible for course: Dr. Lajos Nagy, associate professor, PhD

Lecturer: Dr. Ákos Kuki, associate professor, PhD

Title of course: Chemical Technology II. Code: TTKBE0602_EN	ECTS Credit points: 4
Type of teaching, contact hours - lecture: 3 hours/week - practice: - - laboratory: -	
Evaluation: exam	
Workload (estimated), divided into contact hours: - lecture: 42 hours - practice: - - laboratory: - - home assignment: 38 hours - preparation for the exam: 40 hours Total: 120 hours	
Year, semester: 3 rd year, 1 st semester	
Its prerequisite(s): TTKBE0201_EN, TTKBE0301_EN, TTKBE0401_EN	
Further courses built on it: -	

Topics of course

Basic terms of chemical technology: continuous and batch processing, yield, conversion, efficiency, volume, basic laws of chemical technology. Combustion: burning and combustors. Water processing: production of drinking and process waters, wastewater, wastewater management. Nitrogen industries: synthesis of ammonia and nitric acid. Sulfur industries: production of sulfuric acid. Fertilizers. Electrolysis of brine. Production of alumina, iron and steel. Crude oil and natural gas: genesis (organic and inorganic theories), types, ingredients, mining. Engine fuels, destructive methods (thermic-, catalytic- and hydrocrackig), reforming of gasoline.

Literature

Compulsory:

- Ullmann's Encyclopedia of Industrial Chemistry, Wiley-VCH Verlag GmbH & Co. KGaA., 2002, ISBN: 9783527306732
- J.M. Coulson, J.F. Richardson and R.K. Sinnott: Chemical Engineering, Volume 6., Pergamon Press, 1983.
- G N Pandey: Textbook of Chemical Technology Vol-1, 2, 2006.

Recommended:

- Muhlynov I.: Chemical Technology I-II.

Course objective/intended learning outcomes

a) Knowledge

- He/She has the knowledge to test or measure chemical reactions, systems with scientific methods (including computational) under supervision.
- He/She understands the progress and future trends in chemistry and chemical industries.

b) Abilities

- He/She is able to use the obtained chemical knowledge on the field of chemical, industrial, environmental sciences to solve the actual basic problems and prove the obtained solution.

c) Attitude

- He/She is able to represent his/her own personal scientific ideology toward professional and unprofessional groups.

d) Autonomy and responsibility

- He/She shows responsibility to gain enough experience to be involved in making decision for labs or industrial plants.

Schedule:

1st week

Laws and description of Chemical Technology

2nd week

Purification of water, water treatment

3rd week

Water softening, hardness scales

4th week

Nitrogen industry, steam processing

5th week

Synthesis of ammonia

6th week

Nitric acid production, nitrogen containing fertilizers

7th week

Sulphur industry, sulphuric acid production

8th week

Superphosphate production

9th week

Brine electrolysis, products

10th week

Alumina industry, electrolysis of alumina

11th week

Manufacturing iron, processes in the blast furnace

12th week

Atmospheric distillation of natural oil

13th week

Vacuum distillation of atmospheric residue

14th week

Processing of natural gas

Requirements:

- *for a signature*

Attendance at **lectures** is recommended, but not compulsory.

During the semester there is one test: the end-term test in the 15th week. Students have to sit for the test

- *for a grade*

The exam grade is calculated by the result of end-term test.

The minimum requirement for end-term test is 50%. Based on the score of the test separately, the grade for the test is given according to the following table:

Score	Grade
0-49	fail (1)
50-59	pass (2)
60-74	satisfactory (3)
75-89	good (4)
90-100	excellent (5)

If the score of the test is below 50, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

- *an offered grade:*

It may be offered for students if the grade is at least pass (2).

Person responsible for course: Dr. Lajos Nagy, associate professor, PhD

Lecturer: Dr. Lajos Nagy, associate professor, PhD

Title of course: Chemical Technology II.

Code: TTKBG0602_EN

ECTS Credit points: 2

Type of teaching, contact hours

- lecture: -
- practice: 2 hours/week
- laboratory: -

Evaluation: practice grade

Workload (estimated), divided into contact hours:

- lecture: -
- practice: 28 hours
- laboratory: -
- home assignment: 12 hours
- preparation for the exam: 20 hours

Total: 60 hours
Year, semester: 3 rd year, 1 st semester
Its prerequisite(s): TTKBE0201_EN, TTKBE0301_EN, TTKBE0401_EN
Further courses built on it: -

Topics of course
Basic terms of chemical technology: continuous and batch processing, yield, conversion, efficiency, volume, basic laws of chemical technology. Combustion: burning and combustors. Water processing: production of drinking and process waters, wastewater, wastewater management. Nitrogen industries: synthesis of ammonia and nitric acid. Sulfur industries: production of sulfuric acid. Fertilizers. Electrolysis of brine. Production of alumina, iron and steel. Crude oil and natural gas: genesis (organic and inorganic theories), types, ingredients, mining. Engine fuels, destructive methods (thermic-, catalytic- and hydrocrackig), reforming of gasoline.
Literature
<i>Compulsory:</i> - Ullmann's Encyclopedia of Industrial Chemistry, Wiley-VCH Verlag GmbH & Co. KGaA., 2002, ISBN: 9783527306732 - J.M. Coulson, J.F. Richardson and R.K. Sinnott: Chemical Engineering, Volume 6., Pergamon Press, 1983. - G N Pandey: Textbook of Chemical Technology Vol-1, 2, 2006. <i>Recommended:</i> - Muhlynov I.: Chemical Technology I-II.
Course objective/intended learning outcomes
a) Knowledge - He/She has the knowledge to test or measure chemical reactions, systems with scientific methods (including computational) under supervision. - He/She understands the progress and future trends in chemistry and chemical industries.
b) Abilities - He/She is able to use the obtained chemical knowledge on the field of chemical, industrial, environmental sciences to solve the actual basic problems and prove the obtained solution.
c) Attitude - He/She is able to represent his/her own personal scientific ideology toward professional and unprofessional groups.
d) Autonomy and responsibility - He/She shows responsibility to gain enough experience to be involved in making decision for labs or industrial plants.

Schedule:
<i>1st week</i> Laws and description of Chemical Technology
<i>2nd week</i> Purification of water, water treatment
<i>3rd week</i>

Water softening, hardness scales

4th week

Nitrogen industry, steam processing

5th week

Synthesis of ammonia

6th week

Nitric acid production, nitrogen containing fertilizers

7th week

Sulphur industry, sulphuric acid production

8th week

Superphosphate production

9th week

Brine electrolysis, products

10th week

Alumina industry, electrolysis of alumina

11th week

Manufacturing iron, processes in the blast furnace

12th week

Atmospheric distillation of natural oil

13th week

Vacuum distillation of atmospheric residue

14th week

Processing of natural gas

Requirements:

- for a signature

Participation at **practice classes** is compulsory. A student must attend the practice classes and may not miss more than three times during the semester. In case a student does so, the subject will not be signed and the student must repeat the course. Attendance at practice classes will be recorded by the practice leader. Being late is equivalent with an absence. In case of further absences, a medical certificate needs to be presented. Missed practice classes should be made up for at a later date, to be discussed with the tutor. Active participation is evaluated by the teacher in every class. If a student's behaviour or conduct doesn't meet the requirements of active participation, the teacher may evaluate his/her participation as an absence because of the lack of active participation in class. During the semester there is one test: the end-term test in the 15th week. Students have to sit for the test

- for a grade

The exam grade is calculated by the result of end-term test.

The minimum requirement for end-term test is 50%. Based on the score of the test separately, the grade for the test is given according to the following table:

Score	Grade
0-49	fail (1)
50-59	pass (2)
60-74	satisfactory (3)
75-89	good (4)
90-100	excellent (5)

If the score of the test is below 50, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

-an offered grade:

It may be offered for students if the grade is at least pass (2).

Person responsible for course: Dr. Lajos Nagy, associate professor, PhD

Lecturer: Dr. Lajos Nagy, associate professor, PhD

Title of course: Macromolecular Chemistry

Code: TTKBE0611_EN

ECTS Credit points: 3

Type of teaching, contact hours

- lecture: 2 hours/week
- practice: -
- laboratory: -

Evaluation: exam

Workload (estimated), divided into contact hours:

- lecture: 28 hours
- practice:-
- laboratory: -
- home assignment: 12 hours
- preparation for the exam: 50 hours

Total: 90 hours

Year, semester: 3rd year, 2nd semester

Its prerequisite(s): TTKBE0302_EN

Further courses built on it: TTKBE1213_EN

Topics of course

Principal definitions. Classification of polymers. The most important synthetic polymers. Methods for characterizing polymers. Polymolecularity. Correlation between the structure and properties of polymers. Physical states of polymers. Preparation methods of synthetic polymers and copolymers: radical polymerization and copolymerization, anionic, cationic, living cationic polymerization. Step polymerization: polycondensation and polyaddition.

Literature

Compulsory:

- George Odian: Principles of Polymerization (Wiley, 2004) ISBN: 978-0-471-27400-1
- Leslie H. Sperling: Introduction to Physical Polymer Science (Wiley, 2006) ISBN: 978-0-471-70606-9

Recommended:

- Krzysztof Matyjaszewski, Thomas P. Davis: Handbook of Radical Polymerization (Wiley, 2002) ISBN: 978-0-471-39274-3

Course objective/intended learning outcomes

a) Knowledge

- He/She knows the properties of the most important chemicals, their productions and applications.
- He/She has a mathematical and scientific background to understand processes in chemical and chemistry related industries.

b) Abilities

- He/She capable to apply the learned methods, models and plannings of chemical technology and chemical processes through calculations.
- He/She understands and able to describe the elements of industrial and technological units, their operations including the connectivity options.

c) Attitude

- He/She is open to accept environmentally efficient technologies, and for the application of new, innovative and advanced methods in economy.
- He/She makes effort to improve and apply the practical methods with new results and experiences.

d) Autonomy and responsibility

- Following directions he/she can work without supervision considering all quality and safety rules.
- He/She shares experiences with others to help them.

Schedule:

1st week

Principal definitions. Classification of polymers.

2nd week

Chemical structure, shape and fine structure of polymers.

3rd week

Polymolecularity. Average molecular weights, molecular weight distribution.

4th week

Determination methods for the molecular weight of polymers.

5th week

Physical states of polymers, I.: glass transition temperature, description of amorphous polymers.

6th week

Physical states of polymers, II.: crystallinity of polymers.

7th week

Synthesis of polymers: Radical polymerization I.

8th week

Synthesis of polymers: Radical polymerization II.

9th week

Synthesis of polymers: Types of copolymers, radical copolymerization.

10th week

Synthesis of polymers: Cationic, living cationic polymerization.

11th week

Synthesis of polymers: Anionic polymerization.

12th week

Synthesis of polymers: Coordination polymerization.

13th week

Synthesis of polymers: Step polymerization I.: Polycondensation.

14th week

Synthesis of polymers: Step polymerization II.: Polyaddition.

Requirements:

- for a signature

Attendance at **lectures** is recommended, but not compulsory. Active participation is rewarded by the teacher in every class.

During the semester there is one end-term test in the 15th week for an offered grade (optional). Students have to sit for the tests.

- for a grade

The course ends in an **examination**.

The minimum requirement for the end-term test and the examination respectively is 50%. Based on the score of the test, the grade for the test and the examination is given according to the following table:

<u>Score</u>	<u>Grade</u>
0-49	fail (1)
50-61	pass (2)
62-74	satisfactory (3)
75-87	good (4)
88-100	excellent (5)

If the score the test is below 50, students can take a retake test in conformity with the Education and Examination Rules and Regulations.

-an offered grade:

it may be offered for students if the grade of the end-term test is at least satisfactory (3).

Person responsible for course: Prof. Sándor Kéki, full professor, PhD, DSc

Lecturer: Prof. Sándor Kéki, full professor, PhD, DSc

Title of course: Macromolecular Chemistry

Code: TTKBG0611_EN

ECTS Credit points: 1

Type of teaching, contact hours

- lecture: -
- practice: 1 hour/week
- laboratory: -

Evaluation: practice grade

Workload (estimated), divided into contact hours:

- lecture: -
- practice: 14 hours
- laboratory: -
- home assignment: 6 hours
- preparation for the exam: 10 hours

Total: 30 hours

Year, semester: 3rd year, 2nd semester

Its prerequisite(s): TTKBE0302_EN

Further courses built on it: TTKBE1213_EN

Topics of course

Review of the principal definitions, the classification of polymers. Overview of the most important synthetic polymers, methods for characterizing polymers. Calculation of polymolecularity. Examples for the correlation between the structure and properties of polymers, the physical states of polymers. Overview of the preparation methods of synthetic polymers and copolymers: radical polymerization and copolymerization, anionic, cationic, living cationic polymerization. Step polymerization: polycondensation and polyaddition. Calculation examples for different polymerization methods.

Literature

Compulsory:

- George Odian: Principles of Polymerization (Wiley, 2004) ISBN: 978-0-471-27400-1
- Leslie H. Sperling: Introduction to Physical Polymer Science (Wiley, 2006) ISBN: 978-0-471-70606-9

Recommended:

- Krzysztof Matyjaszewski, Thomas P. Davis: Handbook of Radical Polymerization (Wiley, 2002) ISBN: 978-0-471-39274-3

Course objective/intended learning outcomes

a) Knowledge

- He/She knows the properties of the most important chemicals, their productions and applications.
- He/She has a mathematical and scientific background to understand processes in chemical and chemistry related industries.

b) Abilities

- He/She capable to apply the learned methods, models and plannings of chemical technology and chemical processes through calculations.
- He/She understands and able to describe the elements of industrial and technological units, their operations including the connectivity options.

c) Attitude

- He/She is open to accept environmentally efficient technologies, and for the application of new, innovative and advanced methods in economy.
- He/She makes effort to improve and apply the practical methods with new results and experiences.

d) Autonomy and responsibility

- Following directions he/she can work without supervision considering all quality and safety rules.
- He/She shares experiences with others to help them.

Schedule:

1st week

Review of the principal definitions, the classification of polymers.

2nd week

Chemical structure, shape and fine structure of polymers.

3rd week

Calculation of polymolecularity, average molecular weights, molecular weight distribution.

4th week

Determination methods for the molecular weight of polymers.

5th week

Discussion of the physical states of polymers, I.: glass transition temperature, description of amorphous polymers.

6th week

Discussion of the physical states of polymers, II.: crystallinity of polymers.

7th week

Synthesis of polymers: Radical polymerization I. Discussion, examples.

8th week

Synthesis of polymers: Radical polymerization II. Discussion, examples, calculation.

9th week

Synthesis of polymers: Types of copolymers, radical copolymerization. Determination of the reactivity ratio, distribution of monomers in a copolymer.

10th week

Synthesis of polymers: Cationic, living cationic polymerization. Discussion, examples.

11th week

Synthesis of polymers: Anionic polymerization. Discussion, examples.

12th week

Synthesis of polymers: Coordination polymerization. Discussion, examples.

13th week

Synthesis of polymers: Step polymerization I.: Polycondensation. Discussion, examples, calculation.

14th week

Synthesis of polymers: Step polymerization II.: Polyaddition. Discussion, examples.

Requirements:

- for a signature

Participation at **practice classes** is compulsory. Active participation is rewarded by the teacher in every class. A student must attend the practice classes and may not miss more than three times during the semester. In case a student does so, the subject will not be signed and the student must repeat the course. Attendance at practice classes will be recorded by the practice leader. Being late is equivalent with an absence. In case of further absences, a medical certificate needs to be presented. Missed practice classes should be made up for at a later date, to be discussed with the tutor.

During the semester there is one end-term test in the 15th week for a practice grade. Students have to sit for the tests.

- for a grade

The course ends in an **end-term test**.

The minimum requirement for the end-term test is 50%. Based on the score of the test, the grade for

the test is given according to the following table:

<u>Score</u>	<u>Grade</u>
0-49	fail (1)
50-61	pass (2)
62-74	satisfactory (3)
75-87	good (4)
88-100	excellent (5)

If the score the test is below 50, students can take a retake test in conformity with the Education and Examination Rules and Regulations.

-an offered grade:

it may be offered for students if the grade of the end-term test is at least satisfactory (3).

Person responsible for course: Prof. Sándor Kéki, full professor, PhD, DSc

Lecturer: Prof. Sándor Kéki, full professor, PhD, DSc

Title of course: Environmental Technology Code: TTKBE1114_EN	ECTS Credit points: 3
Type of teaching, contact hours - lecture: 2 hours/week - practice: - - laboratory: -	
Evaluation: exam	
Workload (estimated), divided into contact hours: - lecture: 28 hours - practice: - - laboratory: - - home assignment: 32 hours - preparation for the exam: 30 hours Total: 90 hours	
Year, semester: 3 rd year, 2 nd semester	
Its prerequisite(s): TTKBE1111_EN	
Further courses built on it: -	

Topics of course

Relationship between the humanity and the nature. Sustainable development. Types of municipal and industrial wastes, prevention of their formation. Basics of waste management: landfilling, incineration and other physical and chemical methods. Additive and integrated environmental protection strategies. Treatment technologies of wastes at different states. Pollutants of air, water, and soil, their treatment. Municipal and industrial wastewater treatment. Noise and vibration protection. Renewable energy sources.

Literature

Compulsory:

- D.A. Vallero: Fundamentals of Air Pollution (Academic Press, 2007) ISBN: 9780123736154
- N.L. Nemerow: Industrial Waste Treatment (Butterworth-Heinemann, 2007) ISBN: 9780123724939

Recommended:

- A. Malik, E. Grohmann: Environmental Protection Strategies for Sustainable Development (Springer, 2011), ISBN: 9789400715912

Course objective/intended learning outcomes

a) Knowledge

- He/She knows the basic principles, the planning and controlling options in technology of chemical processes and industrial tasks.
- He/She knows the chemistry and chemical technology related economical, management environmental safety, quality assurance (QC/QA), informatics and intellectual property rules and laws.

b) Abilities

- He/She is able to treat new or unknown system based on the previous studies and experiences, learn and install new technologies and recognize mechanisms related to human health.
- He/She is able to collect and organize, understand information about the health prevention, keeping track on new results, and apply them to make cost and environmentally effective, healthy working areas.

c) Attitude

- He/She is open to accept environmentally efficient technologies, and for the application of new, innovative and advanced methods in economy.
- During everyday work and installation of new technologies he/she always concerned about sustainable development.

d) Autonomy and responsibility

- He/She tends to establish new solutions and technologies.
- Following directions he/she can work without supervision considering all quality and safety rules.

Schedule:

1st week

Relationship between the humanity and the nature. Sustainable development in the industry.

2nd week

Types of industrial wastes, reduction and treatment.

3rd week

Additive and integrated environmental protection strategies. Closed-loop technologies.

4th week

Technologies and methods for the treatment of gaseous wastes.

5th week

Technologies and methods for the treatment of liquid wastes.

6th week

Technologies and methods for the treatment of solid wastes.

7th week

Air pollutants, their effects, prevention, treatment options.

8th week

Water pollutant chemical substances, their effects on the hydrosphere, prevention, treatment options.

9th week

Organic substances as water pollutants, their analytical problems, effects on the living organisms.

10th week

Physical and chemical methods of wastewater treatment.

11th week

Biological methods of wastewater treatment. Sludge treatment.

12th week

Soil pollution, treatment options of different pollutants.

13th week

Renewable energy sources: solar, wind, water, geothermal.

14th week

Noise and vibration protection. Effects of noise on the environment and human health.

Requirements:

- for a signature

Attendance at **lectures** is recommended, but not compulsory. Active participation is rewarded by the teacher in every class.

Students have to **submit an essay** about a given topic as scheduled minimum on a sufficient level.

During the semester there is one end-term test in the 15th week for an offered grade (optional).

Students have to sit for the tests.

- for a grade

The course ends in an **examination**. Based on the average of the grades of the essay and the examination, the exam grade is calculated as an average of them:

The minimum requirement for the end-term test and the examination respectively is 50%. Based on the score of the test, the grade for the test and the examination is given according to the following table:

<u>Score</u>	<u>Grade</u>
0-49	fail (1)
50-61	pass (2)
62-74	satisfactory (3)
75-87	good (4)
88-100	excellent (5)

If the score the test is below 50, students can take a retake test in conformity with the Education and Examination Rules and Regulations.

-an offered grade:

it may be offered for students if the grade of both the essay and the end-term test is at least satisfactory (3). The offered grade is the average of them.

Person responsible for course: Dr. Dávid Rácz, assistant professor, PhD

Lecturer: Dr. Dávid Rácz, assistant professor, PhD

Title of course: Visit in Chemical Industries Code: TTKBX0607_EN	ECTS Credit points: 0
Type of teaching, contact hours 5 days	
Evaluation: signature	
Workload (estimated), divided into contact hours: Total: 40 hours	
Year, semester: 2 nd year, 2 nd semester	
Its prerequisite(s): Parallel registration to TTKBE0601_EN	
Further courses built on it: -	

Topics of course
5 days visit at various factories and companies in the chemical engineering field. The technology processes and equipments are explained to the students by experts. Visited companies are in the field: plastic industry, oil industry, pharmaceuticals, water treatment, waste treatment, power plants, etc.
Literature
-
Course objective/intended learning outcomes
<p>a) Knowledge</p> <ul style="list-style-type: none"> - He/She knows and can use the properties of most important chemicals, lab equipment and basic laboratory processes. He/She can follow and recognize the requirements of safety instructions. - He/She understands the progress and future trends in chemistry and chemical industries. <p>b) Abilities</p> <ul style="list-style-type: none"> - He/She is able to collect and evaluate data on the field of chemistry in order to opening for problems on social, scientific or ethical questions. <p>c) Attitude</p> <ul style="list-style-type: none"> - He/She is ready to discuss problems on the field of chemistry and other science with professionals. - He/She is able to collaborate with other groups and capable to get insight to the chemical aspects of economy and environmental safety. - He/She is committed learn or get insights into new competence or ideology. - He/She is committed learn or get insights into new competence or ideology. <p>d) Autonomy and responsibility</p> <ul style="list-style-type: none"> - He/She is able to operate the most important technological instruments of chemical industries. - He/She stands for his/her opinion or ideology in professional discussions.

Requirements: Attendance at the visits is compulsory.
Person responsible for course: Dr. Ákos Kuki, associate professor, PhD

Title of course: Bachelor thesis I. Code: TTKBL0001_EN	ECTS Credit points: 5
Type of teaching, contact hours - lecture: - - practice: 2 hours/week - laboratory: 3 hours/week	
Evaluation: practical grade	
Workload (estimated), divided into contact hours: - lecture: - - practice: 28 - laboratory: 32 hours - home assignment: 90 hours - preparation for the exam: - Total: 150 hours	
Year, semester: 3 rd year, 1 st semester	
Its prerequisite(s): Completion of 110 credits + special requirements by the supervisor	
Further courses built on it: TTKBL0002	

Topics of course
The course aims at the preparation to solve a problem that can be approached by chemical methods. The student is expected to get the following competences: planning, time management, handling of information (gain and analysis from various sources, such as traditional library, electronic databases, search engines), ability to work alone or in a team, practical application of acquired knowledge, communication in native language both in oral and written ways. The student gets deeper knowledge in methods and procedures of a particular field of chemistry. With the help of the supervisor he/she starts to plan and execute the literature search and experimental work related to the topic of the thesis.
Literature
<i>Provided by the supervisor.</i>
Course objective/intended learning outcomes
a) Knowledge - He/She has knowledge to solve problems in the field of a particular field of chemistry - He/She can understand and communicate professionally on chemical subjects at his/her native language.
b) Abilities - He/She is able to use the previously learned paradigms, theories and laws in the field of chemical, natural science; to plan, execute and evaluate laboratory experiments. - He/She is able to evaluate and discuss the laboratory measurements, and create a report about it.
c) Attitude - He/She is committed learn or get insights into new competence or ideology. - He/She is well aware about his/her propositions and its consequences.
d) Autonomy and responsibility

- He/She takes part in scientific project(s) under supervision.

Schedule:

The student works by following the instructions of the supervisor.

Requirements:

- for a signature

The student have to take part in the research project coordinated by the supervisor.

- for a grade

The work of the student is evaluated by the supervisor considering many aspects, e.g. the quality of the work in the laboratory or industry, the ability to work alone or in a team, the competence for process the literature about the given topic, the problem solving ability and the presentation of the results.

Person responsible for course: Dr. László Somsák, full professor, responsible for the BSc in Chemistry studies

Lecturer: supervisors are staff members of the Institute of Chemistry, UD

Title of course: Bachelor thesis II.

Code: TTKBL0001_EN

ECTS Credit points: 10

Type of teaching, contact hours

- lecture: -

- practice: -

- laboratory: 10 hours/week

Evaluation: practical grade

Workload (estimated), divided into contact hours:

- lecture: -

- practice: -

- laboratory: 140 hours

- home assignment: 160 hours

- preparation for the exam: -

Total: 300 hours

Year, semester: 3rd year, 2nd semester

Its prerequisite(s): TTKBL0001_EN

Further courses built on it:

Topics of course

The student will complete the task of the previous semester by critically evaluating the literature, studying and applying the experimental method(s) to solve the given problem, carrying out the necessary practical work, and summarizing the results in a thesis of 20-30 printed pages. Special and detailed requirements are available at the homepage of the Institute of Chemistry. All these must be realized in a research group of the Institute of Chemistry. Work in an external group is only possible with the consent of the Institute of Chemistry and with a supervisor acknowledged by the

Institute of Chemistry.
Literature
<i>Provided by the supervisor.</i>
Course objective/intended learning outcomes
<p>a) Knowledge</p> <ul style="list-style-type: none"> - He/She has knowledge to solve problems in the field of a particular field of chemistry - He/She can understand and communicate professionally on chemical subjects in a foreign language. <p>b) Abilities</p> <ul style="list-style-type: none"> - He/She is able to use the previously learned paradigms, theories and laws in the field of chemical, natural science; to plan, execute and evaluate laboratory experiments. - He/She is able to evaluate and discuss the laboratory measurements, and create a report about it. <p>c) Attitude</p> <ul style="list-style-type: none"> - He/She is committed learn or get insights into new competence or ideology. - He/She is well aware about his/her propositions and its consequences. <p>d) Autonomy and responsibility</p> <ul style="list-style-type: none"> - He/She takes part in scientific project(s) under supervision.

<p>Schedule:</p> <p><i>The student works by following the instructions of the supervisor.</i></p>
<p>Requirements:</p> <p><i>- for a signature</i> The student have to take part in the research project coordinated by the supervisor.</p> <p><i>- for a grade</i> The work of the student is evaluated by the supervisor considering many aspects, e.g. the quality of the work in the laboratory or industry, the ability to work alone or in a team, the competence for process the literature about the given topic, the problem solving ability and the presentation of the results.</p>
<p>Person responsible for course: Dr. László Somsák, full professor, responsible for the BSc in Chemistry studies</p>
<p>Lecturer: supervisors are staff members of the Institute of Chemistry, UD</p>

Optional Courses

<p>Title of course: Crystallography</p> <p>Code: TTGBE5104_EN</p>	<p>ECTS Credit points: 2</p>
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<p>Type of teaching, contact hours</p> <ul style="list-style-type: none"> - lecture: 2 hours/week - practice:- - laboratory: -
<p>Evaluation: mid-term test, end-term test and written final exam</p>
<p>Workload (estimated), divided into contact hours:</p> <ul style="list-style-type: none"> - lecture: 28 hours - practice:- - laboratory: - - home assignment: 10 - preparation for the exam:30 <p>Total: 68</p>
<p>Year, semester: 1st year, 1st semester</p>
<p>Its prerequisite(s):</p>
<p>Further courses built on it:-</p>

<p>Topics of course</p> <p>Position of crystallography among other fields of science. The definition of space lattice, unit cell and crystallographic axes. Bravais lattices. Unit cells and crystallographic axes in crystal systems. Calculation of Miller indices. Symmetry elements, crystal classes, point groups and space groups. Fundamentals of crystal chemistry and the different types of lattices. Rules of coordination and packing. Lattice defects and element substitutions in the lattice. Physical properties of crystals and their explanation through structural differences.</p> <p>The understanding of constitution of unit cells and symmetry elements will be supported by the in-class study of three dimensional crystal models.</p>
<p>Literature</p> <p><i>Compulsory:</i></p> <p>W. D. Nesse: Introduction to Mineralogy. Oxford University Press. Oxford-New York, 2012 (2nd edition)</p> <p><i>Recommended:</i></p>
<p>Course objective/intended learning outcomes</p> <p>a) Knowledge</p> <ul style="list-style-type: none"> -knows the definition of space lattice, unit cell and crystal cross, the unit cells and crystal systems according to Bravais, - knows and able to identify the simple and combined symmetry elements and crystal forms, -knows the possible combination of the symmetry elements, the point groups and crystal classes, <ul style="list-style-type: none"> -knows the basics of crystal chemistry and the different types of lattices, -knows the most important mechanical, electrical, optical properties of crystals and their

connections to crystal structures.

b) Ability

- able to identify the different crystal systems, can give directions in crystallography, can calculate Miller indexes for lattice planes,
- able to identify the symmetry elements in macroscopic crystals, in crystal lattices and even in chemical molecules,
- able to apply the general rules of crystallography in structure research,
- able to interpret the connection between the crystal lattices and bond types in compounds,
- able to interpret the connection between the physical properties of crystals and their structures.

c) Attitude

- endeavour to completely understand the basic rules in crystallography,
- endeavour to understand the connection between inner structure of crystals and their macroscopic appearance,
- endeavour to understand and identify the symmetry elements,
- endeavour to understand the structure of crystal lattices and their effects on structure and physical/chemical properties of substances,
- endeavour to deeper understand the material structures with the use of gained knowledge in crystallography.

d) Autonomy and responsibility

- accept the scale of values of his/her profession with responsibility,
- cooperates with the experts of other fields of science during his/her work,
- understand the importance of crystallography, especially the symmetry in material structure research,
- able to individually process the scientific literature under the appropriate supervision.

Schedule:

1st week

Subject of crystallography. Properties of crystalline substances, definition of space lattice. Principles of morphology and crystallography.

2nd week

Bravais-type unit cells and crystal systems. Crystal cross in crystallography. Definition of directions, lattice planes and crystal faces. The Miller index.

3rd week

The visible symmetry elements of crystals, simple and combined symmetry elements. The stereographic projection. The translational symmetry.

4th week

Practicing of identification of symmetry elements

5th week

Point groups and the 32 crystal classes. Holohedral, hemihedral and tetrahedral crystal classes.

6th week

Mid-term test. Definition of crystal form. Crystal forms and symmetry elements in triclinic, monoclinic and orthorhombic systems.

7th week

Crystal forms and symmetry elements in trigonal, tetragonal and hexagonal crystal systems

8th week

Crystal forms and symmetry elements in cubic crystal system

9th week

Basics of crystal chemistry. X-ray diffraction and Bragg equation. Types of crystal lattices (atomic, ionic, metallic, molecular lattice). Coordination number, atomic, ionic radii.

10th week

Types of atomic lattices. Metallic lattice and the close packing. Molecular lattices. Properties of ionic lattice substances.

11th week

Isodesmic, anisodesmic and mesodesmic ionic lattices. Structure of silicates. Ortho, ring, chain, sheet and framework silicates.

12th week

Isomorphism and polymorphism. Real lattice structures, lattice defects. Rules of element substitutions. Crystal growth.

13th week

Crystal physics. Cohesion properties. Cleavage and sliding. Mohs-type hardness scale. Thermoelectric and piezoelectric properties. Structural interpretation of physical properties.

14th week

Crystal optics. Isotropic and anisotropic crystals. Birefringency and optical activity. Summary

Requirements:

- for a signature

Participation at **lecture classes** is not compulsory but highly advised.

During the semester there will be two tests, the mid-term test in week 6, and the end-term test in week 15. Students have to sit for the tests.

- for a grade

The course ends with a **writing examination** in the exam period, covering the whole material of the semester. The final grade for the course will be determined according to the followings: it is based

on the average grade of the mid-term test and end-term test in 10 %, and based on the result of written exam in 90 %.

The minimum requirement for the average grade of end-term test and mid-term test and final exam is 50%, respectively. The examination is given according to the following table:

Score	Grade
0-49	fail (1)
50-59	pass (2)
60-72	satisfactory (3)
73-87	good (4)
88-100	excellent (5)

If the score of the test is below 49, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

-an offered grade:

it may be offered for students if the average grade of mid-term test and end-term test is at least satisfactory (3).

Person responsible for course: Dr. Gábor Dobosi, professor, DSc

Lecturer: Dr. Dávid Nagy, assistant lecturer, PhD

Title of course: History of chemistry Code: TTKBE0007_EN	ECTS Credit points: 3
Type of teaching, contact hours - lecture: 2 hours/week - practice: 0 hours/week - laboratory: -	
Evaluation: exam	
Workload (estimated), divided into contact hours: - lecture: 30 hours - practice: 0 hours - laboratory: - - home assignment: 30 hours - preparation for the exam: 30 hours Total: 90 hours	
Year, semester: 3 nd year, 2 nd semester	
Its prerequisite(s): TTKBE0101_EN	
Further courses built on it: - The course is connected to other courses of chemistry teachers (Basics of chemistry teaching, Methods and devices of chemistry teaching)	

Topics of course

The topic of this course is the history of chemical thinking, the philosophical foundations of the science, the thinking systems and the history of discovery and inventions. It also concerns the impact of the development of chemistry on culture, history, the world view and the lifestyle of mankind. The lecture covers the ancient and medieval chemistry (metal processing, cleaning, cosmetics, chemistry of medicines) as well, namely, the age of alchemy. We analyse the conditions of the discovery of gases, the development of the interpretation of chemical reactions, the history of the formation of organic chemical concepts, the formation of a modern chemical industry and the age of modern atom theory, the age of electrochemistry and radiochemistry, and the history of medication development. The historical interpretations help to understand the complex relation between the chemistry and the human culture.

Literature

Compulsory:

- L. Balázs: *History of chemistry I-II. (1996), National Textbook Publisher (Budapest), 1996, p.1-1075. (editors: Oláh Zsuzsa, lector: I. Pais, E. Szilágyi)*

Recommended:

- K. Simonyi (1981): *Cultural history of physics, Publisher: "Gondolat Kiadó", Budapest*

- L. Kovács, D. Csupor, G. Lente, T. Gunda (2011): *100 chemical myths. Publisher: "Akadémiai Kiadó"*

Course objective/intended learning outcomes

a) Knowledge

He/She can critically evaluate and manage the historical contexts and laws of natural processes, natural resources, living and inorganic systems in a historical dimension. He/She is aware of concepts and terminologies that characterize natural processes in different historical periods. He knows the basic qualitative and quantitative relationships and the principles of chemistry and the history of the development of basic chemical methods based on these. He is familiar with the most important verified theories and models of the formation of atoms and molecules based on the scientific results of chemistry and the formation of chemical bonds.

b) Abilities

He/She is able to compare the interpretations and models of anthropogenic chemical processes in the natural and the related contexts. He/She is able to use the previously learned paradigms, theories and laws on the field of chemical, natural science to plan, execute and evaluate laboratory experiments.

He/She is able to collect and evaluate data on the field of chemistry in order to opening for problems on social, scientific or ethical questions.

He/She is able to argue on scientific problems by foreign knowledge.

He/She is able to communicate on the field of chemistry using foreign language(s).

He/She is able to use the previously obtained knowledge on the field of natural and anthropogenic studies to solve practical problems.

c) Attitude

He/She seeks after knowing the relationship between nature (especially the chemical phenomena) and the mankind and he/she seeks after understanding these laws. He/she is opened to wider professional cooperation, receptive to new chemical aspects of economics and environmental protection. He represents authentically the scientific world view and historical changes and can

transmit it to a professional and non-professional audience. He/she is committed to acquiring new competencies and expanding the world view.

d) Autonomy and responsibility

- Even in unexpected decision-making situations he/she is capable of considering complex, fundamental questions from his/her professional field and elaborating them on the basis of the given sources.

Schedule:

1st week: The review of the requirement. Science philosophy. Chemistry knowledge in the prehistoric age.

2nd week: The history of the chemistry in the antiquity (Syria, Arabia, Mezopotámia, Egypt, Asia)

3rd week: Chemistry knowledges in the Greek and a Roman age. The appearance of the alchemy.

4th week: Age of alchemy.

5th week: Develeopment of jatro-chemistry.

6th week: The age of discovery of gases.

7th week: Mixtures, compounds, elements, separation, qualitative analysis, chemical symbols, formules, nominations.

8th week: Development of electrochemistry.

9th week: Development of organic chemistry.

10th week: Development of terminology and language of chemistry

11th week: Chemistry and the turn of the century.

12th week: The history of the discovery of medicines. The history is famous poisons and poisoning.

13th week: Test.

14th week: Evaluation. Declaring of results.

Requirements:

- for a signature

Attendance at **lectures** is recommended, but not compulsory.

- for a offered grade

- During the semester there is an end-term test (70% of the total scores) in the 13th week.
- “work at lecture”: at the beginning of the weekly lecture they can write a test (four questions from the previous lecture) and obtain 4 points, the lecturer will add these point to the end points of term test (10% of the total scores)
- “individual collecting work”: If the students write and send an excellent collecting work

(in themes of the lectures) for the lecturer, she/he can obtain further points (20% of the total scores)

Students can obtain an offered mark, if he/she accept this mark, the examination is not necessary for him/her.

If he/she do not accept the offered mark, the course ends in an writing or oral **examination**.

The minimum requirement for end-term tests and the examination respectively is 60%. Based on the score of the tests separately, the grade for the tests and the examination is given according to the following table:

Score	Grade
0-49	fail (1)
50-59	pass (2)
60-74	satisfactory (3)
75-89	good (4)
90-100	excellent (5)

If the score of any test is below 50%, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

Person responsible for course: Dr. Ágnes Dávid, assistant lecturer, PhD

Lecturer: Dr. Ágnes Dávid, assistant lecturer, PhD

Title of course: Special and dangerous materials. Code: TTKBE0204_EN	ECTS Credit points: 3
Type of teaching, contact hours - lecture: 2 hours/week	
Evaluation: examination	
Workload (estimated), divided into contact hours: - lecture: 28 hours - practice: - - laboratory: - - home assignment: - - preparation for the exam: 56 hours Total: 84 hours: 3 credit x 28 hours	
Year, semester: 2 nd -4 th year, 1 st semesters	
Its prerequisite(s): TTKBE0201_EN, TTKBE0301_EN, TTKBE0401_EN	
Further courses built on it:	

Topics of course
Structure, composition, properties and handling/safe use of special materials, which may represent a personal, social or environmental risk or even a life-threatening danger in case of accidents, war, or illicit use.
Literature

Compulsory:

- 1) Chemical Warfare Agents Chemistry, Pharmacology, Toxicology, and Therapeutics, Edited by James A. Romano, Jr. Brian J. Lukey, Harry Salem, CRC Press, ISBN-13 978-1-4200-4661-8
- 2) High Energy Materials. Propellants, Explosives and Pyrotechnics, Jai Prakash Agrawal, 2010, Wiley-VCH Verlag GmbH & Co. KGaA, Weinheim

Recommended:

- 3) Chemistry of Pyrotechnics, Basic Principles and Theory, 2nd Edition, 2010, CRC Press, ISBN-13: 978-1-4200-1809-7
- 4) The Pleasure Instinct Why We Crave Adventure, Chocolate, Pheromones, and Music, Gene Wallenstein, 2009, John Wiley & Sons, Inc., ISBN 978-0-471-61915-4

Course objective/intended learning outcomes

a) Knowledge

- He/she fundamentally knows major types of dangerous and/or illicit materials, their properties, legal and illegal uses and the danger they represent.

b) Abilities

- He/she is able to recognize the major types of the dangerous materials and to estimate the risk they represent.

- He/she is able to make or take part in a plan to keep the persons/environment safe from dangerous materials. He/she is able to give advice on possible/expected behaviour of the dangerous materials.

c) Attitude

- He/she is open to learn and accept professional, technological improvement and innovation in his/her profession and convey it genuinely.

- He/she makes a decision in complex and unexpected cases by completely taking into account legal and ethical norms.

d) Autonomy and responsibility

- Even in unexpected decision-making situations he/she is capable of considering complex, fundamental questions from his/her professional field and elaborating them on the basis of the given sources.

- He/she confesses and represents the value system of the qualified chemists' or chemical engineers' profession with responsibility. He/she is open to critical remarks which are professionally well-founded.

Schedule:

1st week

Narcotics, hard and soft drugs 1. General properties, groups, addiction, legal state. Treatment of addiction. Cannabis.

2nd week

Narcotics, hard and soft drugs 2. Opium, morphine, heroine, opioids. Treatment of addiction, withdrawal syndroms.

3rd week

Narcotics, hard and soft drugs 3. LSD, mescaline, and related derivatives.

4th week

Narcotics, hard and soft drugs 4. Natural materials: Catinone, harmine, harmaline, bufotenine, ibogaine, ephedrine, LSA, safrole, iso-safrole, myristicyne.

5th week

Narcotics, hard and soft drugs 5. Synthetics 1. Amphetamine and derivatives, Extasy, etc..

6th week

Narcotics, hard and soft drugs 6. Synthetics 2. DON, DOB, STP, designer drugs.

7th week

Chemical weapons 1. Major groups, target organs, toxicity. Tear gases, lachrymators.

8th week

Chemical weapons 2. Blood poisons, lung poisons, vesicants..

9th week

Chemical weapons 3. Nerve gases. Floroorganic poisons.

10th week

Chemical weapons 4. Binary chemical weapons. Incendiaries, flame materials, heat source materials.

11th week

Explosives, pyrotechnics 1. Basic concepts, definitions, modes of action. Deflagration: gun powder. Energetic materials, propellants, high energy polymers.

12th week

Explosives, pyrotechnics 2. Initiators, shock and spark sensitive materials. Blasting caps, detonators. High energy explosives, binary explosives, and their civilian and military uses.

13th week

Explosives, pyrotechnics 3. Basic experimental techniques to determine explosive characteristics and stability of explosives and gun powders. Pyrotechnical materials and devices. Civilian pyrotechnics, fireworks.

14th week

Pheromones. Basic properties, mode of action, role in the behavior control and in the physiological signaling processes. Use of pheromones in the agriculture, and in the animal life. Pheromone-like materials, the Dirty 12.

Requirements:

- *for a signature*

Attendance at **lectures** is recommended, but not compulsory.

- *for a grade*

The course ends in an **examination**.

The minimum requirement for the examination is 50 score. Based on the score, the grade for the examination is given according to the following table:

Score	Grade
0-49	fail (1)
50-62	pass (2)
63-75	satisfactory (3)
76-88	good (4)
89-100	excellent (5)

If the score of any test is below 60, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

Person responsible for course: Dr. István Lázár, associate professor, PhD

Lecturer: Dr. István Lázár, associate professor, PhD

Title of course: Environmental chemistry I.

ECTS Credit points: 3

Code: TTKBE0417_EN	
Type of teaching, contact hours	
<ul style="list-style-type: none"> - lecture: 2 hours/week - practice: - - laboratory: - 	
Evaluation: exam	
Workload (estimated), divided into contact hours:	
<ul style="list-style-type: none"> - lecture: 28 hours - practice: - - laboratory: - - home assignment: - - preparation for the exam: 60 hours Total: 90 hours	
Year, semester: 2 nd /3 rd year, 1 st semester	
Its prerequisite(s): TTKBE0201_EN, TTKBE0301_EN, TTKBE0401_EN	
Further courses built on it: -	

Topics of course
<p>The series of lectures are based on the topics of the chemistry of the environment. First the course reviews the definitions, formation, researching method and relations of environmental chemistry with the other fields of science and economy. Then it deals with the formation of our surroundings (formation of the elements, nuclear processes, nuclear fusion and fission, nuclear energy, formation of the stars, planets and the Earth. After that it deals with the chemistry of the atmosphere (structure and characterization, stratospheric ozone, greenhouse effect, aerosols, types of smog, carbon, oxygen, nitrogen and sulphur cycle). The chemistry of the hydrosphere includes the role of the surface waters and oceans in the geochemical cycle, the comparison of the composition and features of sea water and fresh water, industrial and drinking water, water treatment. At the end the course reviews the chemical processes of the pedosphere and the soil, the role of humus materials, as well as the bio-geochemical cycle of phosphorous and other essential elements.</p>
Literature
<p><i>Compulsory:</i></p> <ul style="list-style-type: none"> - I. Pulford, H. Flowers (2006): Environmental Chemistry at a Glance, Blackwell Publishing. ISBN 978-1405135320 - P. Brimblecombe, J. E. Andrews, T. D. Jickells, P. Liss, B. Reid (2003): An Introduction to Environmental Chemistry, Blackwell Publishing. ISBN 0-632-05905-2 - T. G. Spiro, K. L. Purvis-Roberts, W. M. Stigliani (2011): Chemistry of the Environment, Univ. Sci. Books. ISBN 978-1-891389-70-2 <p><i>Recommended:</i></p> <ul style="list-style-type: none"> - G. W. van Loon, S. J. Duffy (2010): Environmental Chemistry: A global perspective, Oxford Univ. Press. ISBN 9780199228867
Course objective/intended learning outcomes
a) Knowledge

- He/she knows the connection between environmental chemistry and other fields of chemistry as well as the relevant associated sciences.

- He/she fundamentally knows the basic principles and definitions of environmental chemistry.

- He/she knows the main questions and problems of environmental chemistry.

- He/she knows the methods and their theoretical background applied for environmental chemical studies.

b) Abilities

- He/she understands the relations in environmental chemistry, and is able to use the principles in practice.

- He/she is able to collect, record, evaluate and discuss environmental chemical data on the basis of the acquired methods.

- He/she is able to understand the problems of the environmental chemistry, and take part in the professional discussions.

- He/she is able to improve/develop his/her knowledge about environmental chemistry.

c) Attitude

- He/she is open to learn and accept the theories and principles of environmental chemistry.

- He/she recognizes the relationships of the environmental problems with the other fields of chemistry and associated sciences.

- He/she is open to acquire new scientific knowledge in this topic, but decline unestablished, deceptive statements.

- He/she is sensible of the general and global environmental problems, and committed for the chemical principles and methods aiming at sustainable development.

d) Autonomy and responsibility

- He/she is capable of considering with responsibility the risk of chemical processes polluting and loading the environment.

- He/she use the literature of environmental chemistry on his/her own.

Schedule:

1st week

Definitions, development, researching methods and relations of environmental chemistry to the other fields of science and economy. The evolution of the natural environment: the Big Bang theory.

2nd week

The formation of the elements. The possible nuclear chemical processes of the synthesis of nuclei. Hydrogen as an energy source. The formation and geochemical development of the Earth.

3rd week

Energy balance of the Earth. The principles of energy production of atomic fission. The structure of a reactor and a power-plant. Advantages and problems of the nuclear energy-production. The basics of isotope dating.

4th week

Chemical evolution: principles, formation of essential mono- and polimers. The development of the atmosphere and hydrosphere. The hypothesis of the origin of life.

5th week

Composition and regions of the Earth's atmosphere. The chemistry and role of the stratospheric ozone. Environmental problems of ultraviolet radiation.

6th week

Chemical reactions in the atmosphere: formation and reaction of O-, C-, S-, N-containing

compounds.

7th week

The most frequent gas and solid air pollutants, their health hazardous effects, elimination possibilities. Natural and social causes and consequences of the greenhouse effect.

8th week

Types of smog, conditions required for their formation. The self-purification of the atmosphere, generation of acidity in rains.

9th week

The composition of the hydrosphere. Physical and chemical properties of water. The chemistry of aqueous solutions: acid-base and redox equilibria. Solubility of gases, liquids and solids in water.

10th week

Water pollution and water quality. Waste water and its treatment.

11th week

Lithosphere: structure and components (rocks and minerals), soil formation, organic and inorganic components of soil.

12th week

Characterization and main functions of soil. Environmental problems associated with soils (acidification, alkalinity, metal contamination, etc.).

13th week

Biogeochemical cycles: carbon, oxygen, nitrogen.

14th week

Biogeochemical cycles: sulphur, phosphorus and metal ions. Treatment of wastes.

Requirements:

- *for a signature*

Attendance at **lectures** is recommended, but not compulsory.

- *for a grade*

The course ends in an **examination**, which means a written test.

The minimum requirement for the examination is 50%. Based on the score of the test (60), the grade is given according to the following table:

Score	Grade
0-29	fail (1)
30-37	pass (2)
38-45	satisfactory (3)
46-53	good (4)
54-60	excellent (5)

If the score of any test is below 30, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

Person responsible for course: Dr. Mónika Kéri, assistant professor, PhD

Lecturer: Dr. Mónika Kéri, assistant professor, PhD

Title of course: Computational Quantum Chemistry

Code: TTKBG0903_EN

ECTS Credit points: 3

Type of teaching, contact hours - lecture: - - practice: 2 hours/week - laboratory: -
Evaluation: exam
Workload (estimated), divided into contact hours: - lecture: - - practice: 28 hours - laboratory: - - home assignment: 32 hours - preparation for the exam: 30 hours Total: 90 hours
Year, semester: 2 nd / 3 rd year, 2 nd semester
Its prerequisite(s): TTMBE0809_EN, TTMBG0809_EN, TTKBG0911_EN
Further courses built on it: -

Topics of course
- Hartree-Fock Theory - Density Functional Theory - Basis sets - Solvent effect, Polarizable Continuum Model - Geometry optimization - Structural analysis - Calculating energies of chemical reactions
Literature
<i>Compulsory:</i> https://maker.pro/linux/tutorial/basic-linux-commands-for-beginners http://gaussian.com/keywords/ <i>Recommended:</i> http://barrett-group.mcgill.ca/tutorials/Gaussian%20tutorial.pdf
Course objective/intended learning outcomes
a) Knowledge He/She knows the basic qualitative and quantitative chemical principles, and the methods based on it. He/She knows the main models and theories of chemical bonds and molecular structure based on scientific findings. He/She has a basic chemical knowledge on describing simple chemical processes as well as on recognizing, organizing these in practice. He/She has the knowledge to test or measure chemical reactions, systems with scientific methods

(including computational) under supervision.

b) Abilities

He/She is able to evaluate and discuss the calculations, and create a report about it.

He/She is able to collect and evaluate data on the field of chemistry in order to opening for problems on social, scientific or ethical questions.

He/She is able to argue on scientific problems by his/her knowledge.

He/She is able to communicate on the field of chemistry using foreign language(s).

c) Attitude

He/She is ready to discuss problems on the field of chemistry and other science with professionals.

He/She is able to represent his/her own personal scientific ideology toward professional and unprofessional groups.

He/She is committed learn or get insights into new competence or ideology.

He/She is well aware about his/her propositions and its consequences.

d) Autonomy and responsibility

He/She stands for his/her opinion or ideology in professional discussions.

He/She can make reasonable evaluations about his/her own work comparing to others to the same field.

Schedule:

1st week

Basic theory of the Hartree-Fock method: approximations, LCAO-MO theory. Building structures by the GaussView program.

2nd week

Basic Linux commands, using the WinSCP and Putty programs, connecting by SFTP. Using the Gaussian program package, optimizing simple molecules.

3rd week

Geometry optimizations by different basis sets, comparing and calibrating the methods by structural parameters.

4th week

Frequency analysis, calculating Gibbs free energies of simple reactions. Scanning a reaction pathway, finding the transition state, identifying the stationary points of the Potential Energy Surface.

5th week

Basic theory of the post-Hartree-Fock theories. Recalculating the previously studied systems and comparing them to the HF results.

6th week

Solvent effect, using Polarizable Continuum Models to refine the energies.

7th week

Basic theory of the Density Functional Theory. Recalculating the previously studied systems and comparing them to the (post-)HF results.

8th week

Systems with explicit solvent molecules.

9th week

Calculation on more difficult systems: metal complexes and relativistic effects.

10th week

Mid-term exam about calculations by using Gaussian.

11th week

Conformation analysis, more Linux commands.

12th week

Writing simple scripts in b shell.

13th week

Generating input files by scripts.

14th week

Exam of writing scripts in b shell.

Requirements:

- *for a signature*

Attendance is recommended, maximum 3 absences are accepted.

- *for a grade*

Class performance (33%)

Final examination (67%)

Based on the sum of the final practical exam of performing calculations and the class performance the practical grade is calculated.

The final grade is given according to the following table:

Score (%)	Grade
0-49	fail (1)
50-59	pass (2)
60-74	satisfactory (3)
75-89	good (4)
90-100	excellent (5)

If the score of the final grade is below 50%, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

Person responsible for course: Dr. Mihály Purgel, assistant professor, PhD

Lecturer: Dr. Mihály Purgel, assistant professor, PhD (67%)

Dr. Attila Mándi, assistant professor, PhD (33%)

Title of course: Unit Operations I

Code: TTKBG0614_EN

ECTS Credit points: 6

Type of teaching, contact hours

- lecture: 2 hours/week
- practice: 3 hours/week
- laboratory: -

Evaluation: exam
Workload (estimated), divided into contact hours: - lecture: 28 hours - practice: 42 hours - laboratory: - - home assignment: 50 - preparation for the exam: 60 hours Total: 180 hours
Year, semester: 2 nd year, 1 st semester
Its prerequisite(s): TTKBE0201_EN, TTKBE0301_EN, TTKBE0401_EN
Further courses built on it: TTKBG0615_EN

Topics of course
The essence of chemical engineering science. Unit Operations of Chemical Engineering. Basis of chemical engineering thermodynamics of unit operations. Quantities describing the operational unit. Measurement, units and dimensions in chemical engineering. Conversion of units. Conditions of thermal, mechanical and component equilibriums. Transport processes, component, heat and momentum streams. The extended- Damköhler's equation. The classification of operational units. The theory of similitude, dimensional analysis. Flow of fluids, energy and momentum relationships. Pumping of fluids. Pumps, compressors and vacuum pumps. Separation of heterogeneous systems: Sedimentation, filtration, centrifugation, mixing of liquid, gas cleaning.
Literature
<i>Compulsory:</i> McCabe, W.L.; Smith, J.C.; Harriot, P. (1993) Unit Operations Of Chemical Engineering (7th Ed) - McGraw-Hill Richard G. Griskey : Transport phenomena and unit operations: a combined approach, (2002), Wiley, ISBN 0-47 1-43819-7 Christie J Geankoplis : Transport processes and unit operations (1993), 3rd edition, Prentice-Hall, ISBN 0-13-045253-X J. M. Coulson, J. F. Richardson: Chemical Engineering. Volume 1-6. Third Edition. Pergamon Press. Oxford
Course objective/intended learning outcomes
a) Knowledge - He/She knows the basic principles, the planning and controlling options in technology of chemical processes and industrial tasks. - He/She knows the chemical methods for measurements or analysis, their principles and instrumental background, and their applicabilities
b) Abilities - He/She is able to follow and control chemical processes and other technological steps concerning the quality management and quality control.

c) Attitude

- During his/her work he/she committed to apply the quality concerns including the new assurances.
- He/She can collaborate with other people and discuss their opinions in problem solving processes before making new decisions.

d) Autonomy and responsibility

- He/She can manage work and worker resources, follow and control the instruments and measuring units.
- He/She makes decisions according to his/her positions, makes suggestions to qualify his/her colleagues involving their promotions.

Schedule:*1st week*

Definition and classification of unit operations. batch and continuous processes. Flowsheets.

2nd week

Physical quantities, units, dimensions. The SI system. Extensive and intensive quantities. Dimensional and tensorial homogeneity. Scalar-vector-tensor quantities.

3rd week

The fundamental equation of thermodynamics. Conditions of equilibrium, driving force, rate of processes. Degrees of freedom of a chemical system.

4th week

Flows and fluxes. Scalar and vector fields and their derivatives. The Nabla vector, gradient and divergence.

5th week

The general transport equation. Differential and integral form of balance equations valid for one and two phase unit operations. The Damköhler equations. The Onsager theory.

6th week

The mathematical model. Initial and boundary conditions. Balance equations for simple systems: Fourier-I and Fick-I laws.

7th week

Similitude and modelling. Dimensional analysis, dimensionless numbers.

8th week

Mass and energy balances for simple and complex unit operations.

9th week

Flow in unpacked pipes and in pipelines: Fluids in rest, Pascal's law. Navier-Stokes equations. Bernoulli equation. Cavitation. Newtonian and non-Newtonian fluids. Newton's law of viscosity.

10th week

Basic types of fluid flow. Reynolds' experiment. Hagen-Poiseuille equation. Modified Bernoulli equation. Fanning equation. Moody diagram. Energy requirement of fluid transport. Types of pumps.

11th week

Flow near solids, in packed columns: Flow around immersed objects. Interpretation of Reynolds number. Types of flow around spherical particles. Stokes' law for the frictional force. Drag coefficient for laminar, transitional and turbulent regions. Ergun equation. Packed columns, characteristics and types of packings. Methods of flow measurement.

12th week

Basics of filtration. Darcy's law of filtration. Batch filtration using constant pressure, continuous filtration using constant flow rate. Filtration units. Filtration using centrifugal force. Types of centrifuges. Basics of membrane filtration. Concentration polarization.

13th week

Mixing of solids, apparatus. Mixing of fluids. Momentum balance for the agitator. Power requirement of agitation. Fluid mixers.

14th week

Terminal velocity of sedimentation. Stokes' law. Drag coefficient as a function of Reynolds number. Apparatus for settling, dust removers, cyclones.

Requirements:

- for a signature

Attendance at **lectures** is recommended, but not compulsory.

Participation at **practice classes** is compulsory. A student must attend the practice classes and may not miss more than three times during the semester. In case a student does so, the subject will not be signed and the student must repeat the course. A student can't make up any practice with another group. Attendance at practice classes will be recorded by the practice leader. Being late is equivalent with an absence. In case of further absences, a medical certificate needs to be presented. Missed practice classes should be made up for at a later date, to be discussed with the tutor.

During the semester there are two tests: the mid-term test in the 8th week and the end-term test in the 15th week. Students have to sit for the tests.

- for a grade

The course ends in an **examination**.

The minimum requirement for the mid-term and end-term tests and the examination respectively is 60%. Based on the score of the tests separately, the grade for the tests and the examination is given according to the following table:

Score	Grade
0-59	fail (1)
60-69	pass (2)
70-79	satisfactory (3)
80-89	good (4)
90-100	excellent (5)

If the score of any test is below 60, students can take a retake test in conformity with the

EDUCATION AND EXAMINATION RULES AND REGULATIONS.

-an offered grade:

it may be offered for students if the average grade of the mid-term and end-term tests is at least satisfactory (3). The offered grade is the average of them.

Person responsible for course: Dr. Miklós Nagy, associate professor, PhD

Lecturer: Dr. Miklós Nagy, associate professor, PhD

Title of course: Unit Operations II Code: TTKBG0615_EN	ECTS Credit points: 6
Type of teaching, contact hours <ul style="list-style-type: none">- lecture: 2 hours/week- practice: 3 hours/week- laboratory: -	
Evaluation: exam	
Workload (estimated), divided into contact hours: <ul style="list-style-type: none">- lecture: 28 hours- practice: 42 hours- laboratory: -- home assignment: 50- preparation for the exam: 60 hours Total: 180 hours	
Year, semester: 2 nd year, 2 nd semester	
Its prerequisite(s): TTKBG0614_EN	
Further courses built on it: -TTKBG0616_EN	

Topics of course

General characterization of transfer processes. Classification of transfer processes. Heat transfer. General characterization of heat transfer. Heat transfer by convection, conduction and radiation. Application of dimensional analysis to heat-transfer by convection. Heating and cooling. Heat transfer at standard- and changeable temperature difference. Unsteady- and steady state transfer of heat. The logarithmic mean temperature difference. Heat exchangers. Evaporation and crystallization. Evaporators and crystallizers. Cooling and coolers. Mass transfer processes. Mass transfer across a phase boundary, the two-film theory. Common interpretation of the operating line and the equilibrium curve. Mass transfer in the columns, the transfer units. Mass transfer in the cascades, the equilibrium units.

Literature

Compulsory:

McCabe, W.L.; Smith, J.C.; Harriot, P. (1993) Unit Operations Of Chemical Engineering (7th Ed) - McGraw-Hill

[Richard G. Griskey](#): Transport phenomena and unit operations: a combined approach, (2002), Wiley, ISBN 0-47 1-43819-7

[Christie J Geankoplis](#): Transport processes and unit operations (1993), 3rd edition, Prentice-Hall, ISBN 0-13-045253-X

J. M. Coulson, J. F. Richardson: Chemical Engineering. Volume 1-6. Third Edition. Pergamon Press. Oxford

Course objective/intended learning outcomes

a) Knowledge

- He/She knows the basic principles, the planning and controlling options in technology of chemical processes and industrial tasks.
- He/She knows the chemical methods for measurements or analysis, their principles and instrumental background, and their applicabilities

b) Abilities

- He/She is able to follow and control chemical processes and other technological steps concerning the quality management and quality control.

c) Attitude

- During his/her work he/she committed to apply the quality concerns including the new assurances.
- He/She can collaborate with other people and discuss their opinions in problem solving processes before making new decisions.

d) Autonomy and responsibility

- He/She can manage work and worker resources, follow and control the instruments and measuring units.
- He/She makes decisions according to his/her positions, makes suggestions to qualify his/her colleagues involving their promotions.

Schedule:

1st week

Flow near solids, in packed columns: Flow around immersed objects. Interpretation of Reynolds number. Types of flow around spherical particles. Stokes' law for the frictional force. Drag coefficient for laminar, transitional and turbulent regions. Packed columns, characteristics and types of packings. Methods of flow measurement.

2nd week

Fluidization and pneumatic transport. Ergun equation. Geldart classification of powders.

3rd week

The heat equation. Types and calculation of heat transport. Steady state heat conduction in plane pipe walls. Fourier-I equation and thermal insulation.

4th week

Unsteady state heat conduction. Fourier-II equation. Dimensionless numbers for transient heat conduction: Fourier, Biot number and dimensionless temperature. Interpretation of the Heissler chart.

5th week

Boundary layer theory of heat transfer. The Nusselt and Prandtl number.

6th week

Forced convection heat transfer.

7th week

Natural convection heat transfer.

8th week

Radiation heat transfer and solution of complex heat transfer problems

9th week

Heat exchangers. Stationary heat transmission with constant temperature difference through flat and cylindrical wall. Determination of heat flow and thermal resistances.

10th week

Direct and indirect heat exchange. Determination of the power requirement for a stationary recuperative heat exchanger. Temperature-space function of co-current and counter current heat exchangers. Logarithmic mean temperature difference. Types and apparatus of heat exchangers.

11th week

Boiling of liquids. Boiling curves. Critical heat flux of boiling. Leidenfrost effect.

12th week

The aim of evaporation, Calandria, falling film and Robert-type evaporator. Multistage evaporators and their connections.

13th week

Analogies between momentum and heat transfer. Chilton-Colburn analogy.

14th week

Practice.

Requirements:

- for a signature

Attendance at **lectures** is recommended, but not compulsory.

Participation at **practice classes** is compulsory. A student must attend the practice classes and may not miss more than three times during the semester. In case a student does so, the subject will not be signed and the student must repeat the course. A student can't make up any practice with another group. Attendance at practice classes will be recorded by the practice leader. Being late is equivalent with an absence. In case of further absences, a medical certificate needs to be presented. Missed practice classes should be made up for at a later date, to be discussed with the tutor.

During the semester there are two tests: the mid-term test in the 8th week and the end-term test in the 15th week. Students have to sit for the tests.

- for a grade

The course ends in an **examination**.

The minimum requirement for the mid-term and end-term tests and the examination respectively is 60%. Based on the score of the tests separately, the grade for the tests and the examination is given according to the following table:

Score	Grade
0-59	fail (1)
60-69	pass (2)
70-79	satisfactory (3)
80-89	good (4)
90-100	excellent (5)

If the score of any test is below 60, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

-an offered grade:

it may be offered for students if the average grade of the mid-term and end-term tests is at least satisfactory (3). The offered grade is the average of them.

Person responsible for course: Dr. Miklós Nagy, associate professor, PhD

Lecturer: Dr. Miklós Nagy, associate professor, PhD

Title of course: Unit Operations III Code: TTKBE0616_EN	ECTS Credit points: 6
Type of teaching, contact hours - lecture: 2 hours/week - practice: 3 hours/week - laboratory: -	
Evaluation: exam	
Workload (estimated), divided into contact hours: - lecture: 28 hours - practice: 42 hours - laboratory: - - home assignment: 50 - preparation for the exam: 60 hours Total: 180 hours	
Year, semester: 3 nd year, 1 st semester	
Its prerequisite(s): TTKBG0615_EN	
Further courses built on it: -	

Topics of course

Mass transfer processes. Absorption. Evaporation. Distillation. Rectification. Extraction. Adsorption. Drying. Crystallization. Chemical reaction engineering. Chemical reactors. Classification of reactors and choice of reactor type in the industry. Chemical kinetics. Residence time and distribution of residence time. Batch reactors and continuous reactors.

Influence of heat of reaction on reactor type. Isothermal, adiabatic polytrophic reactors. Mechanical operations. Size reduction of solids. Methods of operating crushers: coarse-, intermediate-, fine crushers and colloid mills. Classification of solid particles and settling. Blending of solid particles.

Literature

Compulsory:

McCabe, W.L.; Smith, J.C.; Harriot, P. (1993) Unit Operations Of Chemical Engineering (7th Ed) - McGraw-Hill

[Richard G. Griskey](#): Transport phenomena and unit operations: a combined approach, (2002), Wiley, ISBN 0-47 1-43819-7

[Christie J Geankoplis](#): Transport processes and unit operations (1993), 3rd edition, Prentice-Hall, ISBN 0-13-045253-X

J. M. Coulson, J. F. Richardson: Chemical Engineering. Volume 1-6. Third Edition. Pergamon Press. Oxford

Course objective/intended learning outcomes

a) Knowledge

- He/She knows the basic principles, the planning and controlling options in technology of chemical processes and industrial tasks.
- He/She knows the chemical methods for measurements or analysis, their principles and instrumental background, and their applicabilities

b) Abilities

- He/She is able to follow and control chemical processes and other technological steps concerning the quality management and quality control.

c) Attitude

- During his/her work he/she committed to apply the quality concerns including the new assurances.
- He/She can collaborate with other people and discuss their opinions in problem solving processes before making new decisions.

d) Autonomy and responsibility

- He/She can manage work and worker resources, follow and control the instruments and measuring units.
- He/She makes decisions according to his/her positions, makes suggestions to qualify his/her colleagues involving their promotions.

Schedule:

1st week

Mass transfer theories. Two-film and boundary layer theory of component transfer.

2nd week

Absorption-desorption: Concentration-space diagram of a continuous counter current absorption unit operation. Equation of operating line.

3rd week

Transfer unit and its graphical determination. Chemisorption. Types of absorption-desorption apparatus.

4th week

Thermal separation operations: distillation: Batch and continuous distillation, rectification. The aim of evaporation.

5th week

Operating point. Types and parts of a continuous rectification apparatus. Operating lines of a rectifier. The q-line. Equilibrium stage, its determination using McCabe-Thiele diagram.

6th week

Liquid-liquid extraction. Ternary phase diagram. Distributional diagram of the key component. Batch and continuous extraction. Continuous one-stage mixer-settler extractor. Liquid-solid extraction and its apparatus.

7th week

Crystallization and its phase diagram. Apparatus for crystallization.

8th week

Drying. Types of moisture binding. Rate of drying. Enthalpy of moist air. Types, material-and energy balance of drying apparatus

9th week

Humidification.

10th week

Methods of feed preparation and surface increase: size reduction, sieving, vaporization, homogenization: Crushers and grinders. Energy requirement of size reduction. Screening and classification. Sieve analysis

11th week

Introduction to chemical reactors.

12th week

Classification of reactors based on flow, operation mode, component stream and heat. Operation time, residence time. Concentration-time and concentration-space functions of batch and continuous reactors.

13th week

Heat balance of a reactor. Stability of reactors.

14th week

Practice.

Requirements:

- for a signature

Attendance at **lectures** is recommended, but not compulsory.

Participation at **practice classes** is compulsory. A student must attend the practice classes and may not miss more than three times during the semester. In case a student does so, the subject will not be signed and the student must repeat the course. A student can't make up any practice with another group. Attendance at practice classes will be recorded by the practice leader. Being late is equivalent with an absence. In case of further absences, a medical certificate needs to be presented. Missed practice classes should be made up for at a later date, to be discussed with the tutor.

During the semester there are two tests: the mid-term test in the 8th week and the end-term test in the 15th week. Students have to sit for the tests.

- for a grade

The course ends in an **examination**.

The minimum requirement for the mid-term and end-term tests and the examination respectively is 60%. Based on the score of the tests separately, the grade for the tests and the examination is given according to the following table:

Score	Grade
0-59	fail (1)
60-69	pass (2)
70-79	satisfactory (3)
80-89	good (4)
90-100	excellent (5)

If the score of any test is below 60, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

-an offered grade:

it may be offered for students if the average grade of the mid-term and end-term tests is at least satisfactory (3). The offered grade is the average of them.

Person responsible for course: Dr. Miklós Nagy, associate professor, PhD

Lecturer: Dr. Miklós Nagy, associate professor, PhD

Title of course: Applied Radiochemistry

Code: TTKBE0504_EN

ECTS Credit points: 3

Type of teaching, contact hours

- lecture: 2 hours/week
- practice: -
- laboratory: -

Evaluation: exam

Workload (estimated), divided into contact hours:

- lecture: 28 hours
- practice: -
- laboratory: -
- home assignment: 22 hours
- preparation for the exam: 40 hours

Total: 90 hours
Year, semester: 3 rd year, 1 st semester
Its prerequisite(s): TTKBE0403_EN
Further courses built on it: -

Topics of course
<ul style="list-style-type: none"> - Interaction of radiation with matter and its practical aspects. - Radioactive labeling. - Production of radionuclides. - Chemical, biological, medical applications - Nuclear energy production. - Tools and facilities of isotope laboratories.
Literature
<p><i>Compulsory:</i></p> <ul style="list-style-type: none"> - Kónya, J., Nagy N.M., 2012, 2018. Nuclear and Radiochemistry, Elsevier, Oxford. - Choppin, G.R., Liljenzin, J-O., Rydberg, J. Ekberg, C., 2013. Radiochemistry and Nuclear Chemistry, 4th Edition, Elsevier, Amsterdam. - Kratz, J.-V., Lieser, K.H., 2013. Nuclear and Radiochemistry: Fundamentals and Applications, 3rd Edition, Wiley-VCH Verlag GmbH & Co. KGaA, Weinheim, Germany,
Course objective/intended learning outcomes
<p>a) Knowledge</p> <ul style="list-style-type: none"> - He/She has a basic knowledge on the applications of radioisotopes, including the production of radionuclides, tracer methods, chemical applications, nuclear energy production, operation of isotope laboratories. <p>b) Abilities</p> <ul style="list-style-type: none"> - He/She is able to use the obtained chemical knowledge on the field of the application of radioisotopes, to solve the actual basic problems and prove the obtained solution. - He/She is able to argue on scientific problems by his/her knowledge. <p>c) Attitude</p> <ul style="list-style-type: none"> - He/She is open toward scientific and other postgradual education. - He/She treats the scientific results or intellectual properties with the adequate professional ethics. <p>d) Autonomy and responsibility</p> <ul style="list-style-type: none"> - He/She can make reasonable evaluations about his/her own work comparing to others to the same field. - He/She can evaluate his/her co-workers work's responsibly in both laboratory and industrial environment, and report it to his/her chief.

Schedule:
<i>1st week</i>
Interaction of radiation with matter, general sketch of the applications.
<i>2nd week</i>
Radiotracers, physical chemistry of carrier-free concentrations.
<i>3rd week</i>

Basic rules of tracer studies.

4th week

Selection of radiotracers.

5th week

Preparation of frequently used radiotracers, general methods.

6th week

Preparation of frequently used radiotracers, examples.

7th week

Classification of tracer methods, the role of mixing entropy.

8th week

Tracer studies in physical chemistry: kinetics of exchange reactions, coprecipitation, determination of solubility, diffusion studies, surface area determination.

9th week

Analytical applications: isotope dilution, radiometric titration, activation analysis.

10th week

Nuclear and radioanalytical methods based on radiation-matter interactions.

11th week

Applications of isotopes in chemical industry.

12th week

Tracer studies in medicine.

13th week

New trends in nuclear energy production.

14th week

Operation, tools, and facilities of isotope laboratories.

Requirements:

- for a signature

Attendance at **lectures** is recommended, but not compulsory.

- for a grade

The course ends in an **examination**. Based on the examination, the exam grade is given according to the following table:

Score	Grade
0-59	fail (1)
60-69	pass (2)
70-79	satisfactory (3)
80-89	good (4)
90-100	excellent (5)

If the score of the examination is below 60, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

-an offered grade:

it may be offered for students if they write a test on the 14th week and the score of it is at least 60%.

The offered grade is calculated as the exam grade (see above).

Person responsible for course: Dr. Noémi Nagy, professor, DSc

Lecturer: Dr. Noémi Nagy, professor, DSc

Title of course: NMR Operator Training Practice I. Code: TTKML0004_EN, TTKBL0004_EN	ECTS Credit points: 2
Type of teaching, contact hours - practice: 2 hours/week	
Evaluation: practical exam	
Workload (estimated), divided into contact hours: - practice: 28 hours - preparation for the exam: 32 hours Total: 60 hours:	
Year, semester: 2 nd year, 2 nd semester or 3 rd year, 1 st or 2 nd semester	
Its prerequisite(s): Spectroscopic methods I. TTKBE0503_EN	
Further courses built on it: Advanced NMR practical course TTKMG0530_EN	

Topics of course: practical laboratory course with aim that students would be able to pick up ¹H and ¹³C NMR spectra on the 360MHz high field NMR spectrometer without external help

Literature

Compulsory: P.J. Hore, Nuclear Magnetic Resonance, ISBN 963 19 4426 3

Bruker Topspin 3.x manuals (free download)

Recommended: James Keeler, "Understanding NMR Spectroscopy", 2009, ISBN 0-470-01787-2

Course objective/intended learning outcomes

a) Knowledge: firm knowledge of the basic principles of high resolution NMR spectroscopy. Pulse -Fourier principle of NMR spectroscopy.

b) Abilities: to run Bruker NMR spectrometers using topspin software, pick up 1D ¹H NMR and ¹³C NMR spectra within 1 hr timeframe.

c) Attitude: critical and responsible concerning the obtained NMR spectra, with respect to general quality, accuracy etc.

d) Autonomy and responsibility: The main aim is the independent and autonomous use of the sophisticated superconducting spectrometers. Must pay attention to save the technical conditions of the equipments, including the protection of supercon magnets from accidental failure.

Schedule:

1st week Safety rules in NMR labs. with supercon magnets. Dangers for magnets and human beings. Pulse Fourier measurement principle. Hardware of 360 MHz spectrometer: magnet, probeheads, RF preamplifier, electronic control unit, control PC, manual controls.

2nd week Sample preparation: use of deuterated solvents, quality and cleaning of NMR sample tubes, sample amount and dissolving rules. Positioning the samples before measurement, pneumatic transfer of samples into the magnet. Use of deuterium lock in automatic or manual mode. Lock power, field, phase, gain, finding the lock signal. Optimizing lock parameters avoiding saturation of the deuterium signal.

3rd week Homogenisation of the main magnetic field up to 10^{-9} - 10^{-10} accuracy, using the lock signal amplitude. Sample spinning, use of z-shim coils. Non-spinning shims (x,y) combinations. Changing lock phase. Reading and writing shim files (rsh/wsh). Signs of bad shimming. Indicators of good shims in TMS signal.

4th week Recording proton NMR spectra. Measurement principles: pulse program zg and it's visualisation. Acquisition parameters in eda and ased starting windows. Explanation of important parameters: digital sampling and connection between td, sw, aq parameters. Choice of p1 pulse and d1 relaxation delay for quantitative ¹H-NMR. Real-time FID shimming in gs mode.

5th week Processing proton NMR spectra. Math rules of Fourier transformation with FFT. TD and SI, zero filling. Window functions for S/N enhancement (em) or resolution (gm) enhancement. Phase correction to pure absorption phase - automatic or manual. Baseline correction for accurate integrals. Integration routine and calibration, correction of integrals.

6th week Recording carbon NMR spectra. Pulse programs zgdc and jmod. Explaining the double impact of proton decoupling: removing splittings caused by proton-carbon spin-spin couplings and heteronuclear NOE that improves carbon sensitivity. Explaining the proton channel power and dB scale, and heating effect danger. Exponential line broadening is a must (em) before FT. Explaining and running the jmod spin-echo sequence.

7th week Recording more carbon NMR spectra with gated (zggd) and inverse gated (zgif) sequences. The former for measuring heteronuclear couplings with better sensitivity, the latter for quantitative ¹³C-NMR. Adjusting optimal parameters for carbon NMR. Explaining signal multiplicity of deuterated organic solvents. Peak picking (ppm) of spectra.

8th week Exercising ¹H NMR signal acquisition and processing one by one.

9th week Exercising ¹³C NMR signal acquisition and processing one by one.

10th week Exercising ¹H NMR signal acquisition and processing one by one.

11th week Exercising ¹³C NMR signal acquisition and processing one by one.

12th week Exercising ¹H NMR and ¹³C NMR signal acquisition and processing one by one.

13th week Exercising ¹H NMR and ¹³C NMR signal acquisition and processing one by one.

14th week Exercising ¹H NMR and ¹³C NMR signal acquisition and processing one by one.

Requirements:

- for a signature

Attendance of laboratory exercises is compulsory.

A student must attend the practice classes and may not miss more than two times during the semester. In case a student does so, the subject will not be signed and the student must repeat the course. A student can't make up any practice with another group. Attendance at practice classes will be recorded by the practice leader. Being late is equivalent with an absence. In case of further absences, a medical certificate needs to be presented. Missed practice classes should be made up for at a later date, to be discussed with the tutor. Active participation is evaluated by the teacher in every class. If a student's behaviour or conduct doesn't meet the requirements of active participation, the teacher may evaluate his/her participation as an absence because of the lack of active participation in class.

- *for a grade*

The course ends in an **examination**. The student must produce an ¹H NMR spectrum with quantitative integrals and a ¹³C NMR spectrum with peak list within one hour time limit, without external help. They may ask for tutor help, however this may result in lowering their mark.

- the result of the practical examination may be 1 (failed) 2,3,4,5 (passed)

Person responsible for course: Dr. Batta Gyula, professor, PhD

Lecturer: Dr. Batta Gyula, professor, PhD

Title of course: Biochemistry III Code: TTBBE0304_EN	ECTS Credit points: 3
Type of teaching, contact hours - lecture: 2 hours/week - practice: - - laboratory: -	
Evaluation: exam	
Workload (estimated), divided into contact hours: - lecture: 28 hours - practice: - - laboratory: - - home assignment : - preparation for the exam: 60 hours Total: 60 hours	
Year, semester: 2 nd year, 2 nd semester	
Its prerequisite(s): Biochemistry I	
Further courses built on it: -	

Topics of course

The lectures cover the main features of the protein structures including fibrous proteins and the membrane proteins with their role in transport. There is an insight into the photosynthesis: the light

reactions and the carbon-assimilation reactions. The nucleotide metabolism is summarized. The biosynthesis of macromolecules such as DNA, RNA and protein will also be described. Post-translational modification: N-glycosylation is also mentioned.

Literature

Compulsory: The lecture notes

Recommended:

Nelson D.L., Cox M.M.: Lehninger Principles of Biochemistry (W. H. Freeman Sixth edition, 2012) ISBN-13: 978-14234146.

Berg J.M., Tymoczky J.L., Gatto G.J. and Stryer L.: Biochemistry (W. H. Freeman; Eighth edition, 2015), ISBN-13: 978-1464126109.

Albert B., Bray D. Essential Cell Biology (Fourth edition, Garland Science, 2014) ISBN: 978-0-8153-4454-4.

Course objective/intended learning outcomes

a) Knowledge

- He/she knows the structural and functional features of the proteins including fibrous and membrane proteins.
- She/he knows the principles that govern the photosynthetic processes.
- She/he is also aware of the characteristics of the nucleic acid and protein biosynthesis.

b) Abilities

- He/she is able to understand the function of the different structural form of the proteins.
- He/she is able to understand the fundamentals of the biosynthetic pathways of the macromolecules.
- He/she is able to understand of the complex events at the different stages of the photosynthesis.

c) Attitude

-He/she is open to the contextual observations of the studied area and is motivated to follow the latest scientific theories in that field.

d) Autonomy and responsibility

- He/she is capable of considering complex questions on the studied scientific field on her/his own as well as in a team.
- He/she shows responsibilities in her/his profession.

Schedule:

1st week

The different structural level of proteins. Protein folding and chaperons. Protein misfolding. Structural classification of proteins.

2nd week

Fibrous proteins: α -keratin, fibroin and the structure of collagen fibrils. Structural feature of membrane protein.

3rd week

The role of membrane proteins in transport processes of the cell. Facilitated diffusion by transport proteins. Primary and secondary active transport. The ion selective channels.

4th week

The role, the location and the components of photosynthesis. The light driven electron flow in Photosystem I and II. The function and structure of Cythochrome b_6f complex.

5th week

The synthesis of ATP and NADPH in the light reactions of photosynthesis. The cyclic photophosphorylation. The water splitting complex. Comparing the light reactions of the photosynthesis with the oxidative phosphorylation taking place at the mitochondria.

6th week

Photosynthetic assimilation of carbon dioxide. The function, structure and regulation of Rubisco. The three stages of the Calvin cycle. Photorespiratory reactions and the C₄ pathway.

7th week

Nucleotide Metabolism. The biological function of nucleotides. The pyrimidin *de novo* biosynthesis. The interconversion of nucleoside mono- di- and triphosphates.

8th week

The purin *de novo* biosynthesis. The role of tetrahydrofolate in the nucleotide biosynthesis. The Salvage pathway. The function of ribonucleotide reductase in the generation of deoxyribonucleotides. Degradation of purin and pyrimidine nucleotides.

9th week

The biosynthesis of deoxyribonucleic acid. The helical structure of DNA. The Meselson-Stahl experiment. The stages of replication in prokaryotes. The replication forks. DNA synthesis on the leading and lagging strands.

10th week

The function of the protein factors and enzymes involved in the the processes of replication including primase, DNA polymerases I and III, DNA ligase. Termination of chromosome replication in bacterial cell.

11th week

The biosynthesis of ribonucleic acids in prokaryotes. The function and characteristics of the DNA - dependent RNA polymerase. Transcription initiation, elongation and termination.

12th week

The biosynthesis of ribonucleic acids in eukaryotes. The function of the different RNA polymerases. Assembly of the Initiation Complex. RNA processing: 5' capping and 3' Poly(A) Tail. RNA splicing.

13th week

The biosynthesis of proteins. The genetic code. The structure and the function of tRNA. The components of the ribosome. The stages of the protein biosynthesis. Proofreading on the ribosome. Antibiotics inhibit translation.

14th week

Signal sequences and protein targeting. Protein translocation into the ER. Post-translational modification: N-glycosylation and its function.

Requirements:

- *for a signature*

Attendance at **lectures** is recommended, but not compulsory.

- *for a grade*

The course ends in an **examination**.

The grade for the examination is given according to the following table:

Score	Grade
0-59	fail (1)
60-69	pass (2)
70-79	satisfactory (3)
80-89	good (4)

90-100 excellent (5)
If the score of examination is below 60, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.
Person responsible for course: Dr. Teréz Barna, PhD
Lecturer: Dr. Teréz Barna, PhD

Title of course: Biocolloids Code: TTKBE0405_EN	ECTS Credit points: 3
Type of teaching, contact hours - lecture: 2 hours/week - practice: - - laboratory: -	
Evaluation: exam	
Workload (estimated), divided into contact hours: - lecture: 14 hours - practice: - - laboratory: - - home assignment: 40 hours - preparation for the exam: 40 hours Total: 94 hours	
Year, semester: 2 nd /3 rd year, 2 nd semester	
Its prerequisite(s): TTKBE0402_EN	
Further courses built on it: -	

Topics of course
The goal of this series of lectures is to give knowledge about the relationship between biological sciences and colloid/surface phenomena. A further goal is to deepen colloid chemical knowledge of students about biological phenomena related to colloids. It makes them able to approach biological problems from a colloid chemical perspective and to solve possible problems and tasks in this context.
Literature
<i>Compulsory:</i> - Lecture slides downloadable from the Department's homepage (http://fizkem.unideb.hu)
<i>Recommended:</i> - D. Fennell Evans, Hakan Wennerstrom: The Colloidal Domain: Where Physics, Chemistry and Biology Meet, 2nd Ed. ,Wiley, 1999 - Pashley, R. M.: Applied Colloid & Surface Chemistry. Wiley&Sons, ISBN 0-a470-a86883-aX, 2004

- Cosgrove T.: Colloid science. Blackwell Publishing ISBN:978-a14051-a2673-a1, 2005

Course objective/intended learning outcomes

a) Knowledge

- He/She knows the main models and theories of chemical bonds and molecular structure based on scientific findings.
- He/She has knowledge to solve problems on the field of natural processes, using natural sources, and understanding the chemical background of living and non living systems.

b) Abilities

- He/She is able to use the previously obtained knowledge on the field of natural and anthropogenic studies to solve practical problems.
- He/She is able to argue on scientific problems by his/her knowledge.

c) Attitude

- He/She is ready to discuss problems on the field of chemistry and other science with professionals.
- He/She is able to represent his/her own personal scientific ideology toward professional and unprofessional groups..
- He/She is open toward scientific and other postgradual education.

d) Autonomy and responsibility

- He/She can make reasonable evaluations about his/her own work comparing to others to the same field.
- He/She stands for his/her opinion or ideology in professional discussions.

Schedule:

1st week

Importance of colloidal state in biology. Hypotheses about the origin of life in the past and nowadays. Occurrence of organic matter in space. Hyperresistant organisms and survival under the conditions found in space. Shadow biosphere and "artificial life".

2nd week

Formation of interfaces. Films and membranes. Langmuir-Blodgett films and liquid crystals. Membrane models, structure of the cell membrane.

3rd week

Diffusion and transport phenomena through membranes, osmosis and dialysis. Transport phenomena in living organisms. Function of the kidneys, artificial kidney.

4th week

Adsorption phenomena in biological systems, processes in biotechnology and separation sciences.

5th week

Surface tension and its importance in nature. Motion of striders on the surface of water. Reproduction using surface tension: ballistospores of fungi. Wetting, contact angle, influencing the surface tension. Capillarity, water transport in plants and the transpiration-adhesion-tension-cohesion hypothesis. The importance of capillarity under arid climates. Adhesion to smooth surfaces. Atherosclerosis and interfacial influences leading to atherosclerosis.

6th week

Association colloids, micelles and inverse micelles. Critical micelle concentration and its importance. Detergents and their uses. Biological detergents in the digestion: bile acids. Solubilization with polar molecules. Lung surfactants and their role in breathing.

7th week

Modern instrumental methods in the study of biomacromolecules (ultracentrifugation, electrophoresis, size exclusion chromatography, scanning confocal microscopy, electron microscopy, scanning probe microscopy, surface plasmon resonance, X-ray diffraction, NMR).

8th week

Macromolecules, types and importance of macromolecules. Characterization and importance of dispersity, shape, and conformation.

9th week

Important and interesting biomacromolecules, their properties, importance and uses (*polysaccharides*: cellulose, starch, chitin, etc.; *proteins*: collagen, silk, green fluorescent protein, etc.; *others*: lignin, chlorophylls, haemoglobin, etc.).

10th week

Dispersion colloids in nature. Bioaerosols and smokes. Importance of foams, emulsions, sols and their biological relevance. Making and breaking of dispersions in different biological, medical, pharmaceutical, etc. processes.

11th week

Coherent systems and lyogels. The eye as a natural lyogel system. Biocomposites: structure and formation of bones. A complex disperse system: the soil.

12th week

Electrokinetic effects, precipitation from liquids. Epitaxis. Kidney and bile stones, processes of their formation.

13th week

Flow properties. Biorheology. Rheology of blood and its importance in blood coagulation.

14th week

Nanotechnology and its development. Nanostructures from non-living matter. Natural nanostructures: diatoms and the fine structure of butterfly scales. Nanodevices. Natural nanomotors: kinesins, dyneins, the actomyosin complex. DNS machines, active molecular tweezers.

Requirements:

- *for a signature*

Attendance at **lectures** is recommended, but not compulsory.

- *for a grade*

The course ends in an **examination**. The minimum requirement for the examination is 50%. The grade for the examination is given according to the following table:

Score	Grade
0-49	fail (1)
50-59	pass (2)
60-74	satisfactory (3)
75-89	good (4)
90-100	excellent (5)
If the score of any test is below 50%, students can make a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.	
Person responsible for course: Dr. Levente Novák, assistant professor, PhD	
Lecturer: Dr. Levente Novák, assistant professor, PhD	

Title of course: Colloid Chemistry Code: TTKBE0415_EN	ECTS Credit points: 3
Type of teaching, contact hours - lecture: 2 hours/week - practice: - - laboratory: -	
Evaluation: exam	
Workload (estimated), divided into contact hours: - lecture: 14 hours - practice: - hours - laboratory: - - home assignment: 40 hours - preparation for the exam: 40 hours Total: 94 hours	
Year, semester: 3 rd year, 2 nd semester	
Its prerequisite(s): TTKBE0402_EN	
Further courses built on it: -	

Topics of course
The goal of this series of lectures is to give knowledge to the students about the relation between size and physico-chemical properties. Students are expected to get acquainted with the behaviour of nanosized particles, the role of the interfaces and their possible applications.
Literature
<i>Compulsory:</i> - Lecture slides downloadable from the Department's homepage (http://fizkem.unideb.hu) - Barnes, GT, Gentle, IR: Interfacial Science. Oxford UP. ISBN 0-a19-a927882-a2, 2005 - Pashley, R. M.: Applied Colloid & Surface Chemistry. Wiley&Sons, ISBN 0-a470-a86883-aX, 2004

- Cosgrove T.: Colloid science. Blackwell Publishing ISBN:978-a14051-a2673-a1, 2005

Course objective/intended learning outcomes

a) Knowledge

- He/She knows the main models and theories of chemical bonds and molecular structure based on scientific findings.
- He/She has a knowledge to solve problems on the field of natural processes, using natural sources, and understanding the chemical background of living and non living systems.

b) Abilities

- He/She is able to use the previously obtained knowledge on the field of natural and anthropogenic studies to solve practical problems.
- He/She is able to argue on scientific problems by his/her knowledge.

c) Attitude

- He/She is ready to discuss problems on the field of chemistry and other science with professionals.
- He/She is able to represent his/her own personal scientific ideology toward professional and unprofessional groups.
- He/She is open toward scientific and other postgradual education.

d) Autonomy and responsibility

- He/She can make reasonable evaluations about his/her own work comparing to others to the same field.
- He/She stands for his/her opinion or ideology in professional discussions.

Schedule:

1st week

Introduction. The notion of colloids and the classification of colloid systems. Synthesis of colloids. Relation between colloids and nanotechnology. Average and types of average.

2nd week

Molecular interactions. Quantitative description of electrostatic and van der Waals interactions, their role in the synthesis of colloids. Lennard-Jones potential. Hydrophilic and hydrophobic interactions.

3rd week

Notion and characterization of interfaces. Fluid interfaces. Interfacial phenomena, the concept of surface tension. The Eötvös rule. Laplace pressure, importance of curved surfaces.

4th week

Nonfluid interfaces. Contact angle, wetting and spreading. Adhesion and cohesion. Adsorption at fluid interfaces, the Gibbs isotherm. Langmuir and Langmuir-Blodgett layers.

5th week

Adsorption at solid-liquid interfaces. Adsorption isotherms. Formation of charged interfaces and their significance. Chromatographies.

6th week

Formation of the electrostatic double layer, its structure and description. Comparison of the

Helmholtz, Gouy-Chapman and Stern models. Potentials. Zeta potential.

7th week

Electrokinetic phenomena. Electrophoretic mobility. The phenomenon of electroosmosis and its practical use in capillary electrophoresis.

8th week

Stabilization and destabilization of lyophobic colloids. The Hamaker model. The DLVO theory. Sterical stabilization. Salting out. Destabilization of lyophilic colloids. The technology of butter- and cheese-making.

9th week

Gas-liquid disperse systems. Stability, preparation and importance of aerosols. Stability, preparation and practical use of foams.

10th week

Liquid-liquid disperse systems. Preparation and breaking of emulsions. Emulsifiers, the HLB value.

11th week

Solid-liquid disperse systems. Their preparation, stabilization, kinetic description of their formation.

12th week

Association colloids. Surface activity. Amphiphilic molecules and micelles. Micelle formation, the critical micelle concentration. Surfactants, detergents.

13th week

Types of macromolecular colloids. Macromolecules and plastics. Drug transport and targeted delivery.

14th week

Basics of rheology. Viscosity and its measurement. Viscosity- and flow curves. Basic rheological types. Applications.

Requirements:

- for a signature

Attendance at **lectures** is recommended, but not compulsory.

- for a grade

The course ends in an **examination**. The minimum requirement for the examination is 50%. The grade for the examination is given according to the following table:

Score	Grade
0-49	fail (1)
50-59	pass (2)
60-74	satisfactory (3)
75-89	good (4)
90-100	excellent (5)

If the score of any test is below 50%, students can make a retake test in conformity with the

EDUCATION AND EXAMINATION RULES AND REGULATIONS.

Person responsible for course: Dr. Levente Novák, assistant professor, PhD

Lecturer: Dr. Levente Novák, assistant professor, PhD

Title of course: Plastics and Processing II. Code: TTKBE0711_EN	ECTS Credit points: 2
Type of teaching, contact hours - lecture: - - practice: 2 hours/week - laboratory: -	
Evaluation: exam	
Workload (estimated), divided into contact hours: - lecture: - - practice: 28 hours - laboratory: - - preparation for the tests: 40 hours Total: 68 hours	
Year, semester: 3 rd year, 2 nd semester	
Its prerequisite(s): TTKBE0611_EN	
Further courses built on it: -	

Topics of course
The current situation and future prospects of world and domestic plastics production and use. Production of polyethylene I. (high pressure). Production of polyethylene II. (high pressure tube reactor and medium pressure processes) and its applications. Production of polypropylene, newer technology development. Domestic technologies for production of polypropylene (bulk polymerization and gas phase processes), use of polypropylene. Production of polystyrene (high impact strength and expandable polystyrene) and its use. Possibilities of manufacturing PVC. Home production and use of PVC. Possibilities for producing polyamides. Production and use of polyamide-6. Production and use of polyacrylonitrile. Manufacture and use of polyester fabrics. Additives used in the plastics industry.
Literature
<i>Recommended:</i> 1. Website of MOL Petrochemicals 2. <i>Ullmann's Encyclopedia of Industrial Chemistry</i> , Wiley-VCH Verlag GmbH & Co. KGaA (2002) 3. George Odian: <i>Principles of Polymerization</i> , McGraw-Hill, New York (1983)
Course objective/intended learning outcomes
a) Knowledge - He/She knows the production technologies of the most important polymers.
b) Abilities

- He/She is able to interpret the operations in the entire process of technology
- He/She is able to conduct professional communication from the above area.
- He/She is able to expand / improve its technological know-how in new tasks..

c) Attitude

- He/She is open to gain new knowledge in the subject.
- He/She also asks his colleagues to comply with the rules of accident protection and safety and shows an example of his own work.

d) Autonomy and responsibility

- He/She tends to establish new solutions and technologies.
- He/She with a little professional guidance is able to carry out smaller sub-tasks independently.
- He/She is able to carry out safety training and, in simpler cases he/she is able to evaluate realistically potential hazards independently.

Schedule:

1st week

The current situation and future prospects of world and domestic plastics production and use.

2nd week

Production of polyethylene I. (high pressure).

3rd week

Production of polyethylene II. (high pressure tube reactor and medium pressure processes) and its applications.

4th week

Production of polypropylene, newer technology development.

5th week

Domestic technologies for production of polypropylene (bulk polymerization and gas phase processes), use of polypropylene.

6th week

Production of polystyrene (high impact strength and expandable polystyrene) and its use.

7th week

Possibilities of manufacturing PVC

8th week

Home production and use of PVC.

9th week

Possibilities for producing polyamides. Production and use of polyamide-6.

10th week

R Production and use of polyacrylonitrile.

11th week

Manufacture and use of polyester fabrics.

12th week

Additives used in the plastics industry.

13th week

Consultation and PPT presentations.

14th week

Test and PPT presentations.

Requirements:

Attendance at seminars is compulsory.

The course ends with test for a partial grade. (This test is compulsory!) The minimum requirement for the test is 50%. The grade is given according to the following table:

Score	Grade
0-49	fail (1)
50-59	pass (2)
60-80	satisfactory (3)
81-90	good (4)
91-100	excellent (5)

All the students will deliver a ppt presentation on a subject. They will get a second partial grade. The term grade will be calculated by the following way: 60% of the test result, 40% of the ppt presentation

Person responsible for course: Dr. György Deák, associate professor, PhD

Lecturer: Dr. György Deák, associate professor, PhD

Title of course: Seminar in Organic Chemistry I. Code: TTKBG0311_EN	ECTS Credit points: 1
Type of teaching, contact hours - lecture: - - practice: 1 hour/week - laboratory: -	
Evaluation: term mark	
Workload (estimated), divided into contact hours: - lecture: - - practice: 14 hours - laboratory: - - home assignment: 14 hours - preparation for the exam: - Total: 28 hours	
Year, semester: 1 st year, 2 nd semester	
Its prerequisite(s): General Chemistry I. (lecture) TTKBE0101_EN	
Further courses built on it: -	

Topics of course

- Review the basic of organic chemistry basics
- Types and theories of chemical bonds
- Review the acid-base theories
- Basic concepts of isomerism and stereochemistry.
- Classification of organic chemical reactions.
- Functional groups and the basics of organic nomenclature.

- The structure, nomenclature, synthesis and reactions of alkanes, alkenes, alkynes, mono- and polycyclic, homo- and heteroaromatic hydrocarbons.

Literature

Compulsory:

6. Course material, concept and task collection for lectures, seminars in the e-learning system.

Recommended:

7. J. G. Smith: Organic Chemistry, 5th Edition, 2016, McGraw Hill; ISBN-13: 9780077354725
8. J. McMurry: Organic Chemistry, 8th Edition, 2012, Brooks/Cole; ISBN-13: 9780840054449
9. J. Clayden, N. Greeves, and S. Warren: Organic Chemistry, 2nd Edition, 2012, Oxford University Press; ISBN-13: 9780199270293
10. F. A. Carey: Organic Chemistry, 4th Edition, 2000, The McGraw-Hill Companies; ISBN-13: 9780072905014
11. L. G. Wade: Organic Chemistry, 8th Edition, 2012, Pearson; ISBN-13: 9780321768148
12. T. W. G. Solomons, C. Fryhle, Organic Chemistry, 10th Edition, 2009, Wiley & Sons; ISBN-10: 0470556595
13. H. Meislich, E. K. Meislich, J. Sharefkin: 3000 Solved Problems in Organic Chemistry, 1st Edition, 1994, McGraw-Hill Companies; ISBN-13: 978-0070564244

Course objective/intended learning outcomes

a) Knowledge

He/She knows the basic concepts and theories which are necessary to understand and interpret structure and reactivity of organic compounds (chemical bond, hybridization, resonance theory, isomerism) He/she knows the structure, physical and chemical properties and synthetic methods of saturated, unsaturated and aromatic hydrocarbons and He/She can apply these knowledges to solve chemical problems.

b) Abilities

- He/she knows and is able to apply his/her knowledge to solve simple tasks on the field of hydrocarbons.
- He/she is able to participate in professional communication on the field of structure, and chemical transformation of hydrocarbons.
- He/she is able to expand and/or develop his/her knowledge from the natural products.

c) Attitude

- He/she is open to getting new, scientifically proven knowledge on the subject, but to reject unsubstantiated or possibly misleading claims

d) Autonomy and responsibility

- He/she is able to independently perform the tasks of the course with professional guidance, and he/she can interpret and evaluate of the results obtained.

Schedule:

1st week

Comparison and exercise of representation of organic compounds. Determination of the order (primary, secondary, tertiary, quaternary) of carbon atoms in compounds.

2nd week

The use of resonance structures and hybridization in the interpretation of the structure of organic compounds. Interpretation of electron shift or delocalization phenomena (inductive and mesomeric effect, conjugation and hyperconjugation).

3rd week

Exercise the recognition of organic compounds and functional groups.

4th week

Use of the substitutive and functional class nomenclature in naming hydrocarbons. Practice the names of alkyl groups.

5th week

Exercise of the most important types of organic chemical reactions, recognition of reactive particles (electrophile, nucleophile, radical).

6th week

Exercise the concept of constitution, conformation and configuration. Recognition and differentiation of enantiomers and diastereomers.

7th week

Practice the representation and projection of the organic molecules. The absolute configuration of chiral compounds, Fischer and Cahn-Ingold-Prelog convention.

8th week

Interpretation of radical transformations of alkanes. Statistical and regioselective halogenation of alkanes. Synthesis of alkanes.

9th week

Methods for the synthesis of alkenes, cycloalkenes. Addition reactions of alkenes, regioselectivity and its interpretation in addition reactions.

10th week

Addition reactions of conjugated dienes, partial and complete addition. 1,2- and 1,4- addition and its interpretation based on kinetic and thermodynamic control. Diels-Alder cycloaddition.

11th week

Synthesis of alkynes. Chemical transformations of alkynes: C-H acidity, addition reactions and their significance. The role of acetylene in the chemical industry, coal-based chemical industry.

12th week

Exercise the criteria of aromaticity. Interpretation of aromatic electrophilic substitution reactions.

13th week

The S_{EAr} reactions of substituted benzene derivatives –the reactivity and regioselectivity.

Classification of substituents and interpretation of their effect on reactivity and regioselectivity.

14th week

Reactions of aromatic hydrocarbons containing alkyl residues, interpretation of the stability of benzyl-type reactive intermediates. Most important representatives of polycyclic aromatic hydrocarbons.

Requirements:

The course is recommended in parallel with the lecture Organic Chemistry I. (TTKBE0301_EN).

Evaluation:

- for a signature

Attendance at seminars is **compulsory**. A student may not miss the seminar more than three times during the semester. In case of further absences, a medical certificate needs to be presented. In case a student does not do this, the subject will not be signed, and the student must repeat the course.

The performance of the students in the seminar is verified 4 times in the form of written tests.

- for a grade

The term mark is based on the average of the grades of written tests.

The minimum requirement for the written tests respectively is 50%. Based on the score of the tests separately, the grade for the tests is given according to the following table:

Score	Grade
0-49	fail (1)
50-65	pass (2)
66-80	satisfactory (3)
80-89	good (4)
90-100	excellent (5)

Person responsible for course: László Dr. Juhász, associate professor, PhD, Habil.

Lecturer: László Dr. Juhász, associate professor, PhD, Habil.

Title of course: Seminar in Organic Chemistry II. Code: TTKBG0312_EN	ECTS Credit points: 1
Type of teaching, contact hours - lecture: - - practice: 1 hour/week - laboratory: -	
Evaluation: term mark	
Workload (estimated), divided into contact hours: - lecture: - - practice: 14 hours - laboratory: - - home assignment: 14 hours - preparation for the exam: - Total: 28 hours	
Year, semester: 2 nd year, 1 st semester	
Its prerequisite(s): Inorganic Chemistry I. (lecture) TTKBE0201_EN, Organic Chemistry I. (lect and sem.) TTKBE0301_EN, Physical Chemistry I. (lecture) TTKBE0401_EN	
Further courses built on it: -	

Topics of course

Overview and exercising of the structure, physical, chemical properties of hydrocarbons possessing heteroatoms as halogenated hydrocarbons, organometallic derivatives, alcohols, phenols, ethers and their thio analogues; amines, nitro derivatives, diazonium salts, aldehyde, ketones, carboxylic acids and their derivatives, derivatives of carbonic acid

Literature

mpulsory:

14. Course material, concept and task collection for lectures, seminars in the e-learning system.

commended:

15. J. G. Smith: Organic Chemistry, 5th Edition, 2016, McGraw Hill; ISBN-13: 9780077354725

16. J. McMurry: Organic Chemistry, 8th Edition, 2012, Brooks/Cole; ISBN-13: 9780840054449

17. J. Clayden, N. Greeves, and S. Warren: Organic Chemistry, 2nd Edition, 2012, Oxford University Press; ISBN-13: 9780199270293

18. F. A. Carey: Organic Chemistry, 4th Edition, 2000, The McGraw-Hill Companies; ISBN-13: 9780072905014

19. L. G. Wade: Organic Chemistry, 8th Edition, 2012, Pearson; ISBN-13: 9780321768148

20. T. W. G. Solomons, C. Fryhle, Organic Chemistry, 10th Edition, 2009, Wiley & Sons; ISBN-10: 0470556595

21. H. Meislich, E. K. Meislich, J. Sharefkin: 3000 Solved Problems in Organic Chemistry, 1st Edition, 1994, McGraw-Hill Companies; ISBN-13: 978-0070564244

Course objective/intended learning outcomes

a) Knowledge

- He/she knows the structure, physical and chemical properties and synthetic methods of the most important organic compounds possessing heteroatoms (halogenated hydrocarbons, organometallic derivatives, alcohols, phenols, ethers and their thio analogues; amines, nitro derivatives, diazonium salts, aldehyde, ketones, carboxylic acids and their derivatives, derivatives of carbonic acid). He/she is able to apply his/her knowledge to solve simple tasks on the field of this compounds. He/she knows the application and practical significance of these derivatives., and their applicabilities

b) Abilities

- He/she knows and is able to apply his/her knowledge to solve simple tasks on the field of hydrocarbons.
- He/she is able to participate in professional communication on the field of structure, and chemical transformation of hydrocarbons.
- He/she is able to expand and/or develop his/her knowledge from the natural products.

c) Attitude

- He/she is open to getting new, scientifically proven knowledge on the subject, but to reject unsubstantiated or possibly misleading claims

d) Autonomy and responsibility

- He/she is able to independently perform the tasks of the course with professional guidance, and he/she can interpret and evaluate of the results obtained.

Schedule:

1st week

Practice the classification and synthesis of halogenated hydrocarbons.

2nd week

Practice the elimination and substitution reactions of halogenated hydrocarbons.

3rd week

Practice the preparation of Grignard compounds and their application.

4th week

Preparation of alcohols, ethers, phenols and their thioanalogues. The acid-base properties of alcohols, phenols and their thioanalogues

5th week

Practice the chemical properties of alcohols and phenols, ethers and their thioanalogues.

6th week

Practice the classification of amines and characterization of their bonding systems. Practice the synthetic methodologies of aliphatic and aromatic amines, industrial methods.

7th week

Practice the basicity and chemical transformations of the amines (alkylation, acylation, sulfonamide formation, reaction with nitric acid). Reactions of aromatic rings of anilines.

8th week

Practice the preparation of nitro compounds, diazonium salts. Reactions and practical significance of aromatic diazonium salts.

9th week

Practice the synthetic possibilities of aldehydes and ketones and an overview of their acid-base properties.

10th week

Practice the transformations of aldehydes and ketones. Reactions of the carbonyl group (nucleophilic addition reactions with O-, S-, N- and C-nucleophiles) and reactions on the α -carbon atoms.

11th week

Practice the classification and preparation of carboxylic acids and their derivatives.

12th week

Practice the acid-base properties of carboxylic acids and its derivatives. The acyl nucleophilic substitution and the reductive transformations of carboxylic acid derivatives, transformation of their carbon skeleton.

13th week

Chemical properties of β -dicarboxylic acids, malonester synthesis.

14th week

Chemical properties of β -oxocarboxylic acid derivatives, acetoacetic ester and cyanoacetic ester syntheses.

Requirements:

The course is recommended in parallel with the lecture Organic Chemistry II. (TTKBE0302_EN).

Evaluation:

- for a signature

Attendance at seminars is **compulsory**. A student may not miss the seminar more than three times during the semester. In case of further absences, a medical certificate needs to be presented. In case a student does not do this, the subject will not be signed, and the student must repeat the course.

The performance of the students in the seminar is verified 4 times in the form of written tests.

- for a grade

The term mark is based on the average of the grades of written tests.

The minimum requirement for the written tests respectively is 50%. Based on the score of the

tests separately, the grade for the tests is given according to the following table:

Score	Grade
0-49	fail (1)
50-65	pass (2)
66-80	satisfactory (3)
80-89	good (4)
90-100	excellent (5)

Person responsible for course: László Dr. Juhász, associate professor, PhD, dr. habil.

Lecturer: László Dr. Juhász, associate professor, PhD, dr. habil

Title of course: Advanced seminar in organic chemistry Code: TTKBG0313_EN	ECTS Credit points: 2
Type of teaching, contact hours - lecture: - - practice: 2 hours/week - laboratory: -	
Evaluation: term mark	
Workload (estimated), divided into contact hours: - lecture: - - practice: 28 hours - laboratory: - - home assignment: 21 hours - preparation for the exam: - Total: 49 hours	
Year, semester: 2 nd year, 2 nd semester	
Its prerequisite(s): Organic Chemistry II. (lect .and sem.) TTKBE0302_EN	
Further courses built on it: -	

Topics of course
The aim of the course is to enable students to master the complex organic chemistry problem solving skills, and to be able to apply the knowledge acquired in basic courses in solving complex synthetic tasks and designing syntheses.
Literature
<i>mpulsory:</i> 22. Course material, concept and task collection for lectures, seminars in the e-learning system.
<i>commended:</i> 23. J. G. Smith: Organic Chemistry, 5 th Edition, 2016, McGraw Hill; ISBN-13: 9780077354725 24. J. McMurry: Organic Chemistry, 8 th Edition, 2012, Brooks/Cole; ISBN-13:

9780840054449

25. J. Clayden, N. Greeves, and S. Warren: Organic Chemistry, 2nd Edition, 2012, Oxford University Press; ISBN-13: 9780199270293
26. F. A. Carey: Organic Chemistry, 4th Edition, 2000, The McGraw-Hill Companies; ISBN-13: 9780072905014
27. L. G. Wade: Organic Chemistry, 8th Edition, 2012, Pearson; ISBN-13: 9780321768148
28. T. W. G. Solomons, C. Fryhle, Organic Chemistry, 10th Edition, 2009, Wiley & Sons; ISBN-10: 0470556595
29. H. Meislich, E. K. Meislich, J. Sharefkin: 3000 Solved Problems in Organic Chemistry, 1st Edition, 1994, McGraww-Hill Companies; ISBN-13: 978-0070564244

Course objective/intended learning outcomes

a) Knowledge

- He/she knows the structure, physical and chemical properties organic compounds and he/she is able to apply his/her knowledge to solve complex tasks on the field of this compounds. He/she knows the application and practical significance of these derivatives., and their applicabilities

b) Abilities

- He/she is able to participate in professional communication on the field of structure, and chemical transformation of this field.
- He/she is able to expand and/or develop his/her knowledge from the natural products.

c) Attitude

- He/she is open to getting new, scientifically proven knowledge on the subject, but to reject unsubstantiated or possibly misleading claims

d) Autonomy and responsibility

- He/she is able to independently perform the tasks of the course with professional guidance, and he/she can interpret and evaluate of the results obtained.

Schedule:

1st week

The basics of retrosynthetic analysis, the concept of synthones and retrones. Types of disconnections. Interconversion of functional groups. The use of the method in the exploration of simple synthetic possibilities for compounds.

2nd week

Retrosynthetic analysis of aromatic compounds. Use of the directing and activating/deactivating effects to form the appropriate substituent pattern.

3rd week

Methods for forming C-C bond I. Base catalyzed conversions I. (aldol condensation and its variants).

4th week

Methods for forming C-C bond II. Base catalyzed conversions II. (malonic ester and acetoacetic ester syntheses).

5th week

Methods for forming C-C bond III. Acid catalyzed transformations.

6th week

Methods for forming C-C bond IV. Possibilities for the formation and use of Grignard compounds.

7th week

Methods for forming C-C bond V. Transition metal (Pd, Pt, Ru, Cu, etc.) catalyzed conversions.

8th week

Methods for forming carbon-oxygen and carbon-sulfur bonds.

9th week

Possibilities for forming carbon-nitrogen bonds.

10th week

Reactions suitable for the synthesis of oxo compounds.

11th week

Reactions for the preparation of carboxylic acids and their derivatives.

12th week

Preparation and reactions of amino acids. Peptide synthesis.

13th week

The basic chemical properties of monosaccharides. Protecting Groups. Essential questions of synthesis of di- and oligosaccharides.

14th week

The synthesis of basic heterocycles and their chemical properties.

Requirements:

The course is recommended in parallel with the lecture Organic Chemistry III. (TTKBE0303_EN).

- *for a signature*

Attendance at **lectures** is recommended, but not compulsory.

Attendance at seminars is compulsory. A student may not miss the seminar more than three times during the semester. In case of further absences, a medical certificate needs to be presented. In case a student does not do this, the subject will not be signed, and the student must repeat the course.

- *for a grade*

The course ends in an **examination**.

The exam grade is the result of the written exam.

The minimum requirement for the examination respectively is 50%. The grade for the written exam is given according to the following table:

Score	Grade
0-49	fail (1)
50-62	pass (2)
63-75	satisfactory (3)
76-87	good (4)
88-100	excellent (5)

If the score of any test below 50, students can take a retake test in conformity with the EDUCATION AND EXAMINATION RULES AND REGULATIONS.

Person responsible for course: László Dr. Juhász, associate professor, PhD, dr. habil.

Lecturer: László Dr. Juhász, associate professor, PhD, dr. habil.

STAFF HANDBOOK

Name	<i>Prof. Dr. Gyula Batta</i>
Position	<i>Professor, Structural Chemistry and Biology</i>
Academic career	<i>Initial academic appointment (UD, 1980) Habilitation (UD, 2002)</i>
	<i>doctoral qualification (MTA DSc., 2001) Doctorate, Chemistry (MTA CSc. 1987) MSc., physics (Institution, 1976)</i>
Employment	<i>Professor - UD – from 2007</i>
Research and development projects over the last 5 years	<i>Name of project or research focus: Antifungal proteins: structure and dynamics, Glycopeptide Antibiotics 2013-2018 Partners, if applicable: Innsbruck Medical University, Institute of New Antibiotics, Moscow Amount of financing: 33M HUF (NKFI 110821) , ca. 150M HUF(GINOP)</i>
Industry collaborations over the last 5 years	<i>Project title: NMR Structure determination Partners: TEVA, CF Pharma, GLYCOM</i>
Patents and proprietary rights	<i>N/A</i>
Important publications over the last 5 years	<i>Selected recent publications from a total of approx. (200):</i> <ol style="list-style-type: none"> 1. Anna Huber, Dorottya Hajdu, Doris Bratschun-Khan, Zoltán Gáspári, Mihayl Varbanov, Stéphanie Philippot, Ádám Fizil, András Czajlik, Zoltán Kele, Christoph Sonderegger, László Galgóczy, Andrea Bodor, Florentine Marx, Gyula Batta <i>New antimicrobial potential and structural properties of PAFB: a cationic, cysteine-rich protein from Penicillium chrysogenum Q176</i> 2. Sándor Boros, Zoltán Gáspári, Gyula Batta <i>Accurate NMR determinations of proton–proton distances</i> ANNUAL REPORTS ON NMR SPECTROSCOPY 94: pp. 1-39. (2018) 3. Christoph Sonderegger, Ádám Fizil, Laura Burtscher, Dorottya Hajdu, Alberto Muñoz, Zoltán Gáspári, Nick D Read, Gyula Batta, Florentine Marx: <i>D19S mutation of the cationic, cysteine-rich protein PAF: novel insights into its structural dynamics, thermal unfolding and antifungal function</i> PLOS ONE 12:(1) Paper e0169920. 21 p.

	<p>(2017)</p> <p>4. Fizil Á, Gáspári Z, Barna T, Marx F, Batta G: "Invisible" Conformers of an Antifungal Disulfide Protein Revealed by Constrained Cold and Heat Unfolding, CEST-NMR Experiments, and Molecular Dynamics Calculations. CHEMISTRY-A EUROPEAN JOURNAL 21:(13) pp. 5136-5144. (2015)</p> <p>5. Váradi Györgyi, Tóth Gábor K., Kele Zoltán, Galgóczy László, Fizil Ádám, Batta Gyula: Synthesis of PAF, an Antifungal Protein from P. chrysogenum, by Native Chemical Ligation: Native Disulfide Pattern and Fold Obtained upon Oxidative Refolding, CHEMISTRY-A EUROPEAN JOURNAL 19:(38) pp. 12684-12692. (2013)</p>
Activities in specialist bodies over the last 5 years	<p>Hungarian NMR Society http://www.nmrmb.hu/ - President – from 2018</p> <p>Journal of Antibiotics, Member of Editorial Board</p>

Name	Gábor Bellér PhD
Position	Assistant Professor, UD
Academic career	
	<p>doctoral qualification: PhD, UD, Hungary, 2016</p> <p>Undergraduate degree, subject: Chemistry-english-hungarian special translator MSc, 2010</p>
Employment	<p>Assistant Professor, 2016-, UD</p> <p>Junior Assistant Professor, 2014-2016, UD</p> <p>Junior Research Fellow, 2013-2014, MTA-DE Homogeneous Catalysis and Reaction Mechanisms Research Group</p> <p>Visiting young researcher – University of Cádiz, 2010, 2 months</p>
Research and development projects over the last 5 years	<p>Name of project or research focus: TÁMOP-4.2.2.A-11/1/KONV,</p> <p>EU project: Basic targeted chemical and biological research for the elimination of environmentally hazardous chemicals Period and any other information: 2013-2015</p> <p>Partners, if applicable:</p> <p>Amount of financing:</p>

	<p>Name of project or research focus: GINOP-2.3.2-15-2016-00008,</p> <p>EU project: <i>Chemistry for better life: strategic R&D center at the University of Debrecen.</i></p> <p>Period and any other information: 2016-2020</p> <p>Partners, if applicable:</p> <p>Amount of financing:</p>
Industry collaborations over the last 5 years	
Patents and proprietary rights	
Important publications over the last 5 years	<p>Selected recent publications from a total of approx. (give total number): 7</p> <p>1. Mária Szabó, József Kalmár, Tamás Ditrói, <u>Gábor Bellér</u>, Gábor Lente, Nina Simic, István Fábián Equilibria and kinetics of chromium(VI) speciation in aqueous solution – a comprehensive study from pH 2 to 11 <i>Inorganica Chimica Acta</i>, 2018, 472, 295-301.</p> <p>2. Ferenc, Najóczki, <u>Gábor Bellér</u>, Mária Szabó, István Fábián Substituent effect on the N-oxidation of 1,10-phenanthroline derivatives by peroxomonosulfate ion <i>New Journal of Chemistry</i>, 2017, 41, 9947-9953.</p> <p>3. <u>Gábor Bellér</u>, Gábor Lente, István Fábián Kinetics and mechanism of the autocatalytic oxidation of bis(terpyridine)iron(II) by peroxomonosulfate ion (Oxone) in acidic medium <i>Inorganic Chemistry</i>, 2017, 56, 8270-8277.</p> <p>4. Mária Szabó, <u>Gábor Bellér</u>, József Kalmár, István Fábián The kinetics and mechanism of complex redox reactions in aqueous solution: The tools of the trade <i>Advances in Inorganic Chemistry</i>, 2017, 70, 1-61.</p> <p>5. <u>Gábor Bellér</u>, Mária Szabó, Gábor Lente, István Fábián Formation of 1,10-phenanthroline-N,N'-dioxide under mild conditions: the kinetics and mechanism of the oxidation of 1,10-phenanthroline by peroxomonosulfate ion (Oxone) <i>Journal of Organic Chemistry</i>, 2016, 81, 5345-5353</p>
Activities in specialist bodies over the last 5 years	

Name	<i>Attila Bényei</i>
Position	<i>Associate Professor, teaching physical chemistry, X-ray diffraction, polymorphism</i>
Academic career	<i>Habilitation (University of Debrecen, 2011)</i>
	<i>Doctor of University (University of Debrecen, 1990) PhD in Chemistry (University of Debrecen, 1995) Chemist, English translator (University of Debrecen, 1986)</i>
Employment	<i>Head of Laboratory for X-ray Diffraction, University of Debrecen, from 1995</i>
Research and development projects over the last 5 years	<i>Name of project or research focus: Single crystal X-ray diffraction studies. Determination of solid state structures of organic and organometallic compounds. GINOP-2.3.2-15-2016-00008 and GINOP-2.3.3-15-2016-00004, Participant scientist. Period and any other information: 2015-2019 Partners, if applicable: Amount of financing:</i>
Industry collaborations over the last 5 years	<i>Project title: Single crystal X-ray diffraction studies Partners: Alkaloida Research and Development Ltd.</i>
Patents and proprietary rights	--
Important publications over the last 5 years	<i>Selected recent publications from a total of approx. (110): Author(s): Buglyó P., Kacsir I., Kozsup M., Nagy I., Nagy S., Bényei A.C., Kováts É., Farkas E. Title: Tuning the redox potentials of ternary cobalt(III) complexes containing various hydroxamates Any other information: Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers INORGANICA CHIMICA ACTA 472: pp. 234-242. (2018) 2. Author(s): Buglyo P, Parajdi-Losonczai PL, Benyei AC, Lihi N, Biro L, Farkas E Title: Versatility of Coordination Modes in Complexes of Monohydroxamic Acids with Half-Sandwich Type Ruthenium, Rhodium, Osmium and Iridium Cations Any other information: Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers CHEMISTRYSELECT 2:(26) pp. 8127-8136. (2017)</i>

	<p>3. Author(s): Farkas G, Császár Z, Stágel K, Nemes E, Balogh S, Tóth I, Bényei A, Lendvay G, Bakos J Title: <i>Efficient stereochemical communication in phosphine-amine palladium-complexes: exploration of N-substituent effects in coordination chemistry and catalysis</i> Any other information: Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers JOURNAL OF ORGANOMETALLIC CHEMISTRY 846: pp. 129-140. (2017)</p> <p>4. Author(s): Matyuska F, Szorcsik A, May NV, Dancs Á, Kováts É, Bényei A, Gajda T Title: <i>Tailoring the local environment around metal ions: A solution chemical and structural study of some multidentate tripodal ligands</i> Any other information: Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers DALTON TRANSACTIONS 46:(26) pp. 8626-8642. (2017)</p> <p>5. Author(s): Illyés TZ, Balla S, Bényei A, Kumar AA, Timári I, Kövér KE, Szilágyi L Title: <i>Exploring the Syntheses of Novel Glycomimetics. Carbohydrate Derivatives with Se-S- or Se-Se- Glycosidic Linkages</i> Any other information: Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers CHEMISTRYSELECT 1:(10) pp. 2383-2388. (2016)</p> <p>6. Author(s): Rodríguez-Rodríguez A, Regueiro-Figueroa M, Esteban-Gómez D, Tripier R, Tircsó G, Kálmán FK, Bényei AC, Tóth I, de Blas A, Rodríguez-Blas T, Platas-Iglesias C Title: <i>Complexation of Ln³⁺ Ions with Cyclam Dipicolinates: A Small Bridge that Makes Huge Differences in Structure, Equilibrium, and Kinetic Properties</i> Any other information: Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers INORGANIC CHEMISTRY 55:(5) pp. 2227-2239. (2016)</p>
Activities in specialist bodies over the last 5 years	Organisation - Role – Period International Union of Crystallography, individual

	<i>member, 2004- European Crystallographic Association, SIG13 member 2000-</i>
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Name	<i>Péter Buglyó</i>
Position	<i>associate professor</i>
Academic career	<i>1989 research fellow, Dept. of Inorg. and Anal. Chem., L. Kossuth University 1990 research fellow, Hungarian Academy of Sciences 1992 assistant professor, Dept. of Inorg. and Anal. Chem., L. Kossuth University 1996 postdoctoral fellow, University of British Columbia 1998 assistant professor, Dept. of Inorg. and Anal. Chem., L. Kossuth University 2000 assistant professor, Dept. of Inorg. and Anal. Chem., L. Kossuth University 2006 associate professor, Dept. of Inorg. and Anal. Chem., University of Debrecen</i>
	<i>M.Sc. Chemistry, 1989 Lajos Kossuth University, Debrecen Ph.D. Chemistry, 1993 Lajos Kossuth University, Debrecen Dr.Habil. Chemistry, 2003 University of Debrecen</i>
Employment	<i>1989 research fellow, Dept. of Inorg. and Anal. Chem., L. Kossuth University 1990 research fellow, Hungarian Academy of Sciences 1992 assistant professor, Dept. of Inorg. and Anal. Chem., L. Kossuth University 1996 postdoctoral fellow, University of British Columbia 1998 assistant professor, Dept. of Inorg. and Anal. Chem., L. Kossuth University 2000 assistant professor, Dept. of Inorg. and Anal. Chem., L. Kossuth University 2006 associate professor, Dept. of Inorg. and Anal. Chem., University of Debrecen</i>
Research and development projects over the last 5 years	<i>Chemistry for better life: strategic R&D center at the university of Debrecen (DECHEM), 01/10/2016- 30/09/2020, GINOP-2.3.2-15-2016-00008, 1 984 M HUF, subproject leader</i>

	<p><i>Synthesis and Characterization of bimetallic complexes as potential hypoxia-activated prodrugs with dual anticancer activity, OTKA 112317, 2015-2018, 24.3 M HUF, principal investigator</i></p> <p><i>Synthesis and Characterization of Novel Ruthenium-Hydroxamate Complexes, OTKA 76142, 2009-2013 11.8 M HUF, principal investigator</i></p>
Industry collaborations over the last 5 years	-
Patents and proprietary rights	-
Important publications over the last 5 years	<p><i>Selected recent publications from a total of approx. 81:</i></p> <p>J. Patalenszki, L. Bíró, A. C. Bényei, T. R. Muchova, J. Kasparkova, P. Buglyó, Half-sandwich complexes of ruthenium, osmium, rhodium and iridium with DL-methionine or S-methyl-L-cysteine: A solid state and solution equilibrium study, <i>RSC Advances</i>, 2015, 5, 8094-8107</p> <p>P. L. Parajdi-Losoncz, A. C. Bényei, É. Kováts, I. Timári, T. R. Muchova, J. Kasparkova, P. Buglyó, [(η^6-p-cym)Ru(H₂O)₃]²⁺ Binding Capability of Aminohydroxamates – A Solution and Solid State Study, <i>J. Inorg. Biochem.</i>, 2016, 160, 236–245</p> <p>Zs. Bihari, F. Vultós, C. Fernandes, L. Gano, I. Santos, J. D. G. Correia, P. Buglyó, Synthesis, characterization and biological evaluation of a ⁶⁷Ga-labeled (η^6-Tyr)Ru(η^5-Cp) peptide complex with the HAV motif, <i>J. Inorg. Biochem.</i>, 2016, 160, 189–197</p> <p>E. Farkas, P. Buglyó, Lead(II) Complexes of Amino Acids, Peptides, and Other Related Ligands of Biological Interest, in <i>Lead: Its Effects on Environment and Health</i>, <i>Met. Ions Life Sci.</i> 2017, 17, 201–240</p> <p>D. Sanna, V. Ugone, G. Sciortino, P. Buglyó, Zs. Bihari, P. L. Parajdi-Losoncz, E. Garribba, Complexes of V^{IV}O²⁺ ion with antibacterial quinolones ligands and their interaction with serum proteins, <i>Dalton Trans.</i>, 2018, 47, 2164-2182</p>
Activities in specialist bodies over the last 5 years	<p><i>Secretary of the Chemical Institute, University of Debrecen, 2010-2017</i></p> <p><i>Advisory Board member, COST CM1115, 2013-2016</i></p> <p><i>Elected member of the Physical Chemistry Scientific Committee of the Hungarian Academy of Sciences 2014-2017</i></p> <p><i>Elected secretary of the Physical Chemistry Scientific Committee of the Hungarian Academy of Sciences 2018-</i></p>

Name	<i>Ágnes Dávid</i>
Position	<i>Assistant lecturer</i>
Academic career	
	<p><i>doctoral qualification (Institution of Chemistry, University of Debrecen, 2017)</i></p> <p><i>Doctorate, subject: bioinorganic and coordination chemistry (Institution of Chemistry, University of Debrecen, 2017)</i></p> <p><i>Undergraduate degree, subject: chemist (University of Debrecen, 2012)</i></p> <p><i>chemistry and hungarian literature and grammar teacher (University of Debrecen, 2014)</i></p>
Employment	<i>assistant lecturer (Institution of Chemistry, University of Debrecen, from 2016 until now)</i>
Research and development projects over the last 5 years	<p><i>Name of project or research focus: –</i></p> <p><i>Period and any other information: –</i></p> <p><i>Partners, if applicable: –</i></p> <p><i>Amount of financing: –</i></p>
Industry collaborations over the last 5 years	<p><i>Project title: –</i></p> <p><i>Partners: –</i></p>
Patents and proprietary rights	<i>Title (Year):–</i>
Important publications over the last 5 years	<p><i>Selected recent publications from a total of approx. (give total number): 2</i></p> <p><i>Author(s): Ágnes Dávid, Éva Tünde Hartman, Norbert Lihi, Imre Sóvágó and Katalin Várnagy</i></p> <p><i>Title: Complex formation of nickel(II) and zinc(II) ions with the peptide fragments of rat amylin</i></p> <p><i>Any other information: New J. Chem., 2018, 42, 8277.</i> IF: 3,269 DOI: 10.1039/c8nj90036a</p> <p><i>Author(s): Ágnes Dávid, Csilla Kállay, Daniele Sanna, Norbert Lihi, Imre Sóvágó, Katalin Várnagy</i></p> <p><i>Title: Potentiometric and spectroscopic studies on the copper(II) complexes of rat amylin fragments. The anchoring ability of specific non-coordinating side chains</i></p> <p><i>Any other information: Dalton Trans., 2015, 44, 17091–17099. IF: 4,197 DOI: 10.1039/c1dt10835b</i></p>
Activities in specialist bodies over the last 5 years	–

Name	György Deák PhD
Position	associate professor, deputy head of department, coordinator of Chemical Engineering BSc and MSc.
Academic career	Habilitation: University of Debrecen, 2006 PHD. University of Debrecen, Debrecen, 1997 "University doctor" Kossuth Lajos University, Debrecen, 1981
	<i>Undergraduate degree: Chemistry MSc, Debrecen, 1978</i> <i>Kossuth Lajos Secondary Grammar School, Tiszafüred, 1972</i> <i>Kiss Pál Elementary School, Tiszafüred, 1968</i>
Employment	Associate professor, University of Debrecen, 2007- Assistant professor, Lajos Kossuth University, 1993-2007 Visiting scientist, University of Akron, Akron OH, 1993-1995 Assistant research fellow, Lajos Kossuth University, 1991-1993 Head of the Pilot Plant, Biogal Pharmaceutical Works, Debrecen, 1989-1991 Visiting scientist, University of Brooklyn, NY, 1988-1989 Development engineer, Biogal Pharm Works, Debrecen, 1981-1988 Assistant research fellow, Lajos Kossuth University, 1978-1981
Research and development projects over the last 5 years	OTKA K116465, OTKA K101850, OTKA K109006, TÁMOP-4.2.2.A-11/1/KONV-2012-0036, Smart functional materials, 2012-2015 GINOP-2.3.2-15-2016-00041, Regional Workshop for Excellence in Materials Science, 2017-2020
Industry collaborations over the last 5 years	TEVA Pharmaceutical Works, Debrecen, 2014. Evonik Agroferm ZRt. Nádudvar, 2016. Coloplast, Nyírbátor, 2017. Water Technologies and Solutions, Tatabánya, 2018.
Patents and proprietary rights	1. Majoros I.(20%), Nemes S. (17,5%), Elek S. (13,5%), Mihók M. (12%), Mihók Mné (12%), Zsuga M. (5%), Deák Gy. (5%), Sályi Sz. (5%): Method for the processing of organic mother liquor formed by the preparation of 7-amino-cephalosporic acid. <i>Hungarian patent: HU 199 483 (1990)</i>

	<p>2. Zahoránszky L. (10%), Deák Gy. (10%), Radnai F. (10%), Ockenfusz Cs. (27%), Vas Gy. (10%), Deák L. (10%), Móricz K. (3%), Varga I. (3%), Szőlósi J. (3%), Sipos S. (3%), Sarkadi F (2%), Jóna J. (2%), Szabó L. (2%), Juhász Z. (2%), Ágni Zs. (3%) : Method for the purification of waste solvents or solvent mixtures containing nitrous gases. <i>Hungarian patent: HU 200 283 (1990)</i></p> <p>3. Szánya T. (22%), Hanák L. (18%), Nagyné Árvai E. (14%), Kéri V. (13%), Marton Gy. (5%), Bálint J. (4%), Melczer I. (4%), Deák Gy. (3%), Karácsony E. (3%), Radnai F. (3%), Dencs Béláné (2%), Hajdúfi Cs. (2%), Karczub A. (2%), Makó Gy-né (2%), Strbka A-né (2%), Kelemen J. (1%): Process for the Purification of Cyclosporin A. <i>US patent: 5 382 655 (1995)</i></p> <p>4. Melczer I. (12%), Kéri V. (18%), Dr. Bálint J. (6%), Hógye I. (4%), Dr. Nagy L. (13%), Dr. Deák Gy. (4%), Budai J. (11%), Tamás T. (9%), Miskolczy I. (13%), Lajter J. (3%), Széles L. (7%): Process for removing biomass from broth containing antibiotics <i>Hungarian patent: HU 213 884 (1997)</i></p> <p>5. Szánya T. (22%), Hanák L. (18%), Nagyné Árvai E. (14%), Kéri V. (13%), Marton Gy. (5%), Bálint J. (4%), Melczer I. (4%), Deák Gy. (3%), Karácsony E. (3%), Radnai F. (3%), Dencs Béláné (2%), Hajdúfi Cs. (2%), Karczub A. (2%), Makó Gy-né (2%), Strbka A-né (2%), Kelemen J. (1%): Method of the purification of cyclosporine A. <i>Hungarian patent: HU 213 553 (1997)</i></p> <p>6. Kennedy J. P. (25%), Zsuga M. (25%), Kéki S. (25%), Deák Gy. (25%): Star-Forming Glassy- Rubbery Copolymers <i>US patent 6 130 291 (2000. 10. 10.)</i></p>
<p>Important publications over the last 5 years</p>	<p><i>Selected recent publications from a total of approx. 40:</i></p> <p>[1] Hermenean, A; Stan, M; Ardelean, A; Pilát, L; Mihali, CV; Popescu, C; Nagy, L; Deak, G; Zsuga, M; Keki, S; Bacskay, I; Fenyvesi, F; Costache, M; Dinischiotu, A; Vecsernyes, M: Antioxidant and hepatoprotective activity of milk thistle (<i>Silybum marianum</i> L. Gaertn.) seed oil, <i>Open Life Sciences</i>, 10(1), 225-236 (2015)</p>

	<p>IF: 0,78</p> <p>[2] István Lázár, Helga Fruzsina Bereczki, Sándor Manó, Lajos Daróczi, György Deák, István Fábián, Zoltán Csernátony: Synthesis and study of new functionalized silica aerogel poly(methyl methacrylate) composites for biomedical use, <i>Polymer Composites</i>, 36(2), 348-358 (2015) IF: 1,48</p> <p>[3] Nagy L, Nagy T, Deák Gy, Kuki Á, Purgel M, Narmandakh M, Iván B, Zsuga M, Kéki S, Can non-polar polyisobutylenes be measured by electrospray ionization mass spectrometry? Anion-attachment proved to be an appropriate method, <i>J. Am. Soc. Mass Spectrom.</i>, 2016, 27, 432.</p> <p>[4] Nagy, L; Kuki, A; Deak, G; Purgel, M; Vekony, A; Zsuga, M; Keki, S, Gas-Phase Interaction of Anions with Polyisobutylenes: Collision-Induced Dissociation Study and Quantum Chemical Modeling, <i>J. Phys. Chem. B</i>, 2016, 120, 9195.</p> <p>[5] Kocsis D; Deák G; Kéki S; Godó Z.A; Horváth R: Creep and Quasi-Relaxation Examination of Artificially Aged Plasticized PVC, <i>Journal Of Testing And Evaluation</i>, 45:(4), 1213-1221 (2017). IF: 0,39, Q2</p>
Activities in specialist bodies over the last 5 years	Member of the Committee of Physical Chemistry of the Hungarian Academy of Sciences

Name	Gábor Dobosi, DSc
Position	<i>Professor, Geology</i>
Academiccareer	<i>Professor (UD, Department of Mineralogy and Geology, 2014-)</i> <i>Habilitation (UD, 2012)</i>
	<i>Doctoral qualification:</i> <i>DSc, Hungarian Academy of Sciences, 2004</i> <i>PhD (CSc), Hungarian Academy of Sciences, 1994</i> <i>Doctorate, subject (Institution, year)</i> <i>Undergraduate degree: (subject (Institution, year):</i> <i>Chemistry MSc, ELTE 1975</i> <i>Geology MSc, ELTE 1979</i>
Employment	<i>From assistant research fellow to scientific advisor - Hungarian Academy of Sciences - 1975 - 2014</i> <i>Professor, UD Department of Mineralogy and Geology,</i>

	<i>2014 - present</i>
Research and development projects over the last 5 years	<i>Name of project or research focus: Period and any other information: Partners, if applicable: Amount of financing:</i>
Industry collaborations over the last 5 years	<i>Project title: Partners:</i>
Patents and proprietary rights	<i>Title (Year)</i>
Important publications over the last 5 years	<i>Selected recent publications from a total of approx. (give total number): 14</i> Jankovics MÉ, Taracsák Z, <u>Dobosi G</u> , Embey-Isztin A, Batki A, Harangi Sz, Hauenberger CA Clinopyroxene with diverse origins in alkaline basalts from the western Pannonian Basin: implications from trace element characteristics LITHOS 262: pp. 120-134 (2016) <u>Dobosi G</u> , Harangi Sz Hungarian National Report on IAVCEI (2011-2014) GEOMATIKAI KÖZLEMÉNYEK 18:(1) pp. 137-145. (2015) Downes H, Carter A, Armstrong R, <u>Dobosi G</u> , Embey-Isztin A Lower crustal zircons reveal Neogene metamorphism beneath the Pannonian Basin (Hungary) OPEN GEOSCIENCES 7:(1) pp. 223-233. (2015) Batki A, Pál-Molnár E, <u>Dobosi G</u> , Skelton A Petrogenetic significance of ocellar camptonite dykes in the Ditrău Alkaline Massif, Romania LITHOS 200-201: pp. 181-196. (2014) Embey-Isztin A, <u>Dobosi G</u> , Bodinier J-L, Bosch D, Jenner G A, Pourtales S, Bruguier O Origin and significance of poikilitic and mosaic peridotite xenoliths in the western Pannonian Basin: geochemical and petrological evidences CONTRIBUTIONS TO MINERALOGY AND PETROLOGY 168:(3) pp. 1-16. (2014)
Activities in specialist bodies over the last 5 years	<i>Central European Geology (AK journals), member of the Editorial Board, 2012 - present</i>

Name	Attila Gáspár
Position	Professor
Academic career	Doctor of Science (HAS, 2015) Habilitation (UD, 2005)

	<p>PhD (UD, 1997) chemist, chemistry teacher, translator of English (UD, 1994)</p>
Employment	<p>University of Debrecen, Dept. of Inorganic and Analytical Chemistry (1997-)</p>
Research and development projects over the last 5 years	<p>„Chemistry for better life: strategic R&D center at the University of Debrecen” GINOP-2.3.2-15-2016-00008 project Support: 1.983.995.445 Ft Number of participants: 139 Duration: 2016.10.01-2020.09.30 website for further information: www.dechem.unideb.hu</p> <p>„Integrated Instrumental Infrastructure for Research in Molecular Science and Molecular Medicine (I2M2)” GINOP-2.3.3-15-2016-00004 project Support: 807 759 930 Ft Number of participants: 71 Duration: 2016.09.01-2019.06.30. website for further information: www.i2m2.unideb.hu</p> <p>OTKA K75286, 2008-10-01 2013-10-31 9 Title: Fabrication of new types of microfluidic devices and the study of their capabilities for liquid chromatographic, electrophoretic and electrochromatographic separation</p> <p>OTKA K111932, 2015-01-01 2018-12-31 Title: Analytical separations in microfluidic chips 19701000 HUF</p>
Industry collaborations over the last 5 years	<p>Project title: Capillary electrophoretic analysis of monoclonal antibodies Partner: Richter G. Plc.</p>
Patents and proprietary rights	<p>A.Gáspár, H.Berndt: "Vorrichtung zum Atomisieren von flüssigen Proben", German patent, DE 19944650.4 Gomez, F.A, Gaspar, A., Piyasena, M.E., Magnetically controlled valve for flow manipulation in polymer microfluidic devices, US Patent application, #44,228 2006. Patent No. 2008/0163,946</p>
Important publications over the	<p>A.Kiss, A.Gaspar: Fabrication of a Microfluidic Flame</p>

last 5 years	<p>Atomic Emission Spectrometer: a Flame-on-a-Chip, Anal.Chem., 2018, 90, 5995-6000</p> <p>A.Kecskemeti, A.Gaspar: Particle based liquid chromatographic separations in microfluidic devices, Anal.Chim.Acta, 2018, 1021, 1-19</p> <p>A.Kecskemeti, A.Gaspar: Particle-based immobilized enzymatic reactors in microfluidic chips, Talanta, 2018, 180, 211-228</p> <p>A.Kecskemeti, C.N.Nagy, G.Kallo, E.Csoz, A.Gaspar: The application of a microfluidic reactor including spontaneously adsorbed trypsin for rapid protein digestion of human tear samples, Proteomics Clin. Appl., 2017, 11, 1700055</p> <p>A.Kecskemeti, J.Bako, I.Csarnovics, E.Csoz, A.Gaspar: The application of non-covalently immobilized trypsin in a poly(dimethylsiloxane) microfluidic device for rapid protein digestion, Anal. Bioanal. Chem., 2017, 409, 3573-3585</p> <p>A.Kecskemeti, A.Gaspar: Preparation and characterization of a packed bead immobilized trypsin reactor integrated into a PDMS microfluidic chip for rapid protein digestion, Talanta, 2017, 166, 275–283</p>
Activities in specialist bodies over the last 5 years	<p>MTA Work Committee for Analysis of Organic Materials and Pharmaceuticals (2005-)</p> <p>Hungarian Society for Separation Sciences (2005-)</p> <p>CEEPUS H-076 program (2001-)</p>

Name	<i>Dr. Gyöngyi Gyémánt</i>
Position	<p><i>Teaching area and designation</i></p> <p><i>Biochemistry for pharmacists and chemical engineers.</i></p> <p><i>Biochemistry practicals for biologists and chemists.</i></p> <p><i>Bioanalytical courses: Chromatography, MS, Carbohydrate analysis</i></p>
Academic career	<p><i>Associate Professor of Analytical Chemistry (UD, 2011-)</i></p> <p><i>Habilitation (UD, 2009)</i></p>
	<p><i>PhD in Chemistry, 2002, University of Debrecen</i></p> <p>MSc. in Chemistry, 1983, University of Lajos Kossuth, Debrecen</p> <p>High School Gárdonyi Géza Gimnázium, Eger, 1978</p>
Employment	<p><i>Assistant Professor of Analytical Chemistry,(UD, 2009-2011)</i></p> <p><i>Assistant Professor of Biochemistry, (UD, 2002-2009)</i></p>

	<p><i>Assistant Lecturer, University of Lajos Kossuth, 1994-2002</i></p> <p><i>Analytical engineer, BIOGAL Pharm. Work 1983-1994</i></p>
Research and development projects over the last 5 years	<p><i>Structure and Function of the Antifungal Proteins PAFB and NFAP, OTKA 2014-2017 33M HUF, participant</i></p> <p><i>Analytical separations on microchip, OTKA, 2015-2018, 19.7 M HUF, participant</i></p>
Industry collaborations over the last 5 years	<p><i>Project title: Structure and interaction of CD derivatives and complexes</i></p> <p><i>Partners: Cyclolab 2015-</i></p>
Patents and proprietary rights	<p><i>Use of siderophores in prevention of vascular diseases, production of siderophores and qualifying siderophore containing meat-products 2008</i></p> <p><i>Novel preparation. of polyvinyl-pyrrolidone-iodine complex - and surgical hand cleanser containing it 1988</i></p> <p>.</p>
Important publications over the last 5 years	<p><i>Selected recent publications from a total of approx. 80</i></p> <p><i>Gábor Lehoczki, Lili Kandra, Gyöngyi Gyémánt</i></p> <p><i>The use of starch azure for measurement of alpha-amylase activity Carbohydrate Polymers, 183 (2018) 263-266. IF:4,811</i></p> <p><i>Lehoczki G, Szabó K, Takács I, Kandra L, Gyémánt G</i></p> <p><i>Simple ITC method for activity and inhibition studies on human salivary α-amylase, JOURNAL OF ENZYME INHIBITION AND MEDICINAL CHEMISTRY 31:(6) (2016) 1648-1653. IF: 4,293</i></p> <p><i>Judit R. Homoki, Andrea Nemes, Erika Fazekas, Gyöngyi Gyémánt, Péter Balogh, Ferenc Gál, Jamil Al-Asri, Jérémie Mortier, Gerhard Wolber, László Babinszky and Judit Remenyik, Anthocyanin composition, antioxidant efficiency, and α-amylase inhibitor activity of different Hungarian sour cherry varieties (<i>Prunus cerasus</i> L.), Food Chemistry, 194, (2016) 222-229. IF: 3,391</i></p> <p><i>Jamil Al-Asri, Erika Fazekas, Gábor Lehoczki, Andrej Perdih, Cornelia Görick, Matthias F. Melzig, Gyöngyi Gyémánt, Gerhard Wolber, Jérémie Mortier</i></p> <p><i>From Carbohydrates to Drug-Like Fragments: Rational Development of Novel α-Amylase Inhibitors, Bioorganic and Medicinal Chemistry, 23 (2015) 6725-6732., IF: 2,793</i></p> <p><i>B Bécsi, D Dedinszki, G Gyémánt, C Máthé, G Vasas, B Lontay, F Erdődi, Identification of protein phosphatase interacting proteins from normal and UVA-irradiated HaCaT cell lysates by surface plasmon resonance</i></p>

	<p><i>based binding technique using biotin-microcystin-LR as phosphatase capturing molecule, Journal of Photochemistry and Photobiology B: Biology 138 (2014) 240-248. IF:3,11</i></p> <p><i>E. Fazekas, K. Szabó, L. Kandra, G. Gyémánt, Unexpected mode of action of sweet potato β-amylase on maltooligomer substrates, Biochimica et Biophysica Acta (BBA) - Proteins and Proteomics, 1834, (2013),1976-1981, IF: 3,73</i></p>
Activities in specialist bodies over the last 5 years	<p><i>Ad hoc reviewer for the following journals and organizations:</i></p> <p><i>Journals.</i></p> <ul style="list-style-type: none"> • <i>Acta Biologica Hungarica</i> • <i>Enzyme and Microbial Technology</i> • <i>Fitoterapia</i> • <i>Carbohydrate Research</i> • <i>Reaction Kinetics, Mechanisms and Catalysis.</i> <p><i>Organisations</i></p> <ul style="list-style-type: none"> • <i>Hungarian Scientific Research Fund (OTKA)</i> • <i>NKTH</i>

Name	<i>László Dr. Juhász Ph.D., dr. habil</i>
Position	<i>associate professor, in charge of organic chemistry of chemist, chemical engineer and bioengineer BSc</i>
Academic career	<p><i>education leader of the Chemistry Institute (oct. of 2008. – jan. of 2013.)</i></p> <p><i>Habilitation (University of Debrecen, 2015)</i></p>
	<p><i>doctoral qualification: Ph.D.: (University of Debrecen, 2001)</i></p> <p><i>Undergraduate degree, subject: Chemist M.Sc. (Lajos Kossuth University, 1996)</i></p>
Employment	<p><i>Research assistant, University of Debrecen; Jan. of 2000. – Aug. of 2000.</i></p> <p><i>Assistant lecturer, University of Debrecen; Sept. of 2000. -June. of. 2007.</i></p> <p><i>Assistant professor, University of Debrecen; July of 2007. – Aug.of 2015.</i></p> <p><i>Associate professor, University of Debrecen; from Sept. of 2015</i></p>
Research and development projects over the last 5 years	<p><i>Name of project or research focus: New transformations of monosaccharide derivatives at and around the anomeric centre - researcher</i></p> <p><i>Period and any other information: 08. of 2013. – 08 of 2018.</i></p> <p><i>Partners, if applicable:</i></p> <p><i>Amount of financing: 43 852 000 HUF</i></p>

	<p>Name of project or research focus: Chemistry for better life: strategic R &D center at the University of Debrecen (DECHEM; GINOP-2.3.2- 15-2016- 00008) - researcher</p> <p>Period and any other information: oct. of 2016. – sept of 2020.</p> <p>Partners, if applicable:</p> <p>Amount of financing: 1 983 995 445 HUF</p>
Industry collaborations over the last 5 years	<p>Project title: -</p> <p>Partners: -</p>
Patents and proprietary rights	<p>L. Somsák, É., Bokor, M. Tóth, L. Juhász, K. Czifrák, B. Kónya, S. Kun, A. Páhi, B. Szócs, G. Varga, L. Kóder, K. Nagy, P. Gergely, T. Docsa: Glikogén foszforiláz inhibitorok. P1100602/P1200475 Hungarian patent application. 2011.PCT/HU2012/000116</p>
Important publications over the last 5 years	<p>Selected recent publications from a total of approx. (give total number): 26</p> <p>L. Lázár; L. Juhász; Gy. Batta, A. Borbás, L. Somsák; Unprecedented β-manno type thiodisaccharides with a C-glycosylic function by photoinitiated hydrothiolation of 1-C-substituted glycals, New Journal of Chemistry, 2017, 41, 1284-1292</p> <p>J. József, L. Juhász, T. Z. Illyés, M. Csávás, A. Borbás, L. Somsák; Photoinitiated hydrothiolation of pyranoid exo-glycals: the D-galacto and D-xylo cases, <i>Carbohydrate Research</i>, 2015, 413, 63-69</p> <p>J. Begum, G. Varga, T. Docsa, P. Gergely, J.M. Hayes, L. Juhász, L. Somsák; Computationally motivated synthesis and enzyme kinetic evaluation of N-(β-D-glucopyranosyl)-1,2,4-triazolecarboxamides as glycogen phosphorylase inhibitors <i>MedChemComm</i>, 2015, 6(1), 80-89</p> <p>L. Juhász, G. Varga, A. Sztankovics, Ferenc Béke, T. Docsa, A. Kiss-Szikszej, P. Gergely, J. Kóňa, I. Tvaroška, L. Somsák; Structure-activity relationships of glycogen phosphorylase inhibitor FR258900 and its analogues: a combined synthetic, enzyme kinetic and computational study. <i>ChemPlusChem</i>, 2014, 79(11), 1558 – 1568</p> <p>M. Polyák, G. Varga, B. Szilágyi, L. Juhász, T. Docsa, P.</p>

	<i>Gergely, J. Begum, J. M. Hayes, L. Somsák; Synthesis, enzyme kinetics and computational evaluation of N-(β-D-glucopyranosyl) oxadiazolecarboxamides as glycogen phosphorylase inhibitors; Bioorganic and Medicinal Chemistry, 2013, 21, 5738 – 5747.</i>
Activities in specialist bodies over the last 5 years	<i>Organisation - Role – Period Membership without a specific role need not be mentioned.</i>

Name	<i>Csilla Kállay PhD</i>
Position	<i>Senior research fellow</i>
Academic career	<i>Habilitation (UD, 2017)</i>
	<i>doctoral qualification: PhD (UD, 2006) Doctorate, subject: coordination chemistry (UD, 2002-2005) Undergraduate degree, subject: Chemistry MSc (UD, 2002)</i>
Employment	<i>Assistant research fellow: 2005-2006 (UD) Postdoctoral Fellow: 2007-2011 (UD) Research fellow: 2011 (UD) Senior research fellow: 2012- (MTA-DE Reaction Mechanisms Research Group)</i>
Research and development projects over the last 5 years	<i>Name of project or research focus: OTKA 105156 Environmentally relevant redox reactions: kinetics and mechanism Period and any other information: 2013-2017 (Investigator) Partners, if applicable: Amount of financing: 66.521 MFt</i> <i>Name of project or research focus: OTKA 115480 Coordination and electrochemical background of role of metal ions in neurodegenerative disorders Period and any other information: 2016-2019 (Investigator) Partners, if applicable: Amount of financing: 25.704 MFt</i> <i>Name of project or research focus: OTKA 124983 The redox reactions of strong inorganic oxidants:</i>

	<p>kinetics and mechanism</p> <p><i>Period and any other information: 2017-2021 (Investigator)</i></p> <p><i>Partners, if applicable:</i></p> <p><i>Amount of financing: 47.712 MFt</i></p> <p><i>Name of project or research focus: CNR-MTA bilateral agreement; Peptide based nanostructures as theranostic tools for biological systems</i></p> <p><i>Period and any other information: 2016-2018</i></p> <p><i>Partners, if applicable:</i></p> <p><i>Amount of financing: 3.265 MFt</i></p>
Industry collaborations over the last 5 years	<p><i>Project title: -</i></p> <p><i>Partners: -</i></p>
Patents and proprietary rights	<p><i>Title (Year)</i></p>
Important publications over the last 5 years	<p><i>Selected recent publications from a total of approx. (give total number): 8</i></p> <p><i>Author(s): Norbert Lihi, Márton Lukács, Mária Raics, Györgyi Szunyog, Katalin Várnagy, Csilla Kállay</i></p> <p><i>Title: The effect of carboxylate groups on the complexation of metal ion with oligopeptides – Potentiometric investigation</i></p> <p><i>Any other information:</i></p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: Inorg. Chim. Acta, 472 (2018) 165-173.</i></p> <p><i>Author(s): Gizella Csire, Lajos Nagy, Katalin Várnagy and Csilla Kállay</i></p> <p><i>Title: Copper(II) Interaction with the Human Prion 103-112 Fragment – Coordination and Oxidation</i></p> <p><i>Any other information:</i></p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: J. Inorg. Biochem., 170 (2017) 195–201.</i></p> <p><i>Author(s): Ágnes Dávid, Csilla Kállay, Daniele Sanna, Norbert Lihi, Imre Sóvágó and Katalin Várnagy</i></p> <p><i>Title: Potentiometric and spectroscopic studies on the copper(II) complexes of rat amylin fragments. The anchoring ability of specific non-coordinating side</i></p>

	<p><i>chains</i></p> <p><i>Any other information:</i></p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers:</i></p> <p><i>Dalton Trans., 44 (2015) 17091–17099.</i></p>
Activities in specialist bodies over the last 5 years	-

Name	<i>Ferenc Krisztián Kálmán PhD</i>
Position	<i>Assistant professor</i>
Academic career	
	<p><i>Doctoral qualification: PhD, University of Debrecen (UD), Hungary, 2008</i></p> <p><i>Undergraduate degree: Chemist MSc, UD, 2002</i></p>
Employment	<p>2006-2007: Research assistant (UD, Department of Inorganic and Analytical Chemistry)</p> <p>2010-2013: Postdoctor (UD, Department of Inorganic and Analytical Chemistry, PD-OTKA 83253)</p> <p>2013-2014: Researcher (UD, Department of Inorganic and Analytical Chemistry)</p> <p>2014-2018: Researcher (UD, Department of Inorganic and Analytical Chemistry)</p> <p>2018-: Assistant professor (UD, Department of Physical Chemistry)</p>
Research and development projects over the last 5 years	<p><i>Name of project or research focus: OTKA 109029</i></p> <p><i>Period and any other information: 2013-2017</i></p> <p><i>Partners, if applicable: -</i></p> <p><i>Amount of financing: -</i></p> <p><i>Name of project or research focus: OTKA 120224</i></p> <p><i>Period and any other information: 2016-2020</i></p> <p><i>Partners, if applicable: -</i></p> <p><i>Amount of financing: -</i></p> <p><i>Name of project or research focus: TÁMOP-4.2.2.A-11/1/KONV-2012-0043</i></p> <p><i>Period and any other information: 2013-2015</i></p> <p><i>Partners, if applicable: -</i></p>

	<p><i>Amount of financing: -</i></p> <p><i>Name of project or research focus: GINOP-2.3.2-15-2016-00008</i></p> <p><i>Period and any other information: 2016-2020</i></p> <p><i>Partners, if applicable: -</i></p> <p><i>Amount of financing:-</i></p>
Industry collaborations over the last 5 years	<p><i>Project title: -</i></p> <p><i>Partners: -</i></p>
Patents and proprietary rights	<p>Botár Richárd, Garda Zoltán, Fodor Tamás, Kálmán Ferenc Krisztián, Nagy Viktória, Tircsó Gyula, Tóth Imre: Triaza-ciklononán alapú vegyületek és alkalmazásuk ligandumként Mn(II)-tartalmú MRI kontrasztanyagban: Triaza-cyclononane derivatives and their use as ligands in Mn(II)-based MRI contrast agents Patent number: HU2015000563A2 (2017)</p> <p>Botár Richárd, Garda Zoltán, Fodor Tamás, Kálmán Ferenc Krisztián, Nagy Viktória, Tircsó Gyula, Tóth Imre: New 3,6,9,15-tetraaza-bicyclo [9.3.1]pentadeca-1(14), 11(15), 12-triene based compounds and their application as ligands of essential metal ion based MRI and ⁵²Mn based PET contrast agents Patent number: WO2017089847A1 (2017)</p> <p>Botár Richárd, Garda Zoltán, Fodor Tamás, Kálmán Ferenc Krisztián, Nagy Viktória, Tircsó Gyula, Tóth Imre: New 2,11-diaza-[3.3](2,6)pyridinophane compounds and their application as ligands of essential metal ion based MRI contrast agents and ⁵²Mn based PET contrast agents Patent number: WO2017089848A1 (2017)</p> <p>Botár Richárd, Garda Zoltán, Fodor Tamás, Kálmán Ferenc Krisztián, Nagy Viktória, Tircsó Gyula, Tóth Imre: New 6-oxa-3,9,15-triaza-bicyclo[9.3.1]pentadeca-1(14),11(15),12-triene derivatives based compounds and their application as ligands of essential metal ion based MRI and ⁵²Mn based PET contrast agents Patent number: WO2017089849A1 (2017)</p> <p>Baranyai Zsolt, Garda Zoltán, Kálmán Ferenc Krisztián, Krusper László, Tircsó Gyula, Tóth Imre, Ghiani Simona, Maiocchi Alessandro: Ethylenediaminetetraacetic acid bis(amide) derivatives</p>

	<p>and their respective complexes with Mn(II) ion for use as MRI contrast agent</p> <p>Patent number: WO2016135234 (2016)</p>
<p>Important publications over the last 5 years</p>	<p><i>Selected recent publications from a total of approx. (give total number): 30</i></p> <p><i>Author(s):</i> Ferenc Krisztián Kálmán, Andrea Végh, Martín Regueiro-Figueroa, Éva Jakab Tóth, Carlos Platas-Iglesias, and Gyula Tircsó</p> <p><i>Title:</i> H₄octapa: Highly Stable Complexation of Lanthanide(III) Ions and Copper(II)</p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers:</i> <i>Inorg. Chem.</i>, 2015, 54(5), 2345–2356.</p> <p><i>Author(s):</i> Edit Farkas, Tamás Fodor, Ferenc K. Kálmán, Gyula Tircsó and Imre Tóth</p> <p><i>Title:</i> Equilibrium and Dissociation Kinetics of [Al(1,4,7-triazacyclononane-1,4,7-triacetate)] ([Al(NOTA)]) Complex</p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers:</i> <i>Reaction Kinetics, Mechanisms and Catalysis</i>, 2015, 116(1), 19-33.</p> <p><i>Author(s):</i> Zoltán Garda, Attila Forgács, Quyen N. Do, Ferenc K. Kálmán, Sarolta Timári, Imre Tóth, Zsolt Baranyai, Lorenzo Tei, Zoltán Kovács and Gyula Tircsó</p> <p><i>Title:</i> Physico-chemical properties of Mn²⁺ complexes formed with cis- and trans-DO2A: thermodynamic, electrochemical and kinetic studies</p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers:</i> <i>J. Inorg. Biochem.</i>, 2016, 163, 206–213.</p> <p><i>Author(s):</i> Christian Vanasschen, Enikő Molnár, Gyula Tircsó, Ferenc K. Kálmán, Éva Tóth, Marie Brandt, Heinz H. Coenen, and Bernd Neumaier</p> <p><i>Title:</i> Novel CDTA-based, Bifunctional Chelators for Stable and Inert Mn^{II} Complexation: Synthesis and Physicochemical Characterization</p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers:</i> <i>Inorg. Chem.</i>, 2017, 56(14), 7746–7760.</p>

	<p><i>Author(s):</i> Gyula Tircsó, Zsolt Baranyai, Ferenc Krisztián Kálmán, Zoltán Kovács, Ernő Brücher and Imre Tóth</p> <p><i>Title:</i> Stability of metal complexes, chapter 1.3 in Contrast Agents for MRI: Physical Methods</p> <p>Lability of metal complexes, chapter 1.4 in Contrast Agents for MRI: Physical Methods</p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers:</i> Eds Valerie C. Pierre and Matthew J. Allen RSC publishing, UK, 2018, pp. 40-74.</p>
Activities in specialist bodies over the last 5 years	<p><i>Organisation - Role – Period</i></p> <p><i>Membership without a specific role need not be mentioned.</i></p>

Name	<i>Prof. Sándor Kéki, PhD, DSc</i>
Position	<i>Full Professor, in charge of Chemical Engineering BSc and MSc</i>
Academic career	<p><i>Head of Department (University of Debrecen Department of Applied Chemistry, 2010-)</i></p> <p><i>DSc, Hungarian Academy of Sciences (2008)</i></p> <p><i>Habilitation (University of Debrecen, 2004)</i></p>
	<p><i>Doctoral qualification: PhD, Kossuth Lajos University, Hungary, 1996</i></p> <p><i>Doctorate, subject: Oscillation reactions</i></p> <p><i>Undergraduate degree, subject: Chemistry MSc, Kossuth Lajos University, Hungary, 1989</i></p>
Employment	<p><i>Full Professor, University of Debrecen, Department of Applied Chemistry (2010-)</i></p> <p><i>Associate Professor, University of Debrecen, Department of Applied Chemistry (2004-2010)</i></p> <p><i>Assistant Professor, University of Debrecen, Department of Applied Chemistry (1998-2004)</i></p> <p><i>Research assistant, Kossuth Lajos University, Department of Applied Chemistry (1994-1998)</i></p> <p><i>Research assistant, Kossuth Lajos University, Department of Physical Chemistry (1992-1994)</i></p> <p><i>TMB fellow, Kossuth Lajos University, Department of Physical Chemistry (1989-1992)</i></p>
Research and development projects over the last 5 years	<p><i>Name of projects:</i></p> <p>participant in EFOP-3.6.1-16-2016-00022, EU project:</p>

	<p>Debrecen Venture Catapult Program <i>Period: 2017-</i> group leader in GINOP-2.3.2-15-2016-00041, EU project: Regional Material Science Excellence Workshop - Research Program and Infrastructure <i>Period: 2016-</i> project leader in NKFI K-116465, Hungarian project: Synthesis and characterization of smart fluorescent polymers <i>Period: 2016-</i> project leader in NKFI K-101850, Hungarian project: Mass spectrometry of non-polar polymers <i>Period: 2012-2016</i> project leader in TÁMOP-4.2.2.A-11/1/KONV-2012-0036, EU project: Intelligent Functional Materials: Their mechanical, thermal, electromagnetic, optical properties and their applications, <i>Period: 2013-2015</i></p>
<p>Industry collaborations over the last 5 years</p>	<p><i>Project title: Characterization of low molecular weight pharmaceuticals</i> <i>Partner: Teva Pharmaceutical Works Ltd.</i></p> <p><i>Project title: Characterization of polymers</i> <i>Partner: MOL Petrochemicals</i></p> <p><i>Project title: Characterization of isocyanates</i> <i>Partner: BorsodChem</i></p> <p><i>Project title: Investigation and characterization of polyisobutylenes</i> <i>Partner: Infineum Ltd.</i></p>
<p>Patents and proprietary rights</p>	<p><i>Dynamic star polymers and a method for the synthesis thereof (2000)</i></p>
<p>Important publications over the last 5 years</p>	<p><i>Selected recent publications from a total of approx. (give total number): 154</i> <i>Authors: Miklós Nagy, Dávid Rácz, Zsolt László Nagy, Péter Pál Fehér, József Kalmár, István Fábián, Alexandra Kiss, Miklós Zsuga, Sándor Kéki</i> <i>Title: Solvatochromic isocyanonaphthalene dyes as ligands for silver (I) complexes, their applicability in silver (I) detection and background reduction in</i></p>

	<p>biolabelling</p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers:</i> SENSORS AND ACTUATORS B: CHEMICAL, 255, 2555-2567 (2018)</p> <p><i>Authors:</i> Nagy T, Kuki A, Zsuga M, Keki S</p> <p><i>Title:</i> Mass-Remainder Analysis (MARA): a New Data Mining Tool for Copolymer Characterization</p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers:</i> ANALYTICAL CHEMISTRY, 90, 3892-3897 (2018)</p> <p><i>Authors:</i> Borbála Antal, Ákos Kuki, Lajos Nagy, Tibor Nagy, Miklós Zsuga, Sándor Kéki</p> <p><i>Title:</i> Rapid detection of hazardous chemicals in textiles by direct analysis in real time mass spectrometry (DART-MS)</p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers:</i> ANALYTICAL AND BIOANALYTICAL CHEMISTRY, 408, 5189-5198 (2016)</p> <p><i>Authors:</i> Lajos Nagy, Ákos Kuki, György Deák, Mihály Purgel, Ádam Vékony, Miklós Zsuga, Sándor Kéki</p> <p><i>Title:</i> Gas-Phase Interaction of Anions with Polyisobutylenes: Collision Induced Dissociation Study and Quantum Chemical Modeling</p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers:</i> JOURNAL OF PHYSICAL CHEMISTRY B, 120, 9195-9203 (2016)</p> <p><i>Authors:</i> Katalin Czifrák, Csilla Lakatos, József Karger-Kocsis, Lajos Daróczy, Miklós Zsuga, Sándor Kéki</p> <p><i>Title:</i> One-Pot Synthesis and Characterization of Novel Shape-Memory Poly(ϵ-Caprolactone) Based Polyurethane-Epoxy Co-networks with Diels–Alder Couplings</p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers:</i> POLYMERS, 10, 504 (2018)</p>
Activities in specialist bodies	Hungarian Academy of Sciences, Member of the Committee of Inorganic Chemistry and Materials

over the last 5 years	<p>Science (2011-) Hungarian Academy of Sciences, Chairman of the Regional Committee of Polymer Chemistry (2005-) Faculty of Chemical and Biochemical Engineering, Budapest University of Technology and Economics, Member of the Doctoral and Habilitation Committee (2006-) University of Debrecen, Leader of the Macromolecular and Surface Chemistry program of the Chemistry Doctoral School, (2011-) University of Debrecen, Supervisor of the Chemical Engineering BSc major (2010-)</p>
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Name	<i>János Kerékgyártó, PhD, CSc</i>
Position	<i>Senior research fellow, in charge of Biology BSc</i>
Academic career	<i>CSc, Hungarian Academy of Sciences, Budapest, 1994</i>
	<p><i>Doctoral qualification: PhD, University of Debrecen, 1994</i></p> <p><i>Doctorate, subject: „Chemical synthesis of biologically active oligosaccharides”, 1994</i></p> <p><i>Undergraduate degree, subject: Chemistry MSc, Debrecen, 1981</i></p>
Employment	<p><i>Senior research fellow, Department of Botany, Faculty of Sciences and Technology, University of Debrecen, 2009-present</i></p> <p><i>Senior research fellow, Institute of Biochemistry, Faculty of Sciences and Technology, University of Debrecen, 2000-2009</i></p> <p><i>Senior research fellow, Institute of Biochemistry, Faculty of Sciences, Lajos Kossuth University, Debrecen, 1995-2000</i></p> <p><i>Postdoctoral fellow, Bijvoet Center, Department of Bio-Organic Chemistry, Utrecht University, Utrecht, The Netherlands, 1995</i></p> <p><i>Scientific fellow, Institute of Biochemistry, Faculty of Sciences, Lajos Kossuth University, Debrecen, 1988-1994</i></p> <p><i>Post Graduate Scholarship Student of the Hungarian Academy of Sciences, Institute of Biochemistry, Faculty of Sciences, Lajos Kossuth University, Debrecen, 1985-1988</i></p> <p><i>Young investigator, Bijvoet Center, Department of Bio-</i></p>

	<p><i>Organic Chemistry, Utrecht University, Utrecht, The Netherlands, 1986-1987</i></p> <p><i>Junior research assistant, Institute of Biochemistry, Faculty of Sciences, Lajos Kossuth University, Debrecen, 1981-1984</i></p>
Research and development projects over the last 5 years	<p><i>Name of project or research focus: Functional glycomics</i></p> <p><i>Period and any other information: 2012-present</i></p> <p><i>Partners, if applicable:</i></p> <p><i>MTA-PE Translational Glycomics Group, MUKKI, University of Pannonia, Veszprém, Hungary</i></p> <p><i>Horváth Laboratory of Bioseparation Sciences, University of Debrecen, Hungary</i></p>
Industry collaborations over the last 5 years	<p><i>Partners: GlycOptim Ltd., Debrecen, 2018</i></p> <p><i>Cyclolab Ltd., Budapest, 2016-2017</i></p>
Patents and proprietary rights	<p><i>„Procedure for the synthesis of glycosides of aromatic amines”</i></p> <p><i>Hungarian patent: HU 203 896 (1988)</i></p> <p><i>„Procedure for the synthesis of amino-benzoicacid-N-glycoside derivatives”</i></p> <p><i>Hungarian patent: HU 208 146 (1989)</i></p>
Important publications over the last 5 years	<p><i>Selected recent publications from a total of 37:</i></p> <p><i>1. L. Kalmár, Z. Szurmai, J. Kerékgyártó, A. Guttman, M. Bojstrup, K. Ágoston</i></p> <p><i>Phenyl-2-O-acetyl-3-O-allyl-4-O-benzyl-1-thio-β-D-glucopyranoside, a versatile, orthogonally protected building block</i></p> <p><i>Carbohydr. Chem. Proven Methods, 24, (2015) 187-192.</i></p> <p><i>2. K. Ágoston, Gy. Gyémánt, L. Kalmár, J. Kerékgyártó, Z. Szurmai, B. Döncző, A. Guttman</i></p> <p><i>Synthesis and MALDI-TOF analysis of protected oligosaccharide components of N-glycoproteins</i></p> <p><i>J. Carbohydr. Chem., 33, (2014) 326-343</i></p> <p><i>3. B. Döncző, J. Kerékgyártó, Z. Szurmai, A. Guttman</i></p> <p><i>Glycan microarrays: new angels and strategies</i></p> <p><i>Analyst, 139, (2014) 2650-2657</i></p>

	<p>4. B. Dönczö, L. Kalmár, J. Kerégyártó, Z. Szurmai, A. Guttman <i>Combinatorial glycomics 1: Synthesis options</i> <i>Chem. Listy</i> 107, s352-s354 (2013)</p> <p>5. M. Kerégyártó, A. Fekete, Z. Szurmai, J. Kerégyártó, L. Takács, I. Kurucz, A. Guttman <i>Neoglycoproteins as carbohydrate antigens: synthesis, analysis, and polyclonal antibody response</i> <i>Electrophoresis</i>, 34, (2013) 2379-2386</p>
Activities in specialist bodies over the last 5 years	<i>Head of the Asklepios Scientific Foundation, 1995-2017</i>

Name	<i>Mónika Kéri</i>
Position	<i>Colloid and environmental chemistry Assistant professor</i>
Academic career	<i>Assistant lecturer (University of Debrecen, UD, 2013)</i>
	<p><i>PhD in chemistry (UD, 2015), The behavior of PAMAM dendrimers in aqueous solution and their interaction with small molecules</i></p> <p><i>Environmental scientist with specification in chemistry and English-Hungarian special translator (UD, 2008)</i></p>
Employment	<p><i>Assistant lecturer and assistant professor - UD, Department of Physical Chemistry – 2013-present</i></p> <p><i>Environmental Scientist - DTMP (Debreceni Tudományos Műszaki Park) Kht. and Aquaprofit Zrt. - 2008-2009</i></p>
Research and development projects over the last 5 years	<ul style="list-style-type: none"> - <i>Structures and applications of micro- and mesoporous materials, 2017-21, NKFI 124571, 27.3 M HUF</i> - <i>Preparation and physical chemistry of doped carbon nanoparticles, 2013-18, NKFI 109558, 27.9 M HUF</i> - <i>NMR study of porous materials of environmental importance, 2013-2014, Jedlik Ányos Fellowship, 2.4 M HUF</i> - <i>Study of porous adsorbents with NMR cryoporometry, relaxometry and diffusometry, 2013, ENVIKUT project TÁMOP-4.2.2.A-11/1/KONV-2012-0043</i> - <i>The interaction of functionalized dendrimers with metal ions and small molecules, 2011-2013, Chinese-Hungarian Science and Technology Program, TÉT_10-1-2011-0145</i> - <i>Vanadium-dendrimer catalysts for the oxidation of trichloroethylene (TCE), 2010-2011, CHEMIKUT project</i>

	TÁMOP-4.2.2.-08/1-2008-0012
Industry collaborations over the last 5 years	<ul style="list-style-type: none"> - Elimination method for arsenic sludge from drinking water treatment, Aquaprofit Zrt. - Structure study of colloidal drug carriers by NMR, Nangenex Inc.
Patents and proprietary rights	Elimination method for arsenic sludge from drinking water treatment (2015)
Important publications over the last 5 years	<p>Selected recent publications from a total of approx.: 8</p> <ul style="list-style-type: none"> – Kéri M, Nagy Z et.al. Beware of phosphate: evidence of dendrimer—phosphate interactions <i>Phys. Chem. Chem. Phys.</i> 19: pp. 11540-11548 (2017) – Veres P, Kéri M, et. al. <i>Mechanism of drug release from silica - gelatin aerogel</i> – Relationship between matrix structure and release kinetics <i>Colloids And Surfaces B : Biointerfaces</i> 152: pp. 229-237. (2017) –Kalmár J, Kéri M, et.al. The Pore Network and the Adsorption Characteristics of Mesoporous Silica Aerogel: Adsorption Kinetics on a Timescale of Seconds <i>RSC Advances</i> 5:(130) pp. 107237-107246. (2015) –Kéri M, Peng C, Shi X, Bányai I. <i>NMR characterization of PAMAM-G5.NH2 entrapped atomic and molecular assemblies</i> <i>Journal of Physical Chemistry B</i> 119:(7) pp. 3312-3319.(2015) –Bányai I, Kéri M, Nagy Z, Berka M, Balogh LP <i>Self-diffusion of water and poly(amidoamine) dendrimers in dilute aqueous solutions</i> <i>Soft Matter</i> 9:(5) pp. 1645-1655. (2013)
Activities in specialist bodies over the last 5 years	-

Name	Ákos Kuki
Position	associate professor
Academic career	college assistant lecturer, Teachers' Training College Nyíregyháza, 1992 Habilitation: University of Debrecen, 2015
	PhD, Material Sciences, University of Miskolc, 2002 MSc. (electrical engineer, Technical University of

	Budapest, 1991), MSc (teacher of electrical engineering, Technical University of Budapest, 1994)
Employment	college assistant lecturer, Teachers' Training College Nyíregyháza, 1992-1996 college senior lecturer, Teachers' Training College Nyíregyháza, 1996-2002 college associate professor, Teachers' Training College Nyíregyháza, 2002-2008 college associate professor, University of Debrecen, 2008-2016 associate professor, University of Debrecen, 2016-
Research and development projects over the last 5 years	OTKA K116465, OTKA K101850, OTKA K109006, TÁMOP-4.2.2.A-11/1/KONV-2012-0036, Smart functional materials, 2012-2015 GINOP-2.3.2-15-2016-00041, Regional Workshop for Excellence in Materials Science, 2017-2020 EFOP-3.6.1-16-2016-00022, Debrecen Venture Catapult Program, 2018-
Industry collaborations over the last 5 years	
Patents and proprietary rights	
Important publications over the last 5 years	<i>Selected recent publications from a total of approx. 40:</i> [6] Nagy L, Nagy L, Deák Gy, Kuki Á, Purgel M, Narmandakh M, Iván B, Zsuga M, Kéki S, Can non-polar polyisobutylenes be measured by electrospray ionization mass spectrometry? Anion-attachment proved to be an appropriate method, <i>J. Am. Soc. Mass Spectrom.</i> , 2016 , 27, 432. [7] Nagy, L; Kuki, A; Deak, G; Purgel, M; Vekony, A; Zsuga, M; Keki, S, Gas-Phase Interaction of Anions with Polyisobutylenes: Collision-Induced Dissociation Study and Quantum Chemical Modeling, <i>J. Phys. Chem. B</i> , 2016 , 120, 9195. [8] Antal, B; Kuki, A; Nagy, L; Nagy, T; Zsuga, M; Keki, S, Rapid detection of hazardous chemicals in textiles by direct analysis in real-time mass spectrometry (DART-MS), <i>Anal. Bioanal. Chem.</i> , 2016 , 408, 5189. [9] Kuki, A; Nagy, L; Nagy, T; Zsuga, M; Keki, S., Screening of additives and other chemicals in polyurethanes by direct analysis in real time mass spectrometry (DART-MS), <i>Anal. Bioanal. Chem.</i> ,

	<p>2017, 409, 6149-6162.</p> <p>[10] Nagy, T; Kuki, A; Zsuga, M; Keki, S, Mass-Remainder Analysis (MARA): a New Data Mining Tool for Copolymer Characterization, <i>Anal. Chem.</i>, 2018, 90, 3892-3897.</p>
Activities in specialist bodies over the last 5 years	Member of the Committee of Physical Chemistry of the Hungarian Academy of Sciences

Name	István Lázár, PhD
Position	Associate Professor of Chemistr, Department of Inorganic and Analytical Chemistry
Academic career	
	<p>Candidate of Science (C.Sc.) degree: 1994, Hungarian Academy of Scienc, Coordination Chemistry</p> <p>Doctorate, subject: Kossuth Lajos University, 1987, Boron organic chemistry</p> <p>Undergraduate degree, subject: Chemist, MSc, 1984</p>
Employment	<p>PhD scholarship: Hungarian Academy of Science, 1984-1987</p> <p>Assistant lecturer – Associate professor from 1987 to present: University of Debrecen (formerly Kossuth Lajos University)</p> <p>Research Associate, 1989-1991, 26 months; 1994, 5 months, – The University of Texas at Dallas, US</p>
Research and development projects over the last 5 years	<p>Role: researcher</p> <p>Name of project or research focus: TÁMOP-4.2.2.A-11/1/KONV-2012-0036; Intelligent Functional Materials: Mechanical, thermal, electromagnetic and optical properties and applications</p> <p>Period and any other information: 2012-2016</p> <p>Role: researcher</p> <p>GINOP-2.2.1-15-2017-00068, Gel-based biomatrices for human tissue replacements, and the development of their production technologies</p> <p>Period and any other information: 2017-2021</p> <p>Role: researcher</p> <p>GINOP-2.3.2-15-2016-00041, Regionális Anyagtudományi Kiválósági Műhely - Kutatási Program és Infrastruktúra</p> <p>Period and any other information: 2017-2021</p>

Industry collaborations over the last 5 years	L'Oréal Co.,France: Material Transfer, and Evaluation, 2015, L'Oréal Co.,France: Material Production, and Evaluation, 2016
Patents and proprietary rights	<p>Using polyazamacrocyclic compounds for intracellular measurement of metal ions using MRS; U.S. (1993), US 5188816:</p> <p>Polyazamacrocyclic compounds for complexation of metal ions; U.S. (1994), US 5342606</p> <p>ELJÁRÁS KOMPOZIT SZILIKA ALKOGÉLEK, AEROGÉLEK ÉS XEROGÉLEK ELŐÁLLÍTÁSÁRA, VALAMINT AZ ELJÁRÁS FOLYAMATOS MEGVALÓSÍTÁSÁRA ALKALMAS BERENDEZÉS; Hung. Pat. Appl. (2013), HU 2011000060 A2 20130628</p> <p>Method for the preparation of composite silica alcogels, aerogels and xerogels, apparatus for carrying out the method continuously, and novel composite silica alcogels, aerogels and xerogels; PCT Int. Appl. (2013), WO 2013061104 A2 20130502.</p>
Important publications over the last 5 years	<p>Selected recent publications from a total of: 16</p> <p>1) Author: István Lázár Helga Fruzsina Bereczki Sándor Manó Lajos Daróczi György Deák István Fábián Zoltán Csernátony Title: Synthesis and study of new functionalized silica aerogel poly(methyl methacrylate) composites for biomedical use Polym. Compos. (2015) Volume36, Issue2 Pages 348-358</p> <p>2) Author: István Lázár, József Kalmár, Anca Peter, Anett Szilágyi, Enikő Győri, Tamás Ditrói, István Fábián Title: Photocatalytic performance of highly amorphous titania–silica aerogels with mesopores: The adverse effect of the in situ adsorption of some organic substrates during photodegradation Applied Surface Science 356 (2015) 521–531</p> <p>3) Author: I. Lázár, A. Szilágyi, G. Sáfrán, A. Szegedi, S.</p>

	<p>Stichlautner, K. Lázár Title: Iron oxyhydroxide aerogels and xerogels by controlled hydrolysis of FeCl₃.6H₂O in organic solvents: stages of formation RSC Adv., 2015, 5, 72716</p> <p>4) Author: H. F. Bereczki, L. Daróczy, I.Fábián, I. Lázár Title: Sol-gel synthesis, characterization and catalytic activity of silica aerogels functionalized with copper(II) complexes of cyclen and cyclam Microporous and Mesoporous Materials 234 (2016) 392e400</p> <p>5) Author: P. Veres, M. Kéri, I. Bányai, I. Lázár, I. Fábián, C. Domingo, J. Kalmár Title: Mechanism of drug release from silica-gelatin aerogel—Relationship between matrix structure and release kinetics Colloids and Surfaces B: Biointerfaces 152 (2017) 229–237</p>
Activities in specialist bodies over the last 5 years	

Name	<i>Norbert Lihi PhD</i>
Position	<i>Assistant Research Fellow</i>
Academic career	<i>PhD (2018, University of Debrecen)</i>
	<i>doctoral qualification: PhD, University of Debrecen, Hungary, 2018</i> <i>Doctorate, subject: Department of Inorganic and Analytical Chemistry, University of Debrecen, Hungary, Bioinorganic Chemistry</i> <i>Undergraduate degree, subject: Chemistry MSc, Bioinorganic Chemistry</i>
Employment	<i>Assistant Research Fellow at the MTA-DE Redox and Homogeneous Catalytic Reaction Mechanisms Research Group, University of Debrecen (2017-)</i>
Research and development projects over the last 5 years	<i>Name of project or research focus: TÁMOP-4.2.1/B-09/1/KONV-2010-0007</i> <i>Period and any other information: 2010-2014</i> <i>Name of project or research focus: NKFIH K115480;</i>

	<p><i>Coordination and electrochemical background of role of metal ions in neurodegenerative disorders</i> <i>Period and any other information: 2016-2018</i></p> <p><i>Name of project or research focus: GINOP-2.3.2-15-2016-00008; Chemistry for better life: strategic R&D center at the University of Debrecen</i> <i>Period and any other information: 2016-2020</i></p> <p><i>Name of project or research focus: NKFIH K124983; The redox reactions of strong inorganic oxidants: kinetics and mechanism</i> <i>Period and any other information: 2017-2021</i></p>
Industry collaborations over the last 5 years	<p><i>Project title:</i></p> <p><i>Partners:</i></p>
Patents and proprietary rights	<p><i>Title (Year)</i></p>
Important publications over the last 5 years	<p><i>Selected recent publications from a total of approx. (give total number): 16</i></p> <p><i>Author(s): Sciortino Giuseppe, Lihi Norbert, Czine Tamas, Marechal Jean-Didier, Lledos Agusti, Garribba Eugenio</i></p> <p><i>Title: Accurate prediction of vertical electronic transitions of Ni(II) coordination compounds via time dependent density functional theory.</i></p> <p><i>Any other information: INTERNATIONAL JOURNAL OF QUANTUM CHEMISTRY Ahead of Print:(First published: 20 May 2018) Paper e25655. 14 p. (2018)</i></p> <p><i>Author(s): Ágnes Grenács, Norbert Lihi, Imre Sóvágó, Katalin Várnagy</i></p> <p><i>Title: The influence of penicillamine/cysteine mutation on the metal complexes of peptides.</i></p> <p><i>Any other information: DALTON TRANSACTIONS 46:(39) pp. 13472-13481. (2017)</i></p> <p><i>Author(s): Norbert Lihi, Attila József Godó, Giuseppe Sciortino, Eugenio Garribba, Katalin Várnagy</i></p> <p><i>Title: Tridentate (O,N,O) ligands as potential chelator compounds for iron overload.</i></p> <p><i>Any other information: POLYHEDRON 123: pp. 192-205. (2017)</i></p>

	<p><i>Author(s): Norbert Lihi, Daniele Sanna, István Bányai, Katalin Várnagy, Imre Sóvágó</i></p> <p><i>Title: Unusual binding modes in the copper(II) and palladium(II) complexes of peptides containing both histidyl and cysteinyl residues.</i></p> <p><i>Any other information: NEW JOURNAL OF CHEMISTRY 41:(3) pp. 1372-1379. (2017)</i></p> <p><i>Author(s): Imre Sóvágó, Katalin Várnagy, Norbert Lihi, Ágnes Grenács</i></p> <p><i>Title: Coordinating properties of peptides containing histidyl residues.</i></p> <p><i>Any other information: COORDINATION CHEMISTRY REVIEWS 327-328: pp. 43-54. (2016)</i></p>
Activities in specialist bodies over the last 5 years	<p><i>Organisation - Role – Period</i></p> <p><i>Membership without a specific role need not be mentioned.</i></p>

Name	Attila Mándi PhD
Position	researcher / assistant professor (organic chemistry, stereochemistry, chemical informatics, molecular modeling)
Academic career	Habilitation: -
	<p><i>Ph.D. degree in chemistry, University of Debrecen, 2012; Title of dissertation: Quantum chemical and dynamic investigations on acetal derivatives of carbohydrates and cycloaddition reactions of antibiotic derivatives (supervisor: Prof. Dr. András Lipták)</i></p> <p><i>Master's degree in English-Hungarian translation (minor), University of Debrecen, 2009</i></p> <p><i>Master's degree in chemistry (major) and German-Hungarian translation (minor), University of Debrecen, 2005</i></p>
Employment	<p>Assistant research fellow, researcher, postdoc and assistant professor at the Department of Organic Chemistry, University of Debrecen, 2010 -</p> <p>Assistant research fellow at the Research Group for Carbohydrates of the Hungarian Academy of Sciences at University Debrecen, 2009-2010.</p>
Research and development projects over the last 5 years	<p><i>Name of project or research focus: GINOP-2.3.2-15-2016-00008</i></p> <p><i>Period and any other information: 2016 -</i></p>

	<p><i>Partners, if applicable:</i></p> <p><i>Amount of financing:</i></p> <p><i>Name of project or research focus: NKFI K120181</i></p> <p><i>Period and any other information: 2016 -</i></p> <p><i>Partners, if applicable: Prof. Tibor Kurtán</i></p> <p><i>Amount of financing:</i></p> <p><i>Name of project or research focus: TÉT_12_CN-1-2012-0023</i></p> <p><i>Period and any other information: 2013 -2015</i></p> <p><i>Partners, if applicable: Prof. Tibor Kurtán</i></p> <p><i>Amount of financing:</i></p> <p><i>Name of project or research focus: OTKA K106294</i></p> <p><i>Period and any other information: 2013 -2015</i></p> <p><i>Partners, if applicable: Dr. István Komáromi</i></p> <p><i>Amount of financing:</i></p> <p><i>Name of project or research focus: OTKA K105871</i></p> <p><i>Period and any other information: 2012 -2016</i></p> <p><i>Partners, if applicable: Prof. Tibor Kurtán</i></p> <p><i>Amount of financing:</i></p> <p><i>Major international cooperation partners:</i></p> <p><i>Prof. Peter Proksch, Heinrich-Heine-Universität, Düsseldorf</i></p> <p><i>Prof. Yue-Wei Guo, SIMM, Chinese Academy of Sciences, Shanghai</i></p> <p><i>Prof. Yang Ye, SIMM, Chinese Academy of Sciences, Shanghai</i></p> <p><i>Prof. Wen Zhang, Second Military Medical University, Shanghai</i></p> <p><i>Prof. Changsheng Zhang, South China Sea Institute of Oceanology, Guangzhou</i></p> <p><i>Prof. Bingui Wang, Institute of Oceanology, Chinese Academy of Sciences, Qingdao</i></p> <p><i>Prof. Dehai Li, Ocean University of China, Qingdao</i></p> <p><i>Prof. Mary J. Garson, University of Queensland, Brisbane</i></p>
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	<p><i>Prof. Kenji Monde, Hokkaido University, Sapporo</i> <i>Dr. Valentin Paul Nicu, Vrije Universiteit Amsterdam & University of Sibiu</i></p>
Industry collaborations over the last 5 years	<p><i>Project title: -</i> <i>Partners: -</i></p>
Patents and proprietary rights	-
Important publications over the last 5 years	<p><i>Selected recent publications from a total of approx. (give total number): 113</i></p> <p><i>Author(s): P. Sun, D. X. Xu, A. Mándi, T. Kurtán, T. J. Li, B. Schulz, W. Zhang</i> <i>Title: Structure, Absolute Configuration, and Conformational Study of 12-Membered Macrolides from the Fungus Dendrodochium sp. Associated with the Sea Cucumber Holothuria nobilis Selenka</i> <i>Any other information:</i> <i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: J. Org. Chem. 2013, 78, 7030-7047.</i></p> <p><i>Author(s): Y. Ye, A. Minami, A. Mándi, C. Liu, T. Taniguchi, T. Kuzuyama, K. Monde, K. Gomi, H. Oikawa</i> <i>Title: Genome mining for sesterterpenes using bifunctional terpene synthases reveals a unified intermediate of di/sesterterpenes</i> <i>Any other information:</i> <i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: J. Am. Chem. Soc. 2015, 137, 11846-11853.</i></p> <p><i>Author(s): A. Mándi, I. W. Mudianta, T. Kurtán, M. J. Garson</i> <i>Title: Absolute Configuration and Conformational Study of Psammaplysins A and B from the Balinese Marine Sponge Aplysinella strongylata</i> <i>Any other information:</i> <i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: J. Nat. Prod. 2015, 78, 2051-2056.</i></p> <p><i>Author(s): Z. Tian, P. Sun, Y. Yan, Z. Wu, Q. Zheng, S.</i></p>

	<p>Zhou, H. Zhang, F. Yu, X. Jia, D. Chen, A. Mándi, T. Kurtán, W. Liu</p> <p><i>Title: An Enzymatic [4+2] Cyclization Cascade Creates the Pentacyclic Core of Pyrroindomycins</i></p> <p><i>Any other information:</i></p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: Nat. Chem. Biol. 2015, 11, 259-265.</i></p> <p>Author(s): L. F. Liang, T. Kurtán, A. Mándi, L. G. Yao, J. Li, W. Zhang, Y. W. Guo</p> <p><i>Title: Unprecedented Diterpenoids as a PTP1B Inhibitor from the Hainan Soft Coral Sarcophyton trocheliophorum Marenzeller</i></p> <p><i>Any other information:</i></p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: Org. Lett. 2013, 15, 274-277.</i></p>
Activities in specialist bodies over the last 5 years	-

Name	Zoltán Muzsnay
Position	Head of department, Department of Geometry, Institute of Mathematics
Academic career	vice director, Institute of Mathematics, 2011 – 2017, associate professor, 2006 – present, assistant professor, 1998 – 2006, research fellow, University of Debrecen 1997 – 1998 PhD, Université Paul Sabatier, Toulouse, France, 1997, Habilitation, University of Debrecen, 2007
Employment	associate professor, University of Debrecen, Hungary, 2006 - present, assistant professor, University of Debrecen, Hungary, 1998 - 2006, research fellow, University of Debrecen, 1997 - 1998 Ph.D. student, University Paul Sabatier, France, 1992 - 1997, teaching assistant, Kossuth Lajos University, Debrecen, 1989 - 1991,
Research and development projects over the last 5 years	principal investigator in scientific project: - Hungarian Research and Technology Innovation Fund, 2013-

	<p>2014,</p> <ul style="list-style-type: none"> - Hungarian Scientific Research Fund, (OTKA), 2007-2011, - Öveges József Program, 2006-2007, - Research Projects in Higher Education (FKFP), 2000--2004, <p>researcher in scientific project:</p> <ul style="list-style-type: none"> - National Committee for Technological Development, 1994 -- 1995, - Hungarian Scientific Research Fund (OTKA), 1999 - 2003, - Hungarian Scientific Research Fund (OTKA), 1995 - 1999, - BIOMICS, EU FET FP7 project, 2012 - 2014, - LIE-DIFF-GEOM, European Union FP7, 2007 -- 2013. - High performance computing (TÁMOP 4.2.2.C), 2011 -- 2013. - EFOP-3.6.2-16-2017-00015, - EFOP-3.6.1-16-2016-00022,
Industry collaborations over the last 5 years	-
Patents and proprietary rights	-
Important publications over the last 5 years	<p>Sprays metrizable by Finsler functions of constant flag curvature Differential Geom. Appl. 31 (2013), no. 3, 405-415. (with I. Bucataru)</p> <p>Projectively flat Finsler manifolds with infinite dimensional holonomy Forum Mathematicum, 27, 2, 2015, pp 767–786, (with P.T. Nagy)</p> <p>Metrizable isotropic second-order differential equations and Hilbert's fourth problem Journal of the Aust Math. Soc., 97, 1, 2014, pp 27-47 (with I. Bucataru)</p> <p>Characterization of projective Finsler manifolds of constant curvature having infinite dimensional holonomy group Publ. Math. Debrecen, 84/1-2 (2014), 17–28 (with P.T. Nagy)</p> <p>Finsler 2-manifolds with maximal holonomy group of infinite dimension Differential Geometry and its Applications, Vol. 39, 2015, Pages 1–9, (with P.T. Nagy)</p> <p>Non-existence of Funk functions for Finsler spaces of non-vanishing scalar flag curvature Comptes Rendus Mathematique,</p>

	<p>Volume 354, Issue 6, June 2016, Pages 619-622, (with I. Bucataru)</p> <p>On the projective Finsler metrizable and the integrability of Rapcsák equation Czechoslovak Mathematical Journal, vol. 67, no. 2 (2017), pp. 469-495, (with T. Milkovszki)</p> <p>Invariant Metrizability and Projective Metrizability on Lie Groups and Homogeneous Spaces Mediterranean Journal of Mathematics, December 2016, Volume 13, Issue 6, pp 4567–4580 (with I. Bucataru, T. Milkovszki)</p> <p>Freedom of $h(2)$-variationality and metrizable of sprays Differential Geometry and its Applications, (with S.G. Elgend)</p>
<p>Activities in specialist bodies over the last 5 years</p>	<ul style="list-style-type: none"> - member of the advisory board of János Bolyai Research Fellowship of the Hungarian Academy of Sciences, 2016 - present, - vice-president, scientific section, János Bolyai Mathematical Society, 2015 - present, - editor of the Electronic Mathematical Journal of the János Bolyai Math. Society, 2015 - 2016, - member of the scientific review panel, Hungarian National Research, Development and Innovation Fund, 2016 - 2017, - secretary of the Mathematical Subcommittee of the Academic Committee in Debrecen, 1998 - 2000, - member of public body of Hungarian Academy of Sciences - member of the János Bolyai Mathematical Society , - secretary of the Doctoral Council of Natural Sciences, 2008 - 2017, - member of the Faculty Council, Faculty of Sciences, 2005 - 2008.

Name	<i>Lajos Nagy PhD</i>
Position	<i>Associate Professor</i>

Academic career	Habilitation (<i>UD, 2017</i>)
	<p><i>doctoral qualification: PhD, University of Debrecen, Hungary, 2009</i></p> <p><i>Doctorate, subject: University of Debrecen, Hungary Mass Spectrometry, 2004</i></p> <p><i>Undergraduate degree, subject: MSc in Chemistry, Chemical Engineering BSc</i></p> <p><i>Qualification: Erdey-Grúz Tibor Secondary School of Chemistry, 1998</i></p>
Employment	<p><i>Assistant lecturer – 2009</i></p> <p><i>Assistant professor – 2012</i></p> <p><i>Associate professor – 2017</i></p>
Research and development projects over the last 5 years	<p><i>Name of project or research focus: HURO/0901/058/2.2.2., EU project: Hungary-Romania Cross-Border Co-operation Programme</i></p> <p><i>Period and any other information: 2007-2013</i></p> <p><i>Partners, if applicable:</i></p> <p><i>Amount of financing:</i></p> <p><i>Name of project or research focus: GINOP-2.3.2-15-2016-00041, EU project: Regional Workshop for Excellence in Materials Science - Research Program and Infrastructure</i></p> <p><i>Period and any other information: 2017-2020</i></p> <p><i>Partners, if applicable:</i></p> <p><i>Amount of financing:</i></p> <p><i>Name of project or research focus: NKFIH-101850, National Research, Development and Innovation Office</i></p> <p><i>Title: Soft ionization mass spectrometric study of non-polar polymers.</i></p> <p><i>Period and any other information: 2012-2016</i></p> <p><i>Partners, if applicable:</i></p> <p><i>Amount of financing:</i></p> <p><i>Name of project or research focus: NKFIH-116465, National Research, Development and Innovation Office</i></p> <p><i>Title: Synthesis and characterization of smart fluorescent polymers.</i></p> <p><i>Period and any other information: 2016-2019</i></p> <p><i>Partners, if applicable:</i></p>

	<i>Amount of financing:</i>
Industry collaborations over the last 5 years	<i>Project title:</i> <i>Partners:</i>
Patents and proprietary rights	<i>Title (Year)</i>
Important publications over the last 5 years	<p><i>Selected recent publications from a total of approx. (give total number): 30</i></p> <p><i>Author(s): Nagy L., Kuki Á., Nagy T., Mándi A., Deák G., Nagy M., Batta G., Zsuga M., Kéki S.:</i></p> <p><i>Title: Collision-induced dissociation study of isosilychristin, a constituent of Silymarin</i></p> <p><i>Any other information:</i></p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: Rapid Commun. Mass Spectrom. 2013, 27, 1413-1416.</i></p> <p><i>Author(s): Nagy L., Kuki A., Szabó K., Sipos A., Zsuga M., Kéki S.:</i></p> <p><i>Title: Fragmentation study of noscapine derivatives under electrospray conditions</i></p> <p><i>Any other information:</i></p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: Rapid Communications in Mass Spectrometry, 2014, 28, 822-828.</i></p> <p><i>Author(s): Lajos Nagy, Tibor Nagy, György Deák, Ákos Kuki, Borbála Antal, Miklós Zsuga, Sándor Kéki:</i></p> <p><i>Title: Direct analysis in real time mass spectrometry (DART-MS) of highly non-polar low molecular weight polyisobutylenes</i></p> <p><i>Any other information:</i></p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: Journal of Mass Spectrometry, 2015, 50, 1071–1078.</i></p> <p><i>Author(s): Lajos Nagy; Tibor Nagy; György Deák; Ákos Kuki; Mihály Purgel; Mijid Narmandakh; Béla Iván; Miklós Zsuga; Sándor Kéki:</i></p> <p><i>Title: Can Nonpolar Polyisobutylenes be Measured by Electrospray Ionization Mass Spectrometry? Anion-Attachment Proved to be an Appropriate Method</i></p> <p><i>Any other information:</i></p>

	<p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: Journal of The American Society for Mass Spectrometry, 2016, 27, 432-442.</i></p> <p><i>Author(s): Lajos Nagy, Tibor Nagy, Ákos Kuki, Mihály Purgel, Miklós Zsuga and Sándor Kéki:</i></p> <p><i>Title: Kinetics of Uncatalyzed Reactions of 2,4'- and 4,4'-Diphenylmethane-Diisocyanate with Primary and Secondary Alcohols</i></p> <p><i>Any other information:</i></p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: International Journal of Chemical Kinetics, 2017, 49, 643-655</i></p>
Activities in specialist bodies over the last 5 years	<p><i>Organisation - Role – Period</i></p> <p><i>Membership without a specific role need not be mentioned.</i></p>

Name	<i>Miklós Nagy PhD</i>
Position	<i>Associate Professor</i>
Academic career	<i>Habilitation (UD, 2017)</i>
	<p><i>doctoral qualification: PhD, University of Debrecen, Hungary, 2005</i></p> <p><i>Doctorate, subject: University of Debrecen, Synthesis and characterization of functionalized polymers</i></p> <p><i>Undergraduate degree, subject: Chemistry MSc</i></p> <p><i>Qualification: Chemist, 2000</i></p>
Employment	<p><i>Associate Professor, University of Debrecen, Department of Applied Chemistry 2018-</i></p> <p><i>Assistant Professor, University of Debrecen, Department of Applied Chemistry 2007-2018</i></p> <p><i>Teaching Assistant, University of Debrecen, Department of Applied Chemistry - 2002-2007</i></p>
Research and development projects over the last 5 years	<p><i>Name of projects:</i></p> <p><i>participant in GINOP-2.3.2-15-2016-00041, EU project: Regional Material Science Excellence Workshop - Research Program and Infrastructure</i></p> <p><i>Period: 2016-</i></p> <p><i>participant in TÁMOP-4.2.2.A-11/1/KONV-2012-0036, EU project: Intelligent Functional Materials: Their</i></p>

	<p>mechanical, thermal, electromagnetic, optical properties and their applications, <i>Period: 2013-2015</i></p> <p>participant in NKFI K-116465, Hungarian project: Synthesis and characterization of smart fluorescent polymers <i>Period: 2016-</i></p>
Industry collaborations over the last 5 years	<p><i>Project title: Photopolymerization of epoxidated plant oils</i></p> <p><i>Partners: Polinvent Zrt.</i></p>
Patents and proprietary rights	<i>Title (Year)</i>
Important publications over the last 5 years	<p><i>Selected recent publications from a total of approx. (give total number): 11</i></p> <p><i>Authors: Miklós Nagy, Dávid Rácz, Zsolt László Nagy, Péter Pál Fehér, József Kalmár, István Fábián, Alexandra Kiss, Miklós Zsuga, Sándor Kéki</i></p> <p><i>Title: Solvatochromic isocyanonaphthalene dyes as ligands for silver (I) complexes, their applicability in silver (I) detection and background reduction in biolabelling</i></p> <p><i>Sensors and Actuators B: Chemical, 255, 2555-2567 (2018)</i></p> <p><i>Authors: Miklós Nagy, Dávid Rácz, Zsolt László Nagy, Tibor Nagy, Péter Pál Fehér, Mihály Purgel, Miklós Zsuga, Sándor Kéki</i></p> <p><i>Title: An acrylated isocyanonaphthalene based solvatochromic click reagent: Optical and biolabeling properties and quantum chemical modeling.</i></p> <p><i>Dyes and Pigments, 133, 445-457 (2016)</i></p> <p><i>Authors: Miklós Nagy, Dávid Rácz, Sándor Lajos Kovács, László Lázár, Péter Pál Fehér, Mihály Purgel, Miklós Zsuga, Sándor Kéki</i></p> <p><i>Title: New blue light-emitting isocyanobiphenyl based fluorophores: Their solvatochromic and biolabeling properties</i></p> <p><i>Journal of Photochemistry and Photobiology A: Chemistry, 318, 124-134 (2016)</i></p> <p><i>Authors: Miklós Nagy, Dávid Rácz, László Lázár, Mihály Purgel, Tamás Ditrói, Miklós Zsuga, Sándor Kéki</i></p>

	<p><i>Title:</i> Solvatochromic study of highly fluorescent alkylated-isocyanonaphthalenes, their pi-stacking, hydrogen bonding complexation and quenching with pyridine. ChemPhysChem, 15, 3614-3625 (2014).</p> <p><i>Authors:</i> Dávid Rácz, Miklós Nagy, Attila Mándi, Miklós Zsuga, Sándor Kéki</p> <p><i>Title:</i> Solvatochromic properties of a new isocyanonaphthalene based fluorophore Journal of Photochemistry and Photobiology A: Chemistry, 270, 19-27 (2013).</p>
Activities in specialist bodies over the last 5 years	

Name	<i>Noémi M. Nagy</i>
Position	<i>professor, Nuclear and radiochemistry</i>
Academic career	<p><i>2009 DSc, Hungarian Academy of Sciences</i> <i>2003 habilitation, environmental science, University of Debrecen</i> <i>1993 candidate of chemical science, Hungarian Academy of Sciences, Budapest</i> <i>1987 PhD, Kossuth Lajos University, Debrecen</i> <i>1984 Master of Science (MSc) as a chemist and English-Hungarian translator, Kossuth Lajos University, Debrecen</i></p>
Employment	<p><i>2013- Professor</i> <i>2004- Head of Isotope Laboratory</i> <i>2004-2013 Associate professor, University of Debrecen, Dep. Colloid and Environmental Chemistry, Isotope Laboratory</i> <i>1995-2004 Senior research fellow, Kossuth Lajos University, Isotope Laboratory, Debrecen</i> <i>1988-1994 Research fellow, Kossuth Lajos University, Isotope Laboratory, Debrecen</i> <i>1985-1988 Junior research fellow, Kossuth Lajos University, Isotope Laboratory, Debrecen</i> <i>1984 - 1985 Junior research fellow, Kossuth Lajos University, Dep. Colloid Chemistry, Debrecen</i></p>
Research and development projects over the last 5 years	<p>1. EU and co-financed by the European Regional Development Fund under the project GINOP-2.3.2-15-2016-00008, 2016-2020. (within the Chemical Institute)</p> <p>2. Hungarian National Research, Development, and Innovation Office (NKFIH K 120265), Study of phosphate transport of soils by radioisotopic labeling, 2016-2020. 48 MFt</p>

Industry collaborations over the last 5 years	-
Patents and proprietary rights	-
Important publications over the last 5 years	<p>1. Kónya J., N.M. Nagy: <i>Nuclear and radiochemistry</i>, Elsevier, Oxford, 2nd edition, 2018, eBook ISBN: 9780128136447, Paperback ISBN: 9780128136430, pp 480 o.</p> <p>2. Kovács E.M. , Baradács E.E. , Kónya P. , Kovács-Pálffy P. , Harangi S. , Kuzmann E. , Kónya J. , Nagy N.M. <i>Preparation and structure's analyses of lanthanide (Ln) -exchanged bentonites</i> COLLOIDS AND SURFACES A : PHYSICOCHEMICAL AND ENGINEERING ASPECTS 522: pp. 287-294. (2017)</p> <p>3. Nagy N.M. , Kovács E.M. , Kónya J. <i>Ion exchange isotherms in solid: electrolyte solution systems</i> JOURNAL OF RADIOANALYTICAL AND NUCLEAR CHEMISTRY 308:(3) pp. 1017-1026. (2016)</p> <p>4. Kuzmann E. , Garg V.K. , Singh H. , de Oliveira A.C. , Pati S.S. , Homonnay Z. , Rudolf M. , Molnár Á.M. , Kovács E.M. , Baranyai E. , Kubuki S. , Nagy N.M. , Kónya J. <i>Mössbauer study of pH dependence of iron-intercalation in montmorillonite</i> HYPERFINE INTERACTIONS 237:(1) Paper 106. (2016)</p> <p>5. Kónya J. , Nagy N.M. <i>Determination of water-soluble phosphate content of soil using heterogeneous exchange reaction with ³²P radioactive tracer</i> SOIL & TILLAGE RESEARCH 150: pp. 171-179. (2015)</p>
Activities in specialist bodies over the last 5 years	<p><i>Working Committee of Isotope Application, Hungarian Academy of Sciences, chair, 2018-</i></p> <p><i>EuCheMS Division of Nuclear and Radiochemical Division, Hungarian representative 2016-</i></p> <p><i>Bolyai Grant Committee, Hungarian Academy of Sciences, member of jury, 2014-</i></p>

Name	Levente Novák
Position	assistant professor
Academic career	<p><i>Initial academic appointment: Research Associate, University of Debrecen, 1998-2006</i></p> <p><i>Assistant Professor from 2006</i></p>
	<p><i>doctoral qualification: PhD in Molecular and Cellular Biology, Genetics and Biotechnology, Institut National</i></p>

	<p>des Sciences Appliquées (INSA) de Toulouse, France, 1998</p> <p><i>Doctorate, subject (Institution, year):</i> The metabolic network of <i>Lactococcus lactis</i> NCDO2118: interactions between carbon and nitrogen metabolism, INSA de Toulouse, France, 1994-1997</p> <p><i>Undergraduate degree, subject:</i> DEA in Microbiology, INSA, Toulouse, France, 1994</p> <p>Biologist (specialization in Biotechnology), Lajos Kossuth University (later University of Debrecen), 1993</p>
Employment	<p>Natural History Museum, Budapest, Hungary, 1986-1987</p> <p>Fontesz Ltd., Budapest, Hungary, 1997-1998</p> <p>University of Debrecen, 1998-present</p>
Research and development projects over the last 5 years	<p>1. GINOP-2.3.2- 15-2016- 00008, EU project: Chemistry for better life: strategic R&D centre at the University of Debrecen</p> <p>Period and any other information: 2016-2020</p> <p>Partners, if applicable:</p> <p>Amount of financing: 1 983 995 445 HUF</p> <p>2. Hungarian-Chinese Bilateral Cooperation (2013-2015, Hungarian-Chinese TÉT-2012)</p> <p>Period and any other information: 2013-2015</p> <p>Partners, if applicable: Donghua University, Sghanghai, China</p> <p>Amount of financing: 5 464 000 HUF</p> <p>3. TÁMOP-4.2.2.A-11/1/KONV-2012-0043, EU project: ENVIKUT - Basic Chemical and Biological Research for Elimination of Environmentally Hazardous Chemicals</p> <p>Period and any other information: 2013-2015</p> <p>Partners, if applicable: Nuclear Research Institute of the Hungarian Academy of Sciences</p> <p>Amount of financing: 899 600 000 HUF</p>
Industry collaborations over the last 5 years	<p>TEVA Ltd., Hungary, 2017. Rheological measurements</p>
Patents and proprietary rights	<p>1. Borbély J., Fleischer É., Novák L., Borbély Zs.: Hydrogels from biopolymers. US20050238678A1, Application number: US11074314, Publication date: 2005-10-27</p>

	<p>2. Novák L., Fleischer É., Borbély J.: Nanoparticles from biopolymers. US20060246096A1, Application number: US11406208, Publication date: 2006-11-02</p> <p>3. Novák L., Serra B., Bányai I.: Széles pH tartományban stabilis emulgeátorok és emulgeátorkeverékek, eljárás azok előállítására és alkalmazásuk: Process for preparation and application of emulgeator and mixture of it appliea in a wide range of pH. Lajstromszám: HU2012000589A2, Ügyszám: P1200589, Benyújtás éve: 2012., Közzététel éve: 2014, Benyújtás helye: Magyarország</p>
<p>Important publications over the last 5 years</p>	<p><i>Selected recent publications from a total of 19:</i></p> <p>1. Kéri M., Palcsu L., Túri M., Heim E., Czébely A., <u>Novák L.</u>, Bányai I.: ¹³C NMR analysis of cellulose samples from different preparation methods. <i>Cellulose</i> 22 (4): 2211-2220. (2015)</p> <p>2. Kerékgyártó M., Járvás G., <u>Novák L.</u>, Guttman A.: Activation energy associated with the electromigration of oligosaccharides through viscosity modifier and polymeric additive containing background electrolytes. <i>Electrophoresis</i> 37 (4): 573-578 (2016)</p> <p>3. Németh Z., Molnár Á. P., Fejes B., <u>Novák L.</u>, Karaffa L., Keller N. P., Fekete E.: Growth-phase sterigmatocystin formation on lactose is mediated via low specific growth rates in <i>Aspergillus nidulans</i>. <i>Toxins</i> 8 (12): 354 14 p. (2016)</p> <p>4. Tóth I. Y., Nesztor D., <u>Novák L.</u>, Illés E., Szekeres M., Szabó T., Tombácz E.: Clustering of carboxylated magnetite nanoparticles through polyethylenimine: Covalent versus electrostatic approach. <i>J. Magnetism Magn. Mater.</i> 427: 280-288 (2017)</p> <p>5. Kéri M., Nagy Z., <u>Novák L.</u>, Szarvas E., Balogh P L., Bányai I.: Beware of phosphate: evidence of specific dendrimer— phosphate interactions. <i>Phys. Chem. Chem. Phys.</i> 19: 11540-11548 (2017)</p>
<p>Activities in specialist bodies over the last 5 years</p>	

Name	<i>Mihály Purgel PhD</i>
Position	<i>Senior lecturer; Teaching in Chemistry BSc, Chemistry MSc, Chemical Engineering BSc and Chemical Engineering MSc</i>
Academic career	<i>Research fellow (MTA-DE Homogeneous Catalysis and Reaction Mechanisms Research Group, 2014) Assistant research fellow (MTA-DE Homogeneous Catalysis Research Group, 2009)</i>
	<i>PhD, Structural and dynamic study of the aluminium(III)-, thallium(III)- and lanthanide(III) complexes using DFT calculations and NMR spectroscopy (University of Debrecen, 2011) Chemistry MSc, Study of the intramolecular rearrangement of aluminium(III)-edta by NMR and DFT calculations; Study of Pt-Tl compounds with direct metal-metal bond (University of Debrecen, 2006)</i>
Employment	<i>Senior lecturer – University of Debrecen – 2017 - present</i>
Research and development projects over the last 5 years	<i>Name of project or research focus: GINOP-2.3.2-15-2016-00008 Period and any other information: 2016 – 2020 Partners, if applicable: - Amount of financing: - Name of project or research focus: TÁMOP-4.2.2.A-11/1/KONV Period and any other information: 2013 –2015 Partners, if applicable: - Amount of financing: -</i>
Important publications over the last 5 years	<i>Selected recent publications from a total of approx. (give total number): 20 Author(s): P. P. Fehér, H. Horváth, F. Joó, <u>M. Purgel</u> Title: DFT Study on the Mechanism of Hydrogen Storage Based on the Formate-Bicarbonate Equilibrium Catalyzed by an Ir-NHC Complex: An Elusive Intramolecular C–H Activation Any other information: Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: ACS, Inorg. Chem. 2018, 57:(10), 5903-5914.</i>

	<p><i>Author(s): N. Marozsán, H. Horváth, É. Kováts, A. Udvardy, A. Erdei, <u>M. Purgel</u>, F. Joó</i></p> <p><i>Title: Catalytic racemization of secondary alcohols with new (arene)Ru(II)-NHC and (arene)Ru(II)-NHC-tertiary phosphine complexes</i></p> <p><i>Any other information:</i></p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: Elsevier, Mol. Catal. 2018, 445, 248-256.</i></p> <p><i>Author(s): T. Fodor, I. Banyai, A. Benyei, C. Platas-Iglesias, <u>M. Purgel</u>, G.L. Horvath, L. Zekany, G. Tircso, I. Toth</i></p> <p><i>Title: An Extraordinarily Robust Macrocyclic Complex</i></p> <p><i>Any other information:</i></p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: ACS, Inorg. Chem. 2015, 54:(11), 5426-5437</i></p> <p><i>Author(s): P. P. Fehér, F. Joó, <u>M. Purgel</u></i></p> <p><i>Title: Performance of exchange-correlation functionals on describing ground state geometries and excitations of Alizarin Red S: Effect of complexation and degree of deprotonation</i></p> <p><i>Any other information:</i></p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: Elsevier, Comp. Theor. Chem. 2014, 1045, 113-122.</i></p>
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Name	Dávid Rácz, PhD
Position	Assistant Professor
Academic career	-
	<p><i>doctoral qualification: PhD, University of Debrecen, Hungary, 2015</i></p> <p><i>Doctorate, subject: University of Debrecen, Synthesis and characterization of light-emitting compounds</i></p> <p><i>Undergraduate degree, subject: Chemistry MSc</i></p> <p><i>Qualification: Chemist, 2010</i></p>
Employment	Teaching Assistant, University of Debrecen, Department of Applied Chemistry 2013-2017

	<i>Assistant Professor, University of Debrecen, Department of Applied Chemistry 2017-</i>
Research and development projects over the last 5 years	<p><i>Name of projects:</i> participant in GINOP-2.3.2-15-2016-00041, EU project: Regional Material Science Excellence Workshop - Research Program and Infrastructure <i>Period: 2016-</i></p> <p>participant in TÁMOP-4.2.2.A-11/1/KONV-2012-0036, EU project: Intelligent Functional Materials: Their mechanical, thermal, electromagnetic, optical properties and their applications, <i>Period: 2013-2015</i></p> <p>participant in NKFI K-116465, Hungarian project: Synthesis and characterization of smart fluorescent polymers <i>Period: 2016-</i></p>
Industry collaborations over the last 5 years	-
Patents and proprietary rights	-
Important publications over the last 5 years	<p><i>Selected recent publications from a total of approx. 11</i> <i>Authors:</i> Miklós Nagy, Dávid Rácz, Zsolt László Nagy, Péter Pál Fehér, József Kalmár, István Fábíán, Alexandra Kiss, Miklós Zsuga, Sándor Kéki <i>Title:</i> Solvatochromic isocyanonaphthalene dyes as ligands for silver (I) complexes, their applicability in silver (I) detection and background reduction in biolabelling Sensors and Actuators B: Chemical, 255, 2555-2567 (2018)</p> <p><i>Authors:</i> Miklós Nagy, Dávid Rácz, Zsolt László Nagy, Tibor Nagy, Péter Pál Fehér, Mihály Purgel, Miklós Zsuga, Sándor Kéki <i>Title:</i> An acrylated isocyanonaphthalene based solvatochromic click reagent: Optical and biolabeling properties and quantum chemical modeling. Dyes and Pigments, 133, 445-457 (2016)</p> <p><i>Authors:</i> Miklós Nagy, Dávid Rácz, Sándor Lajos Kovács, László Lázár, Péter Pál Fehér, Mihály Purgel, Miklós Zsuga, Sándor Kéki <i>Title:</i> New blue light-emitting isocyanobiphenyl based</p>

	<p>fluorophores: Their solvatochromic and biolabeling properties Journal of Photochemistry and Photobiology A: Chemistry, 318, 124-134 (2016)</p> <p><i>Authors:</i> Miklós Nagy, Dávid Rácz, László Lázár, Mihály Purgel, Tamás Ditrói, Miklós Zsuga, Sándor Kéki <i>Title:</i> Solvatochromic study of highly fluorescent alkylated-isocyanonaphthalenes, their pi-stacking, hydrogen bonding complexation and quenching with pyridine. ChemPhysChem, 15, 3614-3625 (2014).</p> <p><i>Authors:</i> Dávid Rácz, Miklós Nagy, Attila Mándi, Miklós Zsuga, Sándor Kéki <i>Title:</i> Solvatochromic properties of a new isocyanonaphthalene based fluorophore Journal of Photochemistry and Photobiology A: Chemistry, 270, 19-27 (2013).</p>
Activities in specialist bodies over the last 5 years	-

Sheet of staff handbook

Name	László SOMSÁK
Position	Full professor, Head of the Department of Organic Chemistry, Responsible for Chemistry BSc
Academic career	University of Debrecen (UD) 1978 Habilitation (UD, 1997) Doctor of the Hungarian Academy of Sciences 2002
	doctoral qualification (UD, 1983) Undergraduate degree, subject (UD, 1978, MSc in Chemistry)
Employment	Full professor at the Department of Organic Chemistry of UD since 2003, head of the department since 2015
Research and development projects over the last 5 years	2013-2017 Alexander von Humboldt Stiftung – Institutspartnerschaft (Universität Rostock, Prof. P. Langer – University of Debrecen, Prof. L. Somsák) – (53.990 Euro) Untersuchungen zur Anwendbarkeit moderner Synthesemethoden

	<p>für die Herstellung pharmakologisch relevanter Kohlenhydrat-Derivate</p> <p>2013-2018</p> <p>OTKA K-109450 – PI of research project (43.852 kHUF) New transformations of monosaccharide derivatives at and around the anomeric centre</p> <p>2016-2020</p> <p>GINOP-2.3.2-15-2016-00008 – participant (1.984 MHUF) Chemistry for better life quality</p> <p>2017-2021</p> <p>OTKA FK-125067 – participant (35.664 kHUF) New types and new biological applications of C-glycopyranosyl heterocycles</p>
<p>Industry collaborations over the last 5 years</p>	<p><i>Project title: Custom synthesis of carbohydrate derivatives</i> <i>Partners: Carbosynth Ltd, Oxford, UK</i></p> <p><i>Project title: Custom synthesis of cyclodextrin derivatives</i> <i>Partners: Cyclolab Ltd, Budapest, Hungary</i></p>
<p>Patents and proprietary rights</p>	<p>SOMSÁK L.; BOKOR É.; TÓTH M.; JUHÁSZ L.; CZIFRÁK K.; KÓNYA B.; KUN S.; PÁHI A.; SZŐCS B.; VARGA G.; KÓDER L.-NÉ; NAGY K.-NÉ; GERGELY P.; DOCSA T. Glikogén foszforiláz inhibitorok P1100602 alapszámú magyar szabadalmi bejelentés, 2011. PCT/HU2012/00116 WO2013061105 A3</p> <p>PATONAY T.; SZILVÁSSY Z.; PEITL B.; VÍGH L.; TÖRÖK Zs.; BERÉNYI S.; HORVÁTH I.; BALOGH G.; SOMSÁK L.; KISS A.; DRIMBA L.; MOLNÁRNÉ HATVANI I. Hő-sokk protein (HSP) ko-inducer sajátosságát mutató vegyületek alkalmazása 2-es típusú cukorbetegség kezelésében inzulin-érzékenyítő hatás fokozására P1100709 alapszámú magyar szabadalmi bejelentés, 2011.</p> <p>ANTUS, S.; SOMSÁK, L.; CZAKÓ, Z.; KÓNYA, B.; GERGELY, P.; DOCSA, T.; KÓDER, L.-NÉ; MAGYAR, L.-NÉ Glükóz és</p>

	<p>oxigéntartalmú heterociklusok antidiabetikus hatású konjugátumai P1200626 alapszámú magyar szabadalmi bejelentés, 2012.</p>
<p>Important publications over the last 5 years</p>	<p><i>Selected recent publications from a total of 167:</i></p> <p>KUN, S.; BEGUM, J.; KYRIAKIS, E.; STAMATI, E. C. V.; BARKAS, T. A.; SZENNYES, E.; BOKOR, É.; SZABÓ, E. K.; STRAVODIMOS, G. A.; SIPOS, Á.; DOCSA, T.; GERGELY, P.; MOFFATT, C.; PATRASKAKI, M. S.; KOKOLAKI, M. S.; GKERDI, A.; SKAMNAKI, V. T.; LEONIDAS, D. D.; SOMSÁK, L.; HAYES, J. M. A multidisciplinary study of 3-(β-D-glucopyranosyl)-5-substituted-1,2,4-triazole derivatives as glycogen phosphorylase inhibitors: computation, synthesis, crystallography and kinetics reveal new potent inhibitors <i>Eur. J. Med. Chem.</i>, 2018, <i>147</i>, 266-278.</p> <p>BOKOR, É.; KYRIAKIS, E.; SOLOVOU, T. G. A.; KOPPÁNY, Cs.; KANTSADI, A. L.; SZABÓ, K. E.; SZAKÁCS, A.; STRAVODIMOS, G. A.; DOCSA, T.; SKAMNAKI, V. T.; ZOGRAPHOS, S. E.; GERGELY, P.; LEONIDAS, D. D.; SOMSÁK, L. Nanomolar inhibitors of glycogen phosphorylase based on β-D-glucosaminyl heterocycles: a combined synthetic, enzyme kinetic and protein crystallography study <i>J. Med. Chem.</i>, 2017, <i>60</i>, 9251-9262.</p> <p>KASZÁS, T.; TÓTH, M.; SOMSÁK, L. A new synthesis of C-β-D-glycopyranosylmethyl sulfides by metal-free coupling of anhydro-aldose tosylhydrazones with thiols <i>New J. Chem.</i>, 2017, <i>41</i>, 13871-13880.</p> <p>BOKOR, É.; KUN, S.; GOYARD, D.; TÓTH, M.; PRALY, J.-P.; VIDAL, S.; SOMSÁK, L. C-Glycopyranosyl arenes and hetarenes: Synthetic methods and bioactivity focused on antidiabetic potential <i>Chem. Rev.</i>, 2017, <i>117</i>, 1687-1764.</p> <p>BOKOR, É.; KUN, S.; DOCSA, T.; GERGELY, P.; SOMSÁK, L. 4(5)-Aryl-2-C-glucopyranosyl-imidazoles as new nanomolar glucose analog inhibitors of glycogen phosphorylase <i>ACS Med. Chem. Lett.</i>, 2015, <i>6</i>, 1215-1219.</p>
<p>Activities in specialist bodies over the last 5 years</p>	<p><i>European Chemistry Thematic Network (representative of the University of Debrecen) since 1999</i> <i>Chair of the Working Committee on Carbohydrate Chemistry of the HAS since 2006</i></p>

	<p><i>European Carbohydrate Organization (Hungarian representative) since 2006</i></p> <p><i>Chair of Organic Chemistry Panel of the Hungarian Scientific Research Fund since 2016</i></p>
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Name	<i>Károly Teperics PhD</i>
Position	<i>Senior lecturer, in charge of the following courses: History and Structure of the EU</i>
Academic career	<i>Senior lecturer (UD, Department of Social Geography and Regional Development Planning, 2002)</i>
	<p><i>doctoral qualification: PhD, University of Debrecen, Hungary, 2002</i></p> <p><i>Doctorate, subject: Labour market situation of the diploma holders in Hajdú-Bihar county: the impact of the University of Debrecen on the human resources. - University of Debrecen, 2002</i></p> <p><i>Undergraduate degree, subject:</i></p> <p><i>(1) Teacher of Geography, Kossuth Lajos University, Debrecen, 1987</i></p> <p><i>(2) Teacher of History, Kossuth Lajos University, Debrecen, 1987</i></p>
Employment	<p><i>assistant professor – Department of Social Geography and Regional Development Planning, University of Debrecen – 1993-2002</i></p> <p><i>senior lecturer – Department of Social Geography and Regional Development Planning, University of Debrecen – 2002-</i></p>
Research and development projects over the last 5 years	<p><i>Name of project or research focus: HURO/1001/184/2.3.1. “Cross Border Doctoral Programs Consortium”.</i></p> <p><i>Period and any other information: 2012-2013</i></p> <p><i>Partners, if applicable:</i></p> <p><i>Amount of financing:</i></p> <p><i>Name of project or research focus: OTKA K-101867 “LeaRn: Tanuló régiók Magyarországon: Az elmélettől a valóságig”. [“LeaRn: Learning regions in Hungary: from theory to reality”].</i></p> <p><i>Period and any other information: 2012-2015</i></p> <p><i>Partners, if applicable:</i></p> <p><i>Amount of financing:</i></p> <p><i>Name of project or research focus: MTA-TpKP. “MTA-SZTE Geography Methodology Research Group”.</i></p> <p><i>Period and any other information: 2016-2020</i></p>

	<p><i>Partners, if applicable:</i></p> <p><i>Amount of financing:</i></p>
Industry collaborations over the last 5 years	<p><i>Project title: -</i></p> <p><i>Partners: -</i></p>
Patents and proprietary rights	<p><i>Title (Year) -</i></p>
Important publications over the last 5 years	<p><i>Selected recent publications from a total of approx. (give total number): 44</i></p> <p><i>Author(s): Károly TEPERICS, Klára CZIMRE, Sándor MÁRTON</i></p> <p><i>Title: A tanulóvárosok és régiók megjelenése és társadalmi-gazdasági mutatókkal való kapcsolata Magyarországon.</i></p> <p><i>Any other information:</i></p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: EDUCATIO 25:(2) pp. 245-259. (2016)</i></p> <p><i>Author(s): Tamás KOZMA, Károly TEPERICS</i></p> <p><i>Title: Tanuló régiók Magyarországon [Learning regions in Hungary]</i></p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: In: Régió és oktatás; 11. (Eds: Tamás Kozma, Magdolna Benke, Gábor Erdei, Károly Teperics, Zoltán Tózsér, Ágnes Engler, Veronika Bocsi, Ágnes Dusa, Katalin Kardós, Nóra Veronika Németh, Zoltán Györgyi, Erika Juhász, Edina Márkus, Barbara Szabó, Judt Herczegh, Attila Kenyeres Zoltán, Klára Kovács, Józsed Szabó, Tímea Szűcs, Katalin Forray R., Tibor Cserti Csapó, Borbála Heltai, Tamás Híves, Klára Szilágyiné Czimre, Sándor Márton): Tanuló régiók Magyarországon: Az elmélettől a valóságig. Debrecen: University of Debrecen, CHERD, 2015. pp. 225-294. (ISBN:<u>978-963-318-472-1</u>)</i></p> <p><i>Author(s): Gyula DÉZSI, Gábor KOZMA, Károly TEPERICS</i></p> <p><i>Title: Spatial distribution and development of educational institutions in different historical periods: a Hungarian case study</i></p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: HUMAN GEOGRAPHIES: JOURNAL OF STUDIES AND RESEARCH IN HUMAN GEOGRAPHY 8:(2) pp. 89-</i></p>

	<p>99. (2014)</p> <p><i>Author(s): Gábor ERDEI, Károly TEPERICS</i> <i>Title: Adult learning activities as the catalyst for creating learning region</i> <i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: PROCEDIA - SOCIAL AND BEHAVIORAL SCIENCES 142: pp. 359-366. (2014)</i></p> <p><i>Author(s): Gábor KOZMA, Károly TEPERICS, Zsolt RADICS</i> <i>Title: The changing role of sports in urban development: A case study of Debrecen (Hungary)</i> <i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: INTERNATIONAL JOURNAL OF THE HISTORY OF SPORT 31:(9) pp. 1118-1132. (2014)</i></p>
Activities in specialist bodies over the last 5 years	<p><i>Organisation - Role – Period</i> <i>Membership without a specific role need not be mentioned.</i></p>

Name	<i>Mária Ujhelyi PhD</i>
Position	<i>Associate Professor, head of department, in charge of Leadership and Organization Master programme</i>
Academic career	<i>Head of Department (UD, Faculty of Economics and Business, Institute of Management and Organizational Sciences, Department of Management Sciences 2016-)</i>
	<p><i>2003 PhD Budapest University of Economic Sciences and Public Administration</i> <i>1993 MBA Agricultural University of Debrecen, in cooperation with four EU universities</i> <i>1982 Diploma in agricultural sciences Agricultural University of Debrecen</i></p>
Employment	<p><i>2004- Associate professor – University of Debrecen</i> <i>1999-2003 Assistant professor – University of Debrecen Faculty of Economics and Business Administration</i> <i>1993-1999 Lecturer – Lajos Kossuth University, Department of Economics</i> <i>1989-1993 Research administrator, secretary of the MBA programme – Agricultural University of Debrecen</i> <i>1983-1989 Scientific Training Scholarship – Scientific Qualifying Committee</i></p>

	<i>1982-1983 Assistant – Training Farm of the Agricultural University Debrecen</i>
Research and development projects over the last 5 years	<p><i>Name of project or research focus: EFOP-3.5.1-16-2017-00007 The Development of Dual Training Programmes at Debrecen University</i></p> <p><i>Period and any other information: 2017-2018 participant</i></p> <p><i>Name of project or research focus: Hungary Romania Cross-Border Co-operation Programmes</i></p> <p><i>Period and any other information: 2011-2012 Coordinator and participant</i></p>
Industry collaborations over the last 5 years	<p><i>Project title: -</i></p> <p><i>Partners: -</i></p>
Patents and proprietary rights	<i>Title (Year) -</i>
Important publications over the last 5 years	<p><i>Selected recent publications from a total of approx. (give total number): 23</i></p> <p><i>Author(s): Edit BARIZSNÉ HADHÁZI, Roland FILEP, Mária UJHELYI:</i></p> <p><i>Title: Business Students' Attitudes Toward Corporate Social Responsibility</i></p> <p><i>Any other information:</i></p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: ANNALS OF THE UNIVERSITY OF ORADEA ECONOMIC SCIENCE 28:(1) p. 20. (2018)</i></p> <p><i>Author(s): UJHELYI Mária, KUN András István:</i></p> <p><i>Title: A változásmenedzsment és a tulajdonosi szerkezet összefüggései - empirikus tapasztalatok</i></p> <p><i>Any other information:</i></p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: VEZETÉSTUDOMÁNY 48:(1) pp. 18-29. (2017)</i></p> <p><i>Author(s): UJHELYI Mária, MIHALI Mária, KUN András István, KISS Attila:</i></p> <p><i>Title: A szervezeti kultúra és a szubkultúrák vizsgálata összeolvadás után egy gazdaságtudományi karon</i></p> <p><i>Any other information:</i></p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers:</i></p>

	<p><i>ISKOLAKULTÚRA: PEDAGÓGUSOK SZAKMAI-TUDOMÁNYOS FOLYÓIRATA 26:(9) pp. 23-50. (2016)</i></p> <p><i>Author(s): Maria-Madela ABRUDAN, Tomina Gabriela SĂVEANU, Mirabela MATEI, Mária UJHELYI:</i></p> <p><i>Title: OBSTACLES, REALITIES AND OPPORTUNITIES IN HUMAN RESOURCES MANAGEMENT IN PUBLIC ADMINISTRATION INSTITUTIONS FROM BIHOR COUNTY (ROMANIA) AND HAJDÚ-BIHAR COUNTY (HUNGARY)</i></p> <p><i>Any other information:</i></p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: TRANSYLVANIAN REVIEW OF ADMINISTRATIVE SCIENCES 11:(2) pp. 5-25. (2015)</i></p> <p><i>Author(s): Ujhelyi Mária, Barizsné Hadházi Edit, Kun András István:</i></p> <p><i>Title: ANALYSING ORGANIZATIONAL CHANGES – THE CONNECTION BETWEEN THE SCALE OF CHANGE AND EMPLOYEES’ ATTITUDES</i></p> <p><i>Any other information:</i></p> <p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: ANNALS OF THE UNIVERSITY OF ORADEA ECONOMIC SCIENCE 24:(1) pp. 1191-1198. (2015)</i></p>
<p>Activities in specialist bodies over the last 5 years</p>	<p><i>Organisation - Role – Period</i></p> <p><i>2018- Member of the Management Science Committee of The Hungarian Academy of Science.</i></p>

Name	Balázs Ujvári
Position	Assistant professor
Academic career	<p>PhD physics, University of Debrecen, 2013</p> <p>BA Hungarian Language and Literature, University of Debrecen, 2013</p> <p>MSc Economics, University of Debrecen, 2007</p> <p>BSc Computer Science, University of Debrecen, 2006</p> <p>MSc Physics, University of Debrecen, 2000</p>
Employment	<p>Assistant professor - University of Debrecen - 2014 -</p> <p>Assistant lecturer - University of Debrecen - 2008-</p>

	<p>2014</p> <p>Scientific associate – ATOMKI - 2007-2008</p> <p>Database programmer – NI Hungary - 2004 - 2010</p> <p>Position - Employer - Period</p>
Research and development projects over the last 5 years	<p>CERN-CMS alignment 2008-</p> <p>CERN-CMS DT HV development 2017-</p> <p>BNL-PHENIX data analysis 2014-</p> <p>BNL-sPHENIX SiPM tester 2016-</p>
Industry collaborations over the last 5 years	
Patents and proprietary rights	
Important publications over the last 5 years	<p>Selected recent publications from a total of approx. (give total number): 900</p> <p>Motion of CMS detector and mechanical structures during Magnet Cycles and Stability Periods from 2008 to 2013 as observed by the Link Alignment System NUCLEAR INSTRUMENTS & METHODS IN PHYSICS RESEARCH SECTION A-ACCELERATORS SPECTROMETERS DETECTORS AND ASSOCIATED EQUIPMENT 813: pp. 36-55.(2016)</p> <p>Searches for a heavy scalar boson H decaying to a pair of 125 GeV Higgs bosons hh or for a heavy pseudoscalar boson A decaying to Zh, in the final states with $h \rightarrow \tau\tau$ PHYSICS LETTERS B 755: pp. 217-244. (2016)</p> <p>Particle physics education in Hungary NUCLEAR AND PARTICLE PHYSICS PROCEEDINGS 273-275: pp. 2569-2571. (2016)</p> <p>PHENIX results on low pT direct photons in Au + Au collisions NUCLEAR PHYSICS A 956: pp. 417-420. (2016)</p> <p>The XXV International Conference on Ultrarelativistic Nucleus-Nucleus Collisions: Quark Matter 2015. Kobe, Japan: 2015.09.27 -2015.10.03.</p> <p>Weighing the Neutrino INTERNATIONAL JOURNAL OF MODERN PHYSICS E-NUCLEAR PHYSICS 23:(1) Paper 1450004. 12 p. (2014)</p>
Activities in specialist bodies over the last 5 years	CERN – Knowledge Transfer Office – Hungarian deputy, 2017-

Name	<i>Marietta Vágvölgyi Dr. Tóth, PhD</i>
Position	<i>Associate Professor</i>
Academic career	Habilitation (<i>University of Debrecen, 2018</i>)
	PhD, <i>University of Debrecen, Hungary, 2002</i>

	<p>PhD, subject: <i>Access to C-glycosyl-imine type compounds and C-glycosyl-methylene carbenes and investigation of their properties</i></p> <p>Undergraduate degree, subject: <i>chemistry - chemistry teacher – special translator in English, MSc, Kossuth Lajos University, Debrecen, Hungary, 1998</i></p>
Employment	<p>Associate professor, <i>Department of Pharmaceutical Chemistry, Faculty of Pharmacy, University of Debrecen, 2018. 09. 01. –</i></p> <p>Senior lecturer, <i>Faculty of Pharmacy, University of Debrecen, 2012. 02. 01. – 2018. 08. 31.</i></p> <p>Senior lecturer, <i>Organic Chemistry Department, University of Debrecen, 2008. 10. 01. – 2012. 01. 31.</i></p> <p>OTKA Postdoctor, <i>OTKA postdoctoral research fellowship (OTKA PD45927) supported by the Hungarian Scientific Research Fund, Organic Chemistry Department, University of Debrecen, 2003. 10. 01. – 2008. 09. 30.</i></p> <p>Assistant research fellow, <i>Organic Chemistry Department, University of Debrecen, 2001. 11. 01. – 2003. 09. 30.</i></p>
Research and development projects over the last 5 years	<p><i>Name of project or research focus: GINOP-2.3.2-15-2016-00008, supported by the EU co-financed by the European Regional Development Fund: Chemistry for better life: strategic R&D center at the University of Debrecen, Chemistry of carbohydrates and heterocycles subproject. (Participant)</i></p> <p><i>Period and any other information: 2016-2020</i></p> <p><i>Amount of financing: 1 983 995 445 Ft</i></p> <p><i>Name of project or research focus: Research Project of University of Debrecen, 5N5XBTTDDTOMA320, supported by University of Debrecen, Cross-coupling of and glycoenzyme inhibition by hydrazones of monosaccharides (Principal investigator)</i></p> <p><i>Period and any other information: 2015-2016</i></p> <p><i>Amount of financing: 9.292.900 Ft</i></p> <p><i>Name of project or research focus: Magyar Zoltán Postdoctoral Fellowship, A2-MZPD-13-0263, supported by the European Union and the State of Hungary, co-financed by the European Social Fund in the framework</i></p>

	<p>of TÁMOP-4.2.4.A/ 2-11/1-2012-0001 'National Excellence Program', Investigation of the coupling reactions of sugar tosylhydrazones (Principal investigator)</p> <p>Period and any other information: 2014-2015</p> <p>Name of project or research focus: Grant OTKA, K-109450, supported by the National Research, Development and Innovation Office of Hungary (Hungarian Scientific Research Fund): New transformations of monosaccharide derivatives at and around the anomeric center (Participant)</p> <p>Period and any other information: 2013-2017</p> <p>Amount of financing: 43.852.000 Ft</p>
Industry collaborations over the last 5 years	<p>Project title: -</p> <p>Partners: -</p>
Patents and proprietary rights	<p>László Somsák, Éva Bokor, Marietta Vágvölgyiné Tóth, László Juhász, Katalin Czifrák, Bálint Kónya, Sándor Kun, András Páhi, Béla Szócs, Gergely Varga, Pál Gergely, Tibor Docsa, Lászlóné Kóder, Károlyné Nagy</p> <p>Title: Glycogen phosphorylase inhibitors / Glikogén foszforiláz inhibitorok</p> <p>P1100602/P1200475, 2012, Hungarian patent application</p> <p>Title: Preparation of imidazolyl and triazolyl glycosides as glycogen phosphorylase inhibitors and antitumor agents.</p> <p>WO2013061105A2, 2013, International patent application.</p>
Important publications over the last 5 years	<p>Selected recent publications from a total of approx. (give total number): 33</p> <p>Author(s): L. Somsák, É. Bokor, B. Czibere, K. Czifrák, Cs. Koppány, L. Kulcsár, S. Kun, E. Szilágyi, M. Tóth, T. Docsa, P. Gergely</p> <p>Title: Synthesis of C-xylopyranosyl- and xylopyranosylidene-spiro-heterocycles as potential inhibitors of glycogen phosphorylase</p> <p>Any other information:</p> <p>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: Carbohydr. Res. 399 (2014) 38-48</p>

Author(s): S. Kun, É. Bokor, G. Varga, B. Szócs, A. Páhi, K. Czifrák, M. Tóth, L. Juhász, T. Docsa, P. Gergely, L. Somsák

Title: New synthesis of 3-(β -D-glucopyranosyl)-5-substituted-1,2,4-triazoles, nanomolar inhibitors of glycogen phosphorylase

Any other information:

Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: Eur. J. Med. Chem. 76 (2014) 567-579

Author(s): B. Szócs, É. Bokor, K. E. Szabó, A. Kiss-Szikszai, M. Tóth, L. Somsák

Title: Synthesis of 5-aryl-3-C-glycosyl- and unsymmetrical 3,5-diaryl-1,2,4-triazoles from alkylidene-amidrazones

Any other information:

Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: RSC Advances 54 (2015) 43620-43629

Author(s): L. Lázár, M. Csávás, M. Tóth, L. Somsák, A. Borbás

Title: Thio-click approach to the synthesis of stable glycomimetics

Any other information:

Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: Chem. Pap. 69 (2015) 889-895

Author(s): É. Bokor, S. Kun, D. Goyard, M. Tóth, J.-P. Praly, S. Vidal, L. Somsák

Title: C-Glycopyranosyl Arenes and Hetarenes: Synthetic Methods and Bioactivity Focused on Antidiabetic Potential

Any other information:

Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: Chem. Rev. 117 (2017) 1687-1764

Author(s): T. Kaszás, M. Tóth, S. Kun, L. Somsák

Title: Coupling of anhydro-aldose tosylhydrazones with phenols and carboxylic acids: A new route for the synthesis of C- β -D-glycopyranosylmethyl ethers and esters

Any other information:

	<p><i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: RSC Advances 7 (2017) 10454-10462</i></p> <p><i>Author(s): T. Kaszás, M. Tóth, L. Somsák</i> <i>Title: New synthesis of C-β-D-glycopyranosylmethyl sulfides by metal-free coupling of anhydro-aldose tosylhydrazones with thiols</i> <i>Any other information:</i> <i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: New J. Chem. 41 (2017) 13871-13880</i></p> <p><i>Author(s): T. Kaszás, A. Ivanov, M. Tóth, P. Ehlers, P. Langer, L. Somsák</i> <i>Title: Pd-catalyzed coupling reactions of anhydro-aldose tosylhydrazones with aryl bromides to produce substituted exo-glycals</i> <i>Any other information:</i> <i>Publisher, place of publication, date of publication or name of periodical, volume, issue, page numbers: Carbohydr. Res. (2018) DOI: 10.1016/j.carres.2018.02.010</i></p>
<p>Activities in specialist bodies over the last 5 years</p>	<p><i>Organisation - Role – Period</i> Chemistry BSc State Exam Committee – Secretary – 2008 – 2013, 2017 –</p>